ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

Single Correct Answer Type

1. 3-phenylpropenoic acid is IUPAC name of :
a) Mendaleic acid　b) Pyruvic acid　c) Succinic acid　d) Cinnamic acid

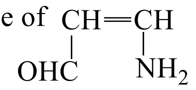
2. How many isomers are possible for the compound having molecular formula $C_3H_5Br_3$?
a) 5　b) 4　c) 6　d) 8

3. The strain in bonds of cyclopropane is :
a) 0°44′　b) 24°44′　c) 9°44′　d) 5°16′

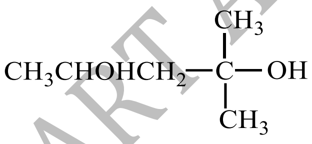
4. Chlorine in vinyl chloride is less reactive because :
a) sp^2-hybridized carbon has more acidic character than sp^3-hybridized carbon
b) C—Cl bond develops partial double bond character
c) Of resonance
d) All of the above are correct

5. The alkene that exhibits geometrical isomerism is
a) Propene　b) 2-methyl propene　c) 2-butene　d) 2-methyl-2-butene

6. Pick out the alkane which differs from the other members of the group
a) 2,2-dimethyl propane　b) Pentane　c) 2-methyl butane　d) 2, 2-dimethyl butane

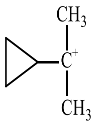
7. The IUPAC name of  $\underset{OHC}{CH}=\underset{NH_2}{CH}$ is :
a) 1-amino prop-2-enal
b) 3-amino prop-2-enal
c) 1-amino-2-formylethene
d) 3-amino-1-oxoprop-2-ene

8. Detection of sulphur in sodium extract is done by
a) Lead acetate　b) Sodium nitroprusside
c) Both (a) and (b)　d) None of these

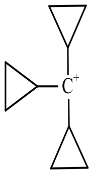
9. The IUPAC name for

$CH_3CHOHCH_2—\overset{CH_3}{\underset{CH_3}{C}}—OH$ is:
a) 1,1-dimethyl-1,2-butanediol
b) 2-methyl-2,4-pentanediol
c) 4-methyl-2,4-pentanediol
d) 1,3,3-dimethyl-1,3-propanediol

10. In the following carbocations, the stability order is :

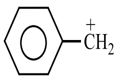
(I) $RCH_2\overset{+}{C}H_3$ (II)



(III)

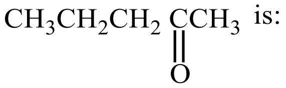


(IV)



a) III > II > IV > I b) IV > I > II > III c) IV > III > II > I d) III > IV > II > I

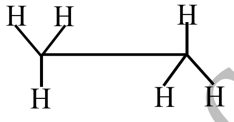
11. The shape of the π electron cloud in acetylene is
    a) Linear b) Planar c) Cylinder d) Doughtnut
12. Acidified sodium fusion extract on addition of ferric chloride solution gives blood red colouration which confirm the presence of
    a) S and Cl b) N and S c) N d) S
13. Conversion of chlorobenzene to phenol involves
    a) Electrophilic substitution b) Nucleophilic substitution
    c) Free radical substitution d) Electrophilic addition
14. In sulphur detection of an organic compound, sodium nitroprusside solution is added to sodium extract. Formation of violet colour is due to
    a) $Na_3Fe(CN)_6$ b) $Na_3[Fe(CN)_5NOS]$ c) $Fe(CNS)_3$ d) None of these
15. The maximum bond energy is present
    a) $C-H$ b) $C-C$ c) $C-N$ d) $C-O$
16. The number of secondary hydrogens in 2, 2-dimethyl butane is
    a) 8 b) 6 c) 4 d) 2
17. The name of the compound,



$CH_3CH_2CH_2\underset{\underset{O}{\|}}{C}CH_3$ is:

    a) 2-pentanone b) Pentanone-2 c) Pentan-2-one d) All are correct
18. Find the non-staggered form(s) of ethane :

a)



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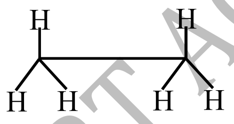

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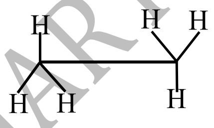
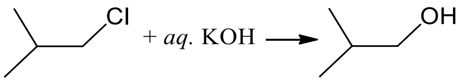


d) None of these

19. With a change in hybridisation of the carbon bearing the charge, the stability of a carbanion increase in the order
    a) $sp < sp^2 < sp^3$ b) $sp < sp^3 < sp^2$ c) $sp^3 < sp^2 < sp$ d) $sp^2 < sp < sp^3$
20. The addition reaction among the following is

a)



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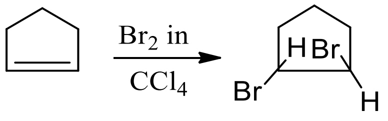
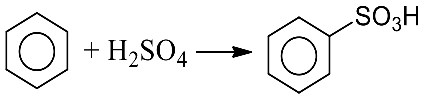
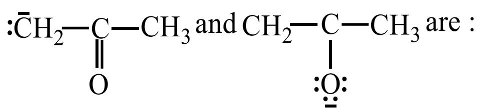


c) 

d) All of the above

21. 

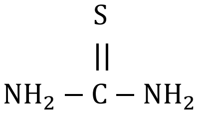
$:\bar{C}H_2-\overset{\|}{\underset{O}{C}}-CH_3$ and $CH_2-\underset{:\ddot{\underset{..}{O}}:}{C}-CH_3$ are :

a) Resonating structures
b) Tautomers
c) Geometrical isomers
d) Optical isomers

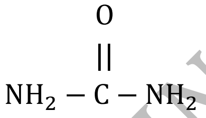
22. The correct definition for organic chemistry is :
a) Chemistry of carbon compounds
b) Chemistry of compounds derived from living organisms
c) Chemistry of hydrocarbons and their derivatives
d) None of the above

23. Which of the organic compounds will give red colour in Lassaigne test?

a) NaCNS

b) 

$NH_2-\overset{\overset{S}{\|}}{C}-NH_2$

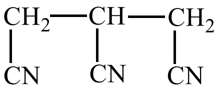
c) 

$NH_2-\overset{\overset{O}{\|}}{C}-NH_2$

d) None of these

24. The compound formed in the positive test for nitrogen with the Lassaigne solution of an organic compound is

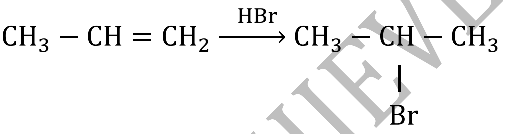
a) $Fe_4[Fe(CN)_6]_3$
b) $Na_3[Fe(CN)_6]$
c) $Fe(CN)_3$
d) $Na_4[Fe(CN)_5NOS]$

25. IUPAC name of the compound,  is :

$\underset{CN}{CH_2}-\underset{CN}{CH}-\underset{CN}{CH_2}$

a) 1,2,3-trieyanopropane
b) Propane-1,2,3-tricarbonitrile
c) 1,2,3-cyanopropane
d) Propane tricarbylamine

26. Which of the following reactions proceeds *via* secondary free radical?

a) 

$CH_3-CH=CH_2\xrightarrow{HBr}CH_3-\underset{Br}{CH}-CH_3$

b) $CH_3-CH=CH_2\xrightarrow[\text{UV light}]{HBr}CH_3-CH_2-CH_2Br$

c) $C_6H_6\xrightarrow{Br_2/FeBr_3}C_6H_5Br$

d) $C_6H_6\xrightarrow[\text{UV ligh}]{Br_2}CH_3-CH_2Br$

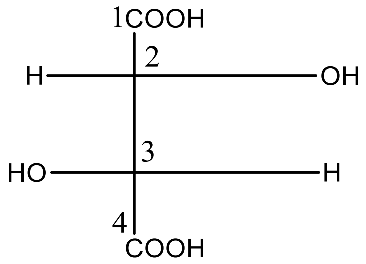
27. The production of an optically active compound from a symmetric molecule without resolution is called :
a) Walden inversion
b) Asymmetric synthesis
c) Partial racemisation
d) None of these

28. Among the following, which one has more than one kind of hybridization?
(i) $CH_3CH_2CH_2CH_3$
(ii) $CH_3CH=CHCH_3$
(iii) $CH_2=CH-CH\equiv CH$
(iv) $CH\equiv CH$

a) (ii) and (iii)
b) (ii) and (i)
c) (iii) and (iv)
d) (iv)

29. The IUPAC name of C_6H_5COCl is
a) Benzoyl chloride
b) Benzene chloro ketone
c) Benzene carbonyl chloride
d) Chloro phenyl ketone

30. In the compound,



Configuration at C_2 and C_3 atoms are

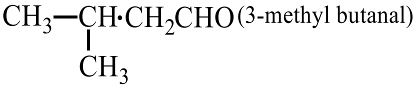
a) S, S b) R, S c) S, R d) R, R

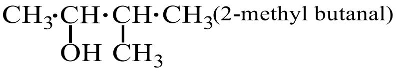
31. The number of isomeric alkenes with molecular formula C_6H_{12} are

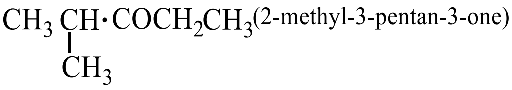
a) 8 b) 10 c) 11 d) 13

32. Which is wrong IUPAC name?

a) $CH_3CH_2CH_2COOCH_2CH_3$ (Ethyl butanoate)

b) 
$CH_3-\underset{\underset{CH_3}{|}}{CH}\cdot CH_2CHO$ (3-methyl butanal)

c) 
$CH_3\cdot\underset{\underset{OH}{|}}{CH}\cdot\underset{\underset{CH_3}{|}}{CH}\cdot CH_3$ (2-methyl butanal)

d) 
$CH_3\,\underset{\underset{CH_3}{|}}{CH}\cdot COCH_2CH_3$ (2-methyl-3-pentan-3-one)

33. Which of the following statements is wrong?

a) In general organic compounds have low m.p. and b.p.

b) Isomerism is common in organic compounds

c) Organic compounds cannot be synthesized in the laboratory

d) The number of organic compound is very large

34. Nitroethane can exhibit one of the following kind of isomerism

a) Metamerism b) Optical activity c) Tautomerism d) Position isomerism

35. Which of the following would show configurational enantiomorphism?

a) NH_3

b) $(CH_3)_3N$

c) Methyl, ethyl, propylamine

d) Methyl, allyl, phenyl, benzyl ammonium iodide

36. Heterolysis of carbon-chlorine bond produces :

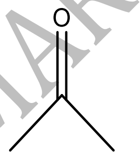
a) Two free radicals

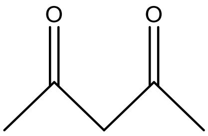
b) Two carbonium ions

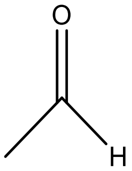
c) Two carbanions

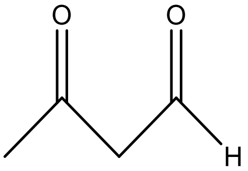
d) One cation and one anion

37. Maximum enol content is in

a) 

b) 

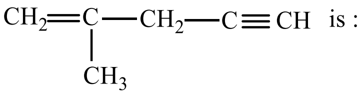
c) 

d) 

38. Which of the following compounds will show metamerism?

a) $CH_3-CO-C_2H_5$ b) $C_2H_5-S-C_2H_5$ c) CH_3-O-CH_3 d) $CH_3-O-C_2H_5$

39. The IUPAC name of the compound,

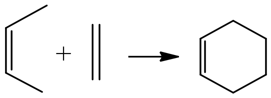

$CH_2=\underset{\underset{CH_3}{|}}{C}-CH_2-C\equiv CH$ is :

a) 2-methylpent-1-en-4-yne

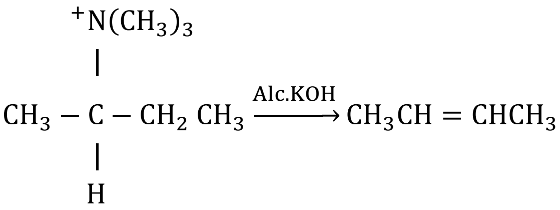
b) 4-methylpent-4-en-1-yne

c) 2-methylpent-2-en-4-yne
d) 4-methylpent-1-en-4-yne

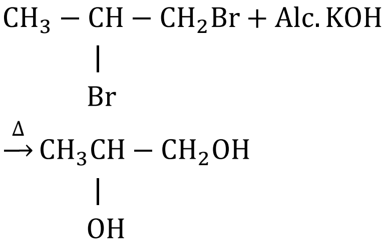
40. Which of the following is elimination reaction

a) 

b) $CH_3CH_2CH_2Cl + aq.\,KOH \longrightarrow CH_3CH_2CH_2OH$

c) 

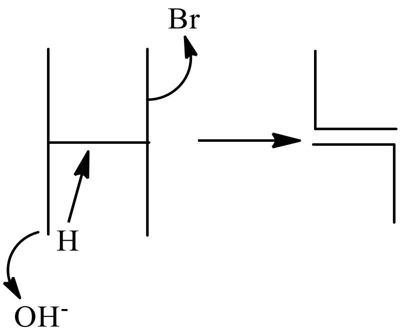
$CH_3-\underset{\underset{H}{|}}{\overset{\overset{^+N(CH_3)_3}{|}}{C}}-CH_2\,CH_3 \xrightarrow{\text{Alc.KOH}} CH_3CH=CHCH_3$

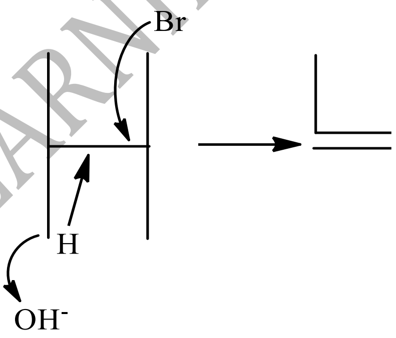
d) 

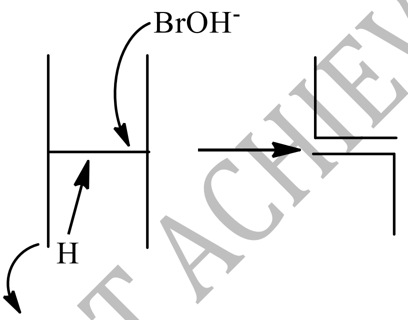
$CH_3-\underset{\underset{Br}{|}}{CH}-CH_2Br+\text{Alc. KOH}$

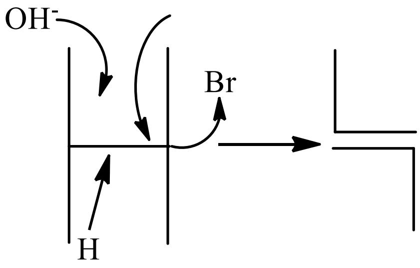
$\xrightarrow{\Delta} CH_3\underset{\underset{OH}{|}}{CH}-CH_2OH$

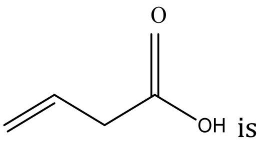
41. Dehydrohalogenation in presence of OH^- is correctly represented by

a) 

b) 

c) 

d) 

42. The IUPAC name of  is

a) But-3-enoic acid b) But-1-enoic acid c) Pent-4-enoic acid d) Prop-2-enoic acid

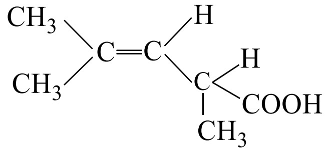
43. On exciting Cl_2 molecules by UV light, we get

a) $Cl^\bullet$ b) Cl^+ c) Cl^- d) All of these

44. Mixture of sugar and common salt is separated by crystallisation by dissolving in

a) H_2O b) C_2H_5OH c) C_5O_6 d) None of these

45. The structure,



shows :

a) Geometrical isomerism

b) Optical isomerism

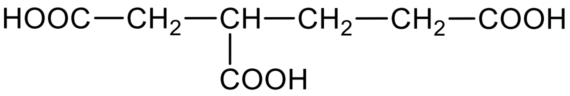
c) Geometrical and optical isomerism

d) Tautomerism

46. The general formula for cycloalkanes is :

a) C_nH_{2n+2} b) C_nH_{2n} c) C_nH_{2n-2} d) C_nH_n

47. The IUPAC name of the compound



$HOOC-CH_2-\underset{\underset{COOH}{|}}{CH}-CH_2-CH_2-COOH$

a) 2(carboxymethyl)-pentane-l,5-dioic acid

b) 3-carboxyhexane-l, 6-dioic acid

c) Butane-l, 2, 4-tricarboxylic acid

d) 4-carboxyhexane-l, 6-dioic acid

48. $Na_2S + Na_2[Fe(CN)_5NO] \rightarrow$ Purple colour. It is due to

a) $Na_4[Fe(CN)_3NOS]$ b) $Na_3[Fe(CN)_5NOS]$ c) $Na_4[Fe(CN)_5NO]$ d) $Na_4[Fe(CN)_5NOS]$

49. The bond that undergoes heterolytic cleavage most easily is

a) $C-O$ b) $C-C$ c) $C-H$ d) $O-H$

50. Increasing order of stability among the three main conformations (*i.e., Eclipse, Anti, Gauche*) of 2-fluoroethanol is

a) *Eclipse, Gauche, Anti* b) *Gauche, Eclipse, Anti* c) *Eclipse, Anti, Gauche* d) *Anti, Gauche, Eclipse*

51. Phosphorus is estimated as

a) Na_3PO_4 b) P_2O_5 c) P_2O_3 d) $Mg_2P_2O_7$

52. The number of asymmetric carbon atoms and the number of optical isomers in $CH_3(CHOH)_2COOH$ are respectively :

a) 3 and 4 b) 1 and 3 c) 2 and 4 d) 2 and 3

53. Species containing carbon with three bonds and an electron are called :

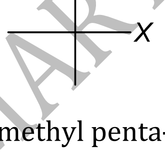
a) Carbenes b) Caarbanions c) Carbocation d) Free radicals

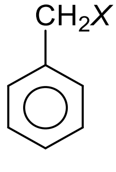
54. Which of the aldehyde is most reactive?

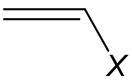
a) C_6H_5-CHO b) CH_3CHO

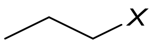
c) HCHO d) All the equally reactive

55. Which of the following cannot show S_N1 reaction?

a) 

b) 

c) 

d) 

56. 3-methyl penta-1,3-diene is :

a) $CH_2 = CH(CH_2)_2CH_3$

b) $CH_2 = CHCH(CH_3)CH_2CH_3$

c) $CH_3CH = C(CH_3)CH = CH_2$

d) $CH_3 - CH = CH(CH_3)_2$

57. Which of the following compounds is optically active?

a) 1 – butanol b) Isopropyl alcohol c) Acetaldehyde d) 2-butanol

58. How many optically active forms are possible for a compound of the formula, $CHO.CHOH.CHOH.CHOH.CH_2OH$?

a) 2 b) 4 c) 3 d) 8

59. "The negative part of the addendum adds on the carbon atom joined to the least number of hydrogen atoms." This statement is called :
a) Markownikoff's rule
b) Peroxide effect
c) Baeyer's strain theory
d) Thiele's theory

60. The total number of isomeric carbocations possible for the formula $C_4H_9^+$ is :
a) 3 b) 4 c) 2 d) 5

61. The correct order for homolytic bond dissociation energies. (ΔHin kcal/mol) for CH_4(A), C_2H_6(B) and CH_3Br(C), under identical experimental conditions
a) C>B>A b) B>C>A c) C>A>B d) A>B>C

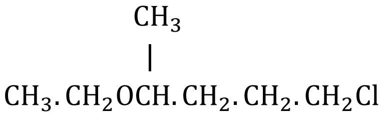
62. The sodium extract of an organic compound on treatment with $FeSO_4$ solution, $FeCl_3$ and HCl gives a red solution. The organic compound contains
a) Both nitrogen and sulphur b) Nitrogen only
c) Sulphur only d) Halogen

63. *d*-tartaric acid and *l*-tartaric acid are :
a) Structureal isomers b) Diastereoisomers c) Tautomers d) Enantiomers

64. Which of the following is a pair of functional isomers?
a) CH_3COCH_3, CH_3CHO b) $C_2H_5CO_2H, CH_3CO_2CH_3$
c) $C_2H_5CO_2H, CHCO_2C_2H_5$ d) CH_3CO_2H, CH_3CHO

65. Which of the following is an optically active compound?
a) Lactic acid b) Chloro acetic acid c) *Meso*-tartaric acid d) Acetic acid

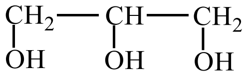
66. Give the correct IUPAC name for



$CH_3.CH_2O\overset{\overset{CH_3}{|}}{C}H.CH_2.CH_2.CH_2Cl$

a) 2-ethoxy-5-chloropentane b) l-chloro-4-ethoxy-4-methylbutane
c) 1-chloro-4-ethoxypentane d) Ethyl-1-chloropentylether

67. The IUPAC name of the compound,



$\underset{OH}{\underset{|}{CH_2}}-\underset{OH}{\underset{|}{CH}}-\underset{OH}{\underset{|}{CH_2}}$ is :

a) 1,2,3-trihydrosypropane
b) 3-hydroxypentane-1,5-diol
c) 1,2,3-hydroxypropane
d) Propane-1,2,3-triol

68. Bond energy ....with the increase in number of lone pairs on the bonded atoms.
a) Decreases b) Increases c) Does not change d) None of these

69. A liquid decomposes at its normal boiling point. It can be purified by
a) Sublimation b) Steam distillation
c) Vacuum distillation d) Fractional distillation

70. On monochlorination of 2-methyl butane, the number of chiral compounds formed are :
a) 2 b) 4 c) 6 d) 8

71. Stability of which intermediate is not governed by hyperconjugation?
a) Carbon cation b) Carbon anion c) Carbon free radical d) None of these

72. The ammonia evolved from the treatment of 0.30g of an organic compound for the estimation of nitrogen was passed in 100mL of 0.1M sulphuric acid. The excess of acid required 20mL of 0.5 M sodium hydroxide solution for complete neutralisation. The organic compound is
a) Acetamide b) Benzamide c) Urea d) Thiourea

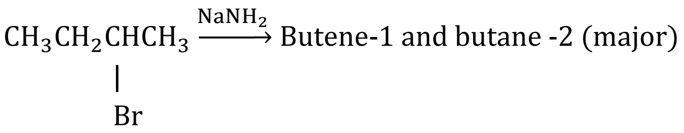
73. Conversion of CH_4 to CH_3Cl is an example of which of the following reaction?
a) Electrophilic substitution b) Free radical addition

c) Nucleophilic substitution   d) Free radical substitution

74. Number of possible isomers of glucose are :

a) 10   b) 14   c) 16   d) 20

75. The reaction



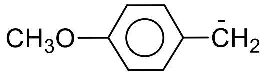
$CH_3CH_2\underset{\underset{Br}{|}}{CH}CH_3 \xrightarrow{NaNH_2}$ Butene-1 and butane -2 (major)

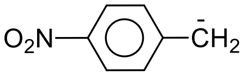
The correct statement (s) are

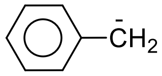
a) 2-butene is Saytzeff product   b) 1-butene is Hofmann (s) product

c) The elimination reaction follows Saytzeff rule   d) All of the above

76. Consider the following carbanions

(I) 

(II) 

(III) 

Correct order of stability is

a) I>II>III   b) III>II>I   c) II>III>I   d) I>III>II

77. The stability of 2,3-dimethyl but-2-ene is more than 2-butene. This can be explained in terms of :

a) Resonance   b) Hyperconjugation   c) Electromeric effect   d) Inductive effect

78. Protin solvent is

a) Diethyl ether   b) n-hexane   c) Acetone   d) Ethanol

79. Addition of Br_2 on *trans*-butene-2 gives :

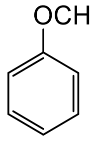
a) A racemic mixture of 2,3-dibromobutane

b) Meso form of 2,3-dibromobutane

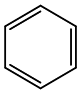
c) Dextro form of 2,3-dibromobutane

d) Laevo form of 2,3-dibromobutane

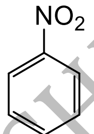
80. Among the following compounds (I-III) the correct order of reaction with electrophilic reagand is



I



II



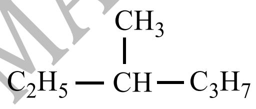
III

a) II>III>I   b) III<I<II   c) I>II>III   d) I=II>III

81. During $AgNO_3$ test for detection of halogens, sodium extract is boiled with few drops of conc. HNO_3 to decompose

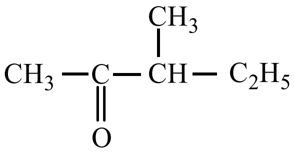
a) NaCN   b) Na_2S c) Both (a) and (b)   d) None of these

82. Which is true about following?



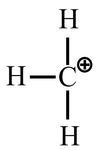
$C_2H_5-\overset{\overset{CH_3}{|}}{CH}-C_3H_7$

(I)



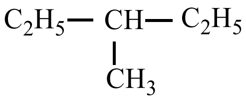
$CH_3-\underset{\underset{O}{\|}}{C}-\overset{\overset{CH_3}{|}}{CH}-C_2H_5$

(II)



$H-\overset{\overset{H}{|}}{\underset{\underset{H}{|}}{C^{\oplus}}}$

(III)



$C_2H_5-\underset{\underset{CH_3}{|}}{CH}-C_2H_5$

(IV)

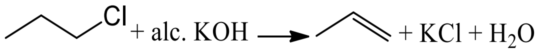
a) Only III is a chiral compound

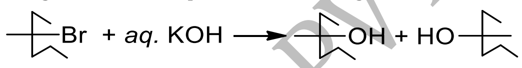
b) Only II and IV are chiral compounds
c) All four are chiral compounds
d) Only I and II are chiral compounds

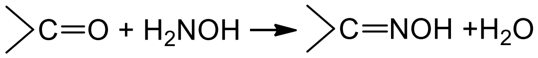
83. How many chiral compounds are possible on monochlorination of 2-methyl butane?
a) 2 b) 4 c) 6 d) 8

84. How many isomers of $C_5H_{11}OH$ will be primary alcohols?
a) 5 b) 4 c) 3 d) 2

85. The epoxide ring consists of which of the following?
a) Three membered ring with two carbon and one oxygen
b) Four membered ring with three carbon and one oxygen
c) Five membered ring with four carbon and one oxygen.
d) Six membered ring with five carbon and one oxygen.

86. The reaction which is not the example of nucleophilic substitution among the following is
a) $CH_3C - Br + CH_3OH \longrightarrow CH_3C - OCH_3 + HBr$
b) $CH_3C - Cl + aq. KOH \longrightarrow CH_3C - OH + KCl$
c) 

d) 

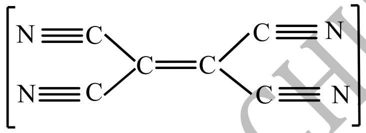
87. Consider the following reaction


Is an example of
a) Substitution b) Elimination c) Addition d) Addition elimination

88. An important chemical method to resolve a racemic mixture makes use of the formation of :
a) *meso* compound b) Enantiomer c) Racemers d) diastereoisomers

89. Red colour complex ion formed on adding $FeCl_3$ to sodium extract when N and S both are present in organic compound is
a) $[Fe(CN)_6]^{4-}$ b) $[Fe(CNS)]^{2+}$ c) $[Fe(CNS)_2]^+$ d) $[Fe(CN)_6]^{3-}$

90. $(I) CH_3CH_2Br \xrightarrow{LAH} C_2H_6 \, and \, (II)$
$(CH_3)_3CBr \xrightarrow{LAH}$ alkene, The reason for this is
a) (I)S_N2(II) E_1 mechanism
b) (I) S_N1, (II) E_2 mechanism
c) (I)S_N1, (II) E_1 mechanism
d) (I) S_N2,(II) E_2 mechanism

91. How many σ and π-bonds are there in the molecule of tetracyanoethylene?


a) 9σ and 9π b) 5σ and 9π c) 9σ and 7π d) 5σ and 8π

92. Which of the following complex formation indicates presence of sulphur in the organic compound when sodium nitroprusside is added to sodium extract of the compound?
a) $Fe_4[Fe(CN)_6]_3$ b) $Na_2[Fe(NO)(CN)_5]$ c) $Fe_4(CNS)_3$ d) $Na_4[Fe(CN)_5NOS]$

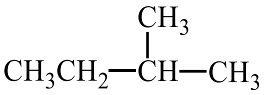
93. Who pointed out the concept hyperconjugation?
a) Nathan and Baker b) Mullikan c) Kekule d) Kolbe

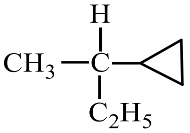
94. Alkyl halide can be converted into alkene by
a) Nucleophilic substitution reaction
b) Elimination reaction
c) Both nucleophilic substitution and elimination reaction
d) Rearrangement

95. The order of reactivities of the following alkyl halides for a S_N2 reaction is :
a) $RF > RCl > RBr > RI$
b) $RF > RBr > RCl > RI$
c) $RCl > RBr > RF > RI$
d) $RI > RBr > RCl > RF$

96. The optically active alkane with lowest molecular weight is :

a) $CH_3CH_2C\equiv CH$

b) 
$CH_3CH_2-CH(CH_3)-CH_3$

c) 

d) $CH_3CH_2{\cdot}CH_2CH_3$

97. Which type of isomerism is most common among ethers?

a) Metamerism b) Functional c) Chain d) Position

98. With a change in hybridisation of the carbon bearing the charge, the stability of a carbanion increase in the order

a) $sp < sp^2 < sp^3$ b) $sp < sp^3 < sp^2$ c) $sp^3 < sp^2 < sp$ d) $sp^2 < sp < sp^3$

99. A molecule is $R_3C—H$. If H is replaced by $Z(R_3C—Z)$ and on doing so electron density on $R_3—C$ part increases, then Z is :

a) Electron attracting group

b) Electron withdrawing group

c) Electron repelling group

d) Either of the above

100. Which of the following compounds are not arranged on order of decreasing reactivity towards electrophilic substitution?

a) Fluorobenzene > chlorobenzene > bromo benzene

b) Phenol> n-propyl benzene> benzoic acid>

c) Chlorotoluene >$para$-nitrotoluene>2-chloro-4-nitro toluene

d) Benzoic acid> phenol>n-propyl benzene

101. A mixture of camphor and benzoic acid can be separated by

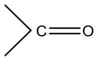
a) Sublimation b) Extraction with a solvent

c) Chemical method d) Fractional crystallisation

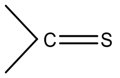
102. Resonance in benzene is accompanied by delocalisation of π-electrons. Each π-electron is attached with :

a) 4 carbon b) 2 carbon c) 3 carbon d) 6 carbon

103. Grignard reagent adds to

a) 

b) $-C\equiv N$

c) 

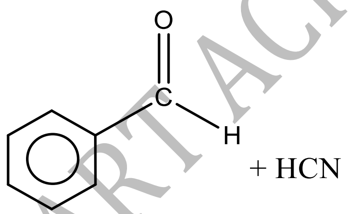
d) All of these

104. Resonance energy is more for

a) C_6H_6 b) Cylohexene

c) Cycloheptene d) Cyclohexa -1,2,3-triene

105. The reaction



is an example of

a) Electrophilic addition b) Electrophilic substitution

c) Nucleophilic substitution d) Nucleophilic addition

106. Which of the following is not chiral?

a) 3-bromopentane

b) 2-hydroxy propanoic acid

c) 2-butanol

d) 2,3-dibromopentane

107. The structures $(CH_3)_3CBr$ and $CH_3[CH_2]_3Br$ represent

a) Chain isomerism b) Position isomerism

c) Chain as well as position isomerism d) Functional isomerism

108. Detection of phosphorus in the compound can be done by its conversion into phosphate. Reagent to identify phosphate ion is
a) Sodium nitroprusside
b) Ammonium molybdate
c) Potassium ferrocyanide
d) Potassium ferricyanide

109. The hemolytic fission of a hydrocarbon results in the formation of:
a) Carbonium ions
b) Free radicals
c) Carbanions
d) Carbenes

110. Which does not have sp^2-hybridised carbon atom?
a) Acetamide
b) Acetic acid
c) Acetonitrile
d) Acetone

111. Methoxy methane and ethanol are
a) Position isomers
b) Chain isomers
c) Functional isomers
d) Optical isomers

112. Which one of the following monoenes does not exhibit geometric isomerism?
a) C_4H_8
b) $C_3\ H_6$
c) C_5H_{10}
d) C_8H_{16}

113. Addition of Br_2 on $cis-$ butene -2 gives :
a) A racemic mixture of 2,3-dibromobutane
b) Meso form of 2,3-dibromobutane
c) Dextro form of 2,3-dibromobutane
d) Laevo form of 2,3-dibromobutane

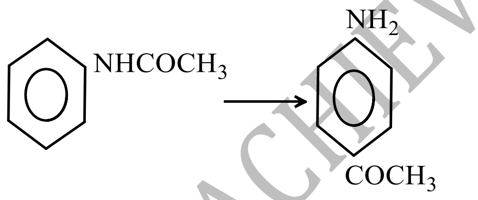
114. When two halogen atoms are attached to two adjacent carbon atoms, the dihaloalkane is called :
a) Alkylidene dihalide
b) Alkane dihalide
c) Alkylene dihalide
d) Alkyl halide

115. The electrophile involved in the sulphonation of benzene is
a) SO_3^+
b) SO_3^{2-}
c) H_3O^+
d) SO_3

116. The number of *meso* forms in the following compound is
$HOOC.CH(CH_3).CH(OH).$
$CH(Cl).CH(OH)CH(CH_3).$
$COOH$
a) 3
b) 4
c) 8
d) 16

117. In Kjeldahl's method, then nitrogen present in the organic compound is quantitatively converted into
a) Gaseous ammonia
b) Ammonium sulphate
c) Ammonium phosphate
d) Ammonia

118. The reaction;



is called :
a) Substitution reaction
b) Elimination reaction
c) Rearrangement reaction
d) None of the above

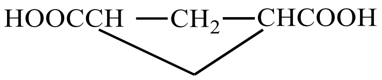
119. The compound which gives the most stable carbonium ion on dehydration is
a) $CH_3CH(CH_3)CH_2OH$
b) $(CH_3)_3COH$
c) $CH_2 = CHCH_2CH_2OH$
d) $CH_3CHOHCH_2-\ CH_3$

120. A similarity between optical and geometrical isomerism is that :
a) Each forms equal number of isomers for a given compound
b) If in a compound, one is present then so is the other
c) Both are included in stereoisomerism
d) They have no similarity

121. Reaction of phenol with chloroforms/sodium hydroxide to give *o*-hydroxy benzaldehyde involves the formation of
a) Dichloro carbene
b) Trichloro carbene
c) Chlorine atoms
d) Chlorine molecules

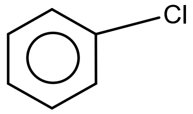
122. Which kind of fission is favoured by sunlight?
a) Heterolytic fission b) Homolytic fission c) Both (a) and (b) d) None of these

123. The stability of the free radicals allyl, benzyl, 3°, 2°, 1° and CH_3 is in the order
a) Benzyl > allyl >3° > 2° > 1° > CH_3 b) Allyl > 3° > benzyl > 2° > 1° > CH_3
c) 3° > 2° > 1° > CH_3>allyl > benzyl d) 3° > 2° > 1° > CH_3 > allyl = benzyl

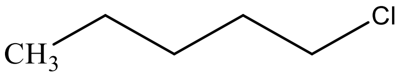
124. Which class of compounds can exhibit geometrical isomerism?
a) $C_6H_5CH = NOH$
b) $CH_3CH = CHCH_3$
c) 

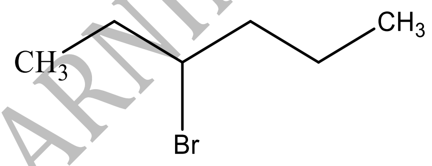
d) All of the above

125. The correct order of increasing basicity of the given conjugate bases (R=CH_3) is
a) $RCO\bar{O} < HC \equiv \bar{C} < \bar{R} < \bar{N}H_2$ b) $\bar{R} < HC \equiv \bar{C} < RCO\bar{O} < \bar{N}H_2$
c) $RCO\bar{O} < NH_2 < HC \equiv \bar{C} < \bar{R}$ d) $RCO\bar{O} < HC \equiv \bar{C} < \bar{N}H_2 < \bar{R}$

126. Which of the following shows S_N1 reaction most readily?
a) 

b) 

c) 

d) 

127. Which of the following compounds is optically active?
a) $(CH_3)_2CHCH_2OH$ b) CH_3CH_2OH c) CCl_2F_2 d) $CH_3CHOHC_2H_5$

128. To which ring size cycloalkanes, Baeyer's strain theory is not valid?
a) 3 carbon b) 4 carbon c) 5 carbon d) ≥ 6 carbon

129. The S_N1 mechanism for substitution reaction by nucleophile is favoured by :
a) Low concentration of nucleophile
b) Weak nature of nucleophile
c) Polar solvent
d) All of the above

130. Which of the following orders is not correct regarding the $-I$ effect of the substitutents?
a) $-I < -Cl < -Br < -F$ b) $-\overset{+}{N}R_3 < -\overset{+}{O}R_2$
c) $-\overset{+}{N}R_2 < -OR < -F$ d) $-SR < -OR < -\overset{+}{O}R_2$

131. Lactic acid shows optical activity in :
a) Solution state b) Liquid state c) Crystalline state d) In all states

132. In cyclopropane, cyclobutane and cyclohexane, the common group is
a) $-\overset{|}{\underset{|}{C}}-$ b) $\overset{|}{\underset{|}{CH_2}}$ c) $-CH_3$ d) $-\overset{|}{\underset{|}{CH}}$

133. Total number of isomeric aldehydes and ketones that can exist with the molecular formula $C_5H_{10}O$:
a) 5 b) 8 c) 6 d) 7

134. Allyl isocyanide has :
a) 9σ and 4π-bonds
b) 8σ and 5π-bonds
c) 9σ, 3π and 2 non-bonded electrons
d) 8σ, 3π and 4 non-bonded electrons

135. $+I$ effect is shown by

a) $-CH_3$ b) $-Br$ c) $-Cl$ d) $-NO_2$

136. $LiAlH_4$ is used as :

a) Oxidizing agent b) Reducing agent c) A mordant d) A water softener

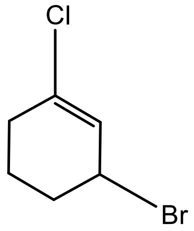
137. 0.765g of an acid gives 0.535g of CO_2 and 0.138 g of H_2O. Then, the ratio of the percentage of carbon and hydrogen is

a) 19:2 b) 18:11 c) 20:17 d) 1:7

138. Which one of the following is the stable structure of cyclohexatriene?

a) Chair form b) Boat form c) Half chair form d) Planar form

139. The IUPAC name of compound shown below is



a) 2-bromo-6- chlorocyclohex-1-ene b) 6-bromo-2-chlorocyclohexene

c) 3-bromo-1-chlorocyclohexene d) 1-bromo-3-chlorocyclohexene

140. Total number of rotational conformers of n-butane are :

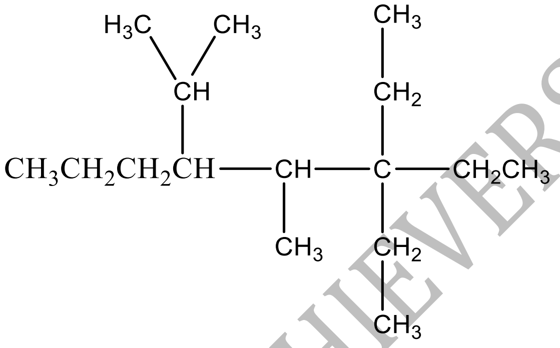
a) 2 b) 6 c) 5 d) 3

141. Sublimation is a process in which a solid

a) Changes into vapour form b) Changes into another allotropic form

c) Changes into liquid form d) None of the above

142. IUPAC name of the compound



$CH_3CH_2CH_2CH(CH(CH_3)_2)-CH(CH_3)-C(CH_2CH_3)(CH_2CH_3)-CH_2CH_3$

a) 5- methyl-4-isopropyl-6, 6'diethyloctane b) 3, 3-dimethyl, 3-ethyl-5- isopropyl octane

c) 3, 3-diethyl-4-methyl-5-(1,1-dimethyl) octane d) 3, 3- diethyl-4-methyl-5-(1'-methylethyl) octane

143. The group named as benzal possesses ……nature.

a) Monovalent b) Bivalent c) Trivalent d) Tetravalent

144. A secondary(2°) carbon is one that is joined to :

a) 1-alkyl group b) 2-alkyl groups c) 3-alkyl groups d) None of these

145. Which type of strain is present in fully eclipsed conformation of butane?

a) Angle strain b) Steric strain c) Both (a) and (b) d) Neither (a) nor (b)

146. 29.5 mg of organic compound containing nitrogen was digested according to Kjeldahl's method and the evolved ammonia was absorbed in 20mL of 0.1 M HCl solution. The excess of the acid required 15mL of 0.1M NaOH solution for complete neutralisation. The percentage of nitrogen in the compound is

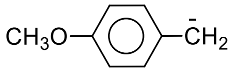
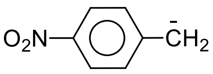
a) 59.0 b) 47.4 c) 23.7 d) 29.5

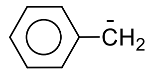
147. The highest electrical conductivity of the following aqueous solutions is of

a) 0.1 M difluoroacetic acid b) 0.1 M fluoroacetic acid

c) 0.1 M chloroacetic acid d) 0.1 M acetic acid

148. Consider the following carbanions

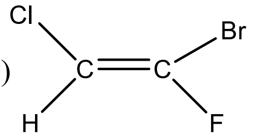
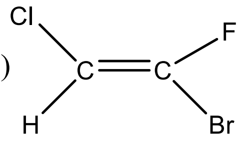
(1)  (2) 

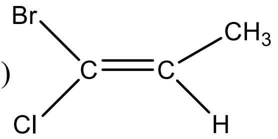
(3) 

Correct order of stability is

a) 1>2>3 b) 3>2>1 c) 2>3>1 d) 1>3>2

149. Which of the following compounds (s) has 'Z' configuration?

(i)  (ii) 

(iii) 

a) (i) only b) (ii) only c) (iii) only d) (i) and (iii)

150. Nucleophiles are :

a) Electron loving b) Electron hating c) Nucleus loving d) Nucleus hating

151. Ethyl acetoacetate exhibits :

a) Optical isomerism

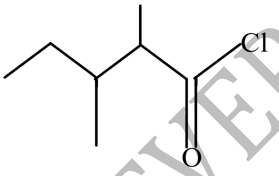
b) Geometrical isomerism

c) Tautomerism

d) enantiomerism

152. The total number of cyclic isomers possible for a hydrocarbon with the molecular formula C_4H_6 is

a) 1 b) 3 c) 5 d) 7

153. The IUPAC name of  is :

a) 2-ethyl-3-methylbutanoyl chloride

b) 2,3-dimethylpentanoyl chloride

c) 3,4-dimethylpentanoyl chloride

d) 1-chloro-l-oxo-2,3-dimethylpentane

154. A molecule of urea can show

a) Chain isomerism b) Position isomerism

c) Geometrical isomerism d) Tautomerism

155. Cyclic hydrocarbon molecule (A) has all the C and H atoms in single plane. All the C–C bonds have same length, less than 1.54 Å but more than 1.34Å. The ∠ (angle) CCC is :

a) 190°28′ b) 100° c) 180° d) 120°

156. The number of π-electrons present in cyclobutadienyl ion, $(C_4H_3)^-$ is :

a) 8 b) 6 c) 4 d) 2

157. Geometrical isomerism is possible in case of

a) Pentene-2 b) Propane c) Pentane d) Ethene

158. The strongest best among the following is :

a) NH_4^+

b) $:NH_3$

c) $:\bar{N}H_2$

d) $:\bar{O}H$

159. Anti-Markownikoff addition of HBr is not observed in :

a) Propene b) Butene-1 c) But-2-ene d) Pent-2-ene

160. The number of 1°, 2° and 3° carbon atoms present in isopentane are respectively :

a) 3, 2, 1 b) 2, 3, 1 c) 3, 1, 1 d) 2, 2,1

161. The restricted rotation about carbon-carbon double bond in 2-butene is due to:

a) Overlap of two p-orbitals

b) Overlap of one p and one sp^2-hybridized orbitals

c) Overlap of two sp^2-hybridized orbitals

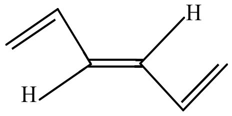
d) Overlap of one s and one sp^2-hybridized orbitals

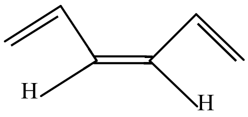
162. Formation of acetylene from ethylene is an example of

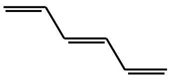
a) Addition reaction b) Substitution reaction

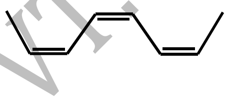
c) Elimination reaction d) Condensation reaction

163. The structure of *cis-bis* (propenyl) ethane is :

a) 

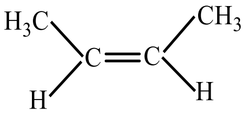
b) 

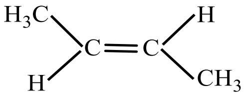
c) 

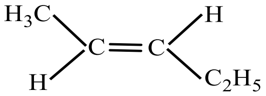
d) 

164. The compound which reacts with HBr obeying Markownikoff's rule is:

a) $CH_2 = CH_2$

b) 

c) 

d) 

165. A molecule of benzene contains :

a) Twelve sigma-bonds and three pi-bonds

b) Eighteen sigma-bonds and three pi-bonds

c) Twelve pi-bonds and three sigma-bonds

d) Six hydrogen-bonds, six sigma-bonds and three pi-bonds

166. Zn—Cu couple used as reducing agent is :

a) Mixture of Zn and Cu powder

b) Copper deposited on granulated zinc

c) Zn deposited on copper fillings

d) A rod half made of copper and half made of zinc

167. Considering the state of hybridization of carbon atoms, find out the molecule among the following which is linear?

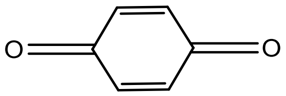
a) $CH_3—CH_2—CH_2—CH_3$

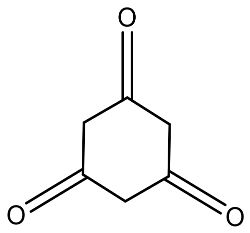
b) $CH_3—CH = CH—CH_3$

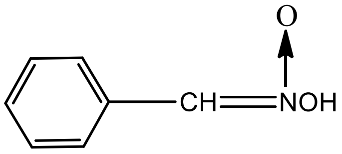
c) $CH_3—C \equiv C—CH_3$

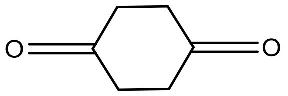
d) $CH_2 = CH—CH_2—C \equiv CH$

168. Tautomerism is not exhibited by

a) 

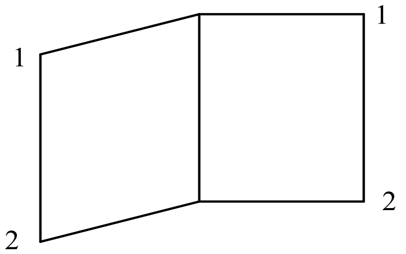
b) 

c) 

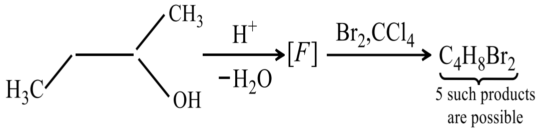
d) 

169. Which of the substance is purified by sublimation?
a) Naphthalene b) Benzoic acid c) Camphor d) All of these

170. The IUPAC name of the following compound is



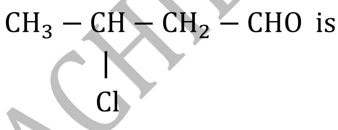
a) Bicyclo [2,2,0] octane b) Bicyclo [0,2,2] hexane c) Bicyclo [2,1,1] hexane d) Bicyclo [2,2,0] hexane

171. 

How many structures of F are possible?
a) 2
b) 5
c) 6
d) 3

172. Example of chlorinolysis is :
a) $CH_2 = CH_2 \xrightarrow{Cl_2} C_2H_4Cl_2$
b) $CCl_4 + H_2O \longrightarrow COCl_2 + 2HCl$
c) $CHCl_3 + 4NaOH \longrightarrow HCOONa + 3NaCl + 2H_2O$
d) $C_3H_8 \xrightarrow{Cl_2} CCl_4 + C_2Cl_6 + 8HCl$

173. The number of optical enantiomorphs of tartaric acid :
a) 3 b) 2 c) 4 d) 1

174. IUPAC name of  is

$CH_3 - \underset{\displaystyle Cl}{\underset{|}{CH}} - CH_2 - CHO$

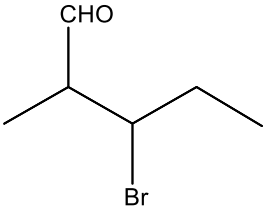
a) 3-chlorobutanol b) 3-chlorobutanaldehyde
c) 3-chlorobutanal d) 2-chlorobutanol

175. 4 g of hydrocarbon on complete combustion gave 12.571 g of CO_2 and 5.143 g of water. What is the empirical formula of the hydrocarbon?
a) CH b) C_2H_3 c) CH_2 d) CH_3

176. The compound which contains all the four 1°, 2°, 3° and 4° carbon atoms is
a) 2, 3-dimethylpentane b) 3-chloro-2, 3-dimethylpentane
c) 2, 3, 4-trimethylpentane d) 3, 3-dimethylpentane

177. Which of the following is useful for making pure water from a solution of salt in water?
a) Filtration b) Distillation c) Chromotography d) Steam distillation

178. Which of the following does not contain chiral carbon atom?
a) Lactic acid b) 2-chlorobutanoic acid c) Tartaric acid d) Succinic acid

179. 

The IUPAC name of is

a) 2-methyl-3-bromohexanal
b) 3-bromo-2-methylbutanal
c) 2-bromo-3-bromobutanal
d) 3-bromo-2-methylpentanal

180. CH_3CH_2Cl undergoes homolytic fission, produces

a) $CH_3\ \dot{C}H_2$ and $\dot{Cl}$
b) $CH_3\ \overset{\oplus}{C}H_2$ and $\overset{\ominus}{Cl}$
c) $CH_3\ \overset{\oplus}{C}H_2$ and $\dot{Cl}$
d) $CH_3\ \dot{C}H_2$ and $\overset{\ominus}{Cl}$

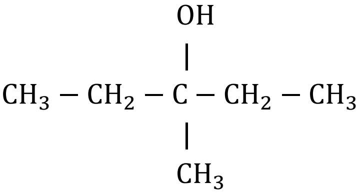
181. Among the following orbital bonds, the angle is minimum between :

a) $sp^3 - sp^3$ bonds
b) p_x and p_y-orbitals
c) H— O— H in water
d) $sp - sp$ bonds

182. Which of the following is the correct order of decreasing S_N2 reactivity? (X=α halogen)

a) $RCH_2X > R_3CX > R_2CHX$
b) $RCH_2X > R_2CHX > R_3CX$
c) $R_3CX > R_2CHX > RCH_2X$
d) $R_2CHX > R_3CX > R_2CH_2X$

183. Write the IUPAC name of



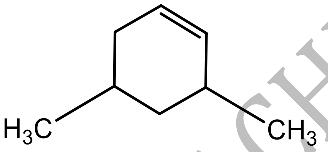
$$\begin{array}{c} \quad\quad\quad\quad OH \\ \quad\quad\quad\quad | \\ CH_3 - CH_2 - C - CH_2 - CH_3 \\ \quad\quad\quad\quad | \\ \quad\quad\quad\quad CH_3 \end{array}$$

a) 3-methylpentane-3-ol
b) 3-hydroxyhexane
c) 3-hydroxy-3-methyl pentane
d) All of the above

184. Polarization of electron in acrolein may be written as

a) $\overset{-\delta}{CH_2}=CH-\overset{+\delta}{CH}=O$
b) $\overset{-\delta}{CH_2}=CH-CH=\overset{+\delta}{O}$
c) $\overset{-\delta}{CH_2}=\overset{+\delta}{CH}-CH=O$
d) $\overset{+\delta}{CH_2}=CH-CH=\overset{-\delta}{O}$

185. IUPAC name of the following compound is



a) 3, 5-dimethylcyclohexene
b) 3, 5-dimethly-1-cyclohexene
c) 1, 5-dimethly-5-cyclohexene
d) 1, 3-dimethyl-5-cyclohexene

186. In this reaction,

$CH_3CHO + HCN \longrightarrow CH_3CH(OH)CN$

$\xrightarrow{H.OH} CH_3CH(OH)COOH$

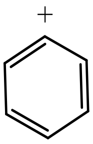
an asymmetric centre is generated. The acid obtained would be

a) 50%D+50%L-isomer
b) 20%D+80%L-isomer
c) D-isomer
d) L-isomer

187. Two crystalline forms of a substance, one being a mirror image of the other are called :

a) Pentane
b) Chain isomers
c) Stereoisomers
d) Functional isomers

188. Which one of the following is an intermediate in the reaction of benzene with CH_3Cl in the presence of anhydrous $AlCl_3$?

a) Cl^+
b) CH_3^-
c) CH_3^+
d) 

189. The number of optical isomers of $CH_3CH(OH)CH(OH)CHO$ is

a) Zero b) 2 c) 3 d) 4

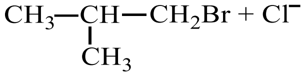
190. The ratio of σ-to π-bonds in benzene is:

a) 2 b) 4 c) 6 d) 8

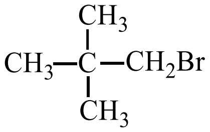
191. In a S_N2 substitution reaction of the type

$R—Br + Cl^- \xrightarrow{DMF} R—Cl + Br^+$

Which one of the following has the highest relative rate?

a) 

$CH_3—\underset{\substack{|\\ CH_3}}{CH}—CH_2Br + Cl^-$

b) 

$CH_3—\overset{\substack{CH_3\\ |}}{\underset{\substack{|\\ CH_3}}{C}}—CH_2Br$

c) CH_3CH_2Br

d) $CH_3—CH_2—CH_2Br$

192. Hyperconjugation is

a) $\sigma-\pi$ delocalisation b) No bond resonance c) $\sigma-\pi$ odd electron d) All of these

193. Which one of the following reactions is a condensation reaction?

a) HCHO ⟶*para*-formaldehyde

b) CH_3CHO ⟶*para*-aldehyde

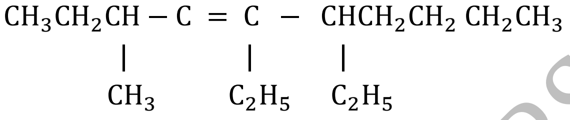
c) CH_3COCH_3⟶ mesityl oxide

d) $CH_2 = CH_2$ ⟶ polyethylene

194. Which group has the maximum-Inductive effect?

a) $-NO_2$ b) $-CN$ c) $-COOH$ d) $-F$

195. The correct IUPAC name of the following compound is



$CH_3CH_2\underset{\substack{|\\ CH_3}}{CH}-C = \underset{\substack{|\\ C_2H_5}}{C} - \underset{\substack{|\\ C_2H_5}}{CH}CH_2CH_2\,CH_2CH_3$

a) 5, 6-dimethyl-8-methyl dec-6-ene b) 6-butyl-5-ethyl-3-methyl oct-4-ene

c) 5, 6-diethyl-3-methyl dec-4-ene d) 2, 4, 5-triethyl non-3-ene

196. Which is incorrect about enantiomorphs?

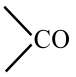
a) They rotate the plane of polarized light in different directions

b) They have mostly identical physical properties

c) They have same configuration

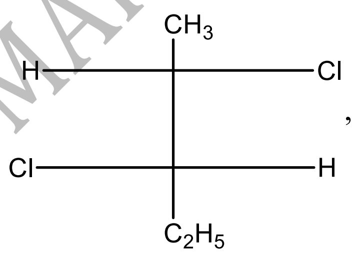
d) They have different biological properties

197. Which one is the seniormost functional group in the nomenclature of an organic compound if it possesses more than one functional group?

a) —CHO b) —COOH c) —OH d) 

$>CO$

198. The absolute configuration of the following



, are

are

a) 2S, 3R b) 2S, 3S c) 2R, 3S d) 2R, 3R

199. Which step is chain termination step in the following mechanism?

(i) $Cl_2 \xrightarrow{hv} Cl^{\bullet} + Cl^{\bullet}$

(ii) $Cl^{\bullet} + CH_4 \longrightarrow \overset{\bullet}{C}H_3 + HCl$

(iii) $\overset{\bullet}{C}H_3 + Cl_2 \longrightarrow CH_3Cl + Cl^{\bullet}$

(iv) $Cl^{\bullet} + \overset{\bullet}{C}H_3 \longrightarrow CH_3Cl$

a) (i) b) (ii) c) (iii) d) (iv)

200. The reaction intermediate produced, by homolytic cleavage of a bond is called

a) Carbene b) Carbocation c) Carbanion d) Free redical

201. Fractional distillation is useful in distillation of

a) Petroleum b) Coal-tar c) Crude alcohol d) All of these

202. Which of the following species is paramagnetic?

a) A carbocation b) A free radical c) A carbanion ion d) All of these

203. Sulphur trioxide is :

a) An electrophile b) A nucleophile c) A homolytic reagent d) A base

204. In Kjeldahl's method, ammonia from 5g of food neutralizes 30 cm^3 of 0.1 N acid. The percentage of nitrogen in the food is

a) 0.84 b) 8.4 c) 16.8 d) 1.68

205. The number of isomeric alkanes having the molecular formula C_5H_{12} is

a) Three b) Five c) Nine d) Thirty two

206. Select the organic compound which was prepared for the first time in laboratory from its elements :

a) Urea b) CH_3COOH c) C_2H_5OH d) None of these

207. Which of the following compounds can exist in optically active form?

a) 1-butanol b) 2-butanol c) 3-pentanol d) 4-heptanol

208. The compound in which carbon uses only its sp^3 hybrid orbitals for bond formation is

a) $(CH_3)_3COH$ b) HCOOH c) CH_3CHO d) $(H_2N)_2CO$

209. How many types of functional group can be present in an amine with the formula C_3H_9N?

a) 1 b) 2 c) 3 d) 4

210. Select the most reactive cycloalkane :

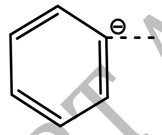
a) Cyclopropane b) Cyclobutane c) Cyclopentane d) Cyclohexane

211. The $-I$ effect is shown by :

a) $-COOH$ b) $-CH_3$ c) $-CH_3CH_2$ d) $-CHR_2$

212. The stability of carbanions in the following ;

(1) $-RC{\equiv}\overset{\ominus}{C}$

(2) 

(3) $R_2C{=}\overset{\ominus}{C}H$

(4) $R_3C-\overset{\ominus}{C}H_2$

is in the order of :

a) (2) > (3) > (4) > (1)

b) (4) > (2) > (3) > (1)

c) (1) > (3) > (2) > (4)

d) (1) > (2) > (3) > (4)

213. Glyoxal is

a) $CH_2OH - CH_2OH$ b) $CHO - CH_2OH$ c) $COOH - CO - COOH$ d) $CHO - CHO$

214. IUPAC name of acraldehyde is

a) But-3-en-1-al b) Propenyl aldehyde

c) But-2-ene-1-al d) Prop-2-en-1-al

215. The IUPAC name of
$CH_3 - C \equiv CH(CH_3)_2$ is
a) 4-methyl-2-pentyne
b) 4, 4-dimethyl-2-butyne
c) methyl isopropyl acetylene
d) 2-methyl-4-pentyne

216. What information is provided by reaction mechanism?
a) The bonds broken and formed
b) The reaction intermediates
c) The relative rates of discrete steps, especially the slowest one
d) All of the above

217. The enolic from of acetone contains
a) 8σ bonds, 2π-bonds and 1 lone pair
b) 9σ bonds, 1π-bonds and 2 lone pairs
c) 9σ bonds, 2π-bonds and 1 lone pair
d) 10σ bonds, 1π-bonds and 1 lone pair

218. Which of the following acids has the smallest dissociation constant?
a) $CH_3CHFCOOH$
b) FCH_2CH_2COOH
c) $BrCH_2CH_2COOH$
d) $CH_3CHBrCOOH$

219. IUPAC name of, $CH_3CH(OH)CH_2CH_2COOH$ is :
a) 4-hydroxypentanoic acid
b) 1-carboxy-3-butanoic acid
c) 1-carboxy-4-butanol
d) 4-carboxy-2-butanol

220. The number of isomers for the compound with the molecular formula $C_2BrClFI$ is
a) 3
b) 4
c) 5
d) 6

221. Among the following the strongest nucleophile is
a) C_2H_5SH
b) CH_3COO^-
c) CH_3NH_2
d) $NCCH_2^-$

222. Which of the following statements is correct?
a) $+I$ group stabilises a carbocation
b) $+I$ group stabilises a carbanion
c) $-I$ group stabilises a carbocation
d) $-I$ group stabilises a free radical

223. Which of the following species does not exert a resonance effect?
a) $C_6H_5NH_2$
b) $C_6H_5\overset{+}{N}H_3$
c) C_6H_5OH
d) C_6H_5Cl

224. The number of different amines corresponding to the formula C_3H_9N is :
a) 2
b) 3
c) 4
d) 5

225. Lactic acid molecule has
a) One chiral carbon atom
b) Two chiral carbon atoms
c) No chiral carbon atom
d) asymmetric molecule

226. The arrangement of atoms that characterises a particular stereoisomer is called :
a) Geometry of isomer
b) Configuration
c) Conformers
d) None of these

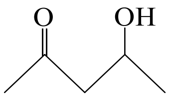
227. An alkane forms isomers if the number of carbon atoms is :
a) ≥ 1
b) ≥ 2
c) ≥ 3
d) ≥ 4

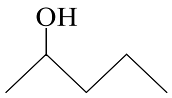
228. Which of the following statements is correct?
a) Desmotropism is another name for tautomerism
b) Allyl carbocation is less stable than isopropyl carbocation
c) $-I$ effect is exhibited by $-\overset{+}{N}H_3$
d) The formula CH_2Cl_2 is non-polar

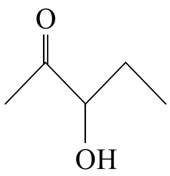
229. The IUPAC name of $CH_3COCH(CH_3)_2$ is
a) Isopropylmethyl ketone
b) 2-methyl-3-butanone
c) 4-methylisopropyl ketone
d) 3-methyl-2-butanone

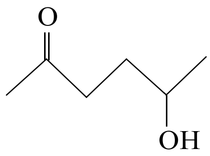
230. Qualitative test of halogens in an organic compound is made by
a) Fleming's test
b) Beilstein test
c) Bayer's test
d) Fehling's test

231. Which one of the following will most readily be dehydrated in acidic conditions?

a) 

b) 

c) 

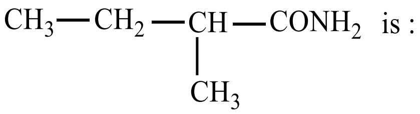
d) 

232. The family to which methoxyethene belongs, is :
a) Hydrocarbon b) Ketone c) Unsaturated ether d) Ester

233. Electrophiles are :
a) Electron loving species
b) Electron hating species
c) Nucleus loving reagents
d) Nucleus hating reagents

234. *Iso*-propyl chloride undergoes hydrolysis by
a) S_N1 mechanism b) S_N2 mechanisms
c) $S_N1$and S_N2 mechanisms d) Neither S_N1 nor S_N2 mechanism

235. The IUPAC name of the compound,

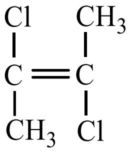


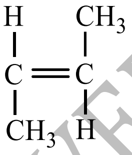
$CH_3-CH_2-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CONH_2$ is :

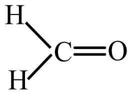
a) 2-ethylbutanamide
b) 2-methylbutanamide
c) 1-amino-2-methylpropane
d) None of the above

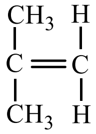
236. Carbon and hydrogen are estimated in organic compounds by
a) Kjeldalhl's method b) Duma's method c) Leibig's method d) Carius method

237. The compound having highest dipole moment is :

a) 

b) 

c) 

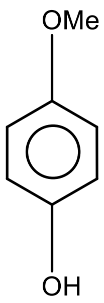
d) 

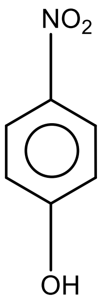
238. A free radical is :
a) Non-existing b) Short lived c) Diamagnetic d) Fairly stable

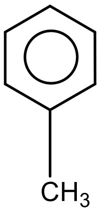
239. In 2-methyl-l-propanol, the hybrid carbons of sp^3, sp^2 and sp are respectively :
a) 3, 2, 1
b) 4, 3, 0
c) 4, 0, 0
d) 1, 2, 3

240. In electrophilic aromatic substitution reaction, the nitro group is *meta* directing because it
a) Decreases electron density at *ortho* and *para* positions
b) Decreases electron density at *meta* position
c) Increases electron density at *meta* postion
d) Increases electron density at *ortho* and *para* positions

241. Given,

$A =$ 

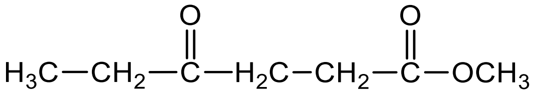
$B =$ 

$C =$ 

The decreasing order of the acidic character is

a) $A>B<C$ b) $B>A>C$ c) $B>C>A$ d) $C>B>A$

242. Give the IUPAC name for,



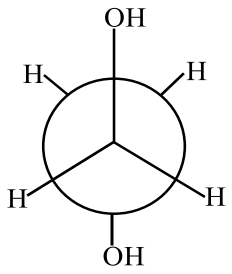
$H_3C-CH_2-\overset{\overset{O}{\|}}{C}-H_2C-CH_2-\overset{\overset{O}{\|}}{C}-OCH_3$

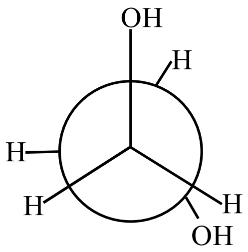
a) Ethyl-4- oxoheptanoate
b) Methyl-4- oxoheptanoate
c) ethyl-4- oxohexanoate
d) Methyl 4- oxohexanoate

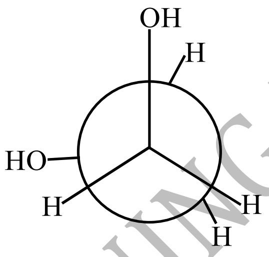
243. The total number of acyclic isomers including the stereoisomers with the molecular formula C_4H_7Cl

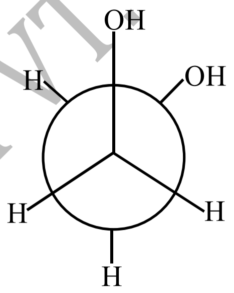
a) 11 b) 12 c) 9 d) 10

244. Which of the following conformers for ethylene glycol is most stable?

a) 

b) 

c) 

d) 

245. Which of the following compounds is resistant to nucleophilic attack by hydroxy ion?

a) Methylacetate b) Acetonitrile c) Acetamide d) Diethyl ether

246. The stabilization due to resonance is maximum in :

a) Cyclohexane b) Cyclohexene c) 1,3-cyclohexadiene d) 1,3,5-cyclohexatriene

247. A mixture of camphor and benzoic acid can be easily separated by

a) Sublimation b) Extraction with solvent
c) Fractional crystallisation d) Chemical method

248. Fractional crystallisations is carried out to separate a mixture of

a) Organic solids mixed with inorganic solids
b) Organic solids slightly soluble in water
c) Organic solids having small difference in their solubilities in suitable solvent
d) Organic solids having great difference in their solubilities in suitable solvent

249. The type of isomerism observed in urea molecule is :

a) Chain b) Position c) Geometrical d) Functional

250. Which of the following intermediate have the complete octet around the carbon atom?

a) Carbonium ion b) Carbanion c) Free radical d) Carbene

251. The name of, $(CH_3)_2$ HC— O— CH_2— CH_2— CH_3 is :

a) Isopropyl propyl ether
b) Dipropyl ether
c) di-isopropyl ether
d) Isopropyl propyl ketone

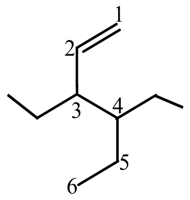
252. A neutral divalent carbon intermediate produced by the removal of two attached atoms is called :

a) Free radical b) Carbanion c) Carbocation ion d) Carbine

253. Which types of isomerism is shown by 2, 3-dichlorobutane?

a) Structural b) Geometric c) Optical d) Diastereo

254. The correct IUPAC name of the compound,



is :

a) 3-(1-ethyl propyl) hex-1-ene
b) 4-Ethyl-3-propyl hex-1-ene
c) 3-Ethyl-4-ethenyl heptane
d) 3-Ethyl-4-propyl hex-5-ene

255. IUPAC name of $(CH_3)_2N-C_2H_5$ is :
a) Dimethyl ethyl amine
b) Dimethylaminomethane
c) Dimethylaminoethane
d) N,N-dimethylethanamine

256. Among the following compounds, the most acidic is

a) p-nitrophenol	b) p-hydroxybenzoic acid
c) o-hydroxybenzoic acid	d) p-toluic acid

257. Electrophiles are :
a) Lewis bases b) Lewis acids c) Amphoteric d) None of these

258. On monochlorination of n-pentane, the number of isomers formed is :
a) 4 b) 3 c) 2 d) 1

259. Cyclohexane is :
a) Aliphatic compound
b) Alicyclic compound
c) Aromatic compound
d) Heterocyclic compound

260. Which of the following is a primary halide?

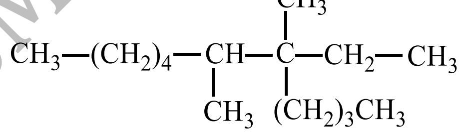
a) Isopropyl iodide	b) Secondary butyl iodide
c) Tertiary butyl bromide	d) Neo hexyl chloride

261. The percentage of $'s'$ character of the hybrid orbital of carbon in ethane, ethane and ethyne respectively are :
a) 25, 33, 50 b) 20, 50, 33 c) 25, 50, 75 d) 33, 66, 99

262. Which is a chiral molecule?
a) CH_3Cl b) CH_2Cl_2 c) $CHBr_3$ d) CHClBrI

263. The stability of a carbonium ion depends upon
a) The bond angle of the attached group
b) The substrate with which it reacts
c) The inductive effect and hyper-conjugative effect of the attached group
d) None of the above

264. The IUPAC name of the compound,



$$CH_3-(CH_2)_4-\underset{\displaystyle CH_3}{\underset{|}{CH}}-\overset{\displaystyle CH_3}{\overset{|}{\underset{\displaystyle (CH_2)_3CH_3}{\underset{|}{C}}}}-CH_2-CH_3$$

is :

a) 3,4-dimethyl-3-n-propylnonane
b) 4-ethyl-4,5-dimethyldecane
c) 6,7-dimethyl-7-n-propylnonane
d) 6,7-dimethyl-7-ethyldecane

265. Bromination of alkanes involves
a) Carbanions b) Carbocations c) Carbenes d) Free radicals

266. The isomeric *cis*-2-butene and *trans*-2-butene can be distinguished on the basis of :
   a) Their physical nature
   b) Their reduction products
   c) The products they give on ozonolysis
   d) The products they give on addition to bromine

267. Lassaigne's test is not used for the detection of
   a) Carbon   b) Halogens   c) Nitrogen   d) Sulphur

268. Consider the following carbocations,

$$\text{(I) } C_6H_5\overset{+}{C}H_2 \quad \text{(II) } C_6H_5CH_2\overset{+}{C}H_2$$

$$\text{(III) } C_6H_5\overset{+}{C}HCH_3 \quad \text{(IV) } C_6H_5\overset{+}{C}(CH_3)_2$$

   a) II<I<III<IV   b) II<III<I<IV   c) III<I<II<IV   d) IV<III<I<II

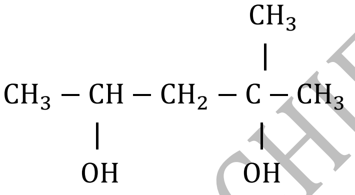
269. The simplest formula of a compound containing 50% of element X (at. wt 10) and 50% of element Y (at. wt. 20) is
   a) XY b) XY_2 c) X_2Y d) X_2Y_2

270. n-pentane, iso-pentane, and neo – pentane are examples for isomers of the type
   a) Geometrical   b) Optical   c) Chain   d) Positional

271. Homolytic fission of C— C bond in ethane gives an intermediate in which carbon is ….hybridized.
   a) sp^3 b) sp^2 c) sp d) sp^2d

272. Pick out the correct statement from the following and choose the correct answer from the codes given below
   I. Hexa-1, 5-diene is a conjugated diene
   II. Prop-1, 2-diene is conjugated diene
   III. Hexa-1, 3-diene is a conjugated diene
   IV. Buta-1, 3-diene is an isolated diene
   V. Prop-1, 2-diene is a cumulative diene
   a) I,II   b) II, III   c) IV, V   d) II, V

273. The IUPAC name of

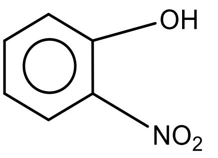


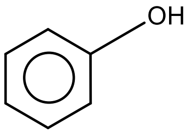
$$CH_3-\underset{OH}{\underset{|}{CH}}-CH_2-\underset{OH}{\underset{|}{\overset{CH_3}{\overset{|}{C}}}}-CH_3$$

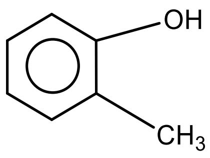
   a) 1, 1-dimethyl-1, 3-butanediol   b) 2-methyl-2, 4-pentanediol
   c) 4-methyl-2, 4-pentanediol   d) l, 3, 3-trimethyl-1, 3-propane diol

274. Which among the following is the correct IUPAC name of isoamylene?
   a) 1-pentene   b) 2-methyl-2-butene   c) 3-methyl-1-butene   d) 2-mythyl-1-butene

275. Which of the following compounds exhibits geometrical isomerism?
   a) C_2H_5Br b) $(CH)_2(COOH)_2$ c) CH_3CHO d) $(CH_2)_2(COOH)_2$

276. Which one of the following compounds, is most acidic?
   a) $Cl-CH_2-CH_2-OH$
   b) 

   c) 

   d) 

277. An organic compound has carbon and hydrogen percentage in the ratio 6:1 and carbon and oxygen percentages in the ratio 3:4. The compound has the empirical formula
   a) C_2H_6O b) CHO_2 c) CH_4O d) CH_2O

278. Among the following the one which does not exhibit functional group isomerism is :
   a) C_2H_6O b) C_3H_8O c) C_4H_{10} d) $C_4H_{10}O$

279. How many structural formulae are possible for $C_5H_{11}Cl$?
a) 6 b) 8 c) 10 d) 12

280. Which one of the following is a secondary alcohol?
a) 2-methyl-1-propanol b) 2-methyl-2-propanol c) 2-butanol d) l-butanol

281. Among the following anions $(a)\ \bar{C}H_3\ (b)\ \bar{N}H_2, (c) OH^-, (d) F^-$ the order of basicity is :
a) $a > b > c > d$ b) $b > a > c > d$ c) $c > b > a > d$ d) $c > a > b > d$

282. Electromeric effect is
a) Permanent effect b) Temporary effect c) Resonance effect d) Inductive effect

283. In the following groups, $-\underset{I}{OAC} -\underset{II}{OMe}-\underset{III}{OSO_2Me}-\underset{IV}{OSO_2CF_3}$ the order of leaving group ability is :
a) I > II > III > IV b) IV > III > I > II c) III > II > I > IV d) II > III > IV > I

284. The angle of rotation of plane of polarized light depends upon :
a) The nature of the light beam
b) The number of the molecules
c) The number of asymmetric carbon atoms in the molecule of the substance
d) All of the above

285. Which of the following shows geometrical isomerism?
a) C_2H_5Br b) $(CH_2)(COOH)_2$ c) $(CH)_2(COOH)_2$ d) C_2H_6

286. Which of the following cannot undergo nucleophilic substitution under ordinary conditions?
a) Chlorobenzene b) *Tert* −butylchloride c) Isopropyl chloride d) None of these

287. The C— C bond length of the following molecules is in the order.
a) $C_2H_6 > C_2H_4 > C_6H_6 > C_2H_2$
b) $C_2H_2 < C_2H_4 < C_6H_6 < C_2H_6$
c) $C_6H_6 > C_2H_2 > C_2H_6 > C_2H_4$
d) $C_2H_4 > C_2H_6 > C_2H_2 > C_6H_6$

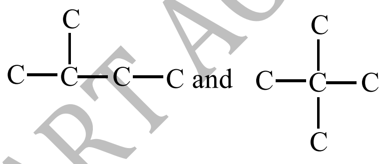
288. Isomerism among compounds due to the migration of a proton is known as :
a) Geometrical b) Optical c) Tautomerism d) Position

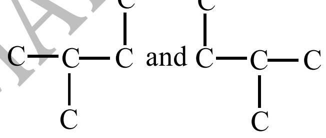
289. Removal of hydrogen atom is easier when it is attached to :
a) 1° carbon b) 2° carbon c) 3° carbon d) Same in all

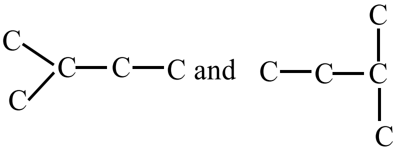
290. The order of stability of carbanions is :
a) $CH_3^- > 1° > 2° > 3°$ b) $3° > 2° > 1° > CH_3^-$ c) $3° > 1° > 2° > CH_3^-$ d) $2° > 3° > 1° > CH_3^-$

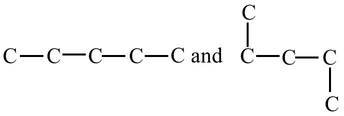
291. Glycerine contains
a) 1° carbon b) 2° carbon
c) 3° carbon d) Both 1° and 2° carbon

292. Which of the following pairs of carbon skeletons in an example of isomerism?

a)


b)


c)


d)


293. In cannizzaro reaction given below

$2\text{PhCHO} \xrightarrow{\text{OH}^{\ominus}} \text{PhCH}_2\text{OH} + \text{PhCO}_2^{\ominus}$

The slowest step is

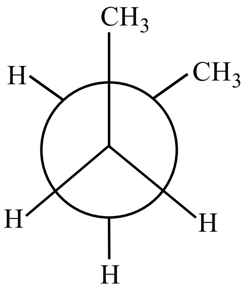
a) The attack of $::OH^{\ominus}$ at the carboxyl group

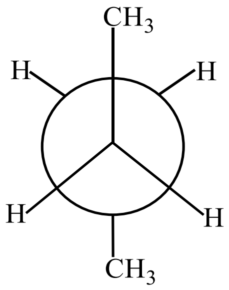
b) The transfer of hydride to the carbonyl group

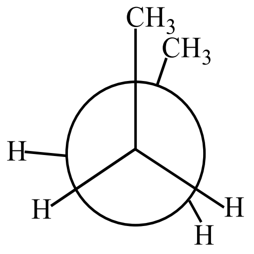
c) The abstraction of proton from the carboxylic group

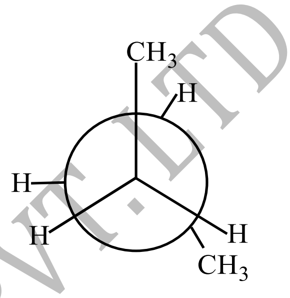
d) The deprotonation of $PhCH_2OH$

294. In the following the most stable conformation of n-butane is :

a)


b)


c)


d)


295. S_N1 mechanism for the reaction, $R - X + \text{KOH} \longrightarrow R\text{OH} + \text{K}X$ follow :

a) Carbocation mechanism

b) Carbanion mechanism

c) Free radical mechanism

d) Either of the above

296. An electrophilic reagent must have

a) A vacant orbital

b) An orbital containing one electron

c) An orbital containing two electrons

d) All completely filled atomic orbitals

297. In which of the following structures the number of sigma bonds are equal to the number of π-bonds?

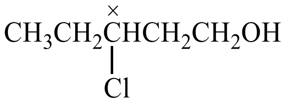
a) 1,2-propadiene

b) 2,3-dicyanobut-2-ene

c) Tetracyanoethylene

d) None of these

298. Which one of the starred carbons is the asymmetric one?

a)


$CH_3CH_2\overset{\times}{\underset{|\atop Cl}{C}}HCH_2CH_2OH$

b) $CH_3CH_2CH_2\overset{\times}{C}H_2CH_2OH$

c) $CH_3\overset{\times}{C}H_2CH_2CH_2CH_2OH$

d) $CH_3CH_2CH_2CH_2\overset{\times}{C}H_2OH$

299. The chemical name of anisole is

a) Ethanoic acid

b) Methoxy benzene

c) Propanone

d) Acetone

300. How many optically active stereomers are possible for butan-2, 3-diol?

a) 1

b) 2

c) 3

d) 4

301. Naphthalene molecule contains :

a) 10π-electrons

b) 8π-electrons

c) 12π-electrons

d) 14π-electrons

302. The first organic compound urea was synthesized in the laboratory by:

a) Kekule

b) Liebig

c) Lavoisier

d) Wöhler

303. In the hydrocarbon

$\underset{6}{C}H_3-\underset{5}{C}H=\underset{4}{C}H-\underset{3}{C}H_2-\underset{2}{C}\equiv\underset{1}{C}H$

The state of hybridization of carbons 1, 3 and 5 are in the following sequence :

a) sp, sp^3, sp^2

b) sp, sp^2, sp^3

c) sp^3, sp^2, sp

d) sp^2, sp, sp^3

304. Which of the following is not a nucleophile?

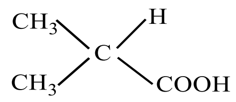
a) BF_3

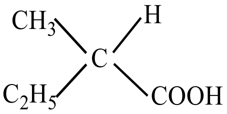
b) NH_3

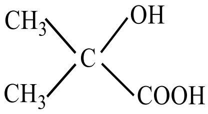
c) CN^-

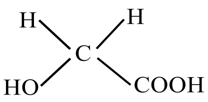
d) OH^-

305. Which of the following compounds can exhibit optical isomerism?

a) 

b) 

c) 

d) 

306. Which is the most stable carbocation?
a) *iso*- propyl cation
b) Triphenylmethyl cation
c) Ethyl cation
d) *n*-propyl cation

307. The correct structure of 4-bromo-3-methyl-but-1-ene.
a) $Br-CH=C(CH_3)_2$
b) $CH_2=CH-CH(CH_3)-CH_2Br$
c) $CH_2=C(CH_3)CH_2CH_2Br$
d) $CH_3-C(CH_3)=CHCH_2-Br$

308. IUPAC name of C_6H_5COCl is :
a) Benzoyl chloride
b) Benzenechloro ketone
c) Benzene carbonyl chloride
d) Chloro phenyl ketone

309. Stability order of... is in order

$$\underset{(I)}{C_6H_5-\overset{+}{C}H_2},\quad \underset{(II)}{CH_2=CH-\overset{+}{C}H_2},\quad \underset{(III)}{(CH_3)_3C^+},\ \underset{(IV)}{CH_2=\overset{+}{C}H}$$

a) IV<III<II<I
b) IV<II<I<III
c) I<II<III<IV
d) IV<I<III<II

310. Relative stabilities of the following carbocations will be in the order

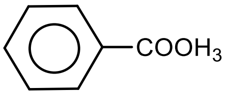
$$\underset{A}{\overset{\oplus}{C}H_3}\qquad \underset{B}{CH_3\,\overset{\oplus}{C}H_2}\qquad \underset{C}{\overset{\oplus}{C}H_2\,OCH_3}$$

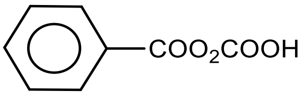
a) C>B>A
b) C<B<A
c) B>C>A
d) C>A>B

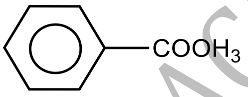
311. Which method is used to separate sugars?
a) Fractional crystallisation
b) Sublimation
c) Chromatography
d) Benedict's reagent

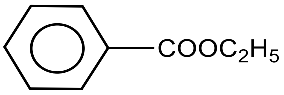
312. Sublimation can't be used for purification of
a) Benzoic acid
b) Camphor
c) Urea
d) Naphthalene

313. Which of the following is phenyl ethanoate?

a) 

b) 

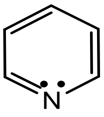
c) 

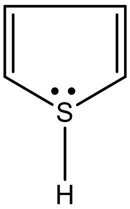
d) 

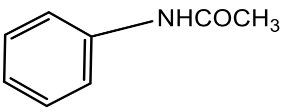
314. Zero inductive effect is shown by :
a) C_6H_5-
b) $-H$
c) CH_3-
d) $Cl-$

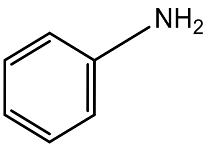
315. Which of the following alkyl halides is used as a methylating agent?
a) C_2H_5Cl
b) C_2H_5Br
c) C_2H_5I
d) CH_3I

316. Which one of the following has the most nucleophilic nitrogen?

a) 

b) 

c) 

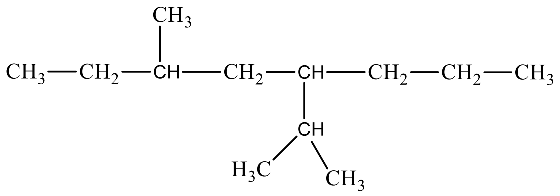
d) 

317. Chlorobenzene is 0, *p*-directing in electrophilic substituting reaction. The directing influence is explained by
a) $+M$ of Ph
b) $+I$ of Cl
c) $+M$ of Cl
d) $-I$ of Ph

318. Which of the following orders regarding relative stability of free radicals is correct?
a) 3°<2°<1° b) 3°>2°>1° c) 1°<2°>3° d) 3°>2°<1°

319. Carbon tetrachloride has no net dipole moment because of :
a) Its planar structure
b) Its regular tetrahedral nature
c) Similar sizes of carbon and chlorine atoms
d) Similar electron affinities of carbon and chlorine

320. IUPAC name of the compound



a) 4-isoprophyl, 6-methyl octane b) 3-methyl, 5-(1-methylethyl) octane
c) 3-methyl, 5-isopropyl octane d) 6-methyl, 4-(1-methylethyl) octane

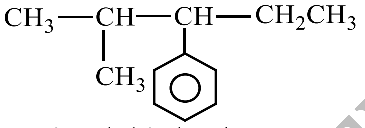
321. The isomers which are interconverted through rotation around a single bond are
a) Conformers b) Diastereomers c) Enantiomers d) Position isomers

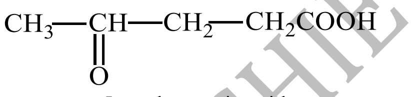
322. The number of optical isomers of pent-3-en-2-ol is :
a) 2 b) 4 c) 8 d) 16

323. Dehydrohalogenation of an alkyl halide is a/an
a) Nucleophilic substitution reaction
b) Elimination reaction
c) Both nucleophilic substitution and elimination reaction
d) Rearrangement

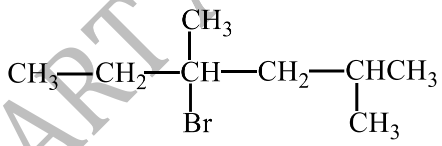
324. The minimum number of carbon atoms which a ketone may contain is :
a) 1 b) 2 c) 3 d) 4

325. Which nomenclature is not according to IUPAC system?

a)

2-methyl-3-phenylpentane

b)

$CH_3-\underset{\underset{O}{||}}{C}H-CH_2-CH_2COOH$
5-oxohexanoic acid

c) $Br-CH_2-CH=CH_2$
1-bromo-prop-2-ene

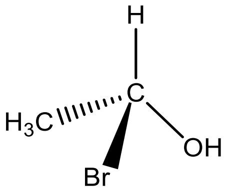
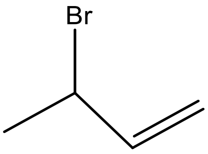
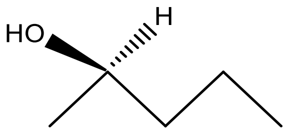
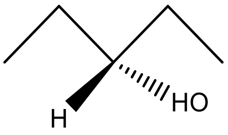
d)

$CH_3-CH_2-\overset{\overset{CH_3}{|}}{\underset{\underset{Br}{|}}{C}}H-CH_2-\underset{\underset{CH_3}{|}}{C}HCH_3$
4-bromo, 2, 4-di-methylhexane

326. What is the formula of tertiary butyl alcohol?
a) $CH_3-CH(CH_3)-CH_2-OH$ b) $CH_3-(CH_2)_2OH$
c) $CH_3-CH(OH)-CH_2-CH_3$ d) $(CH_3)_3.C-OH$

327. The IUPAC name of neopentane is
a) 2-methylbutane b) 2,2-dimethylpropane c) 2-methylpropane d) 2,2-dimethyl butane

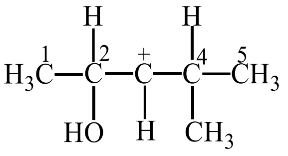
328. Select the strongest bond :
a) $>C-C<$ b) $>C=C<$ c) $-\overset{|}{\underset{|}{C}}-\underset{|}{C}=$ d) $-C\equiv C-$

329. Which of the following molecules is achiral?

a)  b)  c)  d) 

330. An enantiomerically pure acid is treated with racemic mixture of an alcohol having one chiral carbon. The ester formed will be
a) Optically active mixture
b) Pure enantiomer
c) Meso compound
d) Racemic mixture

331. In the following carbocation, H/CH_3 that is most likely to migrate to the positively charged carbon is :



$$\overset{1}{H_3C}-\underset{HO}{\overset{H}{\overset{2}{C}}}-\underset{H}{\overset{+}{C}}-\underset{CH_3}{\overset{H}{\overset{4}{C}}}-\overset{5}{CH_3}$$

a) CH_3 at C–4
b) H at C–4
c) CH_3 at C–2
d) H at C–2

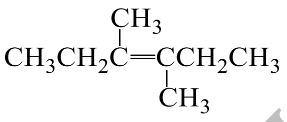
332. The number of π-electrons in benzene molecule is :
a) 3×2
b) 2^3
c) 3×3
d) 3^2

333. Which of the following statements is necessarily true in the case of isomeric organic compounds?
a) They are hydrocarbons
b) They are optically active
c) They yield the same products on complete combustion
d) They have same melting or boiling points

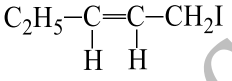
334. *Cis − trans*, isomers generally
a) Contain an asymmetric carbon atom
b) Rotate the plane of polarized light
c) Are enantiomorphs
d) Contain a double bonded carbon atoms

335. Among the following compounds nitrobenzene, benzene, aniline and phenol, the strongest basic behaviour in acid medium is exhibited by :
a) Phenol
b) Aniline
c) Nitrobenzene
d) Benzene

336. Geometrical isomerism is not shown by :

a) 

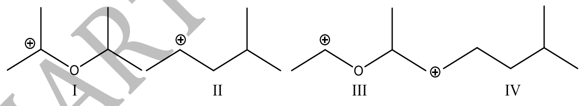
$$CH_3CH_2\overset{CH_3}{C}=\underset{CH_3}{C}CH_2CH_3$$

b) 

$$C_2H_5-\underset{H}{C}=\underset{H}{C}-CH_2I$$

c) $CH_2 = C(Cl)CH_3$
d) $CH_3— CH = CH_2— CH = CH_2$

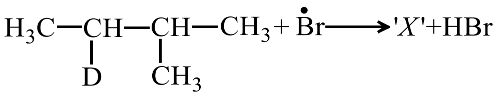
337. The correct stability order for the following species as



a) II>IV>I>III
b) I>II>III>IV
c) II>I>IV>III
d) I>III>II>IV

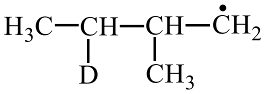
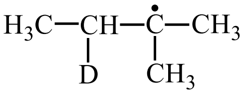
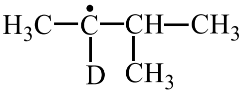
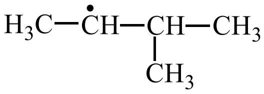
338. An optically active compound is :
a) 1-bromobutane
b) 2-bromobutane
c) 1-bromo-2-methyl propane
d) 2-bromo-2-methyl propane

339. Consider the following reaction,



$$H_3C-\underset{D}{CH}-\underset{CH_3}{CH}-CH_3+\dot{B}r\longrightarrow 'X'+HBr$$

Identify the structure of the major product 'X' :

a) 
$H_3C-\underset{D}{CH}-\underset{CH_3}{CH}-\dot{C}H_2$
b) 
$H_3C-\underset{D}{CH}-\underset{CH_3}{\dot{C}}-CH_3$
c) 
$H_3C-\underset{D}{\dot{C}}-\underset{CH_3}{CH}-CH_3$
d) 
$H_3C-\dot{C}H-\underset{CH_3}{CH}-CH_3$

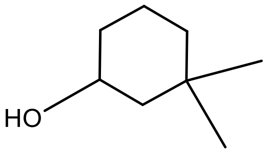
340. In Lassaigen's test, the organic compound is fused with a piece of sodium metal in order to
a) Increase the ionization of the compound
b) Decrease the melting point of the compound
c) Increase the reactivity of the compound
d) Convert the covalent compound into a mixture of ionic compounds

341. Which of the following sodium compound/compounds are formed when an organic compound containing both nitrogen and sulphur is fused with sodium?
a) Cyanide and sulphide
b) Thiocyanate
c) Sulphite and cyanide
d) Nitrate and sulphide

342. The IUPAC name of $CH_3-CH=CH-C\equiv CH$ is
a) Pent-3-en-1-yne
b) Pent-3-en-4-yne
c) Pent-2-en-4-yne
d) Pent-2-en-3-yne

343. The compound having molecular formula $C_4H_{10}O$ can show :
a) Metamerism
b) Functional isomerism
c) Positional isomerism
d) All of these

344. The IUPAC name of the compound



a) 3, 3-dimethyl-1-hydroxy cyclohexane
b) 1, 1-dimethyl-3- hydroxy cyclohexane
c) 3, 3- dimethy-1- cyclohexanol
d) 1,1-dimethyl-3-cyclohexanol

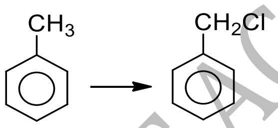
345. In hyperconjugation, the atom involved is :
a) β-H atom
b) α-H atom
c) $\gamma-$ H atom
d) All of these

346. Reactivity of hydrogen atoms attached to different atoms in alkanes has the order :
a) $3° > 1° > 2°$
b) $1° > 2° > 3°$
c) $3° > 2° > 1°$
d) None of these

347. Which has maximum percentage of chlorine?
a) Pyrene
b) PVC
c) Chloral
d) Ethylidene chloride

348. $H_2C=O$ behaves as :
a) Nucleophile
b) Electrophile
c) Both (a) and (b)
d) None of these

349. The most stable carbocation is :
a) $\overset{+}{C}H_3$
b) $CH_3\overset{+}{C}H_2$
c) $(CH_3)_2\overset{+}{C}H$
d) $(CH_3)_3\overset{+}{C}$

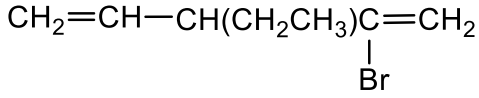
350.


The above reaction proceeds through
a) Free radicals substitution
b) Nucleophilic substitution
c) Electrophilic substitution
d) None of the above

351. Which reaction sequence would be best to prepare 3-chloro-aniline from benzene?
a) Chlorination, nitration, reduction
b) Nitration, chlorination, reduction
c) Nitration, reduction, chlorination
d) Nitration, reduction, acylation, chlorination, hydrolysis

352. Why is light necessary to bring in chlorination reactions of alkane?
a) The dissociation of Cl_2 gives $\dot{Cl}$ free radical
b) The Cl_2 molecule absorbs light to show hemolytic bond fission
c) The formation of $\dot{Cl}$ free radical propagate the chain reaction
d) All of the above

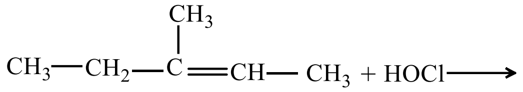
353. IUPAC name of



$CH_2{=}CH{-}CH(CH_2CH_3)C(Br){=}CH_2$

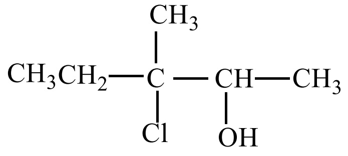
a) 4-bromo-3-ethyl-1,4-pentadiene
b) 2-bromo-3-ethyl-1,4pentadiene
c) 2-bromo-3-ethyl-1-5-pentadiene
d) None of the above

354. The product of reaction,

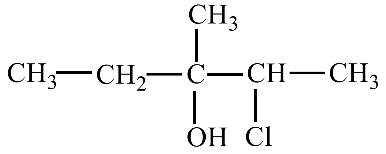


$CH_3{-}CH_2{-}C(CH_3){=}CH{-}CH_3 + HOCl \longrightarrow$

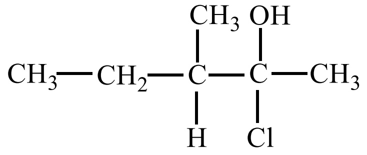
product is :

a) 

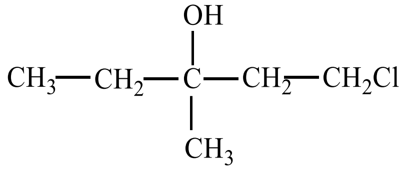
$CH_3CH_2{-}C(CH_3)(Cl){-}CH(OH){-}CH_3$

b) 

$CH_3{-}CH_2{-}C(CH_3)(OH){-}CH(Cl){-}CH_3$

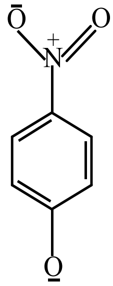
c) 

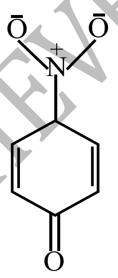
$CH_3{-}CH_2{-}C(CH_3)(H){-}C(OH)(Cl){-}CH_3$

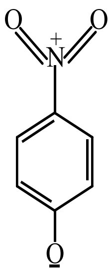
d) 

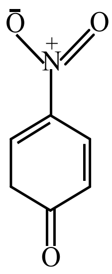
$CH_3{-}CH_2{-}C(OH)(CH_3){-}CH_2{-}CH_2Cl$

355. The most unlikely representation of resonance structures of p-nitrophenoxide ion is :

a) 

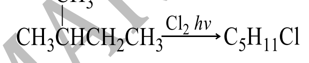
b) 

c) 

d) 

356. For all practical purposes, influence of inductive effect is neglected after :

a) 2nd carbon atom	b) 1st carbon atom	c) 3rd carbon atom	d) None of these

357. 

$CH_3CH(CH_3)CH_2CH_3 \xrightarrow{Cl_2\ h\nu} C_5H_{11}Cl$

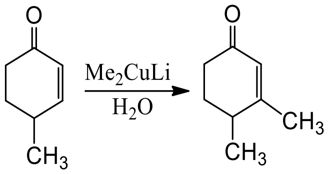
N (isomeric produces) $\xrightarrow[\text{distillation}]{\text{Fractional}}$ M (isomeric product)

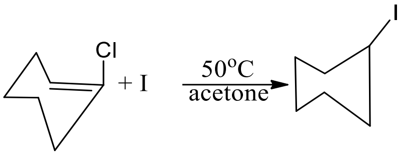
what are the no. of N and M?

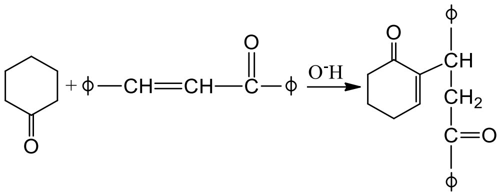
a) 6, 6	b) 6, 4	c) 4, 4	d) 3, 3

358. Geometrical isomers differ in :
a) Position of functional groups
b) Position of atoms
c) Spatial arrangement of atoms
d) Length of carbon chain

359. Which of the following is an example of substitution reaction?

a) 

b) 

c) 

d) None of the above

360. The study of three dimensional structure of molecule is called :
a) Stereochemistry b) Solid state chemistry c) Chirality d) None of these

361. Orbital interaction between the σ-bonds of a substituent group and a neighbouring π-orbital is known as
a) Hyperconjugation
b) Inductive effect
c) Steric effect
d) Electric quadrapole interactions

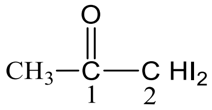
362. The shape of $\bar{\ddot{C}}H_3$ is :
a) Linear b) Planar c) Pyramidal d) None of these

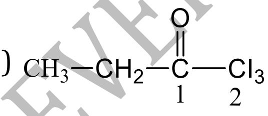
363. Which of the following contains only three pairs of electrons?
a) Carbocation
b) Carbanion
c) Free radical
d) None of these

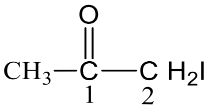
364. 2-hexyne gives *trans*-2-hexene on treatment with
a) Li/NH_3 b) $Pd/BaSO_4$ c) $LiAlH_4$ d) Pt/H_2

365. Lassaigne's test for the detection of nitrogen fails in
a) $H_2N-CO-NHNH_2.HCl$
b) $NH_2-NH_2.HCl$
c) $NH_2-CO-NH_2$
d) $C_6H_5-NH-NH_2.HCl$

366. Which of the following compounds yields most stable carbanion after rupture $(\underset{1}{C}-\underset{2}{C})$ of bond?

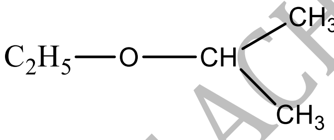
a) 
$CH_3-\underset{1}{\overset{O}{\overset{\|}{C}}}-\underset{2}{C}HI_2$

b) 
$CH_3-CH_2-\underset{1}{\overset{O}{\overset{\|}{C}}}-\underset{2}{C}l_3$

c) 
$CH_3-\underset{1}{\overset{O}{\overset{\|}{C}}}-\underset{2}{C}H_2I$

d) None of these

367. The IUPAC name of C_2H_5-o-CH



is

a) Ethoxy propane b) 1, 1-dimethyl ether c) 2-ethoxy *iso*-propane d) 2-ethoxy propane

368. The relative adsorption of each component of the mixture is expressed in terms of
a) adsorption factor
b) retention factor
c) co-factor
d) sorption factor

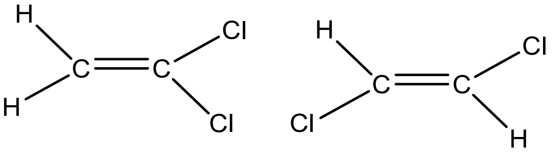
369. Following reaction,
$(CH_3)_3CBr + H_2O \rightarrow (CH_3)_3COH + HBr$
is an example of
a) Elimination reaction
b) Free radical substitution
c) Nucleophilic substitution
d) Electrophilic substitution

370. The most stable carbonium ion among the following is
a) $C_6H_5CH_2\overset{+}{C}H_2$ b) $CH_3\overset{+}{C}H_2$ c) $C_6H_5\overset{+}{C}HC_6H_5$ d) $C_6H_5\overset{+}{C}H_2$

371. *t*-butyl alcohol is

a) 2-methyl propane-2-ol b) 2-methyl propane-1-ol
c) 3-methyl butan-1-ol d) 3-methyl butan-2-ol

372. The following compound differ in



a) Configuration b) Conformation c) Structure d) Chirality

373. A compound containing 80% C and 20% H is likely to be
a) C_6H_6 b) C_2H_6 c) C_2H_4 d) C_2H_2

374. Overlap of which of the following atomic orbitals would be maximum to form the strongest covalent bond?
a) $1s-2s(\sigma)$ b) $1s-2p(\sigma)$ c) $2p-2p(\pi)$ d) $2p-2p(\sigma)$

375. A strong base can abstract an α-hydrogen from :
a) Amine b) Ketone c) Alkane d) Alkene

376. During elimination reactions, the hybrid state of carbon atoms involved in change shows:
a) sp^3 to sp^2 nature
b) sp^2 to sp nature
c) No change in hybridized state
d) Either of the above

377. IUPAC name of C_6H_5CN is :
a) Phenyl nitrile b) Benzene nitrile c) Benzyl nitrile d) Phenyl cyanide

378. Who proposed the tetrahedral mirror image structures to a pair of enantiomers?
a) Kekule b) Wöhler c) van't Hoff d) None of these

379. The S_N1 reactivity of following halides will be in the order
(i) $(CH_3)_3CBr$ (ii)$(C_6H_5)_2CHBr$
(iii)$(C_6H_5)_2C(CH_3)Br$ (iv) $(CH_3)_2CHBr$
(v) C_2H_5Br
a) (v)>(iv)>(i)>(ii)>(iii) b) (ii)>(i)>(iii)>(v)>(iv)
c) (i)>(iii)>(v)>(ii)>(iv) d) (iii)>(ii)>(i)>(iv)>(v)

380. Heterolysis of propane gives :
a) Methyl and ethyl free radicals
b) Methylium cation and ethyl anion
c) Methyl anion and ethylium cation
d) Methylium and ethylium cations

381. Delocalised electrons are present in
a) 1, 3- butadiene b) C_6H_6 c) 1, 3, 5-hexatriene d) All of these

382. Compounds which rotate plane polarised light in clockwise direction are known as :
a) Dextrorotatory
b) Laevorotatory
c) Optically inactive compounds
d) Racemic

383. Carbanions initiate :
a) Addition reactions
b) Substitution reactions
c) Both (a) and (b)
d) None of these

384. Impure glycerine can be purified by
a) Steam distillation b) Simple distillation
c) Vaccum distillation d) Extraction with a solvent

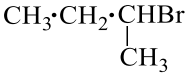
385. IUPAC name of urea is :

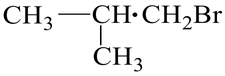
a) Diaminoketone
b) 1-aminoethanamide
c) 1-aminomethanamide
d) aminoacetamide

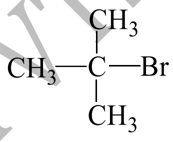
386. Which of the following process is not used for the purification of solid impurities?
a) Distillation b) Sublimation c) Crystallisation d) Vaporisation

387. When the hybridization state of a carbon atom changes from sp^3 to sp^2 and finally to sp, the angle between the hybridized orbitals :
a) Is not affected
b) Increases progressively
c) Decreases considerably
d) Decreases gradually

388. The chief reaction product of reaction in between n-butane and bromine at 130°C is :
a) $CH_3CH_2{\cdot}CH_2{\cdot}CH_2Br$

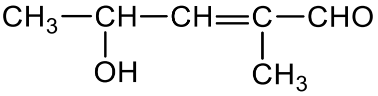
b) 
$CH_3{\cdot}CH_2{\cdot}\underset{\displaystyle CH_3}{\underset{|}{C}}HBr$

c) 
$CH_3-\underset{\displaystyle CH_3}{\underset{|}{C}}H{\cdot}CH_2Br$

d) 
$CH_3-\overset{\displaystyle CH_3}{\overset{|}{\underset{\displaystyle CH_3}{\underset{|}{C}}}}-Br$

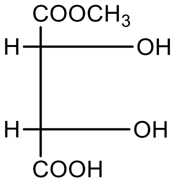
389. Dehydration of alcohol is an example of which type of reaction?
a) Substitution b) Elimination c) Addition d) Rearrangement

390. The IUPAC name of

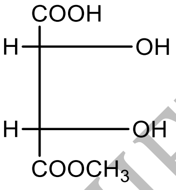

$CH_3-\underset{\displaystyle OH}{\underset{|}{C}}H-CH=\underset{\displaystyle CH_3}{\underset{|}{C}}-CHO$

a) 4-hydroxy-1-methylpentanal
b) 4-hydroxy-4-methylpent-2-en-1-al
c) 2-hydroxy-4-methylpent-2-en-5-al
d) 2-hydroxy-3-methylpent-2-en-5-al

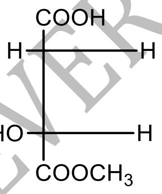
391. The correct statement about the compounds (A), (B) and (C) is



(A)



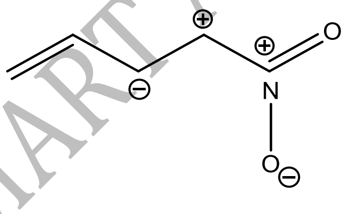
(B)

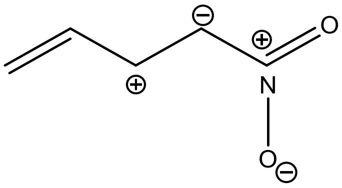


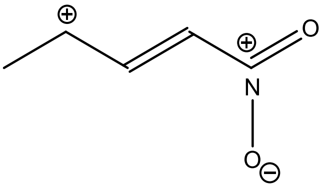
(C)

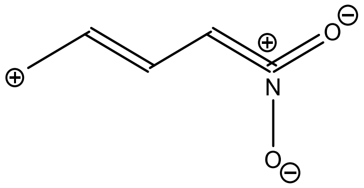
a) (A) and (B) are identical b) (A) and (B) are diastereomers
c) (A) and (C) are enantiomers d) (A) and(B) are enantiomers

392. Among the following the least stable resonance structure is

a) 

b) 

c) 

d) 

393. The organic liquid that mix freely with water is :
a) $CHCl_3$ b) CCl_4 c) CS_2 d) C_2H_5OH

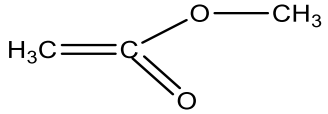
394. The increasing order of +ve I-effect shown by H, CH_3, C_2H_5 and C_3H_7 is :

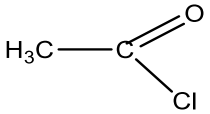
a) $H < CH_3 < C_2H_5 < C_3H_7$
b) $H > CH_3 < C_2H_5 > C_3H_7$
c) $H < C_2H_5 < CH_3 < C_3H_7$
d) None of the above

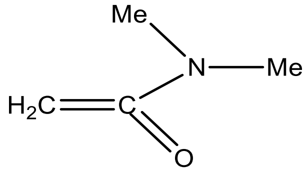
395. The best method for the separation of naphthalene and benzoic acid from their mixture is
a) Chromatography b) Crystallisation c) Distillation d) Sublimation

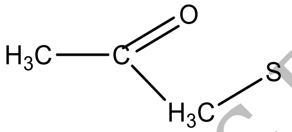
396. The reagent used in dehydrohalogenation process is :
a) Alcoholic KOH b) $NaNH_2$ c) C_2H_5ONa d) All of these

397. The least active electrophile is

a) 

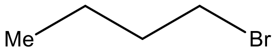
b) 

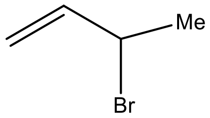
c) 

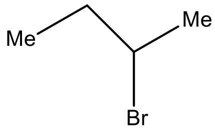
d) 

398. The isomerism which exists between CH_3CHCl_2 and CH_2ClCH_2Cl is :
a) Chain b) Functional c) Positional d) Metamerism

399. Consider the following bromides

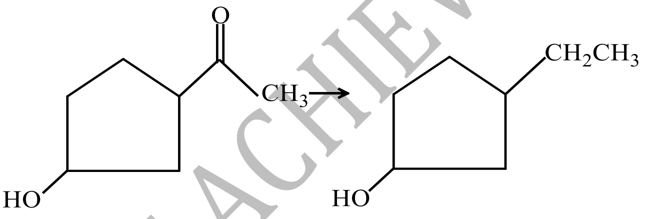
(A) 

(B) 

(C) 

The correct order is S_N1 reactivity is
a) (B)>(C)>(A) b) (B)>(A)>(C) c) (C)>(B)>(A) d) (A)>(B)>(C)

400. The appropriate reagent for the following transformation,



is :

a) Zn(Hg), HCl
b) NH_2NH_2, OH^-
c) H_2/Ni
d) $NaBH_4$

401. Hydride shift from C-2 will give the most stable resonance stabilized carbocation as
a) CH_3 at $C - 4$ b) H at C – 4 c) CH_3 at $C - 2$ d) H at C-2

402. Hyperconjugation involves overlap of the following orbitals
a) σ- σ b) $\sigma - \rho$ c) $p - p$ d) $\pi - \pi$

403. Most stable carbonium ion is
a) $\overset{+}{C}_2H_5$ b) $(CH_3)_3\overset{+}{C}$ c) $(C_6H_5)_3\overset{+}{C}$ d) $C_6H_5\overset{+}{C}H_2$

404. During a nitration of benzene, the attacking electrophile is
a) NO_3^- b) NO_2^- c) NO_2^+ d) HNO_3

405. The (*R*) and (*S*) enantiomers of an optically active compound differ in
a) Their reactivity b) Their optical rotation of plane polarised light

c) Their melting point d) Their solubility in achiral reagents

406. The number of chiral centres in (+) –glucose
a) 4 b) 3 c) 2 d) 1

407. Hydrogen cyanide and hydrogen isocyanide are :
a) Tautomers
b) Positional isomers
c) Metamers
d) Chain isomers

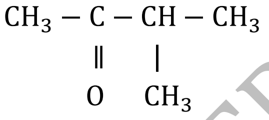
408. Which of the following hydrocarbons is most unsaturated?
a) C_2H_4 b) C_2H_2 c) C_2H_6 d) $CH_3CH = CH_2$

409. Sometimes the behaviour of a compound is explained by assuming that it exists in a world between two or more different possible structures. This phenomenon is called :
a) Isomerism b) Resonance c) Mutarotation d) Allotropism

410. How many primary amines are possible with the formula $C_4H_{11}N$?
a) 1 b) 2 c) 3 d) 4

411. Which one of the following pairs represents stereoisomerism?
a) Geometrical isomerism, position isomerism
b) Geometrical isomerism, conformational isomerism
c) Optical isomerism, geometrical isomerism
d) Optical isomerism, metamerism

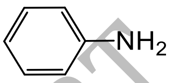
412. The large number of organic compounds is due to:
a) Catenation property of carbon
b) Covalent bond formation
c) Isomerism
d) polymerization

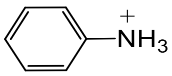
413. The IUPAC name of 
$CH_3 - \underset{\underset{O}{\|}}{C} - \underset{\underset{CH_3}{|}}{CH} - CH_3$ is
a) 2-methyl-3-butanone b) 3-methyl-butan-2-one
c) 3-methyl butanone d) None of these

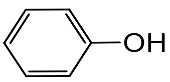
414. Formic acid is a stronger acid than acetic acid. This can be explained using
a) $+M$effect b) $-I$ effect c) $+I$ effect d) $-M$ effect

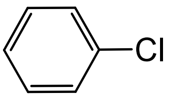
415. The energy of C— C triple bond in acetylene in kcal is :
a) 140 b) 192 c) 60 d) 100

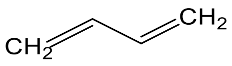
416. In which of the following molecules, the resonance effect is not present?

a) 

b) 

c) 

d) 

417. Which of the following represents the given mode of hybridization $sp^2 - sp^2 - sp - sp$ from left to right?
a) $CH_2 = CH - C \equiv CH$ b) $HC \equiv C - C \equiv N$ c) $CH_2 = C - C = CH_2$ d) 

418. The Lassaigne's extract is boiled with dil HNO_3 before testing for halogens because
a) AGCN is soluble in HNO_3 b) Silver halides are soluble in HNO_3
c) Na_2S and NaCN are decomposed by HNO_3 d) Ag_2S is soluble in HNO_3

419. The tautomeric form which is less stable is called :
a) Anion form b) Cation form c) Labile form d) All of these

420. The effect involving the complete transfer of a shared pair of electrons to one of the atoms joined by a multiple bond at the requirement of attacking reagent is called :
a) Inductive effect b) Mesomeric effect c) Electromeric effect d) None of these

421. Which of the following acids does not exhibit optical isomerism?
a) Lactic acid b) Tartaric acid c) Maleic acid d) α-amino acids

422. Many organic compounds are prepared by using PCl_5 because :
a) OH group of alcohol is easily replaced by chlorine atom
b) Chlorines are added to the unsaturated compounds
c) It removes water from organic compounds
d) Phosphorus atoms are entered in the alcohol

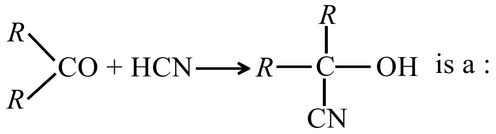
423. Which of the following conformations of cyclohexane is chiral?
a) Twist boat b) Rigid c) Chair d) Boat

424. Which type of isomerism is shown by propanal and propanone?
a) Functional group b) Metamerism c) Tautomerism d) Chain isomerism

425. Identify the product in the given reaction:
$CH_3 - CH = CH_2 + NOCl \longrightarrow$ Product
a) $CH_3CHCl.CH_2.NO$ b) $CH_3CH(NO).CH_2Cl$ c) $CH_3CH_2CH(Cl)(NO)$ d) $CH_2(NO).CH_2.CH_2Cl$

426. A straight chain hydrocarbon has the molecular formula C_8H_{10} . The hybridization for the carbon atoms from one end of the chain to the other are respectively $sp^3, sp^2, sp^2, sp^3, sp^2, sp^2, sp$ and sp. The structural formula of the hydrocarbon would be :
a) $CH_3 - C \equiv C - CH_2 - CH = CH - CH = CH_2$
b) $CH_3 - CH_2 - CH = CH - CH = CH - CH \equiv CH$
c) $CH_3 - CH = CH - CH_2 - C \equiv C - CH = CH_2$
d) $CH_3 - CH = CH - CH_2 - CH = CH - C \equiv CH$

427. Reaction,



$$\begin{matrix} R \\ R \end{matrix}\!\!>\!CO + HCN \longrightarrow R-\underset{\displaystyle CN}{\overset{\displaystyle R}{C}}-OH$$ is a :

a) Electrophilic substitution
b) Nucleophilic substitution
c) Electrophilic addition
d) Nucleophilic addition

428. The total number of acylic iosmers including the stereoisomers (geometrical and optical), with the molecular formula C_4H_7Cl is
a) 12 b) 11 c) 10 d) 9

429. The best method to separate the mixture of *ortho* −and−*para* nitrophenol (1:1) is
a) Vaporisation b) Colour spectrum c) Distillation d) Crystallisation

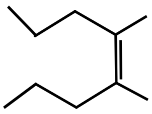
430. Which of the following does not show electrometric effect?
a) Alkenes b) Ethers c) Aldehyde d) Ketones

431. Shifting of electrons of a multiple bond under the influence of a reagent is called :
a) I-effect b) E-effect c) M-effect d) T-effect

432. 0.4 g of a silver salt of a monobasic orgainc acid gave 0.26 g pure silver on ignition. the molecular weight of the acid is (atomic weight of silver=108)
a) 58 b) 37 c) 89 d) 105

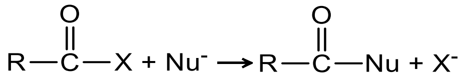
433. The S_N2 mechanism for, $R - X + KOH(aq) \longrightarrow R - OH + KX$ follows with :
a) 100% inversion b) 50% inversion c) 40% inversion d) 30% inversion

434.
IUPAC name of,



is :
a) 4,5-dimethyl oct-4-ene
b) 3,4-dimethyl oct-5-ene
c) 4,5-dimethyl oct-5-ene
d) None of the above

435. The reaction,



$R-\overset{\overset{O}{\|}}{C}-X + Nu^- \longrightarrow R-\overset{\overset{O}{\|}}{C}-Nu + X^-$

is fastest when X is

a) OCOR

b) OC_2H_5

c) NH_2

d) Cl

436. Which pair represents chain isomers?

a) CH_3CHCl_2 and $ClCH_2CH_2Cl$

b) Propyl alcohol and isopropyl alcohol

c) 2-methylbutane and neopentane

d) Diethylether and dipropylether

437. The empirical formula of an acid is CH_2O_2, the probable molecular formula of the acid may be

a) $C_2H_4O_2$ b) $C_3H_6O_4$ c) $C_2H_2O_4$ d) CH_2O_2

438. The number of valence electrons in the excited carbon atom is :

a) Two in s and two in p-orbitals

b) 4 single p-orbitals

c) One in s and three in p-orbitals

d) None of the above

439. A hydrocarbon contains 10.5 g carbon and 1 g hydrogen. Its 2.4 g has 1 L volume at 1 atm and127°C. Hydrocarbon is

a) C_6H_7 b) C_6H_8 c) C_5H_6 d) C_6H_6

440. The number of stereoisomers possible for a compound of the molecular formula $CH_3-CH=CH-CH(OH)-Me$ is

a) 3 b) 2 c) 4 d) 6

441. The structural formula of methyl aminomethane is :

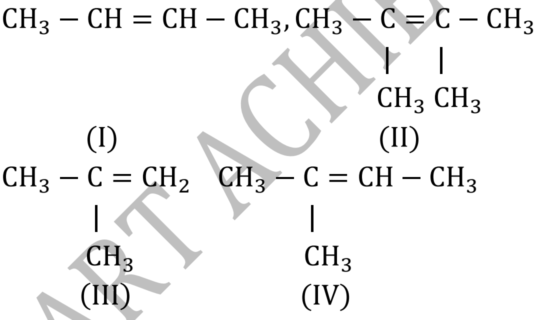
a) $(CH_3)_2CHNH_2$ b) $(CH_3)_3N$ c) $(CH_3)_2NH$ d) CH_3NH_2

442. A mixture of oil and water is separated by

a) Filtration b) Fractional distillation

c) Sublimation d) Using separating funnel

443. The stability of



$CH_3-CH=CH-CH_3$ (I), $CH_3-\underset{\underset{CH_3}{|}}{C}=\underset{\underset{CH_3}{|}}{C}-CH_3$ (II)

$CH_3-\underset{\underset{CH_3}{|}}{C}=CH_2$ (III) $CH_3-\underset{\underset{CH_3}{|}}{C}=CH-CH_3$ (IV)

In the increasing order is

a) III<I<IV<II b) I<II<III<IV c) IV<III<II<I d) II<III<IV<I

444. The number of optically active isomers of tartaric acid are

a) 1 b) 3 c) 4 d) 2

445. The nodal plane in the π-bond of ethene is located in

a) The molecular plane

b) A plane parallel to the molecular plane

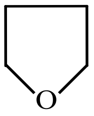
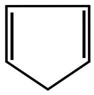
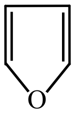
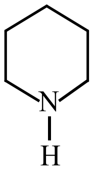
c) A plane perpendicular to the molecular plane which bisects the carbon-carbon sigma bond at right angle

d) A plane perpendicular to the molecular plane which contains the carbon-carbon sigma bond

446. Which of the following has the highest degree of coordination bond?

a) CH_3OH b) $AlCl_3$ c) $(CH_3)_4\overset{+}{N}\overset{-}{\ddot{O}}H$ d) $BF_3O(Et)_2$

447. Which of the following is heterocyclic aromatic species?

a) 
b) 
c) 
d) 

448. A mixture of *o*-nitrophenol and *p*-nitrophenol can be separated by
a) Fractional crystallisation b) Sublimation
c) Chemical separation d) Steam distillation

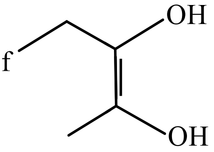
449. The total number of cyclic structural as well as stero isomers possible for a compound with the molecular formula C_5H_{10} is
a) 2 b) 4 c) 6 d) 7

450. State the hybridization of carbon present in triplet carbene
a) sp^3 b) sp^2 c) sp d) None of these

451. Which of the following cannot show electromeric effect ?
a) Alkenes b) Ketones c) Aldehydes d) Ethers

452. Hydride ion transfer takes place in
a) Frankland method b) Wurts reaction
c) Cannizzaro's reaction d) Wolff-Kishner reduction

453. An organic compound contains 29.27% carbon, 5.69 % hydrogen and 65.04% bromine. Its empirical formula is
a) C_3H_5Br b) C_3H_3Br c) $C_2H_4Br_2$ d) C_3H_7Br

454. IUPAC name of  is :
a) but-2-ene-2,3-diol
b) pent-2-ene-2,3-diol
c) 2-methylbut-2-ene-2,3-diol
d) Hex-2-ene-2,3-diol

455. The IUPAC name of $(CH_3)_3C-CH=CH_2$is
a) 1, 1, 1-trimethyl-2-propene b) 3, 3, 3-trimethyl-2-propene
c) 2, 2-dimethyl-3-butene d) 3, 3-dimethyl-1-butene

456. The function of soda lime, a mixture of solid NaOH and solid CaO during decarboxylation of carboxylic acids is :
a) To increase the rate of reaction
b) To decrease the rate of reaction
c) To change the rate of reaction
d) None of the above is correct

457. *t*-butyl chloride reacts with OH^- by S_N1 mechanism and rate $\propto$ [*t*-buty1 chloride]. One of the reasons for this is that
a) Stereochemical inversion takes place
b) *t*- buty1 carbocation is first formed which is more stable
c) The product *t*-butyl alcohol is more stable
d) The intermediate *t*-butyl carbocation is stabilized by solvation

458. Heterolysis of $CH_3CH_2CH_3$ result in formation of
a) $\overset{\oplus}{C}H_3$ and $\overset{-}{C}_2H_5$ b) $\overset{\bullet}{C}H_3$ and $\overset{\bullet}{C}_2H_5$ c) $\overset{-}{C}H_3$ and $\overset{\oplus}{C}_2H_5$ d) $CH_3$and C_2H_5

459. Alkyl c yanide $R-C\equiv N$ and alkyl isocyanides $R-N\overset{=}{\rightarrow}C$ are :

a) Tautomers b) Metamers c) Functional isomers d) Geometrical isomers

460. A racemic mixture is a mixture of :
a) *meso* and its isomers
b) *d*-and its *l*-isomers in equal proportions
c) *d*-and its *l*-isomers in different proportions
d) *meso* and *d*-isomers

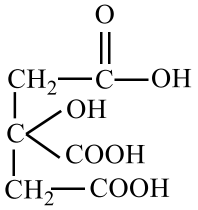
461. A mixture of iodine and sodium chloride can be easily separated by
a) Fractional distillation b) Steam distillation
c) Chromatography d) Sublimation

462. The property by virtue of which a compound can turn the plane of polarization of light is known as :
a) Photolysis b) Phosphorescence c) Optical activity d) polarization

463. Correct order of stability is
a) $HC \equiv \bar{C} > CH_2 = \bar{C}H > CH_3 - \bar{C}H_2$ b) $CH_3 - \bar{C}H_2 > CH_2 = \bar{C}H > CH \equiv \bar{C}$
c) $CH_3 - \bar{C}H_2 > CH \equiv CH \cong CH_2 = \bar{C}H$ d) All are equally stable

464. In the estimation of sulphur in an organic compound, fuming nitric acid is used to convert sulphur into
a) SO_2 b) H_2S c) H_2SO_3 d) H_2SO_4

465. The IUPAC name of compound

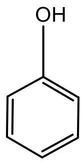


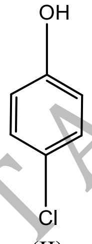
is :
a) 1,2,3-tricarboxy-2,1-propane
b) 3-carboxy-3-hydroxy-1,5-pentanedioic acid
c) 3-hydroxy-3-carboxy-1,5-pentanedioic acid
d) None of the above

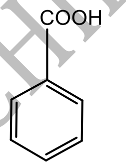
466. Which of the following will be chiral?
a) CH_3CHCl_2 b) $CH_3CHBrCl$ c) CD_2Cl_2 d) CH_2ClBr

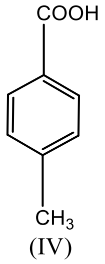
467. In the dehydration reaction $CH_3CONH_2 \xrightarrow{P_2O_5} CH_3C \equiv N$, the hybridization state of carbon change from
a) lsp^3 to sp^2 b) lsp to sp c) lsp^2 to sp d) lsp to sp^3

468. The correct acidity order of the following is


(I)


(II)


(III)


(IV)

a) (III)>(IV)>(II)>(I) b) (IV)>(III)>(I)>(II) c) (III)>(II)>(I)>(IV) d) (II)>(III)>(IV)>(I)

469. Which of the following is not a nucleophile?
a) BF_3 b) CN^- c) OH^- d) NH_3

470. Least stable conformer of cyclohexane is
a) Chair b) Boat c) Twist boat d) Planar hexagon

471. The compound isomeric with acetone is :
a) Propionaldehyde b) Propionic acid c) Ethoxy ethane d) None of these

472. Which of the chloride is less reactive towards hydrolysis?
a) Vinyl chloride b) Allyl chloride c) Ethyl chloride d) *t* −butyl chloride

473. Glycerol is an alcohol which can be classified as
a) Trihydric b) Monohydric c) Dihydric d) Hexahydric

474. Addition of Br_2 on $CH_2 = CH_2$ in presence of NaCl(*aq*.) gives :

a) $CH_2Br.CH_2Br$ b) $CH_2Br.CH_2Cl$ c) $CH_2Br.CH_2OH$ d) All of these

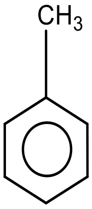
475. The electromeric effect in organic compounds is a :
   a) Temporary effect
   b) Permanent effect
   c) Temporary-permanent effect
   d) None of the above

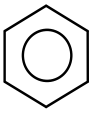
476. The function of boiling the sodium extract with conc. HNO_3 before testing for halogen is
   a) To make the solution acidic b) To make the solution clear
   c) To convert Fe^{2+} to Fe^{3+} d) To destroy CN^- and S^{2-} ions

477. Copper wire test of halogens is known as
   a) Liebig's test b) Lassaigne's test c) Fusion test d) Beilstein's test

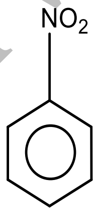
478. Which of the following is singlet carbine?
   a) $(CH_3)_3C^+$ b) $C_2H_5\ddot{C}—H$ c) $CH_3\,\ddot{\bar{C}}\,HCH_3$ d) $CH_2=CH—\overset{+}{C}H_2$

479. Which of the following will be easily nitrated?

a)


b)


c) CH_3NO_2

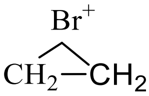
d)


480. Optical isomerism is shown by
   a) Propanol-2 b) Butanol-2 c) Ethanol d) Methanol

481. Williamson's synthesis involves
   a) S_N1 mechanism b) Nucleophilic addition
   c) S_N2 mechanism d) S_E mechanism

482. Free radicals can undergo :
   a) Disproportionation to two species
   b) Rearrangement to a more stable free radical
   c) Decomposition to give another free radical
   d) All of the above are correct

483. During addition of bromine on ethene, the first species formed is

a)


b) $C_2H_4OH^+$ c) $\overset{+}{C}H_2CH_2Br$ d) $C_2H_5^+$

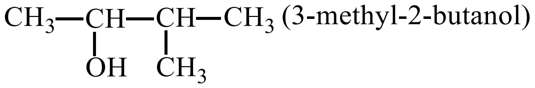
484. Metamers of ethyl propionate are
   a) C_4H_9COOH and $HCOOC_4H_9$ b) C_4H_9COOH and $CH_3COOC_3H_7$
   c) CH_3COOCH_3 and $CH_3COOC_3H_7$ d) $CH_3COOC_3H_7$ and $C_3H_7COOCH_3$

485. Which statement is correct about the hybridization of carbon atoms in,

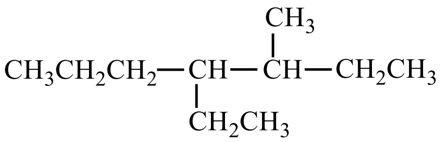
$$H\overset{1}{C}\equiv\overset{2}{C}-\overset{3}{C}\equiv\overset{4}{C}H\,?$$

   a) C_1 and C_4 are sp^2-hybridized
   b) C_2 and C_3 are sp^2-hybridized
   c) All are sp-hybridized
   d) All are sp^2-hybridized

486. Which one is not is IUPAC system?

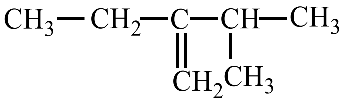
a)


$$\underset{\displaystyle OH}{CH_3—\underset{|}{C}H}—\underset{\underset{\displaystyle CH_3}{|}}{CH}—CH_3 \text{ (3-methyl-2-butanol)}$$

b) 

$CH_3CH_2CH_2-\underset{\underset{CH_2CH_3}{|}}{CH}-\overset{\overset{CH_3}{|}}{CH}-CH_2CH_3$

(3-methyl-4-ethyl heptane)

c) 

$CH_3-CH_2-\underset{\underset{CH_2}{\|}}{C}-\underset{\underset{CH_3}{|}}{CH}-CH_3$

(2-ethyl-3-methyl-but-1-ene)

d) $CH_3-C\equiv C-CH(CH_3)_2$ (4-methyl-2-pentyne)

487. The compound which exhibits optical isomerism is :

a) $CH_3CHOHCH_3$ b) $(CH_3)_2CHCH_2CH_3$ c) $CH_3CHClCH_2CH_3$ d) $CH_3CCl_2CH_2CH_3$

488. Which of the following applies in the reaction $CH_3CHBrCH_2CH_3 \xrightarrow{\text{Alco.KOH}}$?

(I)$CH_3CH=CHCH_3$(Major product)

(II)$CH_2=CHCH_2CH_3$(Minoe product)

a) Hofmann’s rule b) Saytzeff’s rule c) Kharasch effect d) Markownikoff’s rule

489. Homologous compounds have :

a) Same chemical properties

b) Same molecular weight

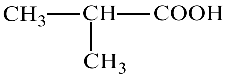
c) Same physical properties

d) Same m.p. and b.p.

490. How many chiral compound are possible on mono chlorination of 2-methyl butane?

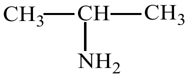
a) 2 b) 4 c) 6 d) 8

491. Which of the following may exist in enantiomorphs?

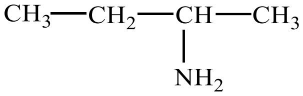
a) 

$CH_3-\underset{\underset{CH_3}{|}}{CH}-COOH$

b) $CH_3=CHCH_2CH_2CH_3$

c) 

$CH_3-\underset{\underset{NH_2}{|}}{CH}-CH_3$

d) 

$CH_3-CH_2-\underset{\underset{NH_2}{|}}{CH}-CH_3$

492. How many isomers are possible for the alkane C_4H_{10}?

a) 3 b) 5 c) 2 d) 4

493. Which of the following IUPAC names is correct?

a) 2-methyl-3-ethylpentane

b) 2-ethyl-3-methylpentane

c) 3-ethyl-2-methylpentane

d) 3-methyl-2-ethylpentane

494. Ethyl acetoacetate shows, which type of isomerism?

a) Chain b) Optical c) Metamerism d) Tautomerism

495. Enol content is highest in

a) Acetone b) Acetophenone c) Acetic acid d) Acetyl acetone

496. The maximum number of stereoisomers possible for 3-hydroxy-2-methyl butanoic acid is :

a) 1 b) 2 c) 3 d) 4

497. Which of the following will exhibit $cis-trans$ isomerism?

a) CH_2Br-CH_2Br b) CBr_3-CH_3 c) $CHBr=CHBr$ d) $CBr_2=CH_2$

498. Which of the following is most reactive towards electrophilic nitration?

a) Toluene b) Benzene c) Benzoic acid d) Nitrobenzene

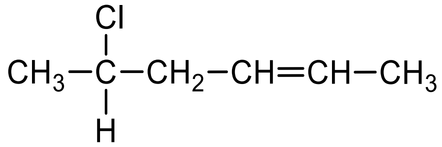
499. The number of structural isomers possible for an organic compound with molecular formula C_5H_{12} is :

a) 5 b) 3 c) 4 d) 2

500. Ethylene dichloride and ethylidine chloride are isomeric compounds. The false statement about these isomers is that they

a) React with alcoholic potash and give the same product
b) Are position isomers
c) Contain the same percentage of chlorine
d) Are both hydrolysed to the same product

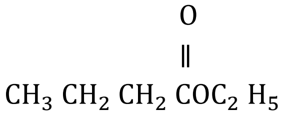
501. The IUPAC name of



$CH_3-\underset{H}{\overset{Cl}{C}}-CH_2-CH=CH-CH_3$

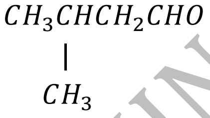
a) 5-chloro-hex-2-ene
b) 2-chloro-hex-5-ene
c) l-chloro-1-methyl-pent-3-ene
d) 5-chloro-5-methyl-pent-2-ene

502. Which of the following compounds has incorrect IUPAC nomenclature?

a) 

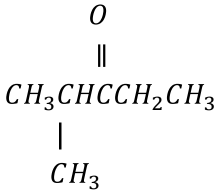
$CH_3\ CH_2\ CH_2\ \overset{\overset{O}{\|}}{C}OC_2\ H_5$

ethylbutanoate

b) 

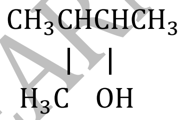
$CH_3\underset{\underset{CH_3}{|}}{CH}CH_2CHO$

3-methyl butanal

c) 

$CH_3\underset{\underset{CH_3}{|}}{CH}\overset{\overset{O}{\|}}{C}CH_2CH_3$

2-methyl-3-pentanone

d) 

$CH_3\underset{\underset{H_3C}{|}}{CH}\underset{\underset{OH}{|}}{CH}CH_3$

2-methyl-3-butanol

503. The IUPAC name for tertiary butyl iodide is

a) 4-iodo butane
b) 2-iodo butane
c) 1-iodo-3-methyl propane
d) 2-iodo-2-methyl propane

504. Geometry of methyl free radical is

a) Pyramidal
b) Planar
c) Tetrahedral
d) Linear

505. Dehydrogenation of ethanol to give ethanal is :

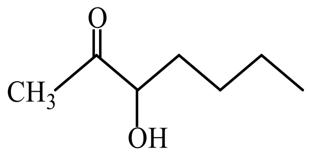
a) Addition reaction
b) α-α elimination reaction
c) α-β elimination reaction
d) α-γ elimination reaction

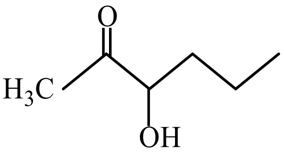
506. The arrangement of decreasing order of stability of

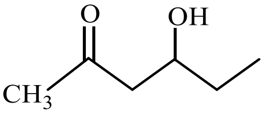
$\dot{C}H_3$, $\dot{C}_2H_5$, $(CH_3)_2\dot{C}H$ and $(CH_3)_3\dot{C}$ free radicals is :

a) $\dot{C}H_3 > \dot{C}_2H_5 > (CH_3)_2\dot{C}H > (CH_3)_3\dot{C}$
b) $(CH_3)_3\dot{C} > (CH_3)_2\dot{C}H > \dot{C}_2H_5 > \dot{C}H_3$
c) $\dot{C}_2H_5 > \dot{C}H_3 > (CH_3)_2\dot{C}H > (CH_3)_3\dot{C}$
d) $(CH_3)_3\dot{C} > (CH_3)_2\dot{C}H > \dot{C}H_3 > \dot{C}_2H_5$

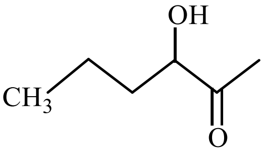
507. Which one of the following compounds will be most readily dehydrated ?

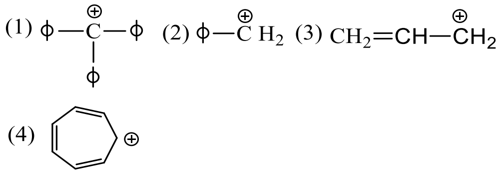
a) 

b) 

c) 

d)



508. 

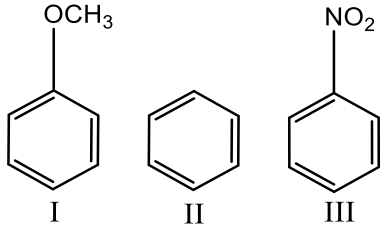
Correct order of stability is

a) 1>4>2>3 b) 1>2>3>4 c) 1>2>4>3 d) 1>3>4>2

509. The organic chloro compound, which shows complete stereochemical inversion during and S_N2 reaction, is

a) CH_3Cl b) $(C_2H_5)_2CHCl$ c) $(CH_3)_3CCl$ d) $(CH_3)_2CHCl$

510. Among the following compounds (I-III) the correct order of reaction with the electrophile is



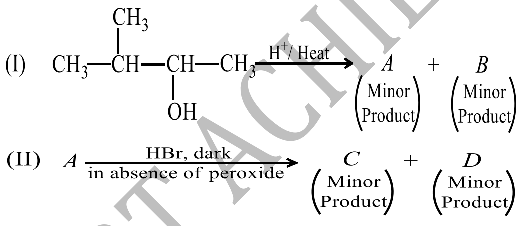
I II III

a) II>III>I b) III<I<II c) I>II>III d) I≈II>III

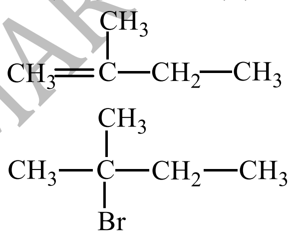
511. Which of the following is an electrophile?

a) $:CCl_2$ b) CO_2 c) H_2O d) NH_3

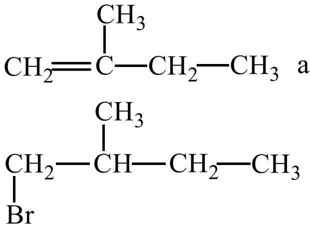
512. In the following reactions,



the major products (A) and (C) are respectively :

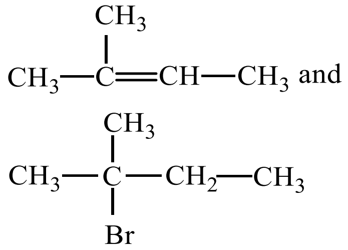
a) 

$CH_3{=}C(CH_3){-}CH_2{-}CH_3$ and $CH_3{-}C(CH_3)(Br){-}CH_2{-}CH_3$

b) 

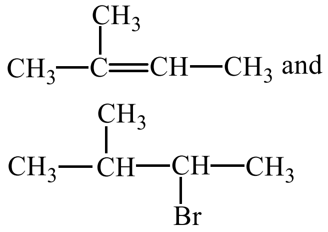
$CH_2{=}C(CH_3){-}CH_2{-}CH_3$ and $CH_2(Br){-}CH(CH_3){-}CH_2{-}CH_3$

c)



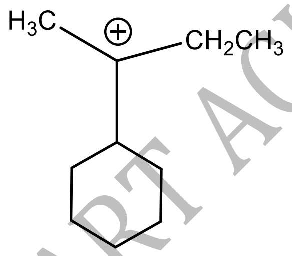
$CH_3-\overset{\overset{CH_3}{|}}{C}=CH-CH_3$ and $CH_3-\underset{\underset{Br}{|}}{\overset{\overset{CH_3}{|}}{C}}-CH_2-CH_3$

d)



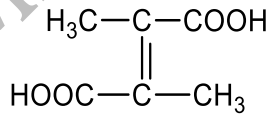
$CH_3-\overset{\overset{CH_3}{|}}{C}=CH-CH_3$ and $CH_3-\overset{\overset{CH_3}{|}}{CH}-\underset{\underset{Br}{|}}{CH}-CH_3$

513. An organic compound having carbon, hydrogen and sulphur contains 4% of sulphur. The minimum molecular weight of the compound is
   a) 500 b) 800 c) 400 d) 100
514. The structure of tertiary butyl carbonium ion is :
   a) Pyramidal b) Trigonal planar c) Tetrahedral d) Square planar
515. A carbonium ion contains :
   a) A + vely charged carbon centre
   b) A –vely charged carbon centre
   c) A carbon with odd electron on it
   d) None of the above
516. The formula of 3-chloro-2,2-dimethylbutane is :
   a) $CH_3CH(CH_3)C(CH_3)_2Cl$ b) $CH_3(CH_3)_2CH_2Cl$ c) $CH_3C(CH_3)_2CH_2Cl$ d) $CH_3CHClC(CH_3)_3$
517. Which shows the easier electrophilic substitution in ring?
   a) N-acetyl aniline b) $C_6H_5NH_3Cl$ c) Aniline d) Nitrobenzene
518. The number of isomeric pentyl alcohols possible is
   a) Two b) Four c) Six d) Eight
519. Naphthalene can be easily purified by
   a) Sublimation b) Crystallisation c) Distillation d) Vaporisation
520. The total number of contributing structures showing hyperconjugation (involving – C – H bonds) for the following carbocation is



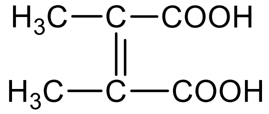
   a) Three b) Five c) Eight d) Six
521. What is the structural formula of fumaric acid?

a)



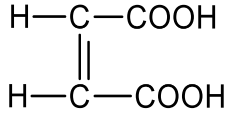
$\begin{array}{c} H_3C-C-COOH \\ \| \\ HOOC-C-CH_3 \end{array}$

b)

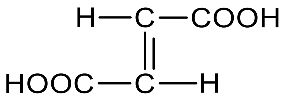


$\begin{array}{c} H_3C-C-COOH \\ \| \\ H_3C-C-COOH \end{array}$

c)



$\begin{array}{c} H-C-COOH \\ \| \\ H-C-COOH \end{array}$

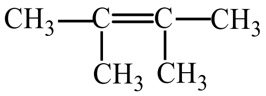
d) 
$\text{H—C—COOH} \,\|\, \text{HOOC—C—H}$

522. The correct structure of dimethylbutyne is :

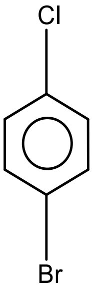
a) $CH_3CH_2—C\equiv C—CH_2CH_3$

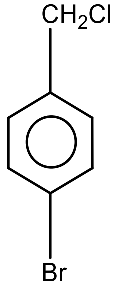
b) $(CH_3)_3C—C—C\equiv CH$

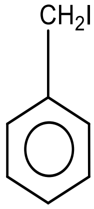
c) $CH_3—C\equiv CCH(CH_3)_2$

d) 
$CH_3—\underset{CH_3}{\underset{|}{C}}=\underset{CH_3}{\underset{|}{C}}—CH_3$

523. Which one of the following gives white precipitate with AgNO_3?

a) 

b) 

c) 

d) None of these

524. The reaction $(CH_3)_3CBr \xrightarrow{H_2O} (CH_3)_3\,C.OH$ is:

a) Elimination reaction

b) Free radical reaction

c) Substitution reaction

d) Displacement reaction

525. Which of the following is free radical?

a) Cl^+ b) Cl^- c) $\dot{Cl}$ d) NO_2

526. An organic compound X having molecular formula $C_6H_7O_2N$ has 6 carbons in a ring system, two double bonds and also a nitro group as substitutent. The X is :

a) Homocyclic but not aromatic

b) Aromatic but not homocyclic

c) Homocyclic and aromatic

d) Heterocyclic

527. Quantitative measurement of nitrogen in an organic compound is done by the method

a) Berthelot method b) Belstein method c) Lassaigne test d) Kjheldahl method

528. During pyrolysis of an alkane, C— C bond breaks faster than the C— H bond because :

a) C— C bond is stronger

b) C— H bond is weaker

c) C— C bond involves π-bond in alkane

d) The bond energy of C— C bond is less than that of C— H bond

529. State of hybridization of carbon atom of carbene in the singlet state is :

a) sp^2 b) sp c) sp^3 d) None of these

530. IUPAC name of $(CH_3)_3CCl$is

a) n-butyl chloride b) 3-chloro butane

c) 2-chloro 2-methyl propane d) t −butyl chloride

531. Most stable carbocation is formed during the heating of which one of the following compound with conc.H_2SO_4?

a) $(CH_3)_3COH$ b) $C_6H_5CH_2OH$ c) $(CH_3)_2CHOH$ d) $CH_3CH(OH)CH_2CH_3$

532. The number of 1° and 2°carbon atoms in n-pentane are respectively :

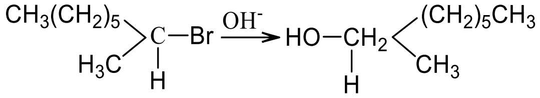
a) 2, 3 b) 3, 2 c) 2, 4 d) 1, 3

533. In benzene, all the C − C bonds are of equal length because of :

a) Isomerism b) Resonance c) Tautomerism d) Inductive effect

534. Nitration of benzene is
a) Electrophilic substitution
b) Electrophilic addition
c) Nucleophilic substitution
d) Nucleophilic addition

535. The following reaction is described as



a) S_E 2 b) S_N 2 c) S_N 1 d) S_N 0

536. Nitrogen containing organic compound when fused with sodium forms
a) Sodium azide b) Sodium cyanide c) Sodamide d) Sodium cyanate

537. Which of the following is strongest nucleophile?
a) Br^- b) $\vdots OH$ c) $\vdots CN$ d) $C_2H_5\ddot{O}\vdots$

538. In Kjeldahl's method for the estimation of nitrogen, the formula used to
a) % of N $= \frac{1.4\,Vw}{N}$ b) % of N $= \frac{1.4\,VN}{w}$ c) % of N $= \frac{VNw}{1.8}$ d) % of N $= \frac{1.4\,wN}{V}$

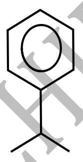
539. The most satisfactory method to separate sugars is to use
a) Fractional crystallisation
b) Chromatography
c) Benedict's reagent
d) Carius method

540. The IUPAC name of an unsymmetrical ether with the molecular formula $C_4H_{10}O$
a) Ethoxy propane b) Methoxy ethane c) Ethoxy ethane d) Methoxy propane

541. S_N1 reaction on optically active substrates mainly gives :
a) Retention in configuration
b) Inversion in configuration
c) Racemic product
d) No product

542. The structures that do not actually exist are called :
a) Tautomers
b) Conformational isomers
c) Canonical structures
d) Optical isomers

543.
The correct IUPAC name of,



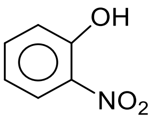
is :

a) Isopropyl benzene b) Cumene c) Phenyl isopropane d) 2-phenyl propane

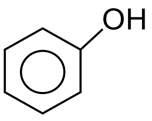
544. When SCN^- is added to an aqueous solution containing $Fe(NO_3)_3$,the complex ion produced is
a) $[Fe(OH_2)_2(SCN)]^{2+}$ b) $[Fe(OH_2)_5(SCN)]^{2+}$ c) $[Fe(OH_2)_8(SCN)]^{2+}$ d) $[Fe(OH_2)(SCN)]^{6+}$

545. Which of the following is the most stable carbocation?
a) $\overset{+}{C}H_3$ b) $R\overset{+}{C}H_2$ c) $R_2\overset{+}{C}H$ d) $R_3\overset{+}{C}$

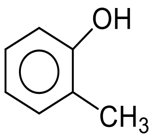
546. Which one of the following compound is most acidic?
a) $Cl-CH_2-CH_2-OH$ b)



c)



d)



547. The number of carbon atoms present in neopentane are :
a) Four 1° carbon, one 4° carbon
b) two 1° carbon, two 2° carbon
c) one 1° carbon, three 4° carbon
d) None of the above is correct

548. The Cl—C—Cl angle in 1,1,2,2-tetrachloroethene and tetrachloromethane respectively will be about :

a) 120° and 109.5° b) 90° and 109.5° c) 109.5° and 90° d) 109.5° and 120°

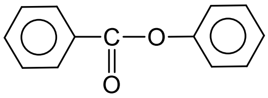
549. $C_6H_5C \equiv N$ and $C_6H_5N \equiv C$ exhibit which type of isomerism?

a) Position b) Functional c) Metamerism d) Dextroisomerism

550. The number of possible enantiomeric pairs that can be produced during monochlorination of 2-methyl butane is :

a) 2 b) 3 c) 4 d) 1

551.



In

Electrophilic substitution occurs at

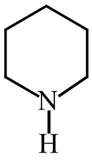
a) *ortho/para* at first ring b) *meta* at first ring

c) *ortho/para* at second ring d) *meta* at second ring

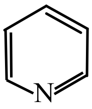
552. In estimation of nitrogen by Duma's method 1.18 g of an orgainc compound gave 224 mL of N_2 at NTP. The percentage of nitrogen in the compound is

a) 20.0 b) 11.8 c) 47.7 d) 23.7

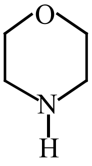
553. In the following compounds,



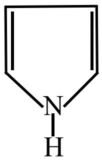
I



II



III



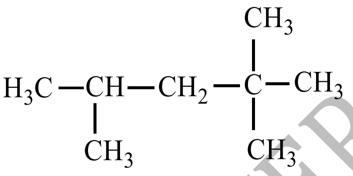
IV

the order of basicity is :

a) IV > I > III > II b) III > I > IV > II c) II > I > III > IV d) I > III > II > IV

554.

In the structure,



$H_3C-\underset{CH_3}{CH}-CH_2-\underset{CH_3}{\overset{CH_3}{C}}-CH_3$

the number of carbons are :

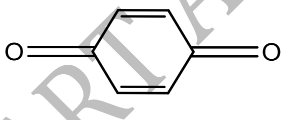
a) One primary, two secondary and one tertiary

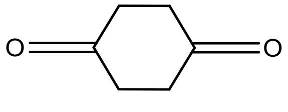
b) Four primary, two tertiary and one secondary

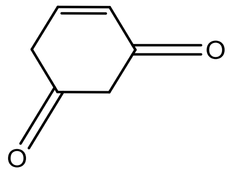
c) One primary, one secondary, one tertiary and one quaternary

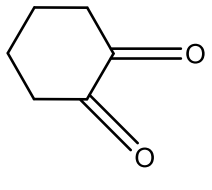
d) Five primary, one secondary, one tertiary and one quaternary

555. Which of the following does not exhibit tautomerism?

a) 

b) 

c) 

d) 

556. Chromatography technique is used for the separation of

a) Small sample of mixture b) Plant pigments

c) Dyestuff d) All of the above

557. CH_3CH_2OH and CH_3OCH_3 are the example of

a) Chain isomerism b) Functional isomerism c) Position isomerism d) Metamerism

558. The number of geometrical isomers in case of a compound with the structure, $CH_3-CH=CH-CH=CH-C_2H_5$ are .

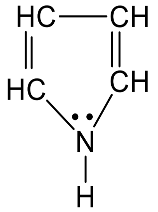
a) Four b) Three c) Two d) five

559. The stability of the carbocation decreases in the order

a) $R_2CH^+ > R_3C^+ > RCH_2^+ > CH_3^+$ b) $R_3C^+ > R_2CH^+ > RCH_2^+ > CH_3^+$

c) $CH_3^+ > R_2CH^+ > RCH_2^+ > R_3C^+$　　d) $CH_3^+ > RCH_2^+ > R_2CH^+ > R_3C^+$

560. How many delocalized π-electrons are there in the compounds

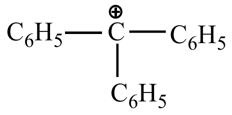


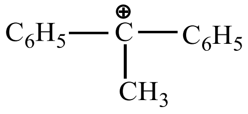
a) 8　　b) 2　　c) 4　　d) 6

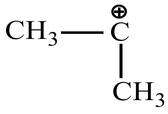
561. What will be the compound if two valencies of carbonyl group are satisfied by two alkyl groups?

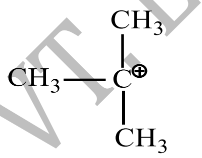
a) Aldehyde　　b) Ketone　　c) Acid　　d) Acidic anhydride

562. The most stable carbocation is :

a) 
$C_6H_5-\underset{C_6H_5}{\overset{\oplus}{C}}-C_6H_5$

b) 
$C_6H_5-\underset{CH_3}{\overset{\oplus}{C}}-C_6H_5$

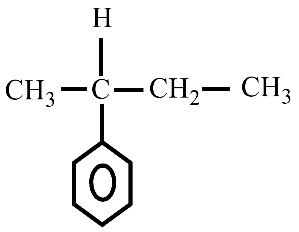
c) 
$CH_3-\underset{CH_3}{\overset{\oplus}{C}}$

d) 
$CH_3-\underset{CH_3}{\overset{CH_3}{C^{\oplus}}}$

563. Which of the following belongs to $-I$ group?

a) $-C_6H_5$　　b) $-CH_3$　　c) $-CH_2CH_3$　　d) $-C(CH_3)_3$

564. IUPAC name of following compound,



is :

a) 2-cyclohexylbutane　　b) 2-phenylbutane　　c) 3-cyclohexylbutane　　d) 3-phenylbutane

565. Which is most commonly used to dry organic liquids?

a) Lithium　　b) Sodium　　c) Potassium　　d) Rubidium

566. In Lassaigne's solution, pink/violet colouration is produced when sodium nitroprusside solution is added. It indicates the presence of

a) Sulphur　　b) Nitrogen　　c) Chlorine　　d) None of these

567. A carbonium ion is formed when a covalent bond between two atoms in an organic compound undergoes :

a) Homolysis　　b) Heterolysis　　c) Cracking　　d) Pyrolysis

568. Racemic mixture is formed by mixing two

a) Isomeric compounds　　b) Chiral compounds

c) *meso* compounds　　d) Enantiomers with chiral carbon

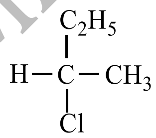
569. In a solution, solvent can be separated from solute by one of the following process

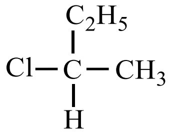
a) Decantation　　b) Filtration　　c) Distillation　　d) Sedimentation

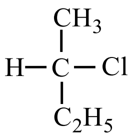
570. Buta-1,3-diene and But-2-yne are :

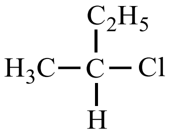
a) Position isomers　　b) Functional isomers　　c) Chain isomers　　d) Tautomers

571. $CH_3-CHCl-CH_2-CH_3$ has a chiral centre. Which one of the following represent its R-configuration?

a) 
$H-\underset{Cl}{\overset{C_2H_5}{C}}-CH_3$

b) 
$Cl-\underset{H}{\overset{C_2H_5}{C}}-CH_3$

c) 
$H-\underset{C_2H_5}{\overset{CH_3}{C}}-Cl$

d) 
$H_3C-\underset{H}{\overset{C_2H_5}{C}}-Cl$

572. Which of the following statements (s) is (are) not true?

a) Carbanions and carbonium ions, usually exist in ion pairs or else solvated
Acidity increases and basicity decreases in going from left to right across a row of Periodic Table

b) $CH_4 < NH_3 < H_2O < HF(acidity)$
$CH_3^- > NH_2^- > OH^- > F^-$(basicity)

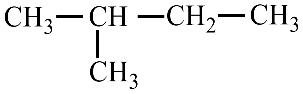
c) $RCOOH$ like $RCOR$ reacts with H_2NOH to give an oxime

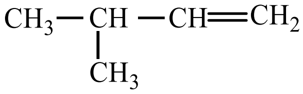
d) Decreasing order of ionizing power of solvents is
$CF_3COOH > HCOOH > H_2O > CH_3COOH > CH_3OH > C_2H_5OH > (CH_3)_2SO > CH_3CN$

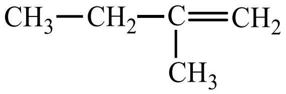
573. The intermediate during the addition of HCl to propene in the presence of peroxide is :
a) $CH_3\dot{C}HCH_2Cl$ b) $CH_3\overset{+}{C}HCH_3$ c) $CH_3CH_2\dot{C}H_2$ d) $CH_3CH_2\overset{+}{C}H_2$

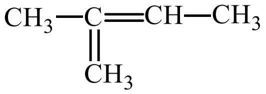
574. Which of the following represents *neo* −pentyl alcohol?
a) $CH_3CH(CH_3)CH_2CH_2OH$ b) $(CH_3)_3C-CH_2OH$ c) $CH_3(CH_2)_3OH$ d) $CH_3CH_2CH(OH)C_2H_5$

575. 2-methyl-2-butene will be represented as :

a) 
$CH_3-\underset{\underset{CH_3}{|}}{CH}-CH_2-CH_3$

b) 
$CH_3-\underset{\underset{CH_3}{|}}{CH}-CH=CH_2$

c) 
$CH_3-CH_2-\underset{\underset{CH_3}{|}}{C}=CH_2$

d) 
$CH_3-\underset{\underset{CH_3}{||}}{C}=CH-CH_3$

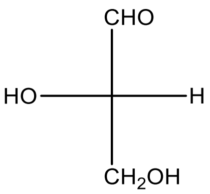
576. The most abundant organic compound in the world is :
a) CH_4 b) Chlorophyll c) Alkaloids d) Cellulose

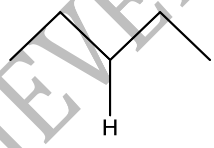
577. The chain initiating species in free radical chlorination of methane is :
a) Cl free radical b) HCl c) CH_3 radical d) Methylene radical

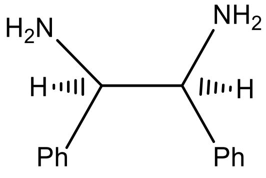
578. Which of the following belongs to $+I$ group?
a) $-OH$ b) $-OCH_3$ c) $-COOH$ d) $-CH_3$

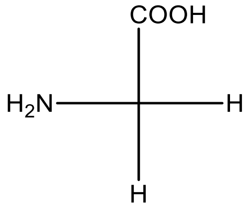
579. Different structures generated due to rotation about, C – C axis, of an organic molecule, are examples of
a) Geometrical isomerism
b) Conformational isomerism
c) Optical isomerism
d) Structural isomerism

580. Which of the following molecules is expected to rotate the plane of plane-polarised light?

a) 

b) 

c) 

d) 

581. Chromatography was discovered by
a) Kekule b) Pauling c) Rutherford d) Tswett

582. Sodium nitroprusside when added to an alkaline solution of sulphide ions produces a colouration
a) Red b) brown c) Blue d) Purple

583. $(CH_3)_4\overset{+}{N}$ is neither an electrophile, nor a nucleophile because it :
a) Does not have electron pair for donation as well as cannot attract electron pair
b) Neither has electron pair available for donation nor can accommodate electron since all shells of N are fully occupied
c) Can act as Lewis acid and base
d) None of the above

584. Isopentane can form four isomeric mono bromo derivatives. How many of them are optically active?
a) 1 b) 2 c) 3 d) None of these

585. Which one of the following does not show resonance?
a) Carbon dioxide b) Benzene c) Nitromethane d) Propane

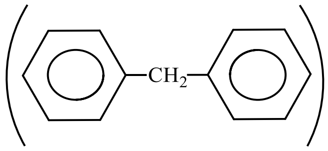
586. Select the organic compounds aliphatic in nature but burn with smoky flame :

a) CCl_4 b) $CHCl_3$ c) $C_6H_5CH_2OH$ d) Both (a) and (b)

587. Which of the following is an electrophilic reagent?

a) RO^- b) BF_3 c) NH_3 d) $R\underset{\cdot\cdot}{\ddot{O}}H$

588. The molecular formula of diphenyl methane is $C_{13}H_{12}$.



How many structural isomers are possible when one of the hydrogen is replaced by a chorine atom?

a) 6 b) 4 c) 8 d) 7

589. A mixture of iron fillings and sulphur cannot be separated by

a) Heating b) Magnet

c) Shaking with CS_2 d) Washing in a current of water

590. Isomers of propionic acid are

a) $HCOOC_2H_5$ and CH_3COOCH_3 b) $HCOOC_2H_5$ and C_3H_7COOH

c) CH_3COOCH_3 and C_3H_7OH d) C_3H_7OH and CH_3COCH_3

591. Reactions involving heterolytic fission are said to proceed *via* :

a) Ionic mechanism b) Polar mechanism c) Both (a) and (b) d) None of these

592. Which of the following orders is true regarding the acidic nature of COOH?

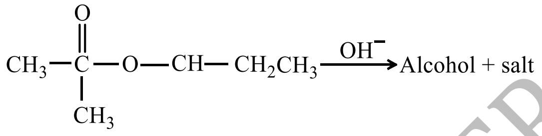
a) Formic acid $>$ acetic acid$>$ propanoic acid b) Formic acid $>$ acetic acid$<$ propanoic acid

c) Formic acid $<$ acetic acid$>$ propanoic acid d) Formic acid $>$ acetic acid$<$ propanoic acid

593. Which behaves both as a nucleophile as well as an electrophile?

a) CH_3OH b) CH_3NH_2 c) CH_3CN d) CH_3Cl

594. Alkaline hydrolysis of an ester (A) gives alcohol and salt



$$CH_3-\overset{\overset{\displaystyle O}{\|}}{C}-O-\underset{\underset{\displaystyle CH_3}{|}}{C}H-CH_2CH_3\xrightarrow{OH^-}\text{Alcohol + salt}$$

The correct statement about the reaction is :

a) In alcohol configuration about chiral carbon atom is retained

b) In alcohol configuration about chiral carbon atom is inverted

c) Alcohol loses optical activity

d) All statement are incorrect

595. In which case the carbon-carbon bond length is same?

a) 2-butene b) Benzene c) 1-butene d) 1-propyne

596. Incorrect statement is

a) Aniline can be purified by steam distillation

b) Beilstein test is not given by fluorine

c) Kjeldahl's method is used for estimation of sulphur

d) Lassaigen's test is used in the qualitative detection of elements in organic compounds

597. The increasing order of stability of the following free radicals are

a) $(CH_3)_2\dot{C}H < (CH_3)_3\dot{C} < (C_6H_5)_2\dot{C}H < (C_6H_5)_3\dot{C}$

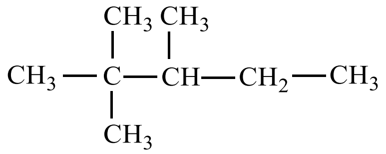
b) $(C_6H_5)_3\dot{C} < (C_6H_5)_2\dot{C}H < (CH_3)_3\dot{C} < (CH_3)_2\dot{C}H$

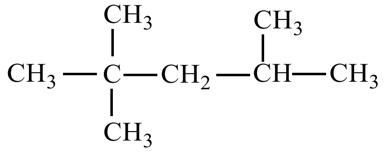
c) $(C_6H_5)_2\dot{C}H < (C_6H_5)_3\dot{C} < (CH_3)_3\dot{C} < (CH_3)_2\dot{C}H$

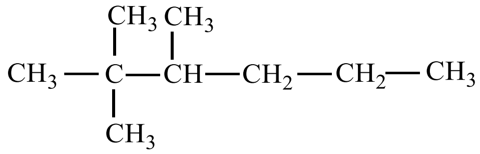
d) $(CH_3)_2\overset{\bullet}{C}H < (CH_3)_3\overset{\bullet}{C} < (C_6H_5)_3\overset{\bullet}{C} < (C_6H_5)_2\overset{\bullet}{C}H$

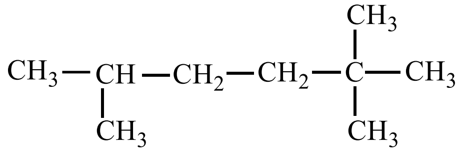
598. Which one of the following explain, why propene undergo electrophilic addition with HBr, but not with HCN?
   a) $Br^-$is better nucleophile than CN^-
   b) HBr being better source of proton as it is stronger acid than HCN
   c) HCN attacks preferentially *via* lone pair of nitrogen
   d) The $C - Br$ bond being stronger is formed easily as compared to $C - CN$ bond

599. The structural formula of 2,2,3-trimethyl hexane is :

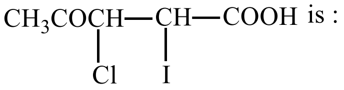
   a) 
   $CH_3-\underset{CH_3}{\overset{CH_3}{C}}-\overset{CH_3}{CH}-CH_2-CH_3$

   b) 
   $CH_3-\underset{CH_3}{\overset{CH_3}{C}}-CH_2-\overset{CH_3}{CH}-CH_3$

   c) 
   $CH_3-\underset{CH_3}{\overset{CH_3}{C}}-\overset{CH_3}{CH}-CH_2-CH_2-CH_3$

   d) 
   $CH_3-\underset{CH_3}{CH}-CH_2-CH_2-\underset{CH_3}{\overset{CH_3}{C}}-CH_3$

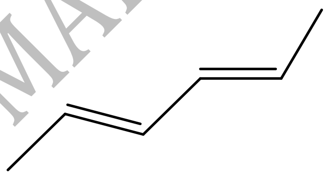
600. The IUPAC name of the compound,


$CH_3CO\underset{Cl}{CH}-\underset{I}{CH}-COOH$ is :

   a) 2-iodo-3-chloro-4-pentanoic acid
   b) 4-oxo-3-chloro-2-iodo pentanoic acid
   c) 4-carboxy-4,3-chloro-2-butanone
   d) 3-chloro-2-iodo-4-oxo-pentanoic acid

601. Select the correct statement about the detection of sulphur in organic compounds
   a) Sulphur present in organic compound on fusion with sodium is converts to $Na_2S_2O_3$
   b) $FeCl_3$ gives purple colour when added to the sodium fusion extract
   c) Sodium nitroprusside is used to detect the presence of sulphur
   d) All of the above

602. The name of the compound



is

a) (2Z,4Z)-2, 4-hexadiene	b) (2Z-, 4E)-2, 4-hexadiene
c) (2E, 4Z)-2, 4-hexadiene	d) (4E, 4Z)-2, 4-hexadiene

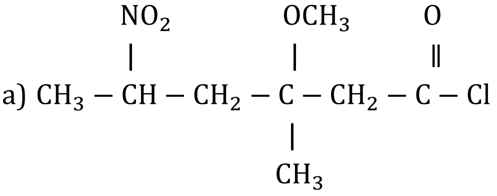
603. Rotation of plane polarized light can be measured by :

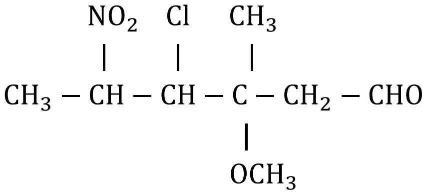
a) Manometer	b) Calorimeter	c) Polarimeter	d) Viscometer

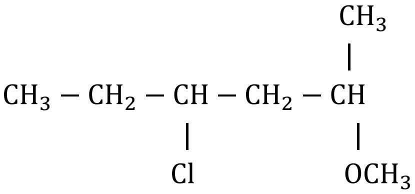
604. A molecule having three different chiral carbon atoms, how many stereoisomers it will have?

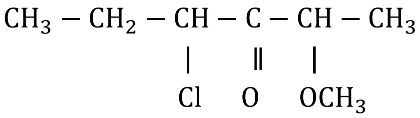
a) 8	b) 3	c) 9	d) 6

605. Correct the structural formula of compound-5-nitro-3-methoxy-3-methylhexanoyl chloride is

a) 
$CH_3-\overset{\overset{NO_2}{|}}{CH}-CH_2-\underset{\underset{CH_3}{|}}{\overset{\overset{OCH_3}{|}}{C}}-CH_2-\overset{\overset{O}{\|}}{C}-Cl$

b) 
$CH_3-\overset{\overset{NO_2}{|}}{CH}-\overset{\overset{Cl}{|}}{CH}-\underset{\underset{OCH_3}{|}}{\overset{\overset{CH_3}{|}}{C}}-CH_2-CHO$

c) 
$CH_3-CH_2-\underset{\underset{Cl}{|}}{CH}-CH_2-\underset{\underset{OCH_3}{|}}{\overset{\overset{CH_3}{|}}{CH}}$

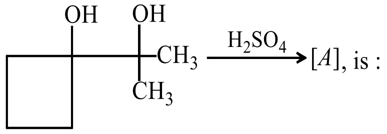
d) 
$CH_3-CH_2-\underset{\underset{Cl}{|}}{CH}-\underset{\underset{O}{\|}}{C}-\underset{\underset{OCH_3}{|}}{CH}-CH_3$

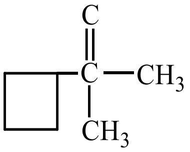
606. Which of the following is the strongest base?
a) Acetamide b) Aniline c) Methylamine d) Dimethylamine

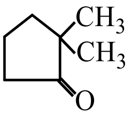
607. The isomeric monosubstitution products theoretically possible for the structure,
$CH_2=HC-CH_2-CH_2-CH=CH_2$ are :
a) 3 b) 2 c) 4 d) 6

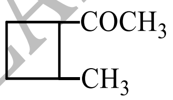
608. Intermediate product formed in the acid catalysed dehydration of n-propyl alcohol is :
a) $CH_3-CH_2-CH_3$ b) $CH_3-CH=CH_2$ c) $CH_3-\overset{+}{C}H-CH_3$ d) $CH_3-CH_2-\overset{+}{C}H_2$

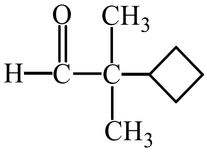
609. Product $[A]$ in the reaction :



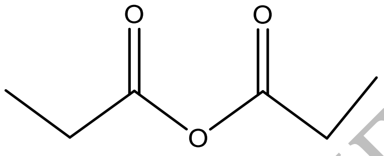
a) 

b) 

c) 

d) 

610. The IUPAC name of the following compound is

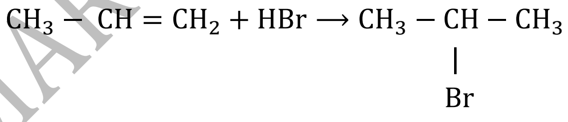


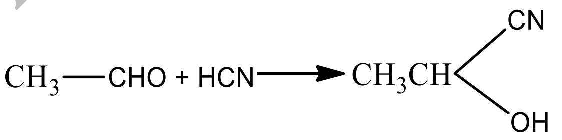
a) Propionic anhydride
b) Dipropanoic anhydride
c) Ethoxy propanoic acid
d) Propanoic anhydride

611. Which of the following has the highest nucleophilicity?
a) F^- b) OH^- c) CH_3^- d) NH_2^-

612. Anthracene is purified by
a) Filtration b) Distillation c) Crystallisation d) Sublimation

613. Which of the following requires radical intermediate?

a) 
$CH_3-CH=CH_2+HBr\longrightarrow CH_3-\underset{\underset{Br}{|}}{CH}-CH_3$

b) 

c) $CH_3-CH=CH_2+HBr\rightarrow CH_3-CH_2-CH_2-Br$

d) $CH_3CHO+NH_2OH\xrightarrow{H^+}CH_3-CH=N-OH$

614. Chiral molecules are those which:
a) Are not superimposable on their mirror images
b) Are superimposable on their mirror images

c) Show geometrical isomerism

d) Are unstable molecules

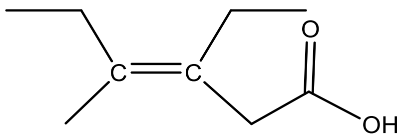
615. Who is called the 'Father of Chemistry'?

a) Faraday b) Priestley c) Rutherford d) Lavoisier

616. With a change in hybridization of the carbon bearing the charge, the stability of a carbanion increase in the order

a) $sp < sp^3 < sp^2$ b) $sp < sp^2 < sp^3$ c) $sp^2 < sp < sp^3$ d) $sp^3 < sp^2 < sp$

617. The correct IUPAC name of the acid



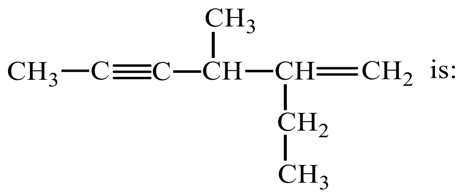
a) Z-3-ethyl-4-methyl hex-3-en-1-oic acid

b) *Z*-3-ethyl-4-methyl hexanoic acid

c) *Z*-3, 4-diethylpent-3-en-1-oic acid

d) *E*-3-ethyl-4-methylhex-4-en-1-oic acid

618. IUPAC name of,



$$CH_3-C\equiv C-\underset{\underset{\underset{CH_3}{|}}{\underset{CH_2}{|}}}{}\overset{\overset{CH_3}{|}}{C}H-\underset{\underset{\underset{CH_3}{|}}{CH_2}}{\underset{|}{C}}H=CH_2$$ is:

a) 2-ethyl-3-methyl-hexa-l-en-4-yne

b) 5-ethyl-4-methyl-hexa-2-yn-5-ene

c) 3-methylene-4-methylhepta-5-yne

d) 5-methylene-5-ethyl-4-methylhepta-2-yne

619. The total number of alkenes possible by dehydromination of 3-bromo-3-cyclopentylhexane using alcoholic KOH is

a) 1 b) 3 c) 5 d) 7

620. The Cl—C—Cl bond angle in dichloro methane will be :

a) > 109°28′ b) < 109°28′ c) 109°28′ d) 120°

621. In the following reaction sequence, the chain initiation steps is :

a) $Cl-Cl \longrightarrow \dot{C}l + \dot{C}l$

b) $CH_4 + \dot{C}l \longrightarrow \dot{C}H_3 + HCl$

c) $\dot{C}H_3 + Cl_2 \longrightarrow CH_3Cl + \dot{C}l$

d) $\dot{C}H_3 + \dot{C}l \longrightarrow CH_3Cl$

622. Amongst the following which of the above are true for S_N2 reaction?

(i) The rate of reaction is independent of the concentration of the nucleophile.

(ii) The nucleophile attacks the carbon atom on the side of the molecule opposite to the group being displaced.

(iii) the reaction proceeds with simultaneous bond formation and bond rupture.

a) (i), (ii) b) (i), (iii) c) (i), (ii), (iii) d) (ii), (iii)

623. Predict the nature of principal product in the reaction,

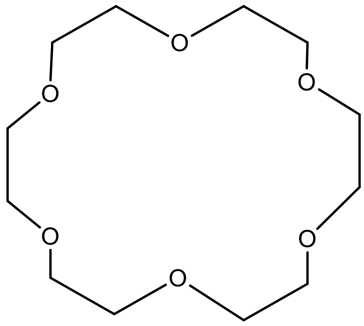
$BrCH_2CH_2CH_2CH_2Br + KOH\ (alc.) \longrightarrow Product:$

a) 1,3-butadiene b) Cyclobutane c) $BrCH_2CH_2CH = CH_2$ d) None of these

624. The most stable carbanion is

a) $CH_3^{\ominus}$ b) $RCH_2^{\ominus}$ c) $R_3C^{\ominus}$ d) $^{\ominus}CH_2CHO$

625. Crown ethers are named as *X*-crown-*Y*. In the following crown ether, *X* and *Y* are respectively



a) 6 and 12 b) 18 and 6 c) 24 and 6 d) 6 and 24

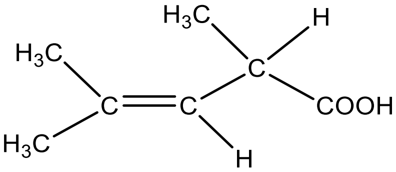
626. Which one of the nitrogen atoms in

$$H_2N - NH - \overset{\overset{\text{O}}{\|}}{C} - NH_2$$ is the most nucleophilic?

I II III

a) III b) I c) II d) All three N atoms

627.



Compound can exhibit

a) Geometrical isomerism b) Tautomerism

c) Optical isomerism d) Geometrical and optical isomerism

628. The order of stability of the alkenes

$R_2\text{C} = \text{C}R_2, R_2\text{C} = \text{CH}R, R_2\text{C} = \text{CH}_2,$

I II III

$R\text{CH} = \text{CH}R,$ $R\text{CH} = \text{CH}_2$

IV V

a) I>II>IV>III>V b) I=II>III>IV>V c) II>I>IV>III>V d) V>IV>III>I>II

629. If a compound on analysis was found to contain $\text{C} = 18.5\%, \text{H} = 1.55\%, \text{Cl} = 55.04\%$ and $\text{O} = 24.81\%$ then its empirical formula is

a) CH_2OCl b) CH_2ClO_2 c) ClCH_2O d) CHClO

630. 2-pentanone and 3-methyl-2-butanone are a pair of ….. isomers.

a) Functional b) Chain c) Positional d) Stereo

631. The number of isomeric ethers with molecular formula $\text{C}_4\text{H}_{10}\text{O}$ is/ are.

a) One b) Two c) Three d) Four

632. Liebig's test is used to estimate

a) H b) C c) C and H Both d) N

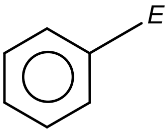
633. Number of isomers possible for C_4H_8O is

a) 3 b) 4 c) 5 d) 6

634. Most stable carbonium ion is

a) $\overset{+}{\text{C}_2}\text{H}_5$ b) $(\text{CH}_3)_3\overset{+}{\text{C}}$ c) $(\text{C}_6\text{H}_5)_3\overset{+}{\text{C}}$ d) $\text{C}_6\text{H}_5\overset{+}{\text{C}}\text{H}_2$

635.

In a compound  electrophilic substitution has occurred. The substitute-E are methyl $-\text{CH}_2\text{Cl}, -\text{CCl}_3$ and$-\text{CHCl}_2$. The correct increasing order towards electrophilic substitution is

a) $-\text{CH}_3 < -C\text{H}_2\text{Cl} < -CHC\text{l}_2 < -CC\text{l}_3$ b) $-\text{CH}_3 < -CHC\text{l}_2 < -\text{CH}_2\text{Cl} < -CC\text{l}_3$

c) $-\text{CCl}_3 < -C\text{H}_2\text{Cl} < -CHC\text{l}_2 < -C\text{H}_3$ d) $-\text{CCl}_3 < -CHC\text{l}_2 < -\text{CH}_2\text{Cl} < -C\text{H}_3$

636. In fructose, the possible optical isomers are

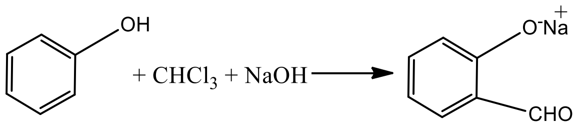
a) 12 b) 8 c) 16 d) 4

637. Which structure can be explained by taking ground state configuration of atom?
a) BeH_2 b) BF_3 c) CH_4 d) H_2O

638. Which one of the following carbanions is the least stable?
a) $CH_3CH_2^-$ b) $HC \equiv C^-$ c) $(C_6H_5)_3C^-$ d) $(CH_3)_3C^-$

639. Which one of the following is the most energetic conformation of cyclohexane?
a) Boat b) Twisted boat c) Chair d) Half chair

640. The energy difference between the chair and boat form of cyclohexane is :
a) 44 kJ mol^{-1} b) 24 kJ mol^{-1} c) 34 kJ mol^{-1} d) 68 kJ mol^{-1}

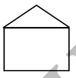
641.


The electrophile involved in the above reaction is

a) Dichloromethyl cation $(\overset{\oplus}{C}HCl_2)$
b) Dichlorocarbene $(:CCl_2)$
c) Trichloromethyl anion $(\bar{C}Cl_3)$
d) Formyl cation $(\overset{\oplus}{C}HO)$

642. Addition of HI on double bond of propene yields isopropyl iodide as major product. It is because the addition proceeds through :
a) More stable carbocation
b) more stable carbanion
c) More stable free radical
d) Homolysis

643. For which of the following parameters the structural isomers C_2H_5OH and CH_3OCH_3 would be expected to have the same values? (Assume ideal behaviour)
a) Heat of vaporisation
b) Vapour pressure at the same temperature
c) Boiling points
d) Gaseous densities at the same temperature and pressure

644.
The IUPAC name of  is :
a) Bicyclo (2,1,0) pentane
b) 1,2-cyclopropyl cyclobutane
c) Cyclopentane (4,3)annulene
d) 1,2-methylene cyclobutane

645. The IUPAC name of neopentane is :
a) 2,2-dimethylpropane b) 2-methylpropane c) 2,2-dimethylbutane d) 2-methylbutane

646. In Carius method of 0.099 g orgainc compound gave 0.287 g AgCl. The percentage of chlorine in the compound will be
a) 28.6 b) 71.7 c) 35.4 d) 64.2

647. The most common type of reaction in aromatic compounds is
a) Elimination reaction
b) Addition reaction
c) Electrophilic substitution reaction
d) Rearrangement reaction

648. Select the correct order of basic nature :
a) $CH_3CH_2^- > CH_2 = CH^- > CH \equiv C^- \longrightarrow OH^-$
b) $CH_3CH_2^- > CH \equiv C^- > CH_2 = CH^- > OH^-$
c) $CH_3CH_2^- > OH^- > CH \equiv C^- > CH_2 = CH^-$
d) $OH^- > CH \equiv C^- > CH_2 = CH^- > CH_3—CH_2^-$

649. Arrange the following carbocations in order of stability

benzyl allyl methyl vinyl
I II III IV

a) IV>III>II>I b) I>II>III>IV c) II>IV>III>I d) III>II>I>IV

650. The prefix name of —SH group in IUPAC system is :
a) Mercapto b) Thiol c) Sulphide d) None of these

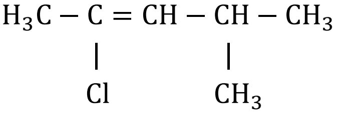
651. The correct name for CH_3COCH_2OH is :
a) 2-keto propanol
b) 1-hydroxy propan-2-one
c) Propan-2-one-1-ol
d) 3-hydroxy propan-2-one

652. The maximum number of alkyl groups in C_8H_{18} is :
a) 6 b) 5 c) 4 d) 2

653. The chlorination of methane to give CCl_4 is an example of
a) Addition b) Elimination c) Substitution d) Chain reaction

654. The number of isomers for the aromatic compound of the formula C_7H_8O is :
a) 2
b) 3
c) 4
d) 5

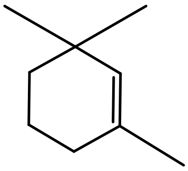
655. The IUPAC name of the compound having the formula $CH \equiv C — CH = CH_2$ is :
a) 1-butene-3-yne b) 3-butene-1-yne c) 1-butyn-3-ene d) But-1-yne-3-ene

656. 

$$H_3C-\underset{\displaystyle Cl}{\underset{|}{C}}=CH-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_3$$

a) 2-chloro-4-methyl-2-pentene b) 4-chloro-2-methyl-3-pentene
c) 4-methyl-2-chloro-2-pentene d) 2-chloro-4, 4-dimethyl-2-butene

657. Give the IUPAC name of the compound



a) 1,1,3-trimethylcyclohex-2-ene b) 1,3,3-trimethylcyclohex-1-ene
c) 1,1,5-trimethylcyclohex-5-ene d) 2,6,6-trimethylcyclohex-1-ene

658. C – H bond energy is about101 kcal/mol for methane, ethane and other alkanes but is only77 kcal/mol for C – H bond of CH_3 in toluene. This is because
a) Of inductive effect due to $-CH_3$ in toluene
b) Of the presence of benzene ring in toluene
c) Of resonance among the structures of benzyl radical in toluene
d) Aromaticity of toluene

659. The reaction $(CH_3)_3C-Br \xrightarrow{H_2O} (CH_3)_3COH$ is
a) Elimination b) Substitution c) Free radical d) Addition

660. The number of structural and configurational isomers of a bromo compound, C_5H_9Br, formed by the addition of HBr to 2-pentyne respectively are :
a) 1 and 2 b) 2 and 4 c) 4 and 2 d) 2 and 1

661. Fractional distillation is a process by which the separation of different fractions from mixture of solution is carried by making use of the following property of the fractions
a) Freezing point b) Boiling point c) Melting point d) Solubility

662. The maximum number of carbon atoms arranged linearly in the molecule, $CH_3-C\equiv C-CH=CH_2$ is
a) 5 b) 4 c) 2 d) 3

663. α-D-glucose and β-D-glucose have a specific rotation of +112° and +19° respectively. In aqueous solution

the rotation becomes +52°. This process is called :

a) Inversion b) Racemization c) Mutarotation d) enolisation

664. Which of the following is an electroplile?

a) Na^+ b) Li^+ c) H^+ d) Ca^{2+}

665. Which of the following is singlet carbine?

a) $CH_3\ddot{C}HCH_3$ b) $C_2H_5\ddot{C}{-}H$ c) $CH_2{=}CH{-}\overset{\oplus}{C}H_2$ d) $(CH_3)_3C^+$

666. Which of the following is the correct order of priority of groups in D-glyceraldehyde?

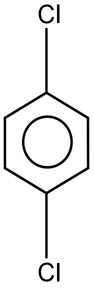
a) OH(1), CHO(2), CH_2OH(3)and H(4) b) OH (1), CH_2OH(2), CHO(3)and H(4)

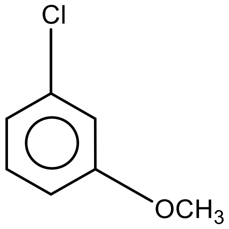
c) CH_2OH(1), CHO(2), OH(3)and H(4) d) CHO(1), OH(2), CH_2OH(3) and H(4)

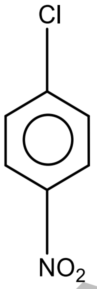
667. *o*-hydroxytoluene and benzyl alcohol are :

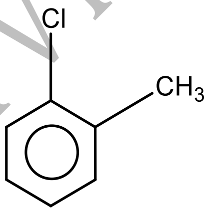
a) Position isomers b) Keto-enol tautomers c) Chain isomers d) None of these

668. Which of the following would react most readily with nucleophiles?

a) 

b) 

c) 

d) 

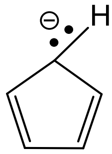
669. Which group has the highest + Inductive effect?

a) CH_3- b) CH_3CH_2- c) $(CH_3)_2CH-$ d) $(CH_3)_3\,C-$

670. The Prussian blue colour obtained during the test of nitrogen by Lassaigne;s test is due to the formation of

a) $Fe[Fe(CN)_6]_3$ b) $Na_4[Fe(CN)_6]$ c) $Fe_3[Fe(CN)_6]_4$ d) $Fe_2[Fe(CN)_6]$

671. How many π-electrons are there in the following structure?



a) 2 b) 4 c) 6 d) 8

672. Which of the following statements is correct?

a) Allyl carbonium ion $(CH_2{=}CH{-}\overset{+}{C}H_2)$ is more stable then propyl carbonium ion

b) Propyl carbonium ion is more stable than the allyl carbonium ion

c) Both are equally stable

d) None of the above

673. Which of the following possesses an *sp*-hybridized carbon in its structure?

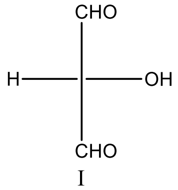
a) $CCl_2=CCl_2$

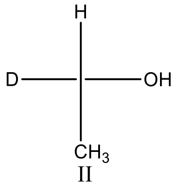
b) $CH_2=C=CH_2$

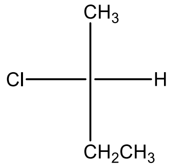
c) $CH_2=CH{-}CH=CH_2$

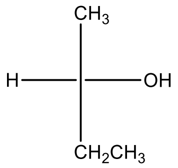
d) $CH_2=CCl{-}CH=CH_2$

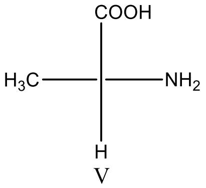
674. Select *R*-isomers from the following


I


II

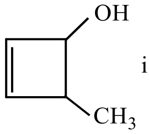

III


IV


V

a) I and III b) II, IV and V c) I, II and III d) II and III

675. Di-chloroacetic acid is a stronger acid than acetic acid.
This due to occurrence of
a) Mesomeric effect b) Hyperconjugation c) Inductive effect d) Steric effect

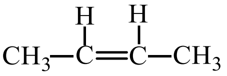
676. The IUPAC name of ,  is :
a) 3-methyl cyclo-1-buten-2-ol
b) 4-methyl cyclo-2-buten-1-ol
c) 4-methyl cyclo-1-buten-3-ol
d) 2-methyl cyclo-3-buten-1-ol

677. The sodium extract of an organic compound on boiling with HNO_3 and addition of ammonium molybdate of solution gives a yellow precipitate. The compound contains
a) Nitrogen b) Phosphorus c) Sulphur d) Chlorine

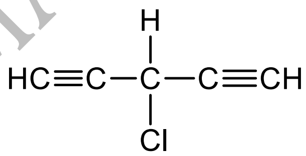
678. Unpaired electron in $\dot{C}H_3$ occupies :
a) sp-hybrid orbital b) sp^3-hybrid orbital c) p-orbital d) sp^2-hybrid orbital

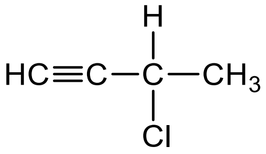
679. In which of the following ways does the hydride ion tend to function?
a) An electrophile b) A nucleophile c) A free radical d) An acid

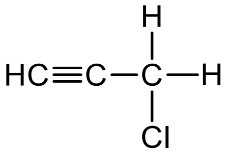
680. The number of isomers of the compound with molecular formula $C_2H_2Br_2$ is
a) 4 b) 3 c) 5 d) 2

681. The treatment of CH_3MgX with $CH_3C \equiv C—H$ produces :
a) CH_4 b) $CH_3—CH = CH_2$ c) $CH_3C \equiv C - CH_3$ d) 
$CH_3-\overset{H}{\overset{|}{C}}=\overset{H}{\overset{|}{C}}-CH_3$

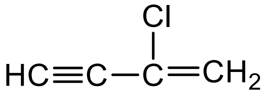
682. The correct decreasing order of priority for the functional groups of organic compounds in the IUPAC system of nomenclature is
a) $-COOH, -SO_3H, -CONH_2, -CHO$ b) $-SO_3H, -COOH, -CONH_2, -CHO$
c) $-CHO, -COOH, SO_3H, -CONH_2$ d) $-CONH_2, -CHO, -SO_3H, -COOH$

683. Which of the following is most likely to show optical isomerism?
a) 
$HC\equiv C-\underset{Cl}{\underset{|}{\overset{H}{\overset{|}{C}}}}-C\equiv CH$

b) 
$HC\equiv C-\underset{Cl}{\underset{|}{\overset{H}{\overset{|}{C}}}}-CH_3$

c) 

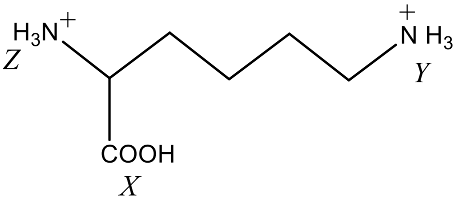
$HC\equiv C-CH_2Cl$

d) 

$HC\equiv C-C(Cl)=CH_2$

684. Which of the following statements is incorrect?
   a) The rate of reaction increases with increase in water concentration in the hydrolysis of tertiary butyl bromide in methanol and water
   b) The relative nucleophilicity in protic solvent is $CN^- > I^- > \bar{O}H > Br^- > Cl^- > F^- > H_2O$
   c) In S_N2 reactions, the order of reactivity of alkyl halides is in the order methyl>primary>secondary> tertiary
   d) S_N2 reaction involves carbonium ions

685. Arrange in order of increasing acidic strength.



a) X>Z>Y b) Z<X>Y c) X>Y>Z d) Z>X>Y

686. For the purification, isolation and separation of organic compounds, the latest technique followed is
   a) Chromatography
   b) Steam distillation
   c) Fractional crystallisation
   d) Sublimation

687. Arrange *p*-toluidine (I) N,N-dimethyl-*p*- toluidine (II) *p*- nitroaniline (III) and aniline (IV) in order of decreasing basicity
   a) I>IV>III>II b) I>II>III>IV c) II>I>IV>III d) III>I>II>IV

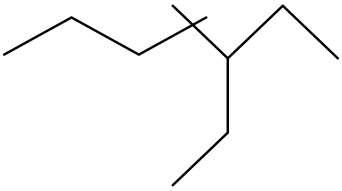
688. The reagent showing addition on alkene against the Markownikoff's rule of :
   a) Br_2 b) H_2S c) HF d) HBr

689. Carbocation can undergo :
   a) Loss of a proton
   b) Addition to multiple bond
   c) Combination with anions
   d) All of the above

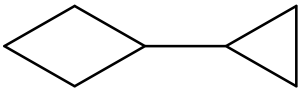
690. Lactic acid is :
   a) Propionic acid
   b) β-hydroxypropanoic acid
   c) α-hydroxypropanoic acid
   d) None of the above

691. Of the following compounds which will have a zero dipole moment?
   a) 1,1-dichloroethylene
   b) *trans*-1,2-dichloroethylene
   c) *cis*-1,2-dichloroethylene
   d) None of the above

692. The IUPAC name of



a) 1,1-diethyl1-2, 2-dimethylpentane
b) 4,4-dimethyl-5,5-diethylpentane
c) 5, 5-diethyl-4,4-dimethylpentane
d) 3-ethyl-4,4-dimethylheptane

693. The correct IUPAC name  is

a) 1- cyclopropyl cyclobutane
b) 1, 1-dicyclobutane
c) 1- cyclobutane-1- cyclopropane
d) None of the above

694. Reactivity towards nucleophilic addition reaction of
I. HCHO II.CH_3CHO
III. CH_3COCH_3 is

a) II>III>I
b) III>II>I
c) I>II>III
d) I>III>II

695. Arrange the following compounds in order of their decreasing reactivity with an electrophile, $E^{\oplus}$.
(A) Chlorobenzene,
(B) 2,4-dinitrochlorobenzene,
(C) p-nitrochlorobenzene

a) C>B>A
b) B>C>A
c) A>C>B
d) A>B>C

696. Isomerism exhibited by acetic acid and methyl formate is :

a) Functional
b) Chain
c) Geometrical
d) Central

697. $C_3H_5Cl + \text{aq. NaOH} \rightarrow C_2H_5OH + NaCl$;
this reaction is

a) Electrophilic substitution of I order
b) Electrophilic substitution of II order
c) Nucleophilic substitution of I order
d) Nucleophilic substitution of II order

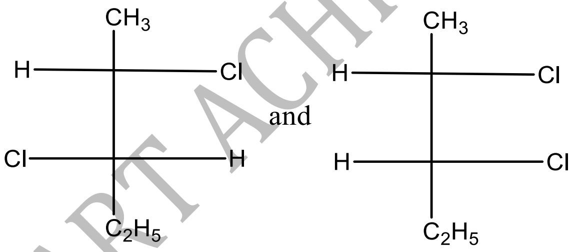
698. In IUPAC suffix name of —COX is :

a) Oyl halide
b) Halo carbonyl
c) Carbamoyl
d) None of these

699. IUPAC name of the compound ,  is :

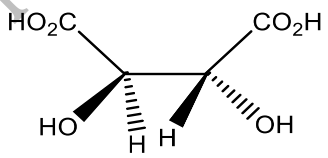
a) but-2-en-l-ol
b) l-hydroxy but-l-ene
c) 4-hydroxy butene-3
d) But-l-en-l-ol

700. The two optical isomers given below, namely



a) Enantiomers
b) Geometrical isomers
c) Diastereomers
d) Structural isomers

701. The absolute configuration of

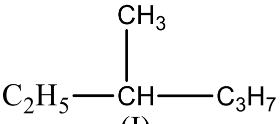


a) S, S
b) R, R
c) R, S
d) S, R

702. Which one of the following compounds is most reactive towards nucleopilic addition?

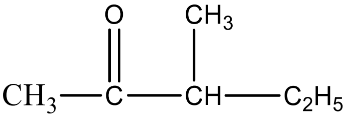
a) CH_3CHO
b) $PhCOCH_3$
c) PhCOPh
d) CH_3COCH_3

703. Among the following four structures I to IV



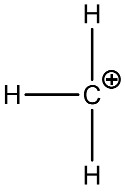
$C_2H_5-\overset{CH_3}{CH}-C_3H_7$

(I)



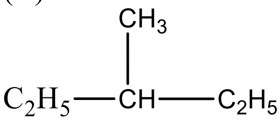
$CH_3-\overset{O}{\overset{\|}{C}}-\overset{CH_3}{CH}-C_2H_5$

(II)



$H-\overset{H}{\underset{H}{C^{\oplus}}}$

(III)



$C_2H_5-\overset{CH_3}{CH}-C_2H_5$

(IV)

it is true that

a) All four are chiral compounds
b) Only I and II are chiral compounds
c) Only III is a chiral compound
d) Only II and IV are chiral compounds

704. Which of the following is the most stable cation?

a) $F_3C-\overset{\oplus}{CH_2}$
b) $(CH_3)_2\,\overset{\oplus}{CH}$
c) $\overset{\oplus}{CH_3}$
d) $\overset{\oplus}{CF_3}$

705. Write the IUPAC name of

$CH_2-O-\underset{CH_3}{CH}-CH_2-CH_3$

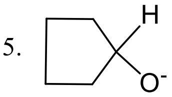
a) 3-methoxy butane
b) 2-methoxy butane
c) 3-methyl-3-methoxy propane
d) Butoxy methane

706. Which of the following species is not electrophilic in nature?

a) $\overset{\oplus}{C}l$
b) BH_3
c) $H_3\overset{\oplus}{O}$
d) $\overset{\oplus}{N}O_2$

707. List the following alkoxide nucleophile in decreasing order of their S_N2 reactivity

1. Me_3CO^- 2. MaO^- 3. $MeCH_2O^-$ 4. Me_2CHO^- 5.



a) 2>3>5>4>1
b) 5>3>2>1>4
c) 1>5>2>3>4
d) 3>5>1>2>3

708. The Beilstein test for organic compound is used to detect

a) Nitrogen
b) Sulphur
c) Carbon
d) Halogens

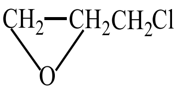
709. Which of the following statements is not characteristic of free radical chain reaction?

a) It gives major product derived from most stable free radical
b) It is usually sensitive to change in solvent polarity
c) It proceeds in three main steps like initiation, propagation and termination
d) It may be initiated by UV light

710. The presence of carbon in an organic compound is detected by heating it with

a) Sodium metal to convert it into NaCN
b) CaO to convert it into CO which burns with a blue flame
c) CuO to convert it into CO_2 which turns lime water milky
d) Cu wire to give a bluish green flame

711. IUPAC name of the compound,

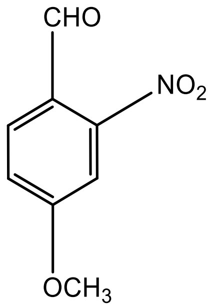


is :

a) 1-choloro-2,3-epoxypropane
b) 3-chloro-1,2-epoxypropane

c) 1-chloroethoxymethane
d) None of the above

712. What is the correct IUPAC name of



a) 4-methoxy-2-nitrobenzaldehyde
b) 4-formyl-3-nitro anisole
c) 4-methoxy-6-nitrobenzaldehyde
d) 2-formyl-5-methoxy nitrobenzene

713. Which one is an elimination reaction?
a) $CH_3CH_3 + Cl_2 \longrightarrow CH_3CH_2Cl + HCl$
b) $CH_3Cl + KOH(aq.) \longrightarrow CH_3OH + KCl$
c) $CH_2 = CH_2 + Br \longrightarrow CH_2BrCH_2Br$
d) $C_2H_5Br + KOH(alc.) \longrightarrow C_2H_4 + KBr + H_2O$

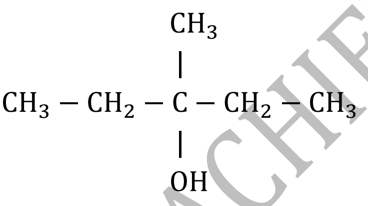
714. Identify the compound that exhibits tautomerism
a) 2-butene
b) Lactic acid
c) 2-pentanone
d) Phenol

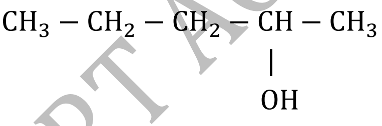
715. Which of the following is an electrophile?
a) H_2O
b) SO_3
c) NH_3
d) ROR

716. The formula of ethanenitrile is :
a) C_2H_5NC
b) C_2H_5CN
c) CH_3CN
d) None of these

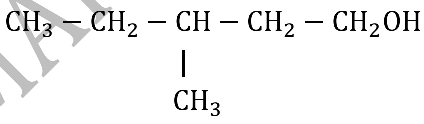
717. Which of the following acids shows stereoisomerism?
a) Oxalic acid
b) Tartaric acid
c) Acetic acid
d) Formic acid

718. Among the following compounds which can be dehydrated very easily is

a)

$CH_3-CH_2-\underset{\underset{OH}{|}}{\overset{\overset{CH_3}{|}}{C}}-CH_2-CH_3$

b)

$CH_3-CH_2-CH_2-\underset{\underset{OH}{|}}{CH}-CH_3$

c) $CH_3-CH_2-CH_2-CH_2-CH_2OH$

d)

$CH_3-CH_2-\underset{\underset{CH_3}{|}}{CH}-CH_2-CH_2OH$

719. Mark the incorrect statement in nitrogen Kjeldahl's method of estimation
a) Nitrogen gas is collected over caustic potash solution
b) Potassium sulphate is used as boiling point elevator of H_2SO_4
c) Copper sulphate or mercury acts as a catalyst
d) Nitrogen is quantitatively decomposed to give ammonium sulphate

720. Which of the following orders is correct regarding the $-I$ effect of the substituents?
a) $-NR_2 > -OR > -F$
b) $-NR_2 < -OR < -F$
c) $-NR_2 > -OR < -F$
d) $-NH_2 < -OR > -F$

721. The ease of dehydrohalogenation of alkyl halide with alcoholic KOH is
a) 3°<2°<1°
b) 3°>2°>1°
c) 3°<2°>1°
d) 3°>2°<1°

722. Lactic acid in which a methyl group, a hydroxyl group, a carboxylic acid group and a hydrogen atom are attached to a central carbon atom shows optical isomerism due to the molecular geometry at the :

a) Carbon atom of the methyl group
b) Carbon atom of the carboxylic acid group
c) Central carbon atom
d) Oxygen of the hydroxyl group

723. Which of the following process is suitable for the purification of aniline?

a) Simple distillation
b) Fractional distillation
c) Fractional crystallisation
d) Steam distillation

724. Maleic and fumaric acids are :

a) Tautomers
b) Geometrical isomers
c) Chain isomers
d) Functional isomers

725. $CH_3Br + Nu^- \longrightarrow CH_3 - Nu + Br^-$ The decreasing order of the rate of the above reaction with nucleophiles (Nu^-) A to D is :

$[Nu^- = (A)PhO^-, (B)AcO^-, (C)HO^-, (D)CH_3O^-]$

a) $D > C > A > B$
b) $D > C > B > A$
c) $A > B > C > D$
d) $B > D > C > A$

726. Which one is least reactive in a nucleophile substitution reaction?

a) CH_3CH_2Cl
b) $CH_2 = CHCH_2Cl$
c) $CH_2 = CHCl$
d) $(CH_3)_3CCl$

727. In methanol solution, bromine reacts with ethylene to yield $BrCH_2CH_2OCH_3$ in addition to 1,2-dibromoethane because :

a) The intermediate carbocation may react with Br^- or CH_3OH
b) The methyl alcohol solvolates the bromine
c) The reaction follows Markownikoff's rule
d) This is a free radical mechanism

728. Number of tertiary carbon atoms in tertiary butyl alcohol is :

a) 1
b) 2
c) Zero
d) 4

729. Which step is chain propagation step in the following mechanism?

(i) $Cl_2 \xrightarrow{hv} Cl^{\bullet} + Cl^{\bullet}$

(ii) $Cl^{\bullet} + CH_4 \longrightarrow \dot{C}H_3 + HCl$

(iii) $Cl^{\bullet} + Cl^{\bullet} \longrightarrow Cl_2$

(iv) $\dot{C}H_3 + Cl^{\bullet} \longrightarrow CH_3Cl$

a) (i)
b) (ii)
c) (iii)
d) (iv)

730. The IUPAC name of the compound $CH_3-N \underset{\rightarrow}{=} C$ is :

a) Ethane nitrile
b) Methane isonitrile
c) Ethane isonitrile
d) None of these

731. IUPAC name of

$CH_3CH_2C(Br) = CH - Cl$ is

a) 2-bromo-1-chloro butene-1
b) 1-chloro-2-bromo butene-1
c) 3-chloro-2-bromo butene-2
d) None of the above

732. Which of the following undergoes nucleophilic substitution exclusively S_N1 mechanism?

a) Benzyl chloride
b) Isopropyl chloride
c) Chlorobenzene
d) Ethyl chloride

733. The sigma bond energy of C—H bond in C_2H_6 is :

a) 99 kcal
b) 140 kcal
c) 200 kcal
d) 60 kcal

734. The general formula $C_nH_{2n}O_2$ could be for open chain

a) Diketones
b) Carboxylic acids
c) Diols
d) Dialdehydes

735. The correct sequence of steps involved in the mechanism of Cannizzaro's reaction is

a) Nucleophilic attack, transfer of $H^-$and transfer of H^+
b) Transfer of H^- , transfer of H^+ and nucleophilic attack
c) Transfer if H^+, nucleophilic attack and transfer of H^-
d) Electrophilic attack by OH^-, transfer of H^+ and transfer of H^-

736. Examine the following statements regarding S_N 2 reaction

(1) The rate of reaction is independents of concentration of nucleophile
(2) The nucleophile attacks the carbon atom on the side of molecule opposite to the group being displaced
(3) The reaction proceeds with simultaneous bond formation and rupture
Which of the above written statements is correct?
a) 1, 2 b) 1, 3 c) 1, 2, 3 d) 2, 3

737. Propanol and propanone are
a) Functional isomers b) Enantiomers c) Chain isomers d) Structural isomers

738. Diastereomers can be separated by :
a) Fractional distillation b) Simple distillation c) Electrophoresis d) All of these

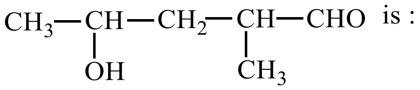
739. Angle strain in cyclopropane is
a) 24°44’ b) 9°44’ c) 44’ d) -5°16’

740. The function of $AlCl_3$ in Friedel-Craft’s reaction is
a) To absorb HCl b) To absorb water c) To produce nucleophile d) To produce electrophile

741. In Kjeldahl’s method of estimation of nitrogen, $CuSO_4$ acts as
a) Oxidising agent b) Reducing agent c) Catalytic agent d) Hydrolysis agent

742. A mixture of acetone and methanol can be separated by
a) Steam distillation b) Vaccum distillation
c) Fractional distillation d) None of these

743. The IUPAC name of,



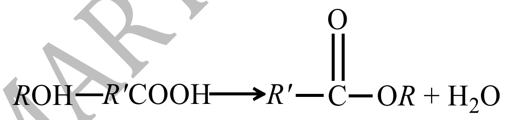
$CH_3-\underset{OH}{\underset{|}{CH}}-CH_2-\underset{CH_3}{\underset{|}{CH}}-CHO$ is :

a) 4-hydroxy-1-methylpentanal
b) 4-hydroxy-2-methylpentanal
c) 3-hydroxy-2-methylpentanal
d) 3-hydroxy-3-methylpentanal

744. The oxygen atom in phenol
a) Exhibits only inductive effect
b) Exhibits only resonance effect
c) Has more dominating resonance effect than inductive effect
d) Has more dominating inductive effect than the resonance effect

745. 2-methylpent-3-enoic acid shows :
a) Optical isomerism
b) Geometrical isomerism
c) Both (a) and (b)
d) None of these

746. In the reaction,



$ROH-R'COOH \longrightarrow R'-\overset{O}{\overset{||}{C}}-OR+H_2O$

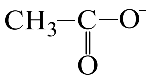
water is formed by the combination of :
a) Hydroxyl of acid with alcoholic hydroxyl hydrogen
b) Hydroxyl of alcohol with carboxylic hydrogen
c) Both the above changes
d) None of the above

747. Pyridine is :
a) An aromatic compound and a primary base
b) A heterocyclic amino compound and a tertiary base
c) An aromatic amino compound and forms salts
d) A cyano derivative of benzene and secondary base

748. The reason for the loss of optical activity of lactic acid when – OH group is changed by H is that

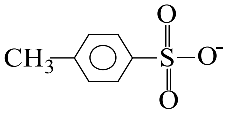
a) Chiral centre of the molecule is destroyed
b) Molecules acquires asymmetry
c) Due to change in configuration
d) Structural changes occurs

749. The correct order of nucleophilicity among the following is :

a) 
$CH_3-C(=O)-O^-$

b) CH_3O^-

c) CN^-

d) 

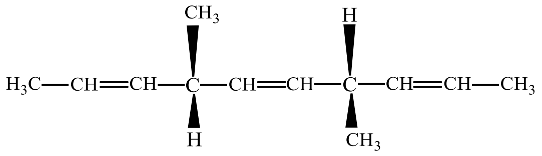
750. Which of the following compounds exhibits rotamers?

a) 2-butene
b) Maleic acid
c) Butane
d) Fumeric acid

751. Ammonia molecule is :

a) A nucleophile
b) An electrophile
c) A homolytic
d) An acid

752. The number of optically active products obtained from the complete ozonolysis of the given compound is :



a) 0
b) 1
c) 2
d) 4

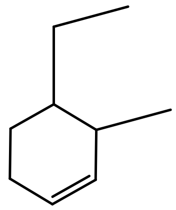
753. The structures,

$CH_3—CH(NH_2)—CH_2—CH_2CH_3$ and

$CH_3—CH_2—CH(NH_2)—CH_2CH_3$ represent:

a) Chain isomers
b) Position isomers
c) Stereo isomers
d) mesomers

754. The systematic (IUPAC) name of the compound with the following structural formula shall be

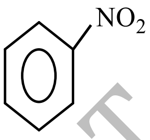


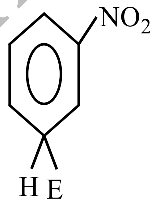
a) 1-ethyl-2-methyl cyclohexene
b) 2-methyl-1-ethyl cyclohexene
c) 3-ethyl-2-methyl cyclohexene
d) 4-ethyl-3-methyl cyclohexene

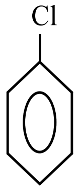
755. 0.5 g of hydrocarbon gave 0.9 g water on combustion. The percentage of carbon hydrocarbon is

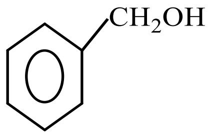
a) 60.6
b) 28.8
c) 80.0
d) 68.6

756. Which one of the following is most reactive towards electrophilic attack?

a) 

b) 

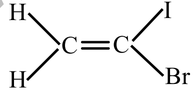
c) 

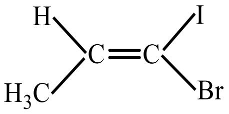
d) 

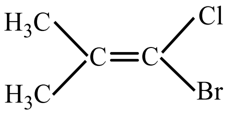
757. Identify, which of the below does not possess any element of symmetry?

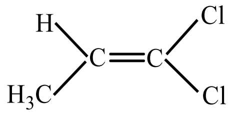
a) (+)(-) tartaric acid
b) Carbon tetrachloride
c) Methane
d) *Meso*-tartaric acid

758. Geometrical isomerism is shown by :

a) 

b) 

c) 

d) 

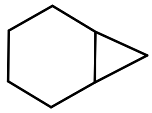
759. When thiourea is heated with metallic sodium, the compound which can't be formed is
a) NaCNS b) NaCN c) Na_2SO_4 d) Na_2S

760. An unknown compound A has a molecular formula C_4H_6. When A is treated with excess of Br_2 a new substance B with formula $C_4H_6Br_4$ is formed. A forms a white ppt. with ammoniacal silver nitrate solution. A may be :
a) But-1-yne b) But-2-yne c) But-1-ene d) But-2-ene

761. The racemisation of optically active compounds is driven by :
a) Entropy b) Enthalpy c) Entropy and enthalpy d) Element of symmetry

762. A cyclic stereoisomer having the molecular formula C_4H_7Cl are classified and tabulated. Find out the correct set of numbers.
Geometrical Optical
a) 6 2 b) 4 2
c) 6 0 d) 4 0

763. The correct name for the following hydrocarbon is



a) Tricycle [4.1.0]heptane b) Bicyclo [5.2.1] heptane
c) Bicyclo [4.1.0] heptane d) Bicyclo [4.1.0] hexane

764. Which of the following is the most stable radical?
a) $\dot{C}H_3$ b) $R\dot{C}H_2$ c) $R_2CH^\bullet$ d) $R_3C^\bullet$

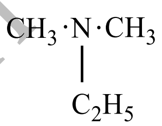
765. The number of 4° carbon atoms in 2,2,4,4-tetramethylpentane :
a) 1 b) 2 c) 3 d) 4

766. Inductive effect involves
a) Delocalisation of σ-electrons b) Displacement of σ-electrons
c) Delocalisation of π −electrons d) Displacement of π-electrons

767. Compounds whose molecules are superimposable on their mirror images even though they contain asymmetric carbon atoms or chiral centres are known as :
a) Enantiomers b) Racemers c) Mesomers d) Conformers

768. Percentage of hydrogen is maximum in .
a) C_2H_4 b) CH_4 c) C_2H_2 d) C_6H_6

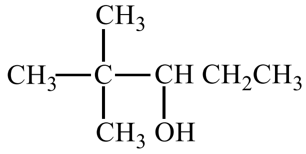
769. Which of the following has most acidic hydrogen?
a) 3-hexanone b) 2, 4-hexanedione c) 2, 4-hexanedione d) 2, 3-hexanedione

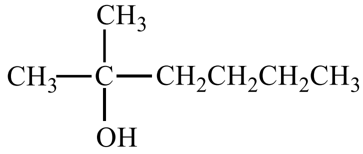
770. IUPAC name of 
$CH_3\cdot\underset{\underset{C_2H_5}{|}}{N}\cdot CH_3$

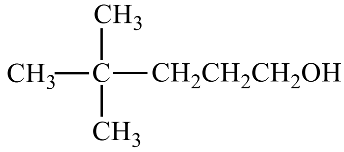
a) N,N- dimethylethanamine
b) N-methyl, N-ethylmethanamine
c) Dimethyl-ethylamine
d) None of the above

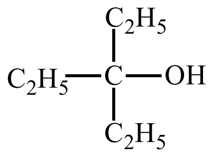
771. Ease of abstraction of hydrogen is greater when attached to :
a) 1° carbon b) 2° carbon c) 3° carbon d) *neo*-carbon

772. *neo*-Heptyl alcohol is correctly represented as :

a) 
$CH_3-C(CH_3)_2-CH(OH)CH_2CH_3$

b) 
$CH_3-C(CH_3)(OH)-CH_2CH_2CH_2CH_3$

c) 
$CH_3-C(CH_3)_2-CH_2CH_2CH_2OH$

d) 
$(C_2H_5)_3C-OH$

773. The IUPAC name for $CH_3COCH(CH_3)_2$ is :
a) 4-methyl isopropyl ketone
b) 3-methyl-2-butanone
c) Isopropyl methyl ketone
d) 2-methyl-3-butanone

774. Steam distillation is based on the fact that vaporisation of organic liquid takes place at
a) Lower temperature than its boiling point
b) Higher temperature than its boiling point
c) Its boiling point
d) Water and organic liquid both undergo distillation

775. IUPAC name of $CH_2 = CH — CH(CH_3)_2$ is :
a) 1,1-dimethyl-2-propene
b) 3-methyl-1-butene
c) 2-vinylpropane
d) 1-isopropyl ethylene

776. The hybridization of central carbon atom in 1,2- propadiene (allene) is
a) sp^3 b) sp^2 c) sp d) None of these

777. The fairly neutral character of CH_3OH is changed to which of the following by adding sodium metal?
a) Acidic b) Neutral c) An electrophile d) A nucleophile

778. The kind of delocalisation involving sigma bond is called
a) Inductive effect b) Hyperconjugation effect
c) Electromeric effect d) Mesomeric effect

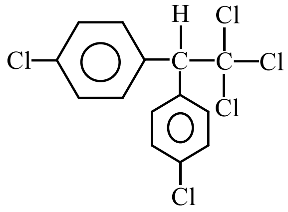
779. In the case of homologous series of alkanes, which one of the following statements is incorrect?
a) The members of the series have the general formula C_nH_{2n+2}, where n is an integer
b) The difference between any two successive members of the series corresponds to 14 unit of relative atomic mass
c) The members of the series are isomers of each other
d) The members of the series have similar chemical properties

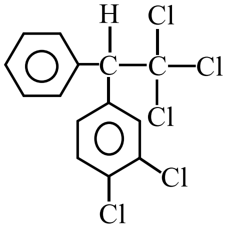
780. Which of the following reagents will be fruitful for separating a mixture of nitrobenzene and aniline?
a) Aq. $NaHCO_3$ b) H_2O c) Aq. HCl d) Aq. NaOH

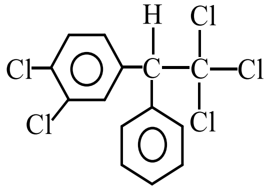
781. The name formic acid was given for HCOOH because it was prepared from :
a) Acetum b) Ant c) Wood d) Oxalis plant

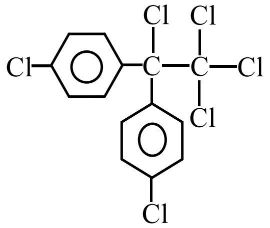
782. 2, 3-dimethyl hexane contains ….. tertiary ….. secondary and ……primary carbon atmos, respectively
a) 2, 2, 1 b) 2, 4, 3 c) 4, 3, 2 d) 3, 2, 4

783. Which one of the following is the correct formula for dichlorodiphenyltrichloro ethane?

a)


b)


c)


d)


784. How many sigma and pi bonds are there in the molecule of di cyano ethane ($\mathrm{CN-CH=CH-CN}$)?
- a) 3 sigma and 3 pi
- b) 5 sigma and 2 pi
- c) 7 sigma and 5 pi
- d) 2 sigma and 3 pi

785. Out of the following, the alkene that exhibits optical isomerism is
- a) 3-methyl-2-pentene
- b) 4-methyl-1-pentene
- c) 3-methyl-1-pentene
- d) 2-methyl-2-pentene

786. The species which use sp^2-hybrid orbitals in its bonding :
- a) $\mathrm{PH_3}$
- b) $\mathrm{NH_3}$
- c) $\mathrm{CH_3^+}$
- d) $\mathrm{CH_4}$

787. Carbanion can undergo :
- a) Rearrangement
- b) Combination with cation
- c) Addition to a carbonyl group
- d) All of the above are correct

788. An organic compound $\mathrm{C_5H_{11}}X$ on dehydrohalogenation gives pentene-2 only. What is halide?
- a) $\mathrm{CH_3CH_2CH}X\mathrm{CH_2CH_3}$
- b) $\mathrm{(CH_3)_2CHCH}X\mathrm{CH_3}$
- c) $\mathrm{CH_3CH_2CH_2CH}X\mathrm{CH_3}$
- d) $\mathrm{CH_3CH_2CH_2CH_2CH_2}X$

789. Percentage composition of an organic compound is as follows

C=10.06, H=0.84, Cl=89.10

Which of the following corresponds to its molecular formula if the vapour density is 60.0?
- a) $\mathrm{CH_3Cl}$
- b) $\mathrm{CHCl_3}$
- c) $\mathrm{CH_2Cl_2}$
- d) None of these

790. Which of the following is most reactive towards nucleophilic substitution reaction?
- a) $\mathrm{CH_2=CH-Cl}$
- b) $\mathrm{C_6H_5Cl}$
- c) $\mathrm{C_6H_5CH_2Cl}$
- d) $\mathrm{ClCH_2-CH=CH_2}$

791. Which among the following statements is correct with respect to the optical isomers?
- a) Enantimoers are non-superimposable mirror images.
- b) Diastereomers are superimposable mirror images.
- c) Enantimoers are superimposable mirror image.
- d) *Meso* forms have no plane of symmetry.

792. Consider thiol anion $(RS^{\ominus})$and alkoxy anion$(RS^{\ominus})$. Which of the following statement is correct?
- a) $RS^{\ominus}$ is less basic and less nucleophilic than $RO^{\ominus}$
- b) $RS^{\ominus}$ is less basic but more nucleophilic than $RO^{\ominus}$

c) $RS^{\ominus}$ is less basic and more nucleophilic than $RO^{\ominus}$
d) $RS^{\ominus}$ is more basic but less nucleophilic than $RO^{\ominus}$

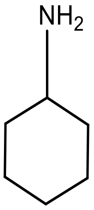
793. The maximum number of alkene isomers for an alkene with molecular formula C_4H_8 is :
a) 2 b) 3 c) 4 d) 5

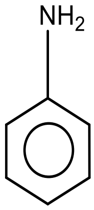
794. The IUPAC name of the compound $(CH_3)_2CH-CH=CH-CHOH-CH_3$ is
a) 5-methyl-hex-3-en-2-ol
b) 2-methyl-hex-3-en-5-ol
c) 2-hydroxy-5-nethyl-3-hexene
d) 5-hydroxy-2-methyl-3-hexene

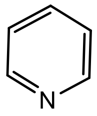
795. The number of isomers in $C_4H_{10}O$ are
a) 7 b) 8 c) 6 d) 5

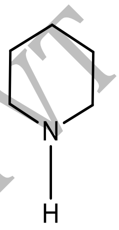
796. Which alkyl halide is preferentially hydrolysed by S_N1 mechanism?
a) $(CH_3)_3C.Cl$ b) $CH_3CH_2CH_2Cl$ c) CH_3CH_2Cl d) CH_3Cl

797. Which of the following is most basic?

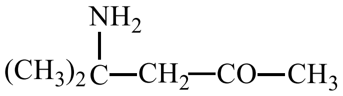
a) 

b) 

c) 

d) 

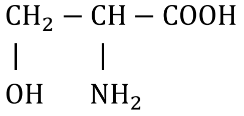
798. The given compound in IUPAC may be called,



$$(CH_3)_2\overset{NH_2}{\overset{|}{C}}-CH_2-CO-CH_3$$

a) Diacetone
b) Acetoneamine
c) Diacetoneamine
d) 4-amino-4-methylpentan-2-one

799. The IUPAC name of the compound,



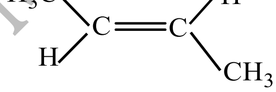
$$\underset{OH}{\underset{|}{CH_2}}-\underset{NH_2}{\underset{|}{CH}}-COOH$$

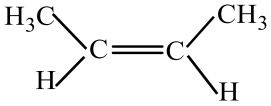
a) 2-amino-3-hydroxy propanoic acid
b) 1-hydroxy-2-amino propan-3-oic acid
c) 1-amino-2-hydroxypropanoic acid
d) 3-hydroxy-2-amino propanoic acid

800. Which of the following compounds is not chiral?
a) 1-chloro-2-methyl pentane
b) 2-chloropentane
c) 1-chloropentane
d) 3-chloro-2-methyl pentane

801. If X is halogen the correct order for S_N2 reactivity is :
a) $R_2CHX > R_3CX > RCH_2X$
b) $RCH_2X > R_3CX > RCH_2X$
c) $RCH_2X > R_2CHX > R_3X$
d) $R_3CX > R_2CHX > RCH_2X$

802. The compound



and


can be distinguished by their :
a) Chlorinated products
b) Products formed by addition of bromine
c) Reaction with H_2/Ni
d) None of the above

803. How many stereoisomers does this molecule have?
$CH_3CH=CHCH_2CHBrCH_3$

a) 6 b) 8 c) 4 d) 2

804. What is the number of possible optical isomers in glucose?

a) 3 b) 4 c) 12 d) 16

805. In which reaction addition takes place according to Markownikoff's rule?

a) $CH_3CH = CHCH_3 + Br \rightarrow$ b) $CH_2 = CH_2 + HBr \rightarrow$

c) $CH_3CH = CH_2 + HBr \rightarrow$ d) $CH_3CH = CH_2 + Br_2 \rightarrow$

806. Presence of halogen in organic compounds can be detected using

a) Leibig's test b) Duma's test c) Kjeldahl test d) Beilstein's test

807. The bond energy for catenation next to carbon is :

a) N b) S c) Si d) P

808. The hydrolysis of alkyl halides by aqueous NaOH is best termed as :

a) Electrophilic substitution reaction

b) Electrophilic addition reaction

c) Nucleophilic addition reaction

d) Nucleophilic substitution reaction

809. Which of the following compounds exhibit stereoisomerism?

a) 3-methyl butyne –1 b) 2-methyl butene –1

c) 2-methyl butanoic acid d) 3-methyl butanoic acid

810. The + I.E.(inductive effect) is shown by :

a) CH_3 b) —OH c) F d) $—C_6H_5$

811. In paper chromatography

a) Mobile phase is liquid and stability phase is solid

b) Mobile phase is solid and stationary phase is liquid

c) Both phases are liquids

d) Both phases are solids

812. Which one of the following is not found in alkenes?

a) Chain isomerism

b) Geometrical isomerism

c) Metamerism

d) Position isomerism

813. Select the correct statement :

a) The prefixes are written before the name of compound

b) The suffixes are written after the name of compound

c) The IUPAC name of a compound is always written as one word

d) All of the above

814. A compound contains 2 dissimilar asymmetric carbon atoms. The number of optically active isomers is :

a) 2 b) 3 c) 4 d) 5

815. The inductive effect

a) Implies the atom's ability to cause bond polarization

b) Increases with increase of distance

c) Implies the transfer of lone pair of electrons from more electronegative atom to the lesser electronegative atom in a molecule

d) Implies the transfer of lone of electrons from lesser electronegative atom to the more electronegative atom in a molecule

816. IUPAC name of the compound, $ClCH_2CH_2COOH$ is :

a) 3-chloropropanoic acid

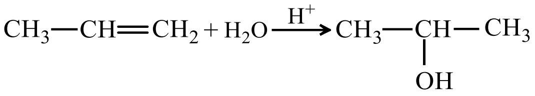
b) 2-chloropropanoic acid

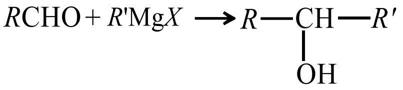
c) 2-chloroethanoic acid

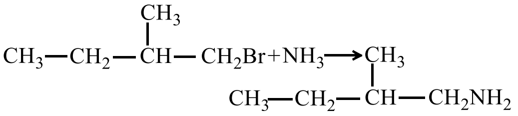
d) Chlorosuccinic acid

817. Which one is a nucleophilic substitution reaction among the following?

a) $CH_3CHO + HCN \longrightarrow CH_3CH(OH)CN$

b) 
$CH_3-CH=CH_2+H_2O \xrightarrow{H^+} CH_3-\underset{OH}{\underset{|}{C}H}-CH_3$

c) 
$RCHO+R'MgX \longrightarrow R-\underset{OH}{\underset{|}{C}H}-R'$

d) 
$CH_3-CH_2-\overset{CH_3}{\overset{|}{C}H}-CH_2Br+NH_3 \longrightarrow CH_3-CH_2-\overset{CH_3}{\overset{|}{C}H}-CH_2NH_2$

818. If there is no rotation of plane polarised light by a compound in a specific solvent, though to be chiral, it means that :
a) It is certainly *meso*
b) It is racemic mixture
c) It is certainly not chiral
d) No such compound

819. Formation of ethylene from acetylene is an example of
a) Elimination reaction b) Substitutions reaction
c) Condensation reaction d) Addition reaction

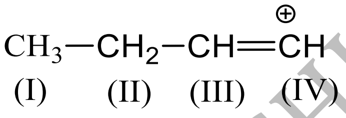
820. Which of the following is nucleophilic addition reaction?
a) Hydrolysis of ethyl chloride by NaOH b) Purification of acetaldehyde by $NaHSO_3$
c) Alkylation of anisol d) Decarboxylation of acetic acid

821. The reagent used in dehalogenation process is :
a) KOH alc. b) Zn dust + alc. c) Na d) KOH(aq)

822. Benzaldoxime exists in how many forms?
a) 1 b) 2 c) 3 d) 4

823. Resonance arises due to the :
a) Migration of atoms
b) Migration of proton
c) Delocalisation of σ-electron
d) Delocalisation of π-electron

824. In the given structure, which carbon atom is most electronegative?


$CH_3-CH_2-CH=\overset{\oplus}{C}H$
(I) (II) (III) (IV)

a) (I) b) (II) c) (III) d) (IV)

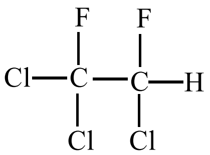
825. The following reactions is an example of …. Reaction.

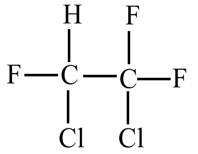
$C_2H_4Br_2 \xrightarrow{\text{Alc.KOH}} C_2H_2$

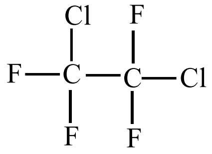
a) Addition b) Dehydrobromination
c) Substitution d) Debromination

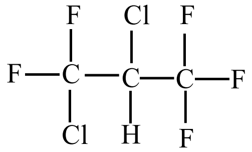
826. Which one of the following pair represents stereoisomerism?
a) Structural and geometrical isomerism
b) Linkage and geometrical isomerism
c) Chain and rotational isomerism
d) Optical and geometrical isomerism

827. Freon-114 is an organic compound. It is chemically called 1,2-dichlorotetrafluoroethane. Its correct structural formula is :

a) 

b) 
$F-\overset{H}{\underset{Cl}{C}}-\overset{F}{\underset{Cl}{C}}-F$

c) 
$F-\overset{Cl}{\underset{F}{C}}-\overset{F}{\underset{F}{C}}-Cl$

d) 
$F-\overset{F}{\underset{Cl}{C}}-\overset{Cl}{\underset{H}{C}}-\overset{F}{\underset{F}{C}}-F$

828. Which of the following compounds is expected to be optically active?
a) $(CH_3)_2CHCHO$ b) $CH_3CH_2CH_2CHO$ c) $CH_3CH_2CHBrCHO$ d) $CH_3CH_2CBr_2CHO$

829. Which of the following is least reactive in a nucleophilic substitution reaction?
a) $(CH_3)_3C-Cl$ b) $CH_2=CHCl$ c) CH_3CH_2Cl d) $CH_2=CHCH_2Cl$

830. During debromination of meso-dibromo-butane, the major compound formed is :
a) n-butane b) l-butene c) *cis*-2-butene d) $trans-2-butene$

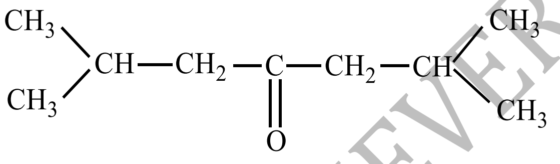
831. What is the empirical formula of a compound having 40% carbon, 6.66% hydrogen and 53.34% oxygen?
a) C_2H_2O b) C_2H_4O c) CH_2O d) CHO

832. Which of the following can act as an nucleophile?
a) BF_3 b) $FeCl_3$ c) $ZnCl_2$ d) C_2H_5MgBr

833. The hybrid orbitals at carbon 2 and 3 in the compound $CH_3CH=CHCH_3$ are :
a) sp^3, sp b) sp^2, sp^2 c) sp, sp d) sp^2, sp

834. The alkyl halide that undergoes S_N1 reaction more radily is
a) Ethyl bromide b) Isopropyl bromide c) Vinyl bromide d) n-propyl bromide

835. The IUPAC name of,

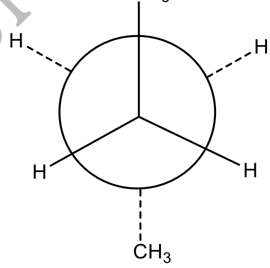


a) 2,4-dimethylhexanone-3
b) 2,6-dimethylheptanone-4
c) 2,6-dimethylhexanone-4
d) 2,6-dimethylheptanone-5

836. In Lassaigne's test, a blue colour is obtained if the organic compound contains nitrogen. The blue colour is due to
a) $K_4[Fe(CN)_6]$ b) $Fe_4[Fe(CN)_6]_3$ c) $Na_3[Fe(CN)_6]$ d) $Cu_2[Fe(CN)_6]$

837. According to Gahn-Ingold-Prelog sequence rules, the correct order of priority for the given group is
a) $-COOH > -CH_2OH > -OH > -CHO$ b) $-COOH > -CHO > -CH_2OH > -OH$
c) $-OH > -CH_2OH > -CHO > -COOH$ d) $-OH > -COOH > -CHO > -CH_2OH$

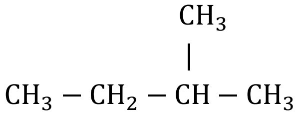
838.

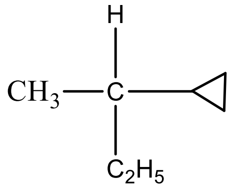


C_2 is rotated anticlockwise 102°C about C_2-C_3 bond. The resulting conformer is
a) Partially eclipsed b) Eclipsed c) gauche d) Staggered

839. Amongst the following compounds, the optically active alkane having lowest molecular mass is

a) $CH_3-CH_2-CH_2-CH_3$

b) 
$CH_3-CH_2-\overset{\overset{CH_3}{|}}{CH}-CH_3$

c) 

d) $CH_3-CH_2-C\equiv CH$

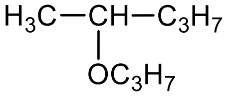
840. How many chiral isomers can be drawn from 2-bromo, 3-chloro butane?

a) 2
b) 3
c) 4
d) 5

841. Glycerol can be separated from spent-lye in soap industry by

a) Steam distillation
b) Fractional distillation
c) Distillation under reduced pressure
d) Ordinary distillation

842. The IUPAC name of


$H_3C-\underset{\underset{OC_3H_7}{|}}{CH}-C_3H_7$

a) 4-propoxy pentane
b) Pentyl-propyl ether
c) 2-propoxy pentane
d) 2-pentoxy propane

843. Correct gradation of basic charactor

a) $NH_3CH_3NH_2 > NF_3$
b) $CH_3NH_2 > NH_3 > NF_3$
c) $NF_3 > CH_3NH_2 > NH_3$
d) $CH_3NH_2 > NF_3 > NH_3$

844. An organic compound contains 49.3% carbon, 6.84% hydrogen and its vapour density is 73. Molecular formula of compound is

a) $C_6H_9O_3$
b) $C_4H_{10}O_2$
c) $C_3H_5O_2$
d) $C_3H_{10}O_2$

845. Vital force theory of the origin of organic compounds was discarded by :

a) Kolbe's synthesis
b) Haber's synthesis
c) Wöhler's synthesis
d) Berthelot's synthesis

846. In the presence of peroxide, hydrogen chloride and hydrogen iodide do not give anti-Markownikoff addition to alkenes because :

a) Both are highly ionic
b) One is oxidizing and the other is reducing
c) One of the steps are exothermic in both the cases
d) All the steps are exothermic in both the cases

847. Which of the following does not have a resonance structure?

a) Benzene
b) Benzaldehyde
c) Acetaldehyde
d) Benzylamine

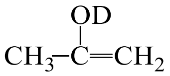
848. Which of the following is the correct order of stability of different conformations of butane?

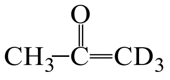
a) Staggered > Gauche > Partially eclipsed > Fully eclipsed
b) Gauche > Staggered > Partially eclipsed > Fully eclipsed
c) Staggered > Fully eclipsed > Partially eclipsed > Gauche
d) None of the above

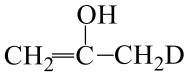
849. Glucose and fructose are :

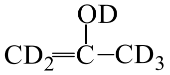
a) Chain isomers
b) Position isomers
c) Functional isomers
d) Optical isomers

850. The enol form of acetone after treatment with D_2O gives :

a) 
$CH_3-\overset{\overset{OD}{|}}{C}=CH_2$

b) 
$CH_3-\overset{\overset{O}{\|}}{C}=CD_3$

c) 
$CH_2=\overset{OH}{\overset{|}{C}}-CH_2D$

d) 
$CD_2=\overset{OD}{\overset{|}{C}}-CD_3$

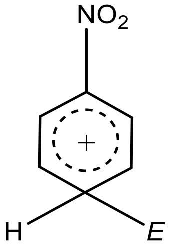
851. Eelipsed and staggered forms of n-butane are called a pair of :
a) Diastereomers b) Conformers c) Isomers d) Enantiomers

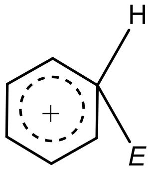
852. Arrange the following in order of increasing dipole moment (I) Toluene (II) m-dichlorobenzene (III) o-dichlorobenzene (IV) p-dichlorobenzene :
a) I < IV < II < III b) IV < I < II < III c) IV < I < III < II d) IV < II < I < III

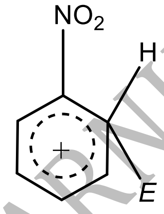
853. In butane, which of the following forms has the lowest energy?
a) Gauche form b) Eclipsed form c) Staggered form d) None of these

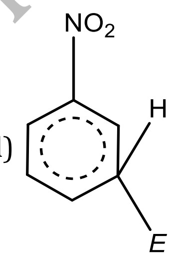
854. Molecular mass of a volatile substances may be obtained by
a) Beilstein method
b) Lassaigne method
c) Victor Mayer's method
d) Leibig's method

855. The electrophile, $E^{\oplus}$ attacks the benzene ring to generate the intermediate σ −complex. Of the following, which σ −complex is of lowest energy?

a) 

b) 

c) 

d) 

856. Which is not deflected by a non-uniform electrostatic field?
a) Water b) Chloroform c) Nitrobenzene d) Hexane

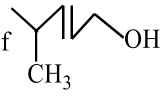
857. The reaction $C_2H_5I + KOH \longrightarrow C_2H_5OH + KI$ is called
a) Hydroxylation substitution
b) Electrophilic substitution
c) Nucleophilic substitution
d) dehydroiodination

858. Correct order of nucleophilicity is
a) $I^- > Br^- > Cl^- > F^-$
b) $F^- > Cl^- > Br^- > I^-$
c) $Cl^- > F^- > Br^- > I^-$
d) $I^- > Cl^- > Br^- > F^-$

859. Due to the presence of an unpaired electron free radicals are
a) Cations b) Anions c) Chemically inactive d) Chemically reactive

860. Which of the following will have *meso* isomers also?
a) 2-hydroxy propanoic acid
b) 2,3-dichlorobutane
c) 2,3-dichloropentane
d) 2-chlorobutane

861. The addition of HBr on butene-2 in presence of peroxide follow the:
a) Electrophilic addition
b) Free radical addition
c) Nucleophilic addition
d) None of these

862. IUPAC name of  is :

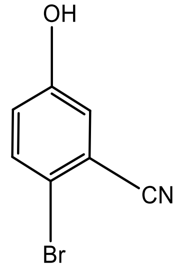
a) 5-methylhexanol b) 2-methylhexanol c) 2-methylhex-3-enol d) 4-methylpent-2-enol

863. In which of the compounds given below there is more than one kind of hybridization $(sp, sp^2, sp^3)$for carbon?
(I)$CH_3CH_2CH_2CH_3$ (II)$CH_3CH = CH - CH_3$
(III)$CH_2 = CH - CH = CH_2$ (IV)$H - C \equiv C - H$

a) (II) and (IV) b) (I) and (IV) c) (II) and (III) d) (II)

864. Which represents nucleophilic aromatic substitution reaction?
a) Reaction of benzene with Cl_2 in sunlight
b) Benzyl bromide hydrolysis
c) Reaction of NaOH with dinitrofluorobenzene
d) Sulphonation of benzene

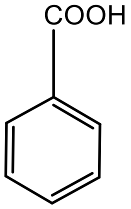
865. The IUPAC name of the following compound, is

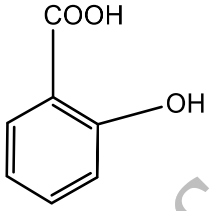


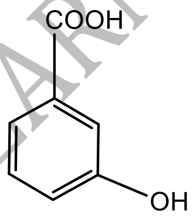
a) 4-bromo-3-cynophenoal
b) 2-bromo-5-hydroxybenzonitrile
c) 2-cyano-4-hydroxybromobenzene
d) 6-bromo-3-hydroxybenzonitrile

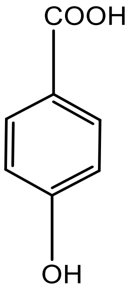
866. Ethoxy ethane and methoxy propane are :
a) Geometrical isomers
b) Optical isomers
c) Functional group isomers
d) Metamers

867. Which of the following aromatic acid is most acidic?

a)


b)


c)


d)


868. The hybridization of carbon in diamond, graphite and acetylene is in the order:
a) sp^3, sp^2, sp b) sp^2, sp^3, sp c) sp, sp^2, sp^3 d) sp^2, sp, sp^3

869. Which is optically active?
a) Isobutyric acid
b) β-chloropropionic acid
c) Propionic acid
d) α-chloropropionic acid

870. Which of the following statement is wrong?
a) Using Lassaigne's test nitrogen and sulphur present in organic compound can be tested
b) Using Beilstein's test the presence of halogen in a compound can be tested
c) In Lassaigne's filtrate the nitrogen present in a organic compound is converted into NaCN
d) In the estimation of carbon, an organic compound is heated with CaO in a combustion tube

871. The reaction, $CH_2 = CHCHO \xrightarrow{HX}$ gives :
a) $CH_3CHXCHO$ b) $CH_2XCHCHO$ c) $CH_2 = CHCHX_2$ d) None of these

872. What kind of isomerism is possible for 1-chloro-2-nitroethene?
a) Functional group isomerism
b) Position isomerism
c) E/Z isomerism
d) Optical isomerism

873. Acetonitrile is
a) CH_3CN b) CH_3COCN c) C_2H_5CN d) C_6H_5CN

874. Formation of cyanohydrin from a ketone is an example of
a) Electrophilic addition
b) Nucleophilic addition
c) Electrophilic substitution
d) Nucleophilic substitution

875. An organic compound which produces a bluish green coloured flame on heating in presence of copper is
a) Chlorobenzene b) Benzaldehyde c) Aniline d) Benzoic acid

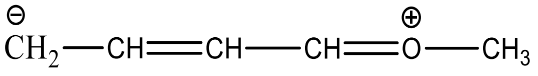
876. The compound abd C—C abd will exist in :
a) 3 forms b) 4 forms c) 5 forms d) 2 forms

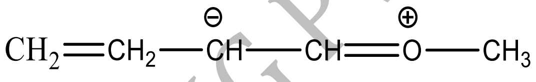
877. Which of the following compounds has the maximum number of π-bonds?
a) $HC \equiv C - CH = CH_2$ b) $CH_2 = CH - CH = CH_2$
c) $CH_3CH_2COCH_3$ d) $C_6H_5 - COOH$

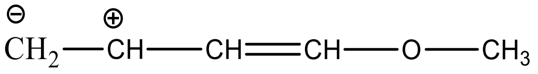
878. The $C - H$ bond distance is longest in
a) C_2H_2 b) C_2H_4 c) C_2H_6 d) $C_2H_2Br_2$

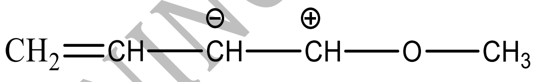
879. The yield in organic reactions is generally poor because the reactions are :
a) Very fast
b) Non-ionic
c) Between covalent compounds
d) Accompanied by side reactions

880. Which of the following resonating structures of 1-methoxy-1, 3-butadiene is least stable?

a) 
$\overset{\ominus}{C}H_2-CH=CH-CH=\overset{\oplus}{O}-CH_3$

b) 
$CH_2=CH_2-\overset{\ominus}{C}H-CH=\overset{\oplus}{O}-CH_3$

c) 
$\overset{\ominus}{C}H_2-\overset{\oplus}{C}H-CH=CH-O-CH_3$

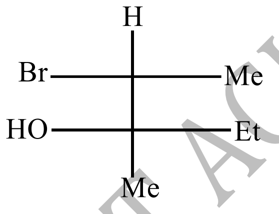
d) 
$CH_2=CH-\overset{\ominus}{C}H-\overset{\oplus}{C}H-O-CH_3$

881. A student named the compound as 1,4-butadiene :
a) The name is correct
b) He committed an error in the selection of carbon chain
c) He committed an error in position of double bond
d) Unpredictable

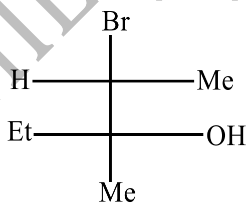
882. The correct IUPAC name of $(C_2H_5)_4$ C is :
a) Tetraethyl methane b) 2-ethylpentane c) 3,3-diethylpentane d) None of these

883. The number of different substitution products possible when ethane is allowed to react with bromine is sunlight are :
a) 9 b) 6 c) 8 d) 5

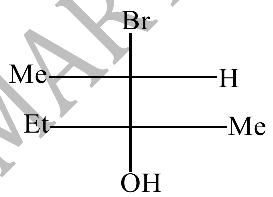
884. Which of the following structures are superimposable?



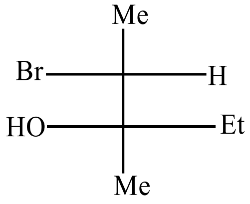
(1)



(2)



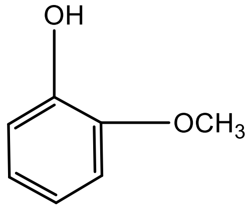
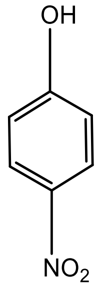
(3)



(4)

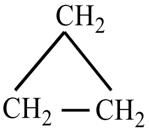
a) 1 and 2 b) 2 and 3 c) 1 and 4 d) 1 and 3

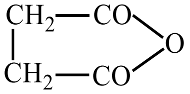
885. Phenol is more acidic than

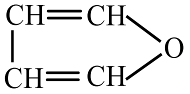
a)  b)  c) C_2H_2 d) Both (a) and (c)

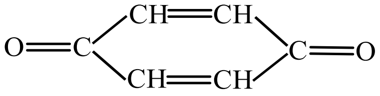
886. During the fusion of an organic compound with sodium metal, nitrogen of the compound is converted into
a) $NaNO_2$ b) $NaNH_2$ c) NaCN d) NaNC

887. The structure representing a heterocyclic compound is :

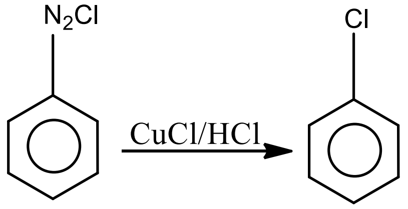
a) 

b) 

c) 

d) 

888. Following reaction is,



a) S_N b) S_E c) El d) EI-CB

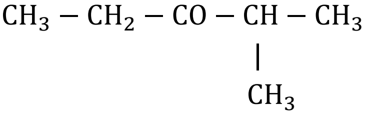
889. Which of the following reactions is an example of nucleophilic substitution reaction?
a) $RX + \text{Mg} \longrightarrow R\text{Mg}X$
b) $RX + \text{KOH} \longrightarrow R\text{OH} + \text{K}X$
c) $2RX + 2\text{Na} \longrightarrow R - R + 2\text{Na}X$
d) $RX + H_2 \longrightarrow RH + HX$

890. How many structural isomers are possible for C_4H_9Cl?
a) 2 b) 4 c) 8 d) 10

891. In which of the following species the central carbon atom is negatively charged?
a) Carbonium ion b) Carbanion c) Carbocation d) Free radicals

892. Select the molecule having only one π-bond :
a) $CH \equiv CH$ b) $CH_2 = CH - CHO$ c) $CH_3 - CH = CH_2$ d) $CH_3 - CH = CHCOOH$

893. Optically active compound among the following is :
a) 2-ethylbutanol-1 b) n-butanol c) 2,2-dimethylbutanol d) 2-methylbutanol-1

894. Which of the following compounds will be most reactive towards nucleophilic addition reaction?
a) $CH_3COCH_2CH_2CH_2CH_3$
b) $CH_3CH_2COCH_2CH_2CH_3$
c) $CH_3CH_2CH_2CH_2CH_2CHO$
d) 

$CH_3 - CH_2 - CO - \underset{\displaystyle CH_3}{\underset{|}{CH}} - CH_3$

895. Lactic acid, $CH_3CH(OH)COOH$ molecule shows :

a) Geometrical isomerism
b) Metamerism
c) Optical isomerism
d) Tautomerism

896. n-pentane and neopentane are :
a) Functional isomers b) Geometrical isomers c) Chain isomers d) Position isomers

897. The IUPAC name of acryldehyde is
a) Prop-2-en-1-al b) Propenylaldehyde c) But-2-en-1-al d) Propenal

898. Due to presence of an unpaired electron, free radicals are
a) Cations b) Anions c) Chemically inactive d) Chemically reactive

899. 2-methylpent-3-ene is a chiral because it has :
a) A centre of symmetry
b) A plane of symmetry
c) Symmetry at C_2 carbon
d) Both centre and a plane of symmetry

900. Which of the following molecules contain asymmetric carbon atom?
a) $CH_3CHClCOOH$ b) CH_3CH_2COOH c) $ClCH_3.CH_2COOH$ d) $Cl_2CHCOOH$

901. Cyclobutane and butene-1 are :
a) Chain isomers b) Position isomers c) Ring-chain isomers d) Metamers

902. Which of the following is not true for carbanions?
a) The carbon carrying the charge has eight valence electrons
b) They are formed by heterolytic fission
c) They are paramagnetic
d) The carbon carrying the charge is sp^3 hydridised

903. Which of the following structures permits *cis-trans* isomerism?
a) $X_2C = CY_2$ b) $XYC = CZ_2$ c) $X_2C = CXY$ d) $XYC = CXY$

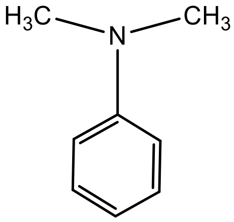
904. Which one of the following compound will show optical iosmerism?
a) $(CH_3)_2 - CH - CH_2 - CH_3$ b) $CH_3 - CHOH - CH_3$
c) $CH_3 - CHCl - CH_2 - CH_3$ d) $CH_3 - CCl_2 - CH_2 - CH_3$

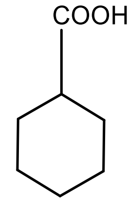
905. The Kolbe's electrolysis proceeds *via*
a) Nucleophilic substitution mechanism b) Electrophilic addition mechanism
c) Free radical mechanism d) Electrophilic substitution reaction

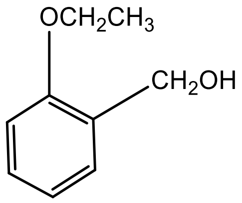
906. Which of the following statements is not correct?
a) Primary carbocation are more stable than secondary ones
b) Secondary free radicals are more stable than primary free radicals
c) Tertiary free radicals are more stable than secondary ones
d) Tertiary carbonium ions are more stable than primary ones

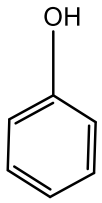
907. Adsorbent is made of ... in TLC
a) Silica gel b) Alumina c) Both (a) and (b) d) None of these

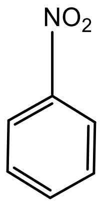
908. Amongst the following, the total number of compounds soluble in aqueous NaOH is

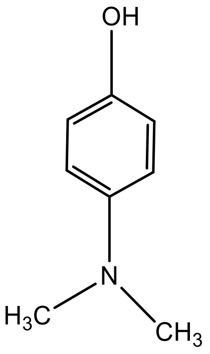


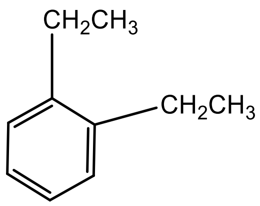


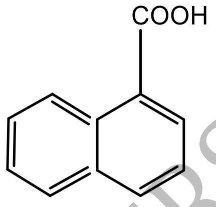












a) 1 b) 2 c) 3 d) 4

909. The ammonia evolved from the treatment of 0.30 g of an organic compound for the estimation of nitrogen was passed in 100 mL of 0.1 M sulphuric acid. The excess of acid required 20 mL of 0.5 M sodium hydroxide solution for complete neutralization. the organic compound is

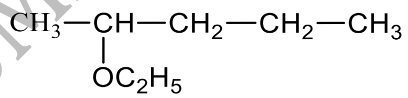
a) acetamide b) benzamide c) urea d) thiourea

910. The structure remaining after one H is removed from hydrocarbon is :

a) Alkyl group b) Alkenyl group c) Alkynyl group d) All of these

911. C_6H_{12} on addition of HBr in presence and in absence of peroxide gives some product.
It is :

a) Hexene-3
b) 2,3-dimethyl butane-2
c) Symmetrical alkene
d) All of these

912. 

$$\underset{\displaystyle OC_2H_5}{\underset{|}{CH_3-CH}}-CH_2-CH_2-CH_3$$

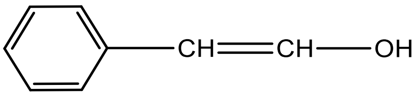
the IUPAC name is

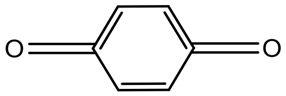
a) 2-ethoxy pentane b) 4-ethoxy pentane c) Pentyl-ethyl ether d) 2-pentoxy ethane

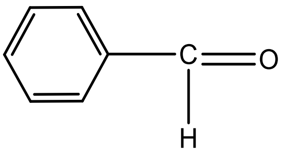
913. A solution of D(+)-2-chloro-2-phenylethane in toluene racemises slowly in the presence of small amount of $SbCl_5$ due to the formation of

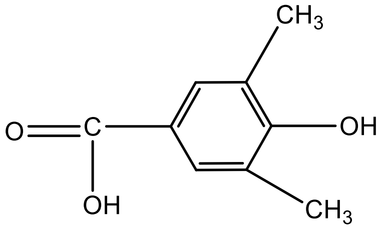
a) Carbanion b) Carbene c) Free radical d) Carbocation

914. Tautomerism is exhibited by

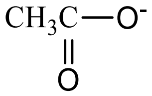
a) 

b) 

c) 

d) 

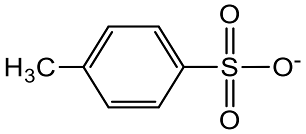
915. The decreasing order of nucleophilicity among the nucleophilies

(A) 

$CH_3\underset{\underset{O}{\|}}{C}-O^-$

(B) CH_3O^-

(C) CN^-

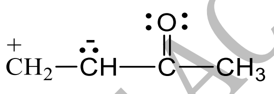
(D) 

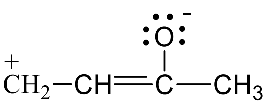
a) (C), (B), (A), (D)
b) (B), (C), (A), (D)
c) (D), (C), (B), (A)
d) (A), (B), (C), (D)

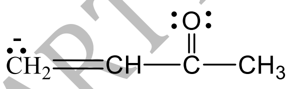
916. Which of the following statements is incorrect?
a) S_N 2 reaction proceeds with inversion
b) S_N1 reaction proceeds with racemisation
c) S_N 2 reaction involves transition state
d) In transition state, one end carries δ^+ and another end carries δ^- charge

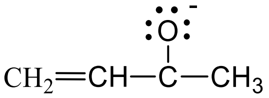
917. The hybridization of carbon atoms in $C-C$ single bond of $HC\equiv C-CH=CH_2$ is
a) sp^3-sp b) sp^3-sp^3 c) sp^2-sp^3 d) $sp-sp^2$

918. One of the stable resonating forms of methyl vinyl ketone is

a) 

b) 

c) 

d) 

919. 5.6 g of an organic compound on burning with excess of oxygen gave 17.6g of CO_2 and 7.2 g of H_2O. The organic compound is
a) C_6H_6 b) C_4H_8 c) C_3H_8 d) CH_3COOH

920. Base strength of,

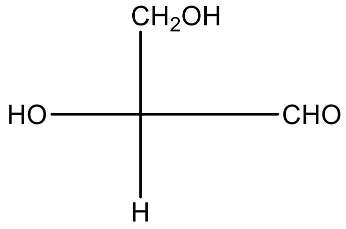
1) $H_3\overset{\ominus}{C}CH_2$

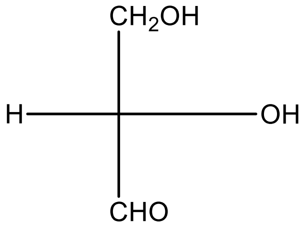
2) $H_2C=\overset{\ominus}{C}H$ and

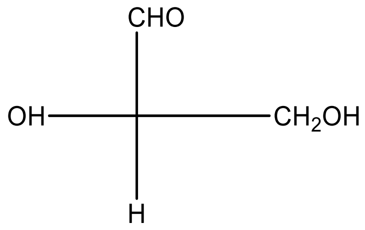
3) $H-C\equiv\overset{\ominus}{C}$ is in the order of :

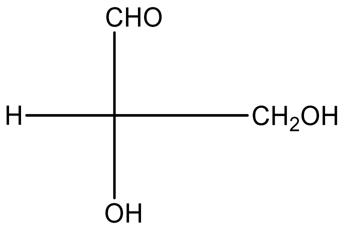
a) (3) > (2) > (1) b) (1) > (3) > (2) c) (1) > (2) > (3) d) (2) > (1) > (3)

921. Which of the following Fischer's projection formula is identical to D-glyceraldehyde?

a) 

b) 

c) 

d) 

922. 1.2g of organic compound of Kjeldahlization liberates ammonia which consumes $30\ cm^3$ of 1N HCl. The percentage of nitrogen in the organic compound is
a) 30 b) 35 c) 46.67 d) 20.8

923. Among the following the dissociation constant is highest for
a) C_6H_5OH b) $C_6H_5CH_2OH$ c) $CH_3-C\equiv CH$ d) $CH_3NH_3^+Cl^-$

924. How much of sulphur is present in an organic compound, if 0.53g of the compound gave 1.158g of $BaSO_4$ on analysis?
a) 10% b) 15% c) 20% d) 30%

925. Which of the following is a dynamic isomerism?
a) Metamerism b) Geometrical isomerism
c) Tautomerism d) Coordinate isomerism

926. Which among the following statements is correct with respect to the optical isomers?
a) Enantiomers are non-superimposable mirror images
b) Diastereomers are superimposable mirror images
c) Enantiomers are superimposable mirror images
d) *Meso* forms have no plane of symmetry

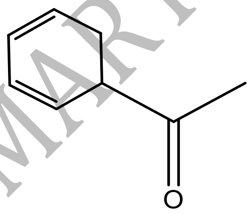
927. The stability order for carbocations given below is :

(I) $R\overset{+}{C}H_2$ (II) $R-\overset{+}{C}=CH_2$ (III) $R-\overset{+}{C}H-CH_3$

a) I < II < III b) III < II < I c) III < I < II d) II < I < III

928. Duma's method involves the determination of nitrogen content in the organic compound in the form of
a) NH_3 b) N_2
c) NaCN d) $(NH_4)_2SO_4$

929. The IUPAC name of



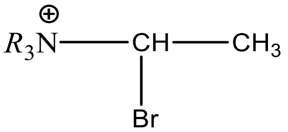
is

a) 1- cyclohexa-2,4-dienylethanone b) 3- cyclohexa-2,4-dienylethanone
c) 1- cyclohexa-3,5-dienylethanone d) 3- cyclohexa-3,5-dienylethanone

930. In the nucleophilic substitution reactions ($S_N2$or S_N1), the reactivity of alkyl halides follows the sequence
a) R-I>R-Br>R-Cl>R-F b) R-Cl>R-F>R-Br>R-I
c) R-F>R-Cl>R-Br>R-I d) R-I>R-F>R-Cl>R-Br

931. $\overset{\oplus}{R_3N}-CH=CH_2 \xrightarrow{HBr}$ product.

Predominant product is

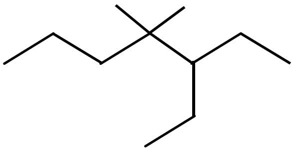
a) 
$\overset{\oplus}{R_3N}-\underset{\underset{Br}{|}}{CH}-CH_3$

b) $R_3N-CH_2-CH_2-Br$

c) $CH_2{=}CH-\overset{\oplus}{N}R_3\ \overset{\ominus}{Br}$

d) No reaction

932. n- pentane and 2-methyl butane are a pair of
a) Enantiomers
b) Stereoisomers
c) Diastereomers
d) Constitutional isomers

933. IUPAC name of compound,  is :
a) 3-ethyl-4,4-dimethyl heptane
b) 1,1-diethyl-2,2-dimethyl pentane
c) 4,4-dimethyl-5,5-diethyl pentane
d) 5,5-diethyl-4, 4-dimethyl pentane

934. IUPAC name of, $(C_2H_5)_2CHCH_2OH$ is :
a) 2-ethylbutanol-1
b) 2-methylpentanol-1
c) 2-ethylpentanol-1
d) 3-ethylbutanol-1

935. Amongst the following the compound that can most readily get sulphonated is
a) Benzene
b) Toluene
c) Nitrobenzene
d) Chlorobenzene

936. In E2 elimination, some compounds follow Hofmann's rule which means:
a) The double bond goes to the most substituted carbon
b) The compound is resistant to elimination
c) No double bond is formed
d) The double bond goes mainly towards the least substituted carbon

937. How many asymmetric carbon atoms are present in
(i) 1, 2-dimethylcyclohexane
(ii) 3-methylcyclopentane and
(iii) 3-methylcyclohexene?
a) Two, one , one
b) One, one, one
c) Two, none, two
d) Two, none One

938. Which of the following is a chiral compound?
a) Hexane
b) n-butane
c) Methane
d) 2,3,4,trimethyl hexane

939. How many structures can compound with molecular formula C_7H_9N have?
a) 5
b) 4
c) 3
d) 2

940. The maximum number of possible optical isomers in 1-bromo-2-methyl cyclobutane is
a) 4
b) 2
c) 8
d) 16

941. Which of the following types of reaction occurs when a substituent has got a double bond with evently distributed π electron cloud?
a) Electrophilic addition
b) Nucleophilic addition
c) Any of the (a) and (b)
d) None of the above

942. Vinyl alcohol and acetaldehyde are :
a) Geometrical isomers
b) Keto-enol tautomers
c) Chain isomers
d) None of these

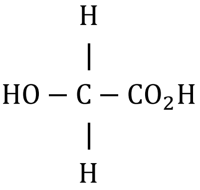
943. 0.25 g of an organic compound on Kjeldahl's analysis gave enough ammonia to just neutralise $10 cm^3$ of 0.5M H_2SO_4. The percentage of nitrogen in the compound is

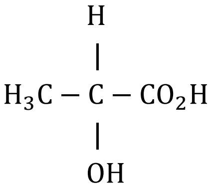
a) 28 b) 56 c) 14 d) 112

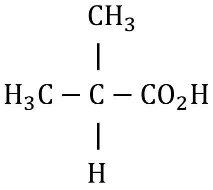
944. Stereoisomers (geometrical or opticals) which are neither superimposable nor mirror image to each other are called :

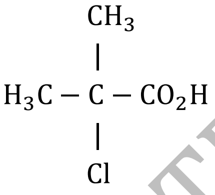
a) Enantiomers b) Mesomers c) Tautomers d) Diastereomers

945. Which one of the following will show optical isomerism?

a) 
$HO-\underset{H}{\overset{H}{C}}-CO_2H$

b) 
$H_3C-\underset{OH}{\overset{H}{C}}-CO_2H$

c) 
$H_3C-\underset{H}{\overset{CH_3}{C}}-CO_2H$

d) 
$H_3C-\underset{Cl}{\overset{CH_3}{C}}-CO_2H$

946. The ion formed by the reaction of HNO_2 and H_2SO_4 is

a) Nitronium ion b) Nitrosonium ion c) Nitrite ion d) Nitrate ion

947. Chloroacetic acid is a stronger acid than acetic acid. This can be explained using

a) $-M$ effect b) $-I$ effect c) $+M$effect d) $+I$ effect

948. The basicity of aniline is less than that of cyclohexylamine. This is due to

a) +R effect of – NH_2 group
b) -I effect of – NH_2 group
c) -R effect of –NH_2 group
d) Hyperconjugation effect

949. The compound is an example of :

a) Aromatic compound
b) Heterocyclic compound
c) Annulene
d) Xanthates

950. Dehydration of alcohol usually goes by

a) E1 mechanism b) E2 mechanism c) E1 cb mechanism d) S_N2 mechanism

951. Geometrical isomerism is possible in

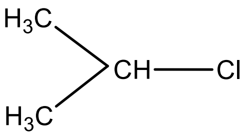
a) Acetone-oxime b) Isobutene c) Acetophenone-oxime d) Benzophenone-oxime

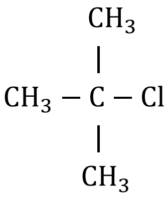
952. Ethers are isomeric with

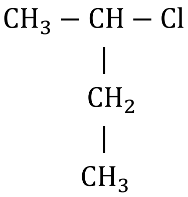
a) Aldehydes
b) Ketones
c) Both aldehydes and ketones
d) Alcohols

953. S_N1 reaction is fastest in

a) CH_3CH_2Cl

b) 

c) 
$CH_3-\underset{CH_3}{\overset{CH_3}{C}}-Cl$

d) 
$CH_3-\underset{\underset{CH_3}{|}}{\underset{CH_2}{\underset{|}{CH}}}-Cl$

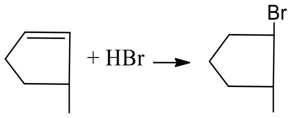
954. Vinyl chloride undergoes

a) Only addition reactions
b) Only elimination reactions
c) Both (a) and (b)
d) Substitution reactions

955. Fischer projection indicates :

a) Horizontal substituents above the plane
b) Vertical substituents above the plane
c) Both horizontal and vertical substituents below the plane
d) Both horizontal and vertical substituents above the plane

956. The reaction,



Is an example of
a) Nucleophilic substitution
b) Electrophilic addition
c) Elimination reaction
d) Nucleophilic addition

957. Acetone and propen-2-ol are :
a) Positional isomers b) Leto-enol tautomers c) Geometrical isomers d) Chain isomers

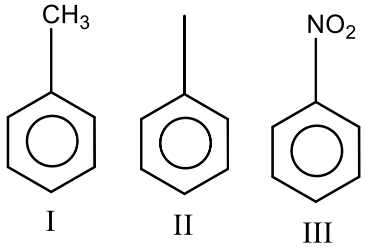
958. The number of stereoisomers obtained by bromination of *trans* − 2 −butene is?
a) 1 b) 2 c) 3 d) 4

959. The compound which forms one monochloro product when treated with chlorine is :
a) *n*-pentane b) Isopentane c) *neo*-pentane d) None of these

960. Reactivity towards nucleophilic addition reaction of
(I)HCHO (II)CH_3CHO (III)$CH_3COCH_3$is
a) II>III>I b) III>II>I c) I>II>III d) I>II<III

961. Maleic acid and fumaric acid are
a) Position isomers b) Geometric isomers c) Enantimoers d) Functional isomers

962. The ease of nitration of the following three hydrocarbons follows the order



a) II=III≈I b) II>III>I c) III>II>I d) I=III>II

963. Which represents the condensed formula for pentanes?
a) $CH_3(CH_2)_3CH_3$ b) $(CH_3)_3CCH_3$ c) $(CH_3)_2CHCH_2CH_3$ d) All of these

964. Which of the substance is purified by sublimation?
a) Benzoic acid b) Camphor c) Naphthalene d) All of these

965. The halogen compound which most readily undergoes nucleophilic substitutions is
a) $CH_2 = CHCl$ b) $CH_3CH = CHCl$
c) $CH_2 = CHC(Cl) = CH_2$ d) $CH_2 = CHCH_2Cl$

966. Which of the following order is correct regarding the acidity of carboxylic acids?
a) $Cl_3CCOOH > Cl_2CHCOOH > ClCH_2COOH$ b) $Cl_3CCOOH > Cl_2CHCOOH < ClCH_2COOH$
c) $Cl_3CCOOH < Cl_2CHCOOH > ClCH_2COOH$ d) $Cl_3CCOOH < Cl_2CHCOOH < ClCH_2COOH$

967. An S_N2 reaction at an asymmetric carbon of a compound always gives
a) A mixture of diastereomers b) A single stereoisomer
c) An enantiomer of the substrate d) A product with opposite optical rotation

968. The IUPAC name of the compound, $CH_3CH = CHC \equiv CH$ is :
a) Pent-4-yn-2-ene b) Pent-3-en-1-yne c) Pent-2-en-4-yne d) Pent-1-yn-3-ene

969. Reaction of methyl bromide with aqueous sodium hydroxide involves
a) Racemisation b) S_N1 mechanism
c) Retention of configuration d) S_N2 mechanism

970. An organic compound X(mol. formula $C_6H_5O_2N$) has six carbons in a ring system, three double bonds and also a nitro group as substituent. X is :
a) Homocyclic but not aromatic
b) Aromatic but not homocyclic
c) Homocyclic and aromatic

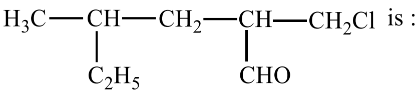
d) heterocyclic

971. The compounds CH_3NH_2 and $CH_3CH_2.NH_2$ are :
a) Isomers b) Isobars c) Homologous d) Allotropes

972. The following compound will undergo electrophilic substitution more readily than benzene
a) Nitrobenzene b) Benzoic acid c) Benzaldehyde d) Phenol

973. Which of the following elements can't be detected by direct tests?
a) N b) O c) S d) Br

974. IUPAC name of ,



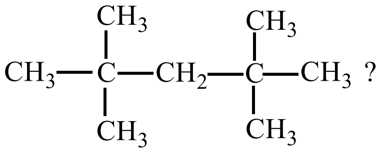
$H_3C-\underset{C_2H_5}{\underset{|}{CH}}-CH_2-\underset{CHO}{\underset{|}{CH}}-CH_2Cl$ is :

a) 2-chloromethyl-4-methyl-hexanal
b) 1-chloro-4-ethyl-2-pentanal
c) 1-chloro-4-methyl-2-hexanal
d) 1-chloro-2-aldo-4-methyl hexane

975. Position isomerism is shown by :
a) o-nitrolhenol and p-nitrophenol
b) Dimethyl ether and ethanol
c) Pentan-2-one and pentan-3-one
d) Acetaldehyde and acetone

976. Formulae of phenyl carbinol and chloral are respectively :
a) $C_6H_5.CH_2CH_2OH$ and $CHCl_2CHO$
b) $C_6H_5CH_2OH$ and CCl_3CHO
c) C_6H_5OH and $CH_2Cl.CHO$
d) C_6H_5CHO and $CHCl_2CHO$

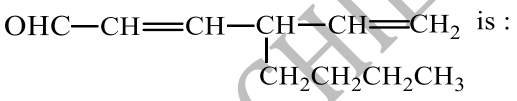
977. How many primary carbon atoms are there in the compound,



$CH_3-\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}-CH_2-\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}-CH_3$?

a) 6 b) 2 c) 4 d) 3

978. IUPAC name of,



$OHC-CH=CH-\underset{CH_2CH_2CH_2CH_3}{\underset{|}{CH}}-CH=CH_2$ is :

a) 4-butyl-2,5-hexadien-l-al
b) 5-vinyloct-3-en-l-al
c) 5-vinyloct-5-en-8-al
d) 3-butyl-1,4-hexadien-6-al

979. The molecular formula of a saturated compound is $C_2H_4Br_2$.
This formula permits the existence of :
a) Functional isomers b) Optical isomers c) Positional isomers d) $cis-trans$ isomers

980. Which of the following solvents are aprotic?
(A)NH_3 (B)SO_2
(C)CH_3CN (D)CH_3CO_2H
a) A,B,C b) A,C,D c) B,C d) A,C

981. The reaction of sodium ethoxide with iodoethane to from diethyl is termed as
a) Electrophilic substitution b) Nuclephilic substitution
c) Electrophilic addition d) Radical substitution

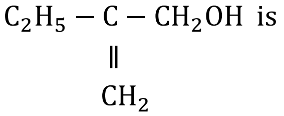
982. The sodium extract of an organic compound on acidification with acetic acid and addition of lead acetate solution gives a black precipitate. The organic compound contains

a) Nitrogen b) Halogen c) Sulphur d) Phosphorus

983. The IUPAC name of the compound $CH_3CONHBr$ is
a) 1-bromoacetamide b) ethanoylbromide c) N-bromoethanamide d) None of these

984. The silver salt of a monobasic acid on ignition gave 60% of Ag. The molecular weight of the acid is
a) 37 b) 57 c) 73 d) 88

985. The IUPAC name of compound,



$C_2H_5 - \underset{\underset{CH_2}{\|}}{C} - CH_2OH$ is

a) 2-ethylprop-2-en-1-ol
b) 2-hydroxymethylbutan -1-0l
c) 2-methylenebutan-1-0l
d) 2-ethyl-3-hydroxyprop-1-ene

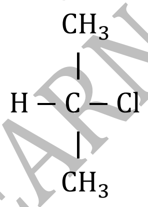
986. How many optically active stereoisomers are possible for butane-2, 3-diol?
a) 0 b) 1 c) 2 d) 3

987. Removal of a hydride ion from a methane molecule will give a:
a) Methyl radical b) Carbonium ion c) Carbanion d) Methyl group

988. Which of the following will have a *meso*-isomer also?
a) 2-chlorobutane
b) 2, 3-dichlorobutane
c) 2, 3-dichloropentane
d) 2-hydroxypropanoic acid

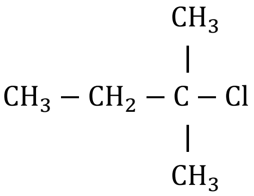
989. Which chlorine atom is more electronegative in the following?

a) $CH_3 - Cl$
b) $CH_3 - CH_2 - Cl$
c)



$H - \underset{\underset{CH_3}{|}}{\overset{\overset{CH_3}{|}}{C}} - Cl$

d)



$CH_3 - CH_2 - \underset{\underset{CH_3}{|}}{\overset{\overset{CH_3}{|}}{C}} - Cl$

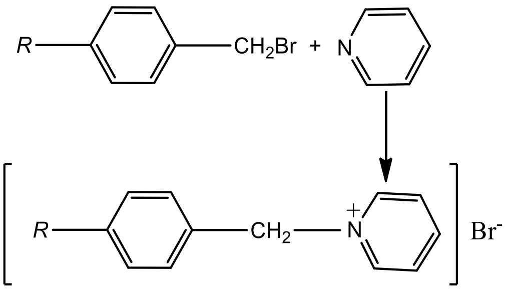
990. The resonating structures :
a) Differ only in the arrangement of electrons
b) Differ in number of paired and unpaired electrons
c) Differ largely in their energy contents
d) Do not lie in the same plane

991. The optical isomers, which are not enantiomers, are called
a) Conformer b) Diastereomer c) Mirror images d) None of these

992. α-D-(+)-glucose and β-D-(+)-glucose are :
a) Enantiomers
b) Conformers
c) Epimers
d) Anomers

993. Which of the following orders is correct regarding the acidity of carboxylic group?
a) $CH_3CH_2CH(Cl)COOH > CH_3CH(Cl)CH_2COOH > ClCH_2CH_2CH_2COOH$
b) $CH_3CH_2CH(Cl)COOH < CH_3CH(Cl)CH_2COOH < ClCH_2CH_2CH_2COOH$
c) $CH_3CH_2CH(Cl)COOH > CH_3CH(Cl)CH_2COOH < ClCH_2CH_2CH_2COOH$
d) $CH_3CH_2CH(Cl)COOH < CH_3CH(Cl)CH_2COOH > ClCH_2CH_2CH_2COOH$

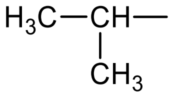
994. The rate of the reaction,

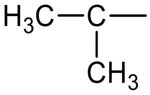


is influenced by the hyper conjugation effect of group *R*. If *R* sequentially is

I. CH_3-

II. CH_3-CH_2-

III. 
$H_3C-\underset{CH_3}{CH}-$

IV. 
$H_3C-\underset{CH_3}{C}-$

the increasing order of speed of the above reaction is

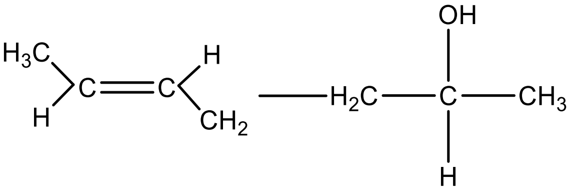
a) IV, III, II, I

b) I, II, III, IV

c) I, IV, III, II

d) III, II, I, IV

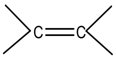
995. The compound, whose stereo-chemical formula is written below, exhibits x geometrical isomers and y optical isomers



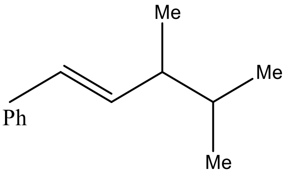
The values of x and y are

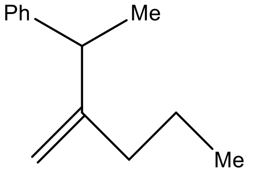
a) 4 and 4 b) 2 and 2 c) 2 and 4 d) 4 and 2

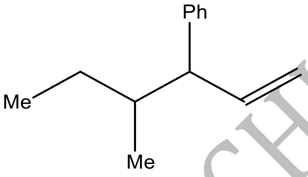
996. Geometrical isomerism is shown by

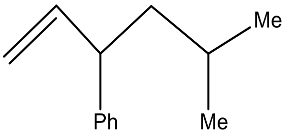
a) $-C\ -C-$ b)  c) $-C\equiv C-$ d) None of these

997. Which one of the following is *s*-butyl phynylvinyl methane?

a) 

b) 

c) 

d) 

998. Arrange the carbanions,
$(CH_3)_3\bar{C}, \bar{C}Cl_3, (CH_3)_2\bar{C}H, C_6H_5\bar{C}H_2$, in order of their decreasing stability

a) $C_6H_5\bar{C}H_2 > \bar{C}Cl_3 > (CH_3)_2\bar{C} > (CH_3)_2\bar{C}H$

b) $(CH_3)_2\bar{C}H > \bar{C}Cl_3 > C_6H_5CH_2 > (CH_3)_3\bar{C}$

c) $\bar{C}Cl_3 > C_6H_5\bar{C}H_2 > (CH_3)_2\bar{C}H > (CH_3)_3\bar{C}$

d) $(CH_3)_3\bar{C} > (CH_3)_2\bar{C}H > \bar{C}H_2 > \bar{C}Cl_3$

999. $RX + I^- \rightarrow R-I+X^-$ is an example of … reaction.

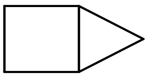
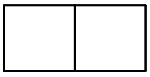
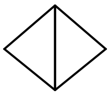
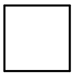
a) Nucleophilic addition

b) Nucleophilic substitution

c) Electrophilic addition

d) Elimination

1000. Bicyclo (1,1,0) butane is

a)  b)  c)  d) 

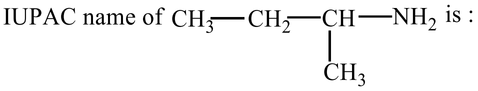
1001. The basic strength of
$CH\equiv\bar{C}, CH_2=\bar{C}H, CH_3\bar{C}H_2$
I II III
Will be in order

I II III

a) I <II<III b) II<III<I c) III<II<I d) III<I<II

1002. Which of the following is most reactive towards elimination reaction?

a) $RCOO^-$ b) CN^- c) NO_3^- d) RO^-

1003. IUPAC name of  $CH_3-CH_2-\underset{CH_3}{\underset{|}{CH}}-NH_2$ is :

a) 1-methyl-1-aminopropane

b) 2-aminobutane

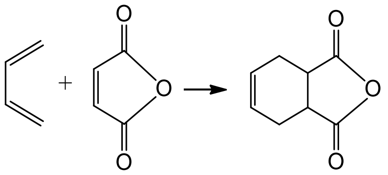
c) 2-methyl-3-aminopropane

d) None of the above

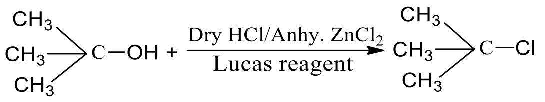
1004. The number of isomeric hexanes is

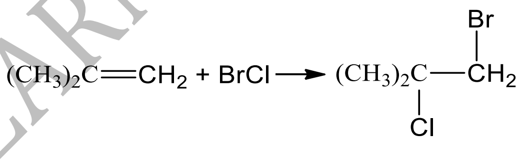
a) 5 b) 2 c) 3 d) 4

1005. The substitution reaction among the following is

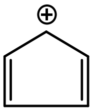
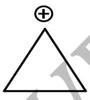
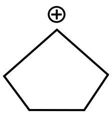
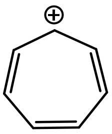
a) 

b) 

c) 

d) 

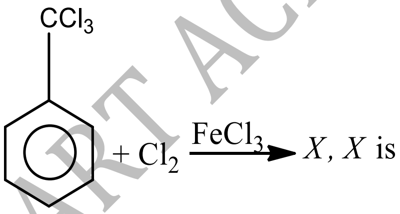
1006. The most stable carbocation is

a) 
b) 
c) 
d) 

1007. Among the following alkenes (I) 1-butene, (II) *cis*-2-butene, (III) *trans*-2-butene the decreasing order of stability is :

a) III > II > I b) III > I > II c) I > II > III d) II > I > III

1008. For the reaction,



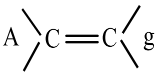
a) Chloro benzene and carbon tetrachloride

b) *meta* chloro benzotrichloride

c) *ortho*, *para* chloro benzotrichloride

d) None of the above

1009. Which of the following statements is not correct?

a)  group is made up of 4σ– bond and 2π–bonds

b) A σ-bond is stronger than π-bond

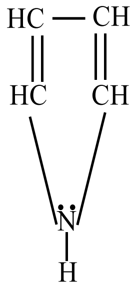
c) A σ-bond can exist independently of π-bond

d) A double bond is stronger than a single bond

1010. The number of sp^3- hybrid carbons in 2-butyne is :

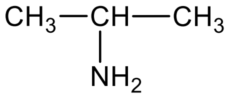
a) 4　b) 3　c) 2　d) 1

1011. How many π-electrons are there in following?



a) 2　b) 4　c) 6　d) 8

1012. IUPAC name of



$CH_3-\underset{NH_2}{\underset{|}{C}H}-CH_3$

is

a) Dimethyl amine

b) 2-amino propane

c) Isopropylamine

d) 2-propanamine

1013. An organic compound having molecular mass 60 is found to contain C=20%, H=6.67% and N=46.67% while rest is oxygen. On heating it gives NH_3 along with a solid residue. The solid residue gives violet colour with alkaline copper sulphate solution. The compound is

a) $CH_3CH_2CONH_2$　b) $(NH_2)_2CO$

c) CH_3CONH_2　d) CH_3NCO

1014. How many chiral carbon atoms are present in 2, 3, 4- trichloropentane?

a) 4　b) 1　c) 2　d) 3

1015. Which one of the following compounds is most polar?

a) CH_2I_2　b) CH_2F_2　c) CH_2Cl_2　d) CH_2Br_2

1016. Geometrical isomerism is not shown by

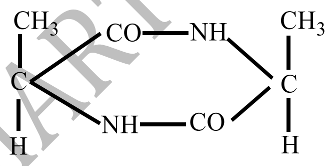
a) 1, 1-dichloro-1-pentene　b) 1,2-dichloro-1-pentene

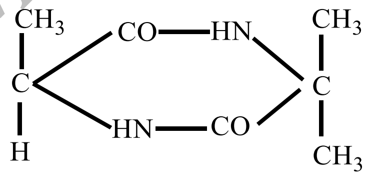
c) 1, 3-dichloro-2-pentene　d) 1, 4-dichloro-2-pentene

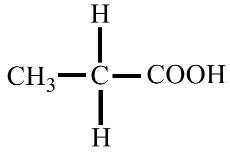
1017. The change in optical rotation with time of freshly prepared solution of sugar is known as:

a) Specific rotation　b) Inversion　c) Rotatory motion　d) Mutarotation

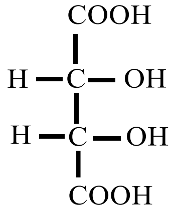
1018. Which of the following does not show stereo isomerism?

a) 

b) 

c) 

$CH_3-\overset{H}{\overset{|}{\underset{H}{\underset{|}{C}}}}-COOH$

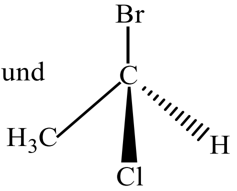
d) 

1019. One of the following compounds exhibit geometrical iosmerism

a) $CH_3CH_2CH_2CH_3$
b) $CH_3 - HC(CH_3) - H(C)CH_3 - CH_3$
c) $CH_3 - HC(CH_3) - CH_3$
d) $CH_3CH = CH - CH_3$

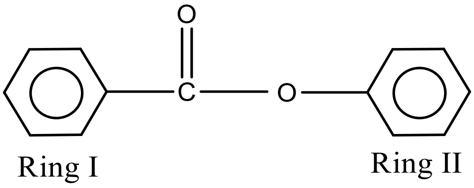
1020. Which one of the following shows functional isomerism?

a) C_2H_4
b) C_3H_6
c) C_2H_5OH
d) CH_2Cl_2

1021. The chirality of the compound 

a) R
b) S
c) Z
d) I

1022. In the compound



electrophilic substitution occurs at

a) *ortho*/*para* position at ring I
b) *meta* position at ring I
c) *ortho*/*para* position at ring II
d) *meta* position at ring II

1023. In the reaction of phenol with chloroform and aqueous solution of NaOH at 70°C, the electrophile attacking the ring is

a) $CHCl_3$
b) $CHCl_2$
c) $:CCl_2$
d) $COCl_2$

1024. The production of an optically active compound from a symmetric molecule without resolution in termed as :

a) Walden inversion
b) Partial racemisation
c) Asymmetric synthesis
d) Partial resolution

1025. An organic compound on heating with CuO produces CO_2 but no water. The organic compound may be

a) Carbon tetrachloride
b) Chloroform
c) Methane
d) Ethyl iodide

1026. Which of the following statement is not applicable to Beilstein test?

a) Green or bluish green flame is due to the formation of volatile cupric halides
b) It does not tell us to which halogen is present in the organic compound
c) It is very sensitive test can be easily performed
d) It is a sure test for the presence of halogen

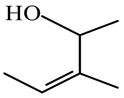
1027. Essential oils can be isolated by

a) Crystallization
b) Steam distillation
c) Sublimation
d) Distillation

1028. Mesomeric effect involves delocalisation of

a) Pi-electrons b) Sigma electrons c) Protons d) None of these

102 The IUPAC name of the compound,
9.



is :

a) 1,2-dimethyl-2-butenol
b) 3-methylpent-3-en-2-ol
c) 3,4-dimethyl-2-buten-4-ol
d) 2,3-dimethyl-3-pentenol

103 Which of the following species is paramagnetic in nature?
0.

a) Carbonium ion b) Free radical c) Carbene d) Nitrene

103 Isobutyl chloride is :
1.

a) $CH_3CH_2CH_2CH_2Cl$ b) $(CH_3)_2CHCH_2Cl$ c) $CH_3CH_2CHClCH_3$ d) $(CH_3)_3\ C-Cl$

103 How many isomers will C_3H_6 have?
2.

a) 1 b) 2 c) Zero d) 4

103 Which one of the following compounds is capable of existing in a *meso* form?
3.

a) 3, 3-dibromopentane b) 4-bromo-2-pentanol
c) 3-bromo-2-pentanol d) 2, 3-dibromopentane

103 Geometrical isomerism is caused :
4.

a) By restricted rotation around $C=C$ bond
b) By the presence of one asymmetric carbon atom
c) Due to different groups attached to the same functional group
d) By swing of hydrogen atom between two polyvalent atoms

103 Lassaigne's test is used for the detection of
5.

a) Carbon only b) Hydrogen only
c) Oxygen only d) Nitrogen, sulphur and halogens

103 Which of the following is arranged according to the nature indicated?
6.

a) Electrophile
b) Electrophile
c) Electrophile – CH_3OH, N_3^-. Nucleophile – NO_2^+, Br^+
d) Electrophile – Br^+, N_3^-, Nucleophile – CH_3OH,

103 The number of possible alkynes with molecular formula C_5H_8 is
7.

a) 3 b) 4 c) 5 d) 6

103 Example of geometrical isomerism is
8.

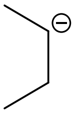
a) 2-butanol b) 2-butene c) Butanal d) 2-butyne

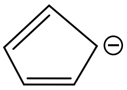
103 Which of the following is an example of elimination reaction?
9.

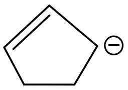
a) Chlorination of methane
b) Dehydration of ethanol
c) Nitration of benzene
d) Hydroxylation of ethylene

104 The order of stability of the following carbanion is

0. I: $\overset{\ominus}{R}CH_2$







I II III IV

a) I>II>III>IV b) I>III>II>IV c) IV>III>II>I d) III>IV>I>II

1041. Which nomenclature in IUPAC is not correct?

a) Pentyne-3 b) Pentyne-2 c) Hexyne-3 d) None of these

1042. The C—C bond angle in cyclopropane is :

a) 60° b) 120° c) 109°28′ d) 180°

1043. Absolute alcohol cannot be obtained by simple fractional distillation because

a) Pure C_2H_5OH is unstable
b) C_2H_5OH forms hydrogen bonds with water
c) Boiling point of C_2H_5OH is very close to that of water
d) Constant boiling azeotropic mixture is formed with water

1044. Racemic compound has

a) Equimolar mixture of enantiomers b) 1:1 mixture of enantiomer and diastereomer
c) 1:1 mixture of diastereomers d) 1:2 mixture of enantiomers

1045. In case of a homologous series each member differs from the preceding or the succeeding member by :

a) a CH_2 group b) a CH_3 group c) Two hydrogen atoms d) Four hydrogen atoms

1046. Vaccum distillation is used to purify liquids which

a) Are highly volatile b) Are explosive in nature
c) Decompose below their boiling points d) Have high boiling point

1047. The number of isomeric structures for C_2H_7N would be :

a) 4 b) 3 c) 2 d) 1

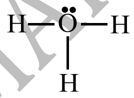
1048. The IUPAC name of CH_3-CH_2-CHO

a) Propanal-1 b) 2-methylbutanal c) Butanal-1 d) Pentanal-1

1049. Which of the following is the weakest base?

a) Ethyl amine b) Ammonia c) Dimethyl amine d) Methyl amine

1050. The structure which has positive charge on the oxygen atom :

a) 
$H-\ddot{O}(H)-H$

b) $CH_3-\ddot{\underset{..}{O}}-CH_3$ c) $CH_3-\ddot{\underset{..}{O}}\colon$ d) $H-\ddot{\underset{..}{O}}-\ddot{\underset{..}{O}}-H$

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

: ANSWER KEY :

1)	d	2)	a	3)	b	4)	c	165)	a	166)	b	167)	c	168)	a
5)	c	6)	d	7)	b	8)	c	169)	d	170)	d	171)	d	172)	d
9)	b	10)	a	11)	c	12)	c	173)	b	174)	c	175)	c	176)	b
13)	b	14)	b	15)	a	16)	d	177)	b	178)	d	179)	d	180)	a
17)	d	18)	b	19)	c	20)	b	181)	b	182)	b	183)	a	184)	d
21)	a	22)	c	23)	b	24)	a	185)	a	186)	a	187)	c	188)	c
25)	b	26)	b	27)	b	28)	a	189)	d	190)	b	191)	c	192)	d
29)	a	30)	d	31)	d	32)	c	193)	c	194)	a	195)	c	196)	c
33)	c	34)	c	35)	d	36)	d	197)	b	198)	b	199)	d	200)	d
37)	b	38)	b	39)	a	40)	c	201)	d	202)	b	203)	a	204)	a
41)	b	42)	a	43)	a	44)	b	205)	a	206)	b	207)	b	208)	a
45)	b	46)	b	47)	b	48)	d	209)	c	210)	a	211)	a	212)	d
49)	d	50)	c	51)	d	52)	c	213)	d	214)	d	215)	a	216)	d
53)	d	54)	c	55)	d	56)	c	217)	b	218)	c	219)	a	220)	d
57)	d	58)	d	59)	a	60)	b	221)	a	222)	a	223)	b	224)	c
61)	b	62)	a	63)	d	64)	b	225)	a	226)	b	227)	d	228)	a
65)	a	66)	a	67)	d	68)	c	229)	d	230)	b	231)	a	232)	c
69)	c	70)	b	71)	b	72)	c	233)	a	234)	c	235)	b	236)	c
73)	d	74)	c	75)	d	76)	c	237)	c	238)	b	239)	c	240)	c
77)	b	78)	d	79)	b	80)	c	241)	c	242)	d	243)	b	244)	d
81)	c	82)	d	83)	b	84)	b	245)	d	246)	d	247)	d	248)	c
85)	a	86)	c	87)	d	88)	d	249)	d	250)	b	251)	a	252)	d
89)	b	90)	a	91)	a	92)	d	253)	c	254)	b	255)	d	256)	c
93)	a	94)	b	95)	d	96)	c	257)	b	258)	b	259)	b	260)	d
97)	a	98)	c	99)	c	100)	d	261)	a	262)	d	263)	c	264)	b
101)	c	102)	d	103)	d	104)	a	265)	d	266)	d	267)	a	268)	a
105)	d	106)	a	107)	c	108)	b	269)	c	270)	c	271)	a	272)	c
109)	b	110)	c	111)	c	112)	b	273)	b	274)	c	275)	b	276)	b
113)	a	114)	c	115)	d	116)	b	277)	d	278)	c	279)	b	280)	c
117)	d	118)	c	119)	b	120)	c	281)	a	282)	b	283)	b	284)	d
121)	a	122)	b	123)	a	124)	d	285)	c	286)	a	287)	b	288)	c
125)	d	126)	b	127)	d	128)	d	289)	c	290)	a	291)	d	292)	a
129)	d	130)	c	131)	d	132)	b	293)	b	294)	b	295)	a	296)	a
133)	d	134)	c	135)	a	136)	b	297)	c	298)	a	299)	b	300)	b
137)	a	138)	d	139)	c	140)	c	301)	a	302)	d	303)	a	304)	a
141)	a	142)	d	143)	b	144)	b	305)	b	306)	b	307)	b	308)	c
145)	c	146)	c	147)	a	148)	c	309)	a	310)	a	311)	c	312)	c
149)	d	150)	c	151)	c	152)	c	313)	c	314)	b	315)	d	316)	a
153)	b	154)	d	155)	d	156)	c	317)	c	318)	b	319)	b	320)	b
157)	a	158)	c	159)	c	160)	c	321)	a	322)	a	323)	b	324)	c
161)	a	162)	c	163)	d	164)	d	325)	c	326)	d	327)	b	328)	d

329)	d	330)	a	331)	d	332)	a
333)	c	334)	d	335)	b	336)	c
337)	d	338)	b	339)	b	340)	d
341)	b	342)	c	343)	d	344)	c
345)	b	346)	c	347)	a	348)	b
349)	d	350)	a	351)	b	352)	d
353)	b	354)	b	355)	c	356)	a
357)	b	358)	c	359)	a	360)	b
361)	a	362)	c	363)	a	364)	a
365)	b	366)	b	367)	d	368)	b
369)	c	370)	c	371)	a	372)	c
373)	b	374)	b	375)	b	376)	d
377)	b	378)	c	379)	d	380)	c
381)	d	382)	a	383)	c	384)	c
385)	c	386)	a	387)	b	388)	b
389)	b	390)	b	391)	d	392)	a
393)	d	394)	a	395)	b	396)	d
397)	c	398)	c	399)	a	400)	b
401)	d	402)	b	403)	c	404)	c
405)	b	406)	a	407)	a	408)	b
409)	b	410)	d	411)	c	412)	a
413)	b	414)	b	415)	b	416)	b
417)	a	418)	c	419)	c	420)	c
421)	c	422)	a	423)	c	424)	a
425)	a	426)	d	427)	d	428)	a
429)	c	430)	b	431)	b	432)	a
433)	a	434)	a	435)	d	436)	c
437)	d	438)	c	439)	a	440)	c
441)	c	442)	d	443)	a	444)	d
445)	c	446)	d	447)	c	448)	d
449)	c	450)	c	451)	d	452)	c
453)	d	454)	b	455)	d	456)	b
457)	b	458)	c	459)	c	460)	b
461)	d	462)	c	463)	a	464)	d
465)	b	466)	b	467)	c	468)	a
469)	a	470)	d	471)	a	472)	a
473)	a	474)	d	475)	a	476)	d
477)	d	478)	b	479)	a	480)	b
481)	c	482)	d	483)	a	484)	d
485)	c	486)	b	487)	c	488)	b
489)	a	490)	a	491)	d	492)	c
493)	c	494)	d	495)	d	496)	d
497)	c	498)	a	499)	b	500)	d
501)	a	502)	d	503)	d	504)	b
505)	b	506)	b	507)	c	508)	a
509)	a	510)	c	511)	a	512)	c
513)	b	514)	b	515)	a	516)	d
517)	c	518)	d	519)	a	520)	d
521)	d	522)	b	523)	b	524)	c
525)	c	526)	a	527)	d	528)	d
529)	a	530)	c	531)	b	532)	a

533)	b	534)	a	535)	b	536)	b
537)	d	538)	b	539)	a	540)	d
541)	c	542)	c	543)	d	544)	b
545)	d	546)	b	547)	a	548)	a
549)	b	550)	a	551)	c	552)	d
553)	d	554)	d	555)	a	556)	d
557)	b	558)	a	559)	b	560)	d
561)	b	562)	a	563)	a	564)	b
565)	b	566)	a	567)	b	568)	b
569)	c	570)	b	571)	b	572)	c
573)	b	574)	b	575)	d	576)	d
577)	a	578)	d	579)	b	580)	a
581)	d	582)	d	583)	b	584)	b
585)	d	586)	d	587)	b	588)	b
589)	a	590)	a	591)	c	592)	a
593)	c	594)	a	595)	b	596)	c
597)	a	598)	b	599)	c	600)	d
601)	c	602)	a	603)	c	604)	a
605)	a	606)	d	607)	a	608)	c
609)	b	610)	d	611)	c	612)	d
613)	c	614)	a	615)	d	616)	d
617)	c	618)	a	619)	c	620)	c
621)	a	622)	d	623)	a	624)	d
625)	b	626)	b	627)	c	628)	a
629)	d	630)	b	631)	c	632)	c
633)	d	634)	c	635)	d	636)	b
637)	d	638)	d	639)	d	640)	a
641)	b	642)	a	643)	d	644)	a
645)	a	646)	b	647)	c	648)	a
649)	b	650)	a	651)	b	652)	a
653)	d	654)	d	655)	a	656)	a
657)	b	658)	c	659)	b	660)	b
661)	b	662)	c	663)	c	664)	c
665)	b	666)	a	667)	b	668)	c
669)	d	670)	a	671)	c	672)	a
673)	b	674)	c	675)	c	676)	b
677)	b	678)	b	679)	b	680)	b
681)	a	682)	a	683)	b	684)	d
685)	a	686)	a	687)	c	688)	d
689)	d	690)	c	691)	b	692)	d
693)	a	694)	c	695)	c	696)	a
697)	d	698)	a	699)	d	700)	c
701)	b	702)	a	703)	b	704)	b
705)	b	706)	c	707)	a	708)	d
709)	b	710)	c	711)	b	712)	a
713)	d	714)	c	715)	b	716)	c
717)	b	718)	a	719)	a	720)	b
721)	b	722)	c	723)	d	724)	b
725)	b	726)	c	727)	a	728)	a
729)	b	730)	b	731)	a	732)	a
733)	a	734)	b	735)	a	736)	d

737)	a	738)	a	739)	a	740)	d	897)	a	898)	d	899)	c	900)	a
741)	c	742)	c	743)	b	744)	c	901)	c	902)	c	903)	d	904)	c
745)	c	746)	a	747)	b	748)	a	905)	c	906)	a	907)	c	908)	d
749)	c	750)	a	751)	d	752)	a	909)	c	910)	d	911)	d	912)	a
753)	b	754)	d	755)	c	756)	b	913)	d	914)	a	915)	b	916)	d
757)	a	758)	b	759)	c	760)	a	917)	d	918)	b	919)	b	920)	c
761)	a	762)	a	763)	c	764)	d	921)	b	922)	b	923)	c	924)	d
765)	b	766)	b	767)	c	768)	b	925)	c	926)	a	927)	a	928)	b
769)	b	770)	a	771)	c	772)	c	929)	a	930)	a	931)	b	932)	d
773)	b	774)	a	775)	b	776)	c	933)	a	934)	a	935)	b	936)	d
777)	d	778)	a	779)	c	780)	c	937)	a	938)	d	939)	a	940)	a
781)	b	782)	a	783)	a	784)	c	941)	a	942)	b	943)	b	944)	d
785)	c	786)	c	787)	d	788)	a	945)	b	946)	b	947)	b	948)	a
789)	b	790)	d	791)	a	792)	b	949)	c	950)	a	951)	c	952)	d
793)	c	794)	a	795)	a	796)	a	953)	c	954)	c	955)	a	956)	b
797)	d	798)	d	799)	a	800)	c	957)	b	958)	a	959)	c	960)	c
801)	c	802)	b	803)	d	804)	d	961)	b	962)	b	963)	d	964)	d
805)	c	806)	d	807)	c	808)	d	965)	d	966)	a	967)	b	968)	b
809)	c	810)	a	811)	c	812)	c	969)	d	970)	c	971)	c	972)	d
813)	d	814)	c	815)	a	816)	a	973)	b	974)	a	975)	a	976)	b
817)	d	818)	a	819)	d	820)	b	977)	a	978)	a	979)	c	980)	a
821)	b	822)	b	823)	d	824)	d	981)	b	982)	c	983)	c	984)	c
825)	b	826)	d	827)	c	828)	c	985)	a	986)	d	987)	b	988)	b
829)	b	830)	d	831)	c	832)	d	989)	d	990)	a	991)	b	992)	d
833)	b	834)	c	835)	b	836)	b	993)	a	994)	a	995)	b	996)	b
837)	d	838)	c	839)	c	840)	c	997)	c	998)	c	999)	b	1000)	c
841)	c	842)	c	843)	b	844)	a	1001)	a	1002)	d	1003)	b	1004)	a
845)	c	846)	c	847)	c	848)	a	1005)	c	1006)	d	1007)	a	1008)	b
849)	c	850)	a	851)	b	852)	b	1009)	a	1010)	c	1011)	c	1012)	d
853)	c	854)	c	855)	b	856)	d	1013)	b	1014)	c	1015)	b	1016)	a
857)	c	858)	a	859)	d	860)	b	1017)	b	1018)	c	1019)	d	1020)	c
861)	b	862)	d	863)	d	864)	c	1021)	a	1022)	c	1023)	c	1024)	c
865)	b	866)	d	867)	b	868)	a	1025)	a	1026)	d	1027)	b	1028)	a
869)	d	870)	d	871)	b	872)	c	1029)	b	1030)	b	1031)	b	1032)	b
873)	a	874)	b	875)	d	876)	a	1033)	c	1034)	a	1035)	d	1036)	a
877)	d	878)	c	879)	d	880)	c	1037)	a	1038)	b	1039)	b	1040)	d
881)	c	882)	c	883)	a	884)	d	1041)	a	1042)	a	1043)	d	1044)	a
885)	d	886)	c	887)	c	888)	a	1045)	a	1046)	c	1047)	c	1048)	a
889)	b	890)	b	891)	b	892)	c	1049)	b	1050)	a				
893)	d	894)	c	895)	c	896)	c								

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

: HINTS AND SOLUTIONS :

1 **(d)**

$C_6H_5—CH = CHCOOH$ is cinnamic acid.

2 **(a)**

Draw position and chain isomers.

3 **(b)**

Strain $= \frac{1}{2}$[Normal valence angel-valence angel]

$$= \frac{1}{2}\,[109°28' - 60°] = 24°44'.$$

6 **(d)**

2, 2-dimethyl butane is 6-carbon hydrocarbon (C_6H_{14})

CH_3
|
$CH_3 - C - CH_2 - CH_3$
|
CH_3

Rest all are the chain isomers of pentane (C_5H_{12}).

CH_3
|
$H_3C - C - CH_3$
|
CH_3

2,2-dimethyl propane
(*neo*-pentane)

CH_3
|
$H_3C - CH - CH_2 - CH_2$

2- methyl butane
(*iso* –pentane)

$CH_3CH_2CH_2CH_2CH_3$
n-pentane

7 **(b)**

Follow IUPAC rules.

8 **(c)**

Detection of sulphur in sodium extract is done by lead acetate and sodium nitroprusside

$Na_2S + (CH_3COO)_2Pb \rightarrow PbS + 2CH_3COONa$
lead acetate black ppt.

$Na_2S + Na_2[Fe(CN)_5NO] \rightarrow Na_4[Fe(CN)_5NOS]$
Sodium nitroprusside sodium thio nitroprusside
(purple colour)

9 **(b)**

–do –

10 **(a)**

Cyclopropyl methyl carbocations are more stable than benzyl carbocations due to conjugation between bent orbitals of cyclopropyl group.

11 **(c)**

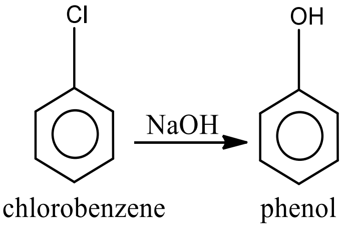
The shape of π electron cloud in acetylene in cylindrical

12 **(c)**

Acidified sodium fusion extract on addition of ferric chloride solution gives blood red colouration, which confirms the presence of N and S.

$3NaCNS + aq.\,FeCl_3 \longrightarrow Fe(CNS)_3 + 3NaCl$
red
(ferric thiocyanide)

13 **(b)**



In this process one group is replaced by other, hence, it is a substitution process and both the leaving and attacking groups are nucleophilic, therefore it is an example of nucleophilic substitution reaction.

14 **(b)**

(i) $Fe(CNS)_3$ is red in colour and is formed when both N and S are present in organic compound

(ii) $Na_2S + Na_2[Fe(CN)_5NO] \longrightarrow [NaFe(C)_5NOS]$
from organic compound sodium nitroprusside violet colour

15 **(a)**

Bond C – H C – C C –

N C – O

Bond energy (in kcal mol^{-1}) 97 83 73 86

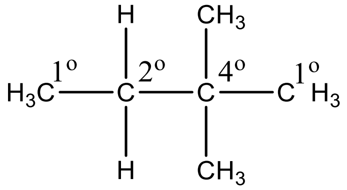
16 **(d)**

1° carbon is attached to one carbon atom.

2° carbon is attached to two carbon atms.

3° carbon is attached to three carbon atoms.

The hydrogen attached to 2° carbon atom are 2°.



∴ It has one 2° carbon atom and two 2°hydrogen atoms.

17 **(d)**

–do –

18 **(b)**

Non-staggered means eclipsed form.

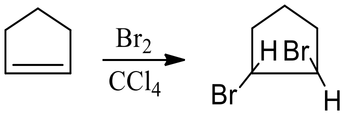
19 **(c)**

Carbanion is electron rich species. Stability of carbanion increases with increase in s –character of hybrid orbitals of carbon bearing the charge.

∴ $sp^3 < sp^2 < sp$

(25%s-character) (33%s-character) (50%s-character)

20 **(b)**



It is example of addition reaction

21 **(a)**

It is a fact.

22 **(c)**

It is the latest modified definition of organic chemistry.

23 **(b)**

In the lassaigne test, if organic compound consists of both N and S, then a red colour is obtained on adding aqueous $FeCl_3$ To sodium extract. $NH_2-\overset{\overset{S}{||}}{C}-NH_2$ contains both N and S hence, it will give red colour in Lassaigne test

$NaCNS + FeCl_3 \longrightarrow [Fe(SCN)]Cl_2 + NaCl$

Blood red colour

24 **(a)**

If nitrogen is present in organic compound then sodium extract contains NaCN.

$Na + C + N \xrightarrow{\text{Fuse}} NaCN$

$FeSO_4 + 6NaCN \rightarrow \underset{(A)}{Na_4[Fe(CN)_6]} + Na_2SO_4$

A changes to Prussian blue $Fe_4[Fe(CN)_6]_3$ on reaction with $FeCl_3$.

$4FeCl_3 + 3Na_4[Fe(CN)_6]$

$\rightarrow Fe_4[Fe(CN)_6]_3 + 12NaCl$

25 **(b)**

Follow IUPAC rules.

27 **(b)**

It is the definition of asymmetric synthesis.

28 **(a)**

(i) has sp^3; (ii) has sp^2, sp^3; (iii) has sp^2, sp; (iv)

31 **(d)**

The number of isomeric alkenes with molecular formula $C_6H_{12}$are 13.

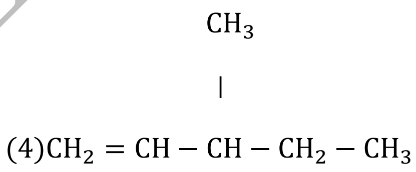
(1) $CH_2 = CH - CH_2 - CH_2 - CH_2 - CH_3$

(2) $CH_3 - CH = CH - CH_2 - CH_2 - CH_3$

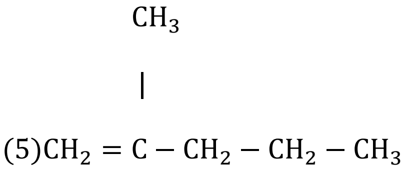
(cis and trans)

(3) $CH_3 - CH_2 - CH = CH - CH_2 - CH_3$

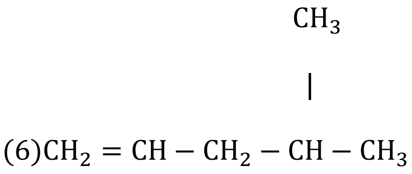
(cis and trans)



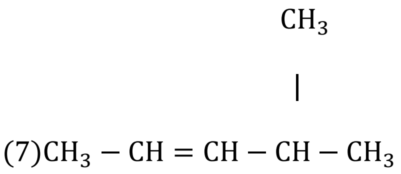
(4)$CH_2 = CH - \overset{\overset{CH_3}{|}}{CH} - CH_2 - CH_3$



(5)$CH_2 = \overset{\overset{CH_3}{|}}{C} - CH_2 - CH_2 - CH_3$

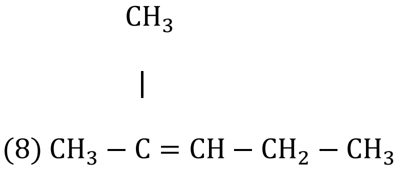


(6)$CH_2 = CH - CH_2 - \overset{\overset{CH_3}{|}}{CH} - CH_3$

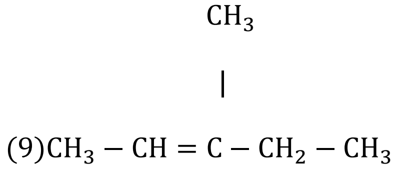


(7)$CH_3 - CH = CH - \overset{\overset{CH_3}{|}}{CH} - CH_3$

(cis and trans)



(8) $CH_3 - \overset{\overset{CH_3}{|}}{C} = CH - CH_2 - CH_3$



$$(9)\ CH_3 - CH = \overset{\displaystyle CH_3}{\overset{|}{C}} - CH_2 - CH_3$$

(*cis and trans*)

32 **(c)**

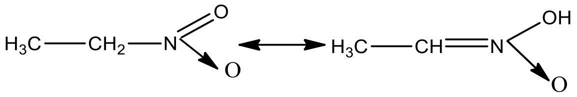
It is 3-methyl butan-2-ol.

33 **(c)**

This statement is not true now.

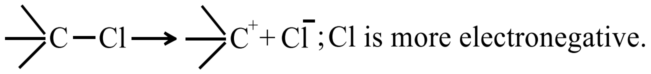
34 **(c)**

Nitroalkanes exhibit tautomerism. In it, α-H-atom is labile and form nitrolic acid.



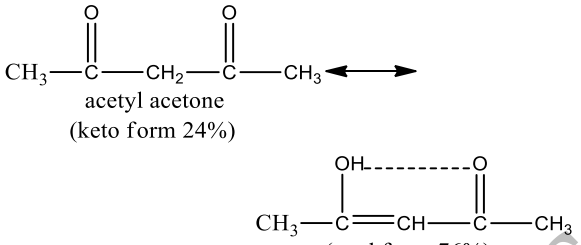
36 **(d)**

Heterolysis involves the bond fission in a manner when either of the two atoms involved in bond fission retains the shared pair of electron, producing +ve and –ve ions, *e.g.*



37 **(b)**

Keto and enol forms are inter convertable. The enol content will be maximum when enol form is stabilised by hydrogen bonding.



acetyl acetone (keto form 24%)

(enol form 76%)

In acetyl acetone, the enol form is stabilised by H-bonding, hence it has more enol content than other.

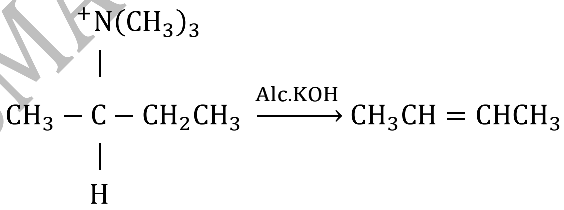
38 **(b)**

Compounds having bivalent functional group (like C=O, –O–, –S – etc) with atleast 4 carbon atoms (in case of ether and thioether) or atleast 5 carbon atoms (in case of ketones) exhibit metamerism. Hence, $C_2H_5 - S - C_2H_5$ will show metamerism.

39 **(a)**

Follow IUPAC rules.

40 **(c)**

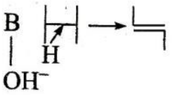


$$CH_3 - \underset{\underset{\displaystyle H}{|}}{\overset{\overset{\displaystyle {}^{+}N(CH_3)_3}{|}}{C}} - CH_2CH_3 \xrightarrow{\text{Alc.KOH}} CH_3CH = CHCH_3$$

is an example of elimination reaction

41 **(b)**

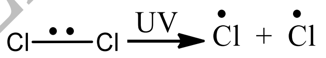
The dehydrohalogenation in presence of $OH^-$is correctly represented by



In this mechanism the base OH^- removes a proton from the β carbon.

43 **(a)**

Covalent bonds are cleaved in homolytic way in presence of UV light. It results in formation of free radical.



chlorine free radicals

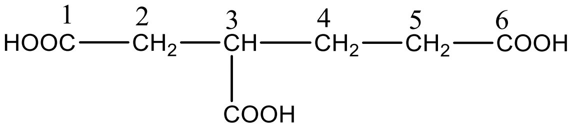
45 **(b)**

One asymmetric carbon atom is present.

46 **(b)**

Both alkene and cyclo alkane have general formula C_nH_{2n}.

47 **(b)**



IUPAC name of the above compound is 3-carboxyhexane-l, 6-dioic acid.

48 **(d)**

$Na_2S + Na_2[Fe(CN)_5NO] \longrightarrow Na_4[Fe(CN)_5NOS]$

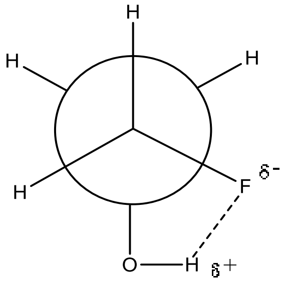
Sodium nitroprusside Solution sodium thio nitro prusside (purple colure)

49 **(d)**

Greater the difference in electronegativity of bonded atoms easier will be heterolytic cleavage

50 **(c)**

$HO - CH_2 - CH_2 - F$



Gauche conformation is comparatively more stable due to hydrogen linkage in between F and H (at O-atom), hence order is *Eclipse*, *Anti* (staggered), *Gauche*.

51 **(d)**

Phosphorous is estimated as $Mg_2P_2O_7$

$P \xrightarrow[\Delta]{HNO_3} H_3PO_4$

$H_3PO_4 + Mg^{2+} + NH_4OH \rightarrow MgNH_4PO_4$

$2MgNH_4PO_4 \rightarrow Mg_2P_2O_7 + H_2O + 2NH_3$

$\% \text{ of P} = \frac{62\times\text{wt.of } Mg_2P_2O_7\times 100}{222\times w}$

52 **(c)**

$n = 2$ and $a = 2^n = 2^2 = 4$.

54 **(c)**

Among carbonyl compounds, reactivity decrease with increase in alkyl groups as alkyl groups ($+I$ effects) decrease positive character on C-atom. Thus, the correct order of reactivity is

$HCHO > CH_3CHO > C_6H_5CHO$

55 **(d)**

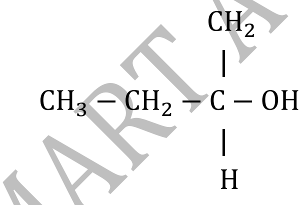
Primary and secondary alkyl halides gives S_N2 reaction

56 **(c)**

Follow IUPAC rules.

57 **(d)**

2-butanol is optically active as it contain as it contain chiral carbon atom.



$$CH_3 - CH_2 - \overset{\overset{\displaystyle CH_2}{|}}{\underset{\underset{\displaystyle H}{|}}{C}} - OH$$

58 **(d)**

$a = 2^n$; $n = 3$.

59 **(a)**

It is Markownikoff's rule.

61 **(b)**

The order of homolytic bond dissociation energies of CH_4, C_2H_6 and CH_3Br is as

$CH_4 > C_2H_6 > CH_3Br$

ΔH (kcal/mol) 105 100 70

63 **(d)**

Each d form has its l form and the pair is known as enantiomer.

64 **(b)**

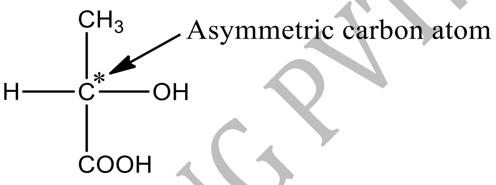
Carboxylic acid and esters show functional group isomerism. When two compounds have same molecular formula but different functional groups, then functional isomerism arises.

e.g.,

C_2H_5COOH and CH_3COOCH_3

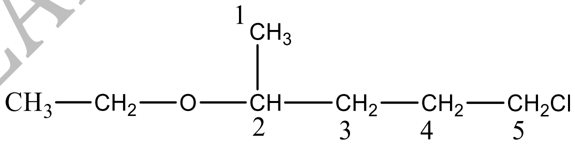
65 **(a)**

One asymmetric carbon atom is present in a lactic acid molecule. Hence, it is an optically active compound.



66 **(a)**

According to IUPAC system, ether are named as alkoxy alkanes. The larger alkyl group froms the parent chain while lower alkyl group is taken ethereal oxygen and forms a part of alkoxy group.



2-etoxy-5-chloropentane

67 **(d)**

Follow IUPAC rules.

68 **(c)**

It is a fact.

69 **(c)**

A liquid, which decomposes at its normal boiling point can be purified by vacuum distillation.

71 **(b)**

Stability of carbanion is not governed by hyperconjugation. Its stability depends on the $+I$ or $-I$ group

72 **(c)**

Let unreacted $0.1M(= 0.2N)H_2SO_4 = V'mL$

$\therefore$ 20 mL of 0.5M NaOH

=V'mL of 02 N H_2SO_4

$\therefore 20 \times 0.5 = V' \times 0.2$

$\therefore \quad V' = 50mL$

Used $H_2SO_4 = 100 - 50 = 50mL$

$$\%\text{Nitrogen} = \frac{1.4\text{ NV}}{w}$$

where, N=normality of H_2SO_4

V=volume of H_2SO_4 used

$$\therefore \%\text{ nitrogen} = \frac{1.4 \times 0.2 \times 50}{0.30}$$

$= 46.67\%$

% of nitrogen in

(a) $CH_3CONH_2 = \frac{14\times100}{59} = 23.73\%$

(b) $C_6H_5CONH_2 = \frac{14\times100}{122} = 11.48\%$

(c) $NH_2CONH_2 = \frac{28\times100}{60} = 46.67\%$

(d) $NH_2CSNH_2 = \frac{28\times100}{76} = 36.84\%$

Therefore, the organic compound is urea.

73 **(d)**

$CH_4 + Cl_2 \xrightarrow{hv} CH_3Cl$

This is an example of free radical substitution reaction

$$Cl{-}Cl \xrightarrow{hv} Cl^{\bullet} + Cl^{\bullet}$$
$$CH_4 + Cl^{\bullet} \longrightarrow CH_3^{\bullet} + HCl$$
$$CH_3^{\bullet} + Cl{-}Cl \longrightarrow CH_3Cl + Cl^{\bullet}$$
$$Cl^{\bullet} + Cl^{\bullet} \longrightarrow Cl_2$$

74 **(c)**

Glucose has four dissimilar asymmetric carbon atoms; $a = 2^4$.

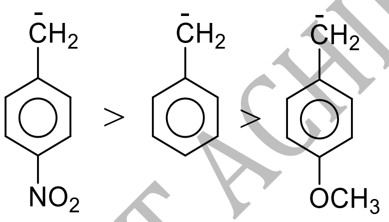
75 **(d)**

The elimination takes place according to Saytzeff rule. The most substituted alkene (butane-2) is called Saytzeff product whereas less substituted alkene (butane-1) is called Hofmann product

76 **(c)**

$-NO_2$ group shows $-M$ effect while CH_3O- group shows $+M$ effect. ($-M$effect stabilizes an anion)

Hence, the order of stability is



(II) (III) (I)

77 **(b)**

The former possesses 12α-H atom whereas, later possesses six α-H atom. More is the no. of α-H atom, more is delocalisation and more is stability.

78 **(d)**

$H_2O, ROH, R-COOH$ etc are protic solvents because they are polar in nature and contain a hydrogen directly bonded to oxygen while other are aprotic solvents as they do not have a hydrogen bonded directly to oxygen. They are especially favourable for S_N1 reactions. While aprotic solvents cannot have hydrogen bond to the nucleophile because they does not have hydrogen bonded to nitrogen or oxygen. They are favourable for S_N2 reactions.

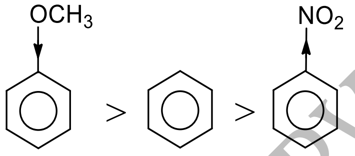
79 **(b)**

Follow mechanism of addition reactions.

80 **(c)**

Methoxy group is electron releasing group it increases electron density of benzene nucleus

$-NO_2$ group is electron withdrawing group, it decreases the electron density of benzene nucleus. Thus, the order of reaction with electrophilic regent is



81 **(c)**

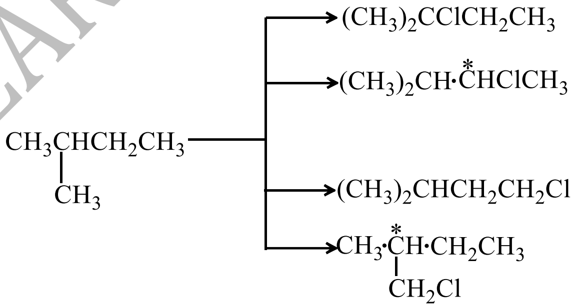
$NaCN + HNO_3 \overset{\Delta}{\rightarrow} HCN\uparrow + NaNO_2$

$Na_2S + HNO_3 \overset{\Delta}{\rightarrow} H_2S\uparrow + 2NaNO_3$

82 **(d)**

The central carbon in I and II is asymmetric.

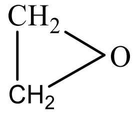
83 **(b)**



Thus, out of four isomers only two have chiral carbon. Each have two isomers.

85 **(a)**

The structural formula of epoxide is



It consists three membered ring with two carbon and one oxygen.

86 **(c)**

$CH_3-CH_2-CH_2-Cl + \text{alc.}KOH \longrightarrow CH_3-CH = CH_2$

It is an example of elimination reaction

87 **(d)**

In the reaction

$>C=O + H_2NOH \longrightarrow >C=NOH + H_2O$

Both addition and elimination takes place simultaneously. Thus, the reaction is addition elimination

88 **(d)**

Resolution of racemic mixture involves the formation of distereoisomers.

89 **(b)**

When N and S both are present in the organic compound, then a red colour complex ion of $[Fe(CNS)]^{2+}$ is formed on adding $FeCl_3$ to sodium extract

$NaCNS + FeCl_3 \rightarrow [Fe(SCN)]Cl_2 + NaCl$

blood red colour

90 **(a)**

$$CH_3CH_2Br \xrightarrow[LAH]{H-H} CH_3-CH_3+HBr$$

S_{N^2}

CH_3

|

$CH_3-C-Br \xrightarrow[LAH]{} CH_3-C=CH_2+HBr$

| (E_2) |

CH_3 CH_3

91 **(a)**

Count σ and π-bonds.

92 **(d)**

The formation of sodium thionitroprusside (blue) shows the presence of sulphur.

$Na_2S + Na_2[Fe(NO)(CN)_5] \longrightarrow Na_4[Fe(CN)_5NOS]$

Sodium sodium nitroprusside sodium thionitroprusside purple

extract

93 **(a)**

It is a fact.

94 **(b)**

$R-CH_2-CH_2X+KOH(alc.) \longrightarrow R-CH=CH_2+KX+H_2O$

Alkyl halid undergo β-elimination to form alkene.

95 **(d)**

The rate of reaction follows the order : $RI > RBr > RCl > RF$; whether it obeys S_N1 or S_N2 mechanism due to steric hindrance of alkyl group.

96 **(c)**

Only this is optically active due to central carbon being asymmetric.

97 **(a)**

Ethers show metamerism. Metamerism arises when a polyvalent functional group $\left(e.g., -O-, >C=O \text{ etc.}\right)$ is attached to different alkyl groups but the molecular formula remains same e.g.,

$C_2H_5-O-C_2H_5$ and $CH_3-O-C_3H_7$

98 **(c)**

Carbanion is electron rich species. Stability of carbanion increases with increase in s −character of hybrid orbitals of carbon bearing the charge.

$\therefore \quad sp^3 \quad < sp^2 \quad < sp$

(25%s-character) (33%s-character) (50%s-character)

99 **(c)**

The Z repells electrons and thus, electron density increases on R_3C part.

100 **(d)**

$\because$ −COOH group is a deactivating group

$\therefore$ Benzoic acid is less reactive towards electrophilic substitution.

So, benzoic acid> phenol>*n*-propyl benzene is not arranged correctly.

101 **(c)**

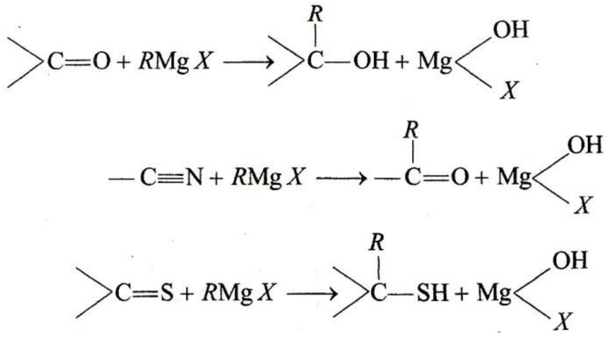
Aqueous solution of $NaHCO_3$ can be used to separate benzoic acid from its mixture with camphor. Benzoic acid form water soluble sodium benzoate with $NaHCO_3$.

102 **(d)**

Each π-electron is delocalised over six carbon atoms in ring.

103 **(d)**

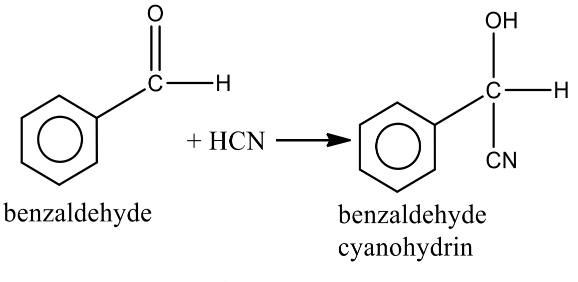
Grignard reagent reacts with >C=O, $-C\equiv N$, $>C=S$ as follows



104 **(a)**

Benzene is the most stable and we know that resonance energy is a direct measure of the stability of a molecule

105 **(d)**

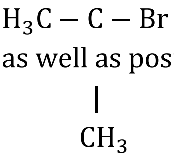


It is an example of nucleophilic addition reaction.

107 **(c)**

CH_3

|


$H_3C-\underset{\underset{CH_3}{|}}{C}-Br$ and $CH_3CH_2CH_2CH_2Br$ are chain as well as position isomers.

108 **(b)**

Detection of phosphorus in the organic compound can be done by its conversion into phosphate. The ammonium molybdate is used to identify phosphate ion

$Na_3PO_4 + 3HNO_3 \rightarrow H_3PO_4\ 3NaNO_3$

$H_3PO_4 + \underset{\text{ammonium modybdate}}{12(NH_4)_2MoO_4} + 21HNO_3 \xrightarrow{\Delta}$

$\underset{\text{yellow ppt.}}{(NH_4)_3PO_4.12MoO_3} + 21NH_4NO_3 + 12H_2O$

109 **(b)**

Homolytic bond fission is one in which each entity involved in bond formation retains its electron involved in shared pair of electron to form free radicals.

110 **(c)**

CH_3CN has sp^3 and sp-hybridised carbon atom.

111 **(c)**

CH_3OCH_3,	C_2H_5OH
Methoxy methane	Ethanol (ether)
(alcohol)	functional group
(C_2H_6O)	(C_2H_6O)molecular formula

In methoxy methane and ethanol both molecular formula is same but functional groups are different, so they are functional isomers.

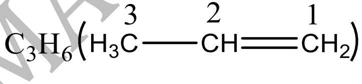
112 **(b)**

The main conditions for exhibiting geometrical isomerism are

(i) Presence of double bond.

(ii) Presence of different groups on same double bonded carbon.

(iii) Presence of at least one similar group on adjacent double bonded carbon atoms.


$C_3H_6\left(\overset{3}{H_3C}-\overset{2}{CH}=\overset{1}{CH_2}\right)$ does not exhibit geometric isomerism due to presence of same group on double bonded carbon atom (C_1).

113 **(a)**

Follow mechanism of addition reaction.

114 **(c)**

Vicinal or alkylene dihalides.

115 **(d)**

The electrophile involved in the sulphonation of benzene is SO_3.

$2H_2SO_4 \rightarrow SO_3 + H_3O^+ + HSO_4^-$

116 **(b)**

Number of meso structures in compound having odd number of chiral carbon atoms and symmetrical molecule $=2^{\left(\frac{n}{2}-\frac{1}{2}\right)}$

Given, $n = 5$

$\therefore$ Number of *meso* forms $=2^{\frac{5}{2}-\frac{1}{2}} = 2^2 = 4$

117 **(d)**

In Kjeldahl's method, the nitrogen is quantitatively converted into ammonia by heating with NaOH

$C + H + N \xrightarrow{\Delta} (NH_4)SO_4 + CO_2 + H_2O$

(from organic compound)

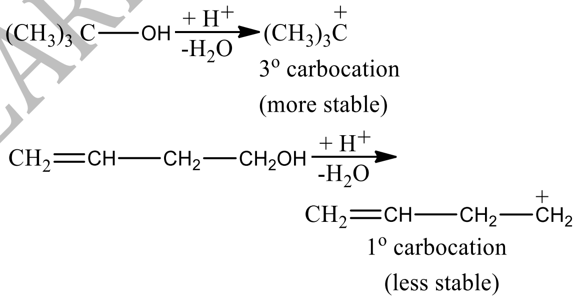
$(NH_4)_2SO_4 + 2NaOH \rightarrow Na_2SO_4 + 2NH_3 + 2H_2O$

In Dumas method nitrogen present in organic compound is quantitatively converted into N_2.

118 **(c)**

It is a fact.

119 **(b)**



Increasing order of stability of carbocation.

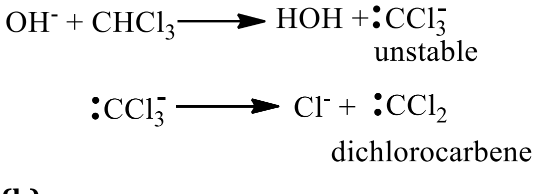
1°carbocation <2° carbocation<3°carbocation

120 **(c)**

Both geometrical and optical isomerism are included in stereoisomerism.

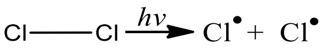
121 **(a)**

Phenol reacts with chloform and NaOH to give *o*-hydroxy benzaldehyde or salicylaldehyde. In this reaction dichlorocarbene $(:CCl_2)$ electrophile is generated. This reaction is called Reimer-Tiemann reaction.



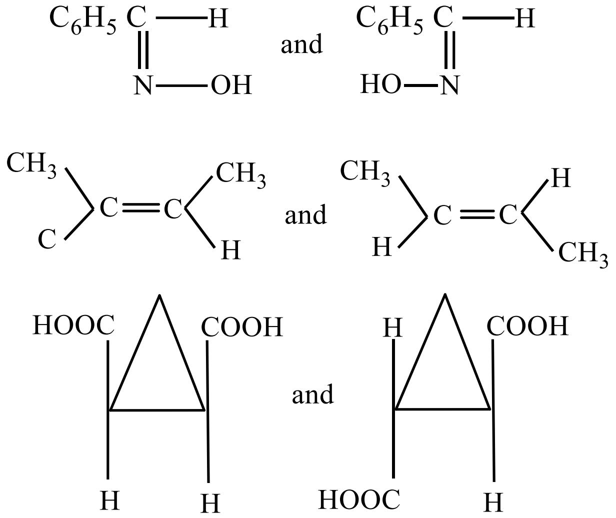
122 **(b)**

Homolytic fission is favoured by sunlight. In it, each bonded atom takes away its shared electrons and thus free radicals are produced.

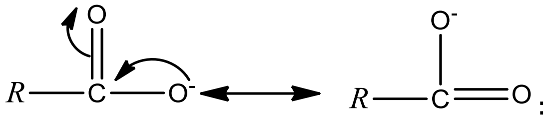


124 **(d)**

All show geometrical isomerism :



125 **(d)**



:In carboxylate ion, the negative charge is present on oxygen, a most electronegative element here, thus it is resonance stabilised.

$HC \equiv C^-$: Carbon is sp-hybridised so its electronegativity is increased higher relative to nitrogen.

$\bar{N}H_2$: Nitrogen is more electronegative than sp^3-hybridised C-atom. From the above discussion, it is clear that the order of the stability of conjugated bases is as

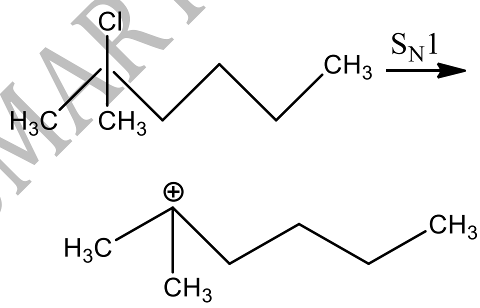
$$RCOO^- > HC \equiv C^- > \bar{N}H_2 > R^-$$

and higher is the stability of conjugated bases, lower will be basic character. Hence, the order of basic character is as

$$RCOO^- < HC \equiv C^- < \bar{N}H_2 < R^-$$

126 **(b)**

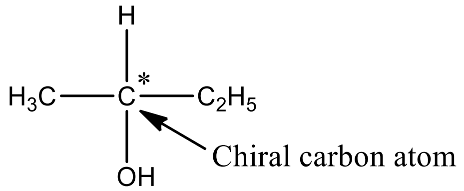
S_N1 Reaction is most favourable for tertiary substance.



3° carbocation
(most sable)

127 **(d)**

$CH_3CHOHC_2H_5$ is optically active because it has chiral C^*-atom



128 **(d)**

The closed ring cycloalkanes beyond five carbon atoms has puckered ring structure maintaining tetrahedral nature or stainless rings, *e.g.*, cyclo hexane has chair and boat form.

129 **(d)**

These are characteristics of S_N1 mechanism.

130 **(c)**

$-I$ power of groups in decreasing order with respect to the reference H

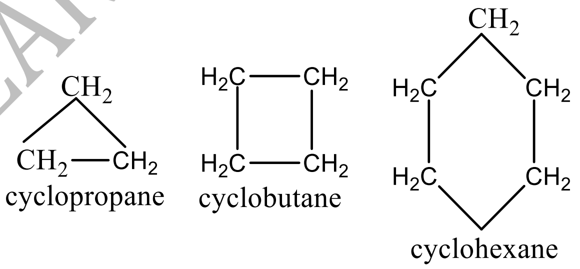
$$NO_2 > CHO > COOR > F > Cl > Br > I > OH > OR > NH_2$$

131 **(d)**

Asymmetry is present in all states.

132 **(b)**

The structure of cyclopropane, cyclobutane and cyclo hexane are as



Hence, the common group in cyclopropane, cyclobutane and cyclohexane is $>CH_2$ group.

133 **(d)**

Draw all possible isomers.

135 **(a)**

+I effect is shown by $-CH_3$ while $-$I effect is shown by $-Br$, $-Cl$ and $-NO_2$.

136 **(b)**

It is a strong reducing agent.

137 **(a)**

% of C= $\frac{12}{44}\times\frac{0.535}{0.765}\times 100 = 19.07$

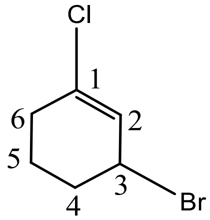
%of H $= \frac{2}{18}\times\frac{0.138}{0.765}\times 100 = 2.004$

Ratio of % of C:H=19:2 (approx.)

138 **(d)**

Benzene has planar structure.

139 **(c)**



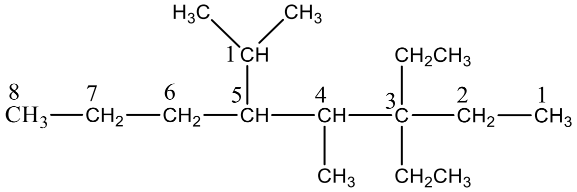
Unsaturation (double bond) is given priority over halogen, then lowest set of locants. So, the correct IUPAC name is 3-bromo-1-chlorocyclohexene.

140 **(c)**

Follow conformation; The conformers for n-butane are two gauche, two eclipsed and one anti.

142 **(d)**

Select longest possible carbon atom chain, number it and name compound according to IUPAC, rules.



3, 3-diethyl-4-methyl-5-(1'-methyl ethyl)-octane

143 **(b)**

Benzal is C_6H_5CH group.

144 **(b)**

A 2° carbon is one of which two valencies are attached to carbon atom.

145 **(c)**

Eclipsed conformation of butane contain angle and steric strain both. Follow conformation.

146 **(c)**

Weight of organic compound =29.5mg

$$NH_3 + HCl \longrightarrow Na_4Cl$$

$$HCl \quad + \quad NaOH \longrightarrow NaCl + H_2O$$

(remaining) $15 \times 0.1\ M$

=1.5 mmol

Total millimole of HCl=2

Millimole used by $NH_3 = 2 - 1.5 = 0.5$

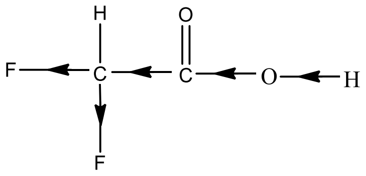
Weight of NH_3=0.5 ×17mg=8.5mg

Weight of nitrogen $= \frac{14}{17} \times 8.5mg = 7mg$

$$\%\ \text{Nitrogen} = \frac{7}{29.5} \times 1100 = 23.7\%$$

147 **(a)**

Fluoro group causes negative inductive effect increasing ionisation, thus 0.1M difluoroacetic acid has highest electrical conductivity.

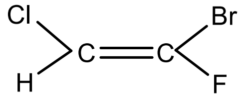


148 **(c)**

$-NO_2$ group shows – M effect white CH_3O –group shows $+M$ effect (– M effect stabilises an anion)

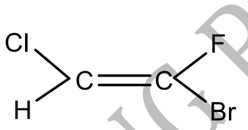
149 **(d)**

When the groups with higher priority ($i.e.$, with high atomic number) are present on same side of double bond, then the configuration is Z but when present on opposite side of double bond, the configuration is E.

(i) 

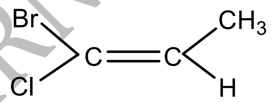
(Z)

(Priority : Cl > H and Br > F)

(ii) 

(E)

(Priority : Cl > H and Br > F)

(iii) 

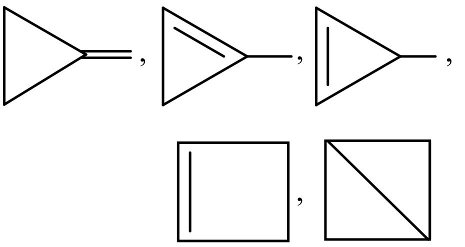
(Z)

150 **(c)**

Nucleophiles are electron rich species and can donate lone pair of electron to carbocation or any +ve centre.

152 **(c)**

C_4H_6 can have five cyclic isomers.



153 **(b)**

–do –

154 **(d)**

Urea shows tautomerism as

$$\underset{\text{keto form}}{NH_2-\overset{\overset{O}{\|}}{C}-NH_2} \leftrightharpoons \underset{\text{enolic form}}{NH=\overset{\overset{OH}{\|}}{C}-NH_2}$$

155 **(d)**

The conditions given are for C_6H_6.

156 **(c)**

Count π-bonds. Delocalisation is not possible.

157 **(a)**

Pentene-2 exhibits *cis* and *trans*-isomerism.

159 **(c)**

The rule is valid for unsymmetrical alkene.

160 **(c)**

$(CH_3)_2CHCH_2CH_3$ is isopentane.

161 **(a)**

π-bond in molecule give rise to hindered rotation.

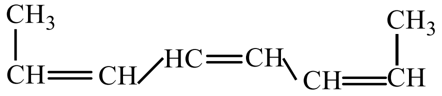
162 **(c)**

$CH_2 = CH_2 \xrightarrow{-H_2} CH \equiv CH$

Conversion of ethylene into acetylene is a example of elimination reaction.

163 **(d)**

The two propenyl group attached to 1,2-position of carbon in *cis*-form.



164 **(d)**

Markownikoff's rule is obeyed during addition of unsymmetrical addendum on unsymmetrical alkene.

165 **(a)**

Benzene has 6 C—C and 6 C—H σ-bonds and 3 C = C π-bonds.

166 **(b)**

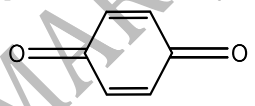
It is a fact.

167 **(c)**

Alkynes are linear due to *sp*-*sp* hybridized carbon.

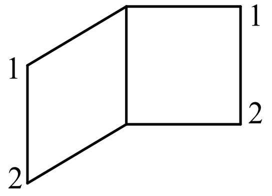
168 **(a)**

For keto-enol isomerism a compound should have at least one α-hydrogen atom with respect to ketone group or in other words for tautomerism presence of α −hydrogen atom is essential.



does not exhibited. tautomerism due to absence of α-hydroegn atom.

170 **(d)**



Hence, correct IUPAC name is bicyclo [2,2,0] hexane.

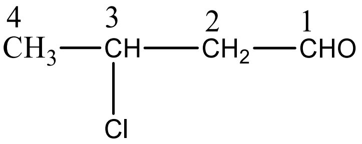
172 **(d)**

Chlorinolysis involves substitution reactions by chlorine.

173 **(b)**

$a = 2^{n-1}$; where n is no. of asymmetric carbon; when molecule possesses symmetry.

174 **(c)**



$\overset{4}{C}H_3-\underset{Cl}{\overset{3}{C}H}-\overset{2}{C}H_2-\overset{1}{C}HO$

3-chlorobutanal

∵ The order is priority is $-CHO > -Cl$.

175 **(c)**

$$\% \text{ C} = \frac{12}{44} \times \frac{12.517}{4.0} \times 100 = 85.7$$

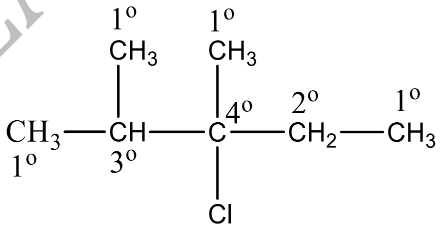
$$\% \text{ H} = \frac{2}{18} \times \frac{5.143}{4.0} \times 100 = 14.3$$

The mole ratio of C to H is $\frac{85.7}{12} : \frac{14.3}{1}$

$= 7.14 : 14.3 = 1 : 2 = CH_2$

176 **(b)**

3-chloro-2, 3-dimethyl pentane contains all the four 1°,2°, 3° and 4° carbon atoms.

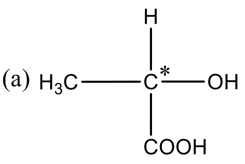


$CH_3-\underset{CH_3}{CH}-\underset{Cl}{\overset{CH_3}{C}}-CH_2-CH_3$

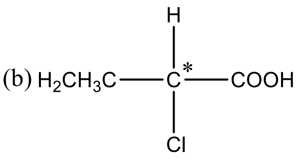
3-chloro-2-3-dimethyl pentane

178 **(d)**

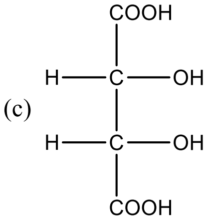
The carbon, four valencies of which are satisfied by four different groups, is termed as chiral carbon atom. The structures of the given compounds are as



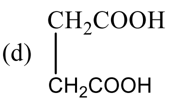
lactic acid



2-chlorobutanoic acid



tartaric acid



succinic acid

(where, $\overset{*}{C}$ = =chiral carbon atom)

Hence, succinic acid does not contain any chiral carbon atom.

180 **(a)**

In homolysis, the covalent bond is broken in such a way that each resulting species known as free radical.

$$CH_3CH_2-Cl \xrightarrow[\text{fission}]{\text{Homolytic}} CH_3\dot{C}H_2 + \dot{C}l$$

181 **(b)**

p-orbitals are at 90°to each other.

182 **(b)**

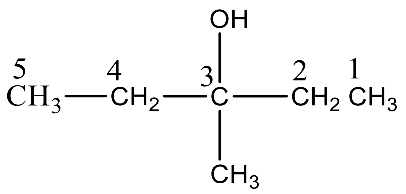
S_N2 reactions are greatly controlled by steric factor.

$R-CH_2-X$ R_2CH-X R_3C-X

1° 2° 3°

S_N2 reactivity decreases as bulkyness of alkyl group increases.

183 **(a)**



3-methyl pentane 3-ol

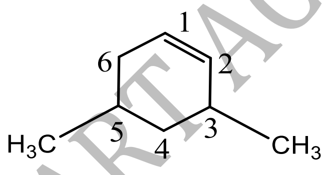
Hydroxy is used when – OH group is written in prefix. So, choice (b) and (c) are wrong.

184 **(d)**

Due to $-R$ effect of – CHO group, oxygen carries – δ charge while the terminal carbon carries + δ, ie,

$$\overset{+\delta}{CH}=CH-CH=\overset{-\delta}{O}$$

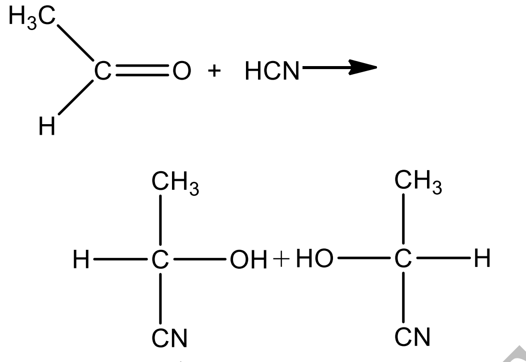
185 **(a)**



Its IUPAC name is 3,5-dimethylcyclohexene.

186 **(a)**

Lactic acid obtained in the given reaction is an optically active compound due to the presence of chiral C-atom. It exists as d and l forms whose ratio 1:1.



50% L-isomer 50% D-isomer

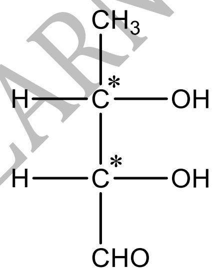
187 **(c)**

The mirror-image isomerism is a class of stereoisomerism and are included in optical isomerism.

188 **(c)**

CH_3^+ acts as an intermediate in the given reaction (Friedel Craft's alkylation). It is an example of electrophilic aromatic substitution. In this reaction CH_3^+ is electrophile.

189 **(d)**



Possible number of optical isomers = 2^n

$= 2^2 = 4$

190 **(b)**

C_6H_6 has 12σ and 3π-bonds.

191 **(c)**

For S_N2 mechanism, there should be least steric hinderance.

193 **(c)**

Rest all are polymerisation.

194 **(a)**

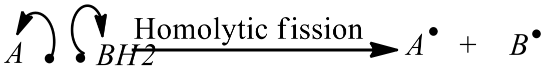
The increasing order of inductive effect is: $-F < -COOH < -CN < -NO_2$.

197 **(b)**

$-COOH$ is on top in preference table.

200 **(d)**

In homolytic cleavage, covalent bond is cleaved in such a way that each atom takes its shared electrons with itself and free radicals are formed.



free radicals

201 **(d)**

Fractional distillation is used for the separation of crude petroleum into various fractions like coal-tar, crude alcohol and petroleum

202 **(b)**

Free radical has unpaired electron.

203 **(a)**

SO_3 can accept lone pair of electron in d-subshell.

204 **(a)**

From Kjeldahl's method,

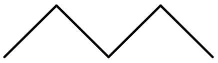
Percentage of nitrogen

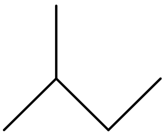
$$=\frac{1.4\times N\times V}{W}=\frac{1.4\times 0.1\times 30}{5}$$

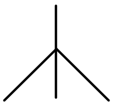
$= 0.84\%$

205 **(a)**

The isomers alkanes having the molecular formula C_5H_{12} are as

(i) 
n-pentane

(ii) 
iso-pentane
(2-methyl butane)

(iii) 
neo-pentane
(2,2-dimethyl propane)

206 **(b)**

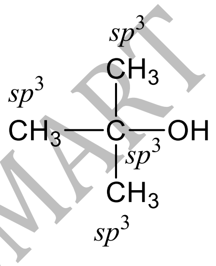
Just after few years when Wöhler prepared urea from KCNO and $(NH_4)_2SO_4$, Kolbe prepared acetic acid in laboratory from its element and gave final blow to Vital force theory.

207 **(b)**

Due to the presence of asymmetric carbon atom , $e.$g.,

$CH_3CH_2\cdot\overset{\otimes}{C}HOHCH_3.$

208 **(a)**



All bonds are σ- bonds hence C uses only its sp^3-hybrid orbitals. In all other compounds there is one $C = O$ double bond, therefore, this carbon is sp^2- hybridized

210 **(a)**

Follow Baeyer's strain theory for stability of cycloalkane.

211 **(a)**

— COOH is electron withdrawing group.

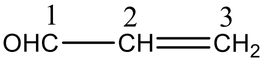
213 **(d)**

CHO
| is called glyoxal
CHO

214 **(d)**

IUPAC name of acraldehyde

$(CH_2 = CH.CHO)$ is:

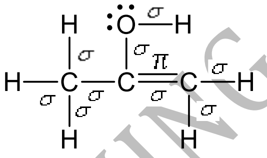


prop-2-ene-1-al

216 **(d)**

These are characteristics known from mechanism of reaction.

217 **(b)**



enolic form of acetone

9σ bonds, 1πbond, 2 lone pairs

218 **(c)**

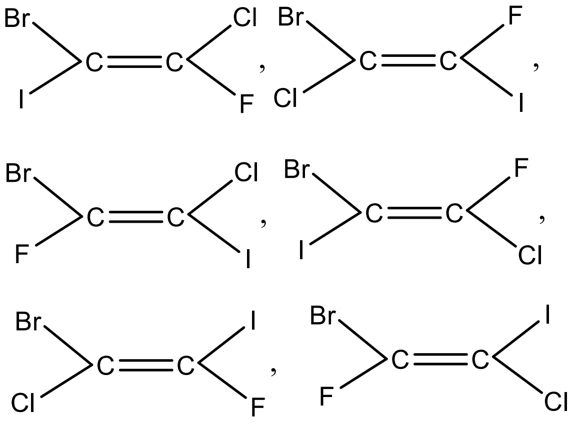
$BrCH_2CH_2COOH$ is the weakest acid and have lowest dissociation constant because. I.E. of Br is lesser than F and is far away from – COOH group.

219 **(a)**

Follow IUPAC rules.

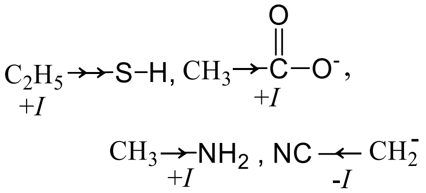
220 **(d)**

Molecular formula C_2BrClFl six isomers are possible.



221 **(a)**

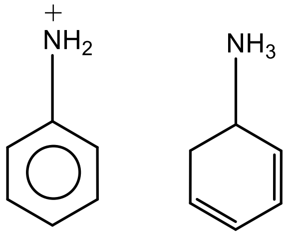
Nucleophilies are those substances which can donate a pair of electrons. They can be neutral or negatively charged. The nucleophilic power depends on the tendency of species to donate the electrons. Due to the presence of $+I$ effect it increases. Hence, higher the $+I$ effect, higher the nucleophilic power. The $+I$ effect of ethyl is greater than $+I$ effect of methyl group



$C_2H_5 \rightarrow\rightarrow S{-}H,\ CH_3 \rightarrow \overset{O}{\overset{\|}{C}}{-}O^-,\ CH_3 \rightarrow NH_2,\ NC \leftarrow CH_2^-$

223 **(b)**

Among the given species $\overset{+}{C_6H_5NH_3}$ does not exert a resonance effect.



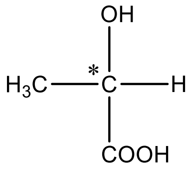
(I) (II) Structure II is not possible because in it nitrogen contains 10 valence electrons.

224 **(c)**

Draw all structures.

225 **(a)**

Chiral carbon atom has all four different groups attached to it.



lactic acid

∴ It has one asymmetric or chiral carbon atom.

226 **(b)**

Different spatial arrangement of atoms leads to its configuration.

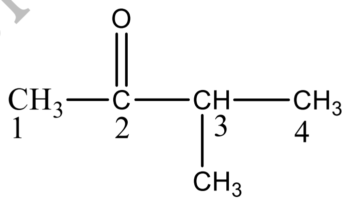
227 **(d)**

Butane and isobutane and all higher alkanes show isomerism.

228 **(a)**

Desmo (bond), tropism (turn). Thus, desmotropism, *i.e.*, isomerism arised due to turning of bond was the name given to tautomerism.

229 **(d)**



3-methyl butan-2-one
or 3-methyl 2-butanone

O

$\|$

Keto $(-C-)$ functional group is given priority.

230 **(b)**

Halogens can be identified in organic compounds by Beilstein-test.

231 **(a)**

It undergoes dehydration easily as the product obtained is conjugated and more stable.

232 **(c)**

Methoxy ethene is $CH_3O—CH = CH_2$; an unsaturated ether.

233 **(a)**

Electrophiles are electron deficient species which can share lone pair of electron with carboanion and are thus, called Lewis acids.

234 **(c)**

Iso −propyl chloride is a 2° halide and 2° halides can undergo hydrolysis either by S_N1 or S_N2 mechanism depending upon the nature of solvent used.

235 **(b)**

Follow IUPAC rules.

236 **(c)**

Carbon and hydrogen are estimated in organic compounds by Liebig's method

$C + 2CuO \xrightarrow{\Delta} 2Cu + CO_2$

$2H + CuO \xrightarrow{\Delta} Cu + H_2O$

Percentage of carbon and hydrogen is calculated from the weight of CO_2 and H_2 produced

237 **(c)**

μ is more for (c) then (d).

238 **(b)**

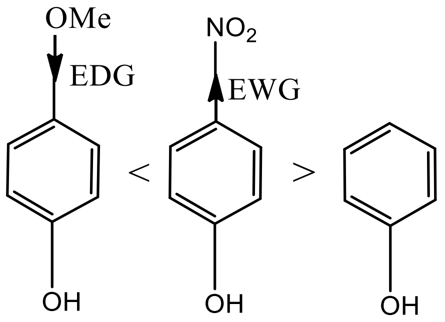
Free radical state is a transient state and thus, has short life.

240 **(c)**

When nitro group is present in the benzene nucleus, it withdraws electrons from *o* and *p*-positions. Thus, the electron density at the *o* and *p* −positions decreases. *m*-positions become positions of comparatively higher electron density and therefore, electrophilic attack occurs at *m*-positions.

241 **(c)**

According to Lewis, electron acceptor compounds are called acids. Therefore, compounds having tendency to accept electrons will be more acidic. The correct order of acidic character is as follows:



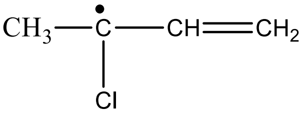
243 **(b)**

C_4H_7Cl is a monochloro derivative of C_4H_8 which itself exists in three isomeric forms.

(i) $CH_3-CH_2-CH=CH_2$: Its possible monochloro derivatives are

$CH_3-CH_2-CH=CH-Cl$

2 isomers *cis* and *trans* forms



$CH_3-\dot{C}(Cl)-CH=CH_2$

Optically active (exists in two forms)

$ClCH_2-CH_2-CH=CH_2$(one form)

Cl

|

$H_3C-CH_2-C=CH_2$ (one form)

(ii) $CH_3-CH=CH-CH_3$: Its possible monochloro derivatives are

$CH_3-CH=C-CH_3$

|

Cl

Exists in two geometrical forms

$CH_3-CH=CH-CH_2Cl$

Exists in two geometrical forms

(iii)$CH_3-C=CH_2$: Its possible monochloro derivatives are

|

CH_3

$CH_3-C=CH-Cl$, (Only one form)

|

CH_3

$ClCH_2-C=CH_2$(only one form)

|

CH_3

Thus, the total acylic isomeric forms of C_4H_7Cl are 12.

244 **(d)**

It shows intramolecular H-bonding.

245 **(d)**

Diethyl ether is resistant to nucleophilic attack by hydroxyl ion.

246 **(d)**

C_6H_6 has more canonical forms.

247 **(d)**

Chemical methods are based upon the distinguishing chemical properties of one class of organic compounds from the other. for example camphor and benzoic acid

248 **(c)**

Fractional crystalliation is used to purify organic solids which dissolve in a particular solvent. But their rate of solubility is different

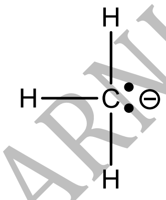
249 **(d)**

NH_4CNO is functional isomer of urea.

250 **(b)**

Carbanion $\overset{\ominus}{(CH_3)}$

Here, the carbon atom carries a negative charge with lone pair of electrons and has eight electrons in outermost orbit and complete its octet.



Reactions in which carbanions are formed as intermediate are said to proceed by a "Carbanion mechanism".

Carbanion is sp^3 hybridised, three sp^3 hybrid orbitals form covalent bonds with three atoms while the fourth sp^3 hybrid orbital has a non-bonding pair of electrons. It is pyramidal in shape as similar to NH_3.

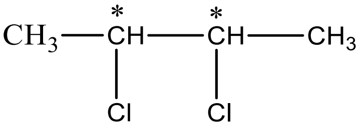
251 **(a)**

Ether group(—O—)has propyl and isopropyl group on its two sides.

252 **(d)**

Follow carbenes.

253 **(c)**



There are two chiral C-atoms (*)

Thus, optical isomerism is possible.

254 **(b)**

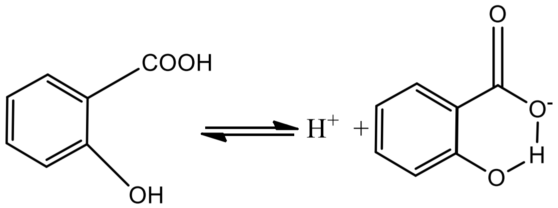
Follow IUPAC rules.

255 **(d)**

−do −

256 **(c)**

A monosubstituted benzoic acid is stronger than a monosubstituted phenol as former being a carboxylic acid. Among the given substituted benzoic acid, *ortho* − hydroxy acid is strongest acid although – OH causes electron donation by resonance effect which tends to decreases acid strength. It is due to very high stabilisation of conjugate base by intramolecular H-bond which outweight the electron donating resonance effect of – OH.



The overall order of acid-strength of given four acids is *ortho*-hydroxybenzoic acid ($pK_a =$ 2.98) > Toluic acid $pk_a = 4.37$) > p −hydroxybenzoic acid ($pk_a = 4.58$) > p −nitrophenol ($pk_a = 7.15$).

257 **(b)**

Electrophiles are electron pair acceptors.

258 **(b)**

$CH_3CH_2CHClCH_2CH_3$; $CH_3CHClCH_2CH_2CH_3$; $CH_2ClCH_2CH_2CH_2CH_3$.

259 **(b)**

Benzene and all its derivatives along with heteroaromatics possess aromatic nature. Rest all possess aliphatic nature.

260 **(d)**

Neohexyl chloride is a primary halide as in it, Cl-atom is attached to a primary carbon.

$$CH_3-\overset{\overset{\large CH_3}{|}}{\underset{\underset{\large CH_3}{|}}{C}}-CH_2-CH_2Cl$$

261 **(a)**

Ethane, ethene and ethyne have sp^3, sp^2 and sp-hybridization respectively.

262 **(d)**

Carbon is asymmetric as all its valencies are attached to different groups.

264 **(b)**

Follow IUPAC rules.

265 **(d)**

Bromination of alkanes in the presence of sunlight involves the formation of free radical, e.g.,

$$CH_4 \xrightarrow[hv]{Br_2} CH_3Br$$

Mechanism

Initiation

$$Br—Br \xrightarrow{hv} \dot{Br} + \dot{Br}$$

Propagation

$$CH_4 + \dot{Br} \longrightarrow \dot{CH_3} + HBr$$

$$\dot{CH_3} + Br—Br \longrightarrow CH_3Br + \dot{Br}$$

Termination

$$\dot{Br} + \dot{Br} \longrightarrow Br_2$$

266 **(d)**

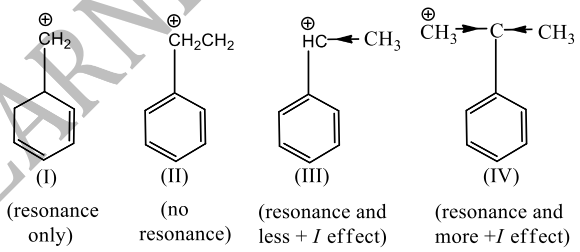
The two butene give different products on addition of Br_2-*cis* butene gives racemic mixture whereas *trans* butene gives *meso* form of 2,3,3-dibromo butene.

267 **(a)**

Lassaigne's test is used for the detection of halogens, nitrogen and sulphur.

268 **(a)**

Resonance and inductive effect decide stability of carbocations.



∴Correct order of stability is

II<I<III<IV

269 **(c)**

Atom	At mass (a)	% (b)	$\frac{b}{a}$	Ratio
X	10	50	$\frac{50}{10} = 5$	2
Y	20	50	$\frac{50}{20} = 2.5$	1

Hence, empirical formula = X_2Y

270 **(c)**

The compounds which differ in the nature of carbon chain are called chain isomers, e.g.,

$$CH_3-CH_2-CH_2-CH_2-CH_3$$

n-pentane

$$CH_3-\overset{\overset{\large CH_3}{|}}{CH}-CH_2-CH_3$$

iso-pentane

$$H_3C-\overset{\overset{\large CH_3}{|}}{\underset{\underset{\large CH_3}{|}}{C}}-CH_3$$

neo-pentane

272 **(c)**

In conjugated diene alternate single and double bonds are present while in cumulative diene, double bonds are present at adjacent positions.

I. $\overset{1}{H_2C}=\overset{2}{C}H\overset{3}{C}H_2\,\overset{4}{C}H_2\,\overset{5}{C}H=\overset{6}{C}H_2$

hexa-1, 5-diene

⇒ It is an isolated diene.

II. $\overset{1}{H_2C}=\overset{2}{C}=\overset{3}{C}H_2$

prop-1, 2-diene

⇒ It is a cumulative diene.

III. $\overset{1}{H_2C}=\overset{2}{C}H-\overset{3}{C}H=\overset{4}{C}H\overset{5}{C}H_2\,\overset{6}{C}H_3$

hexa-1, 3-diene

⇒ It is a conjugated diene.

IV. $\overset{1}{H_2C}=\overset{2}{C}H-\overset{3}{C}H=\overset{4}{C}H_2$

buta-1, 3-diene

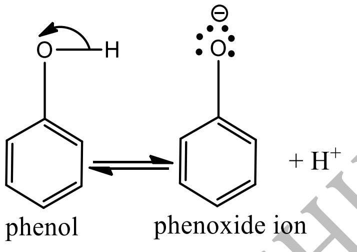
⇒ It is a conjugated diene.

V. Prop -1, 2-diene is a cumulative diene.

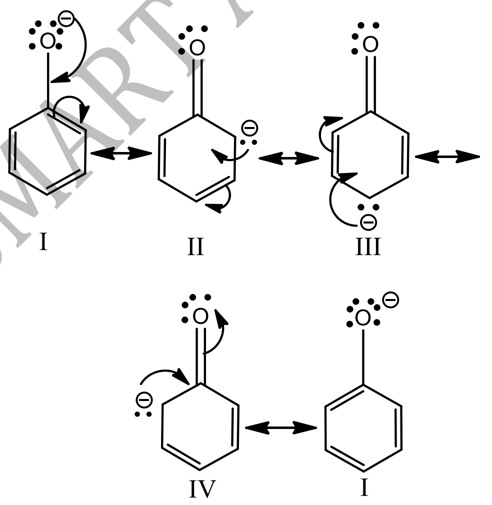
Hence, statement 3 and 5 are correct.

276 **(b)**

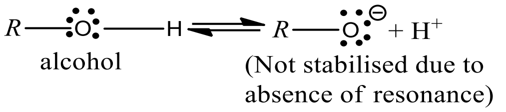
Phenols are much more acidic than alcohol due to the stabilisation of phenoxide ion resonance.



Phenoxide ion is stabilised due to following resonating structures.



While, in alcohols



(Not stabilised due to absence of resonance)

ortho nitrophenol is most acidic because in – NO_2 electron attracting group is attached to *ortho* position which helps in stabilising the negative charge on the oxygen of phenoxide ion. Hence, due to this reason acidic character of phenol is increased, while on attachment of – CH_3 group (electron donating group) acidic strength of phenol is decreased in cresol due to the destabilisation of phenoxide ion.

277 **(d)**

$C:H:O = 6:1:8$

$= \frac{6}{15}\times 100 : \frac{1}{15}\times 100 : \frac{8}{15}\times 100$

$40:6.67:53.3$

$= \frac{40}{12} : \frac{6.67}{1} : \frac{53.3}{16}$

$= 1:2:1$ ie, CH_2O

278 **(c)**

Alkanes never show functional isomerism, metamerism, tautomerism and geometrical isomerism.

279 **(b)**

Draw all possible chain and position isomers.

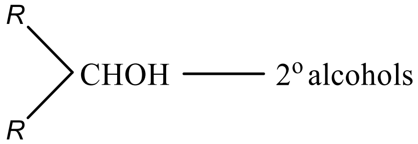
280 **(c)**

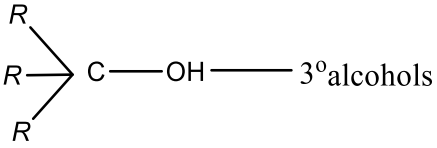
2-butanol has following structure

$$\begin{array}{c} \quad OH \\ \quad | \\ CH_3-CH-CH_2-CH_3 \end{array}$$

2° carbon

$R-CH_2OH-1°$ alcohol





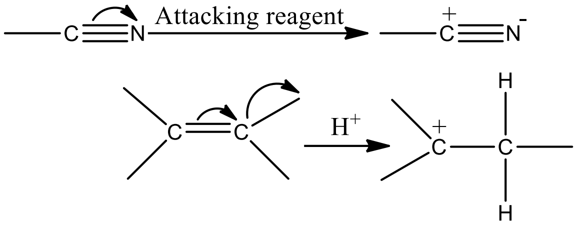
281 **(a)**

Stronger is acid, weaker is its conjugate base or weaker is nucleophilicity.

The acidic order $HF > H_2O > NH_3 > CH_4$.

282 **(b)**

Electromeric effect occurs only in the presence of attacking reagent. It operates in the molecules having multiple bonds. Since, it exists only on the demand of attacking reagent, it is a temporary effect. *e.g.,*



283 **(b)**

Leaving group ability depends upon basicity of group.

284 **(d)**

Read optical activity.

285 **(c)**

The compounds must fulfill two conditions to show geometrical isomerism.

(i) The compound should have at least one C=C.

(ii) The two groups attached to same carbon must be different.

Out of given choices only (c) fulfill both conditions and shows geometrical isomerism.

$H - C - COOH$

$\|$

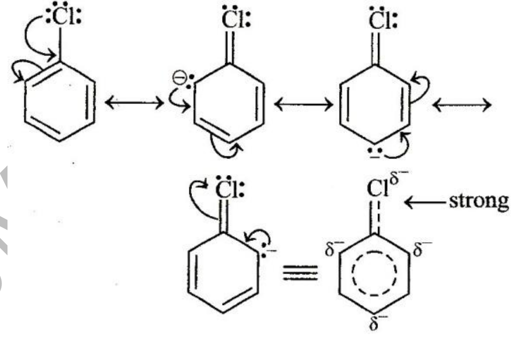
$H - C - COOH$

(c)

2-butene-1,4-dioic acid

286 **(a)**

$C - Cl$ bond is aryl chloride is stable due to delocalisation of electron by resonance. Also $C - Cl$ bond possess a double bond character like vinyl chloride, hence S_N reactions are not possible in chlorobenzene under ordinary conditions.



287 **(b)**

$C - C, C = C$ and $C \equiv C$ bond length are 1.54 Å, 1.34Å and 1.20Å respectively. In benzene $C = C$ is 1.40Å.

288 **(c)**

Definition of tautomerism.

289 **(c)**

The reactivity order is $3°H > 2°H > 1°H$.

290 **(a)**

The +ve inductive effect of CH_3 group on carbanions intensifies negative charge on C^- centre and thus, 3° carbanion is more reactive.

291 **(d)**

Glycerine contains

$^{\alpha}CH_2 - OH \leftarrow 1°$ alcohol

|

$^{\beta}CH - OH \leftarrow 2°$alcohol

|

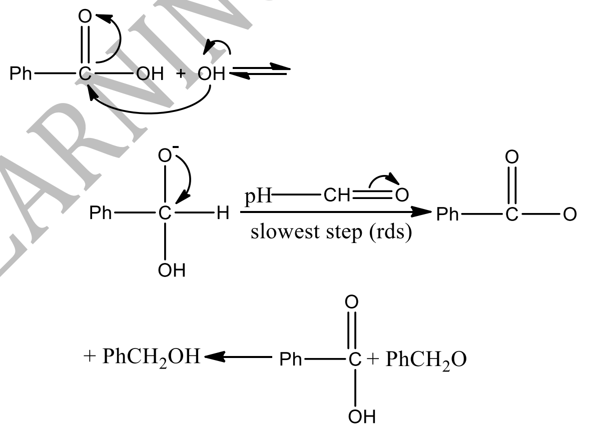
$^{\alpha}CH_2 - OH \leftarrow 1°$ alcohol

292 **(a)**

In rest all carbon chain is same.

293 **(b)**

In Cannizzaro reaction the transfer ofH^- to another carbonyl group is difficult and slowest step. (rate determining step or key step)



294 **(b)**

Anti conformation is the most stable form of n-butane (Bulky groups far apart).

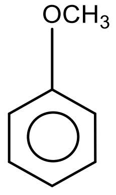
295 **(a)**

$R - X \longrightarrow R^+ + X^- \xrightarrow{OH^-} R - OH.$

298 **(a)**

It has altogether different groups.

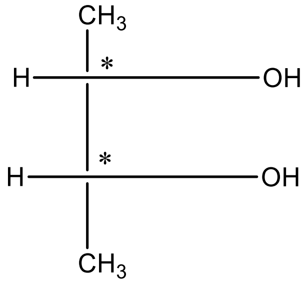
299 **(b)**



anisole

It is an ether and the name of ether is given as alkoxy alkane. So, its name is methoxy benzene.

300 **(b)**



butane-2-3-diol

Where C^*=asymmetric C atom
It is a symmetrical molecule, so the number of optically active stereomers=2^{n-1}
(n=number of asymmetric C atom)
$= 2^{2-1}$
$= 2^{-1} = 2$

302 **(d)**
Wöhler prepared urea from KCNO and $(NH_4)_2SO_4$
.

304 **(a)**
Electron donors having lone pair of electrons are nucleophile.
(i) $BF_3$is not nucleophile because it does not have lone pair of electrons. It is infact Lewis acid because it accepts pair of electron
(ii)NH_3, CN^- and OH^-all have lone pair of electrons, so they are nucleophiles.

305 **(b)**
Due to asymmetric carbon atom.

306 **(b)**
Aryl carbonium ions are more stable than alkyl carbonium ions. The order of stability of carbocation is
Triphenyl methyl >Diphenyl methyl> Benzyl>Allyl>3°>2°>1°> methyl carbocation.

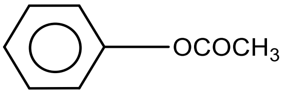
308 **(c)**
Follow IUPAC rules.

310 **(a)**
The dispersal of the charge stabilises the carbocation. More the number of alkyl groups; the greater the dispersal of positive charge and therefore, more the stability of carbocation, thus $C_2H_5^+ > CH_3^+, O - CH_3$ is also an electron donating group, thus it will increase the stability of carbocation, hence, the correct order of stability is C>B>A.

311 **(c)**
Chromatography method is used to separate sugars.

312 **(c)**
Only urea does not sublime while naphthalene, camphor and benzoic acid do

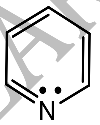
313 **(c)**
Esters are named by prefixing the name of the alkyl or aryl group (of OR' part) before the name of the parent acid and changing the suffix ic acid to ate. Hence, the structure of phenyl ehtanoate is



314 **(b)**
Inductive effect of groups is measured with respect to H.

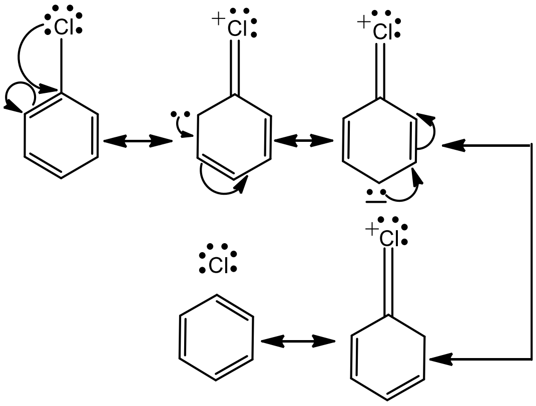
315 **(d)**
Methyl halides are methylating agents.

316 **(a)**
Nucleophiles are the species which have excess of electrons. Among the given species, the lone pair of nitrogen of pyrrole is involved in delocalisation of the ring, thus, are not available for donation. In aniline, the lone pair is involved in conjugation with the π-electrons of the ring while in pyridine, these are relatively free for donation. Thus, nitrogen of pyridine is most nucleophilic.



(phenyl and – $COCH_3$ both are electron withdrawing groups, thus decreases the nuleophilicity of nitrogen).

317 **(c)**
Chlorobenzene is *o, p* directing in electrophilic substitution reaction. The directing influence is explained by $+M$ of Cl atom



$+M$ effect of chlorine atom

318 **(b)**
Free radicals are electrons deficient compounds. Alkyl groups are electron donor groups and they increase the stability of free radicals.
∴ More the number of alkyl groups, more will be stability of free radicals.
∴ 3°>2°>1° is the correct order of stability of free

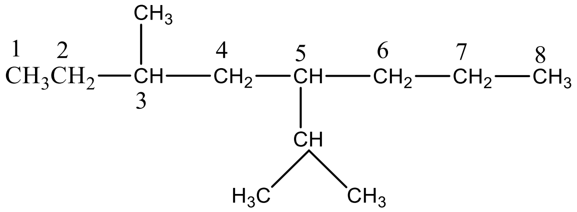
radicals.

319 **(b)**

This give rise to net resultant of four $C-Cl$ vectors equal to zero.

320 **(b)**

IUPAC name of compound.



Straight chain which contains large number of side chains taken as parent chain and counting starts from that side where the side chain is nearest.

3-methyl-5(1-methyl ethyl) octane.

321 **(a)**

The isomerism which arises due to rotation about a C-C is called conformational isomerism and the isomers are called conformational isomers or rotational isomers or conformers.

322 **(a)**

$CH_3CH=CH_2CHOHCH_3$ has one asymmetric carbon.

323 **(b)**

(i) In nucleophilic substitution reaction more powerful nucleophile replaces weaker nucleophile.

(ii) In rearrangement reaction atoms replace their position within molecule.

(iii) In elimination reaction small molecules (e.g., H_2O, NH_3) are lost.

$$R-CH_2CH_2Cl+KOH\,(alc.)\xrightarrow{\Delta}RCH=CH_2+KCl+H_2O$$

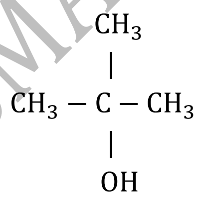
$\because$ KCl and H_2O molecules are lost during reaction.

$\therefore$ It is an elimination reaction.

324 **(c)**

CH_3COCH_3 is simplest ketone.

326 **(d)**

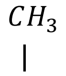


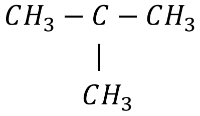
$$\begin{array}{c} CH_3 \\ | \\ CH_3-C-CH_3 \\ | \\ OH \end{array}$$

is the formula of tertiary butyl alcohol as in it – OH group is attached to tertiary carbon.

327 **(b)**

The neopentane:





$$\begin{array}{c} CH_3 \\ | \\ CH_3-C-CH_3 \\ | \\ CH_3 \end{array}$$

IUPAC name:2,2-dimethyl propane

328 **(d)**

Triple bond possesses maximum bond energy.

329 **(d)**

If molecule having asymmetric carbon atom and is not superimoposable on its mirror image then it is chiral while if it is superimposable on its mirror image, it is achiral.

330 **(a)**

(a) When optically active acid reacts with racemic mixture of an alcohol, it forms two types of isomeric esters. In each, the configuration of the chiral centre of acid will remain the same.

So, the mixture will be optically active.

331 **(d)**

Due to H^- shift from C_2 to C_3. Driving force is conjugation from oxygen. Also bulky gps hinders in hydride shift.

332 **(a)**

C_6H_6 has six delocalized π-electrons.

333 **(c)**

Due to same molecular formula.

334 **(d)**

Cis − trans isomers generally contain double bonded carbon atoms.

335 **(b)**

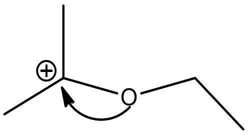
Due to the presence of lone pair on N atom.

336 **(c)**

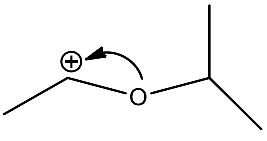
A molecule having doubly bonded carbon atoms shows geometrical isomerism only if both the doubly bonded carbon have altogether different group, *i.e.*, $_{ba}C\equiv C_{ab}$ or $_{ab}C=C_{ac}$ or $_{dc}C=C_{ab}$.

337 **(d)**

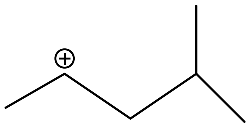
I>III>II>IV



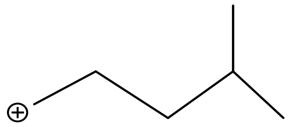
$+M$ effect, hyperconjugation (6Hs) and $+I$ effect (2 Me-groups)



$+M$ effect, hyperconjugation (3Hs) and $+I$ effect (1-Me group)



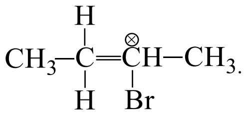
(2° carbocation) Hyperconjugation (5Hs) and +*I* effect



(1° carbocation) Hyperconjugation (2Hs) and +*I* effect

338 **(b)**

2-bromobutane has asymmetric carbon atom.



$\mathrm{CH_3{-}\underset{H}{\overset{H}{C}}{=}\underset{Br}{\overset{\otimes}{C}}H{-}CH_3.}$

339 **(b)**

$\dot{\mathrm{Br}}$ is less reactive and more selective and thus, formation of 3° free radical will be the major product.

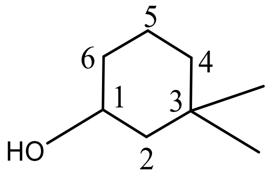
341 **(b)**

When organic compound containing both nitrogen and sulphur is fused with sodium, sodium thiocyanate is formed

343 **(d)**

Alcohols show position isomerism; Ethers show metamerism; Alcohols and ethers show functional isomerism.

344 **(c)**



Carbon with – OH group is given C_1 thus it is 3, 3-dimethyl-1-cyclohexanol.

345 **(b)**

—do—

346 **(c)**

The reactivity order for H atom is 3°H > 2°H > 1°H.

347 **(a)**

Pyrene is CCl_4.

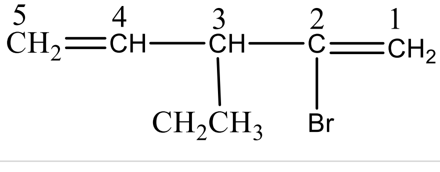
349 **(d)**

3° carbonium ions are more stable as the +ve inductive effect disperses +ve charge on carbon atom.

352 **(d)**

Follow mechanism of free radical substitution.

353 **(b)**



$\overset{5}{\mathrm{CH_2}}{=}\overset{4}{\mathrm{CH}}{-}\underset{\mathrm{CH_2CH_3}}{\overset{3}{\mathrm{CH}}}{-}\underset{\mathrm{Br}}{\overset{2}{\mathrm{C}}}{=}\overset{1}{\mathrm{CH_2}}$

The IUPAC name of this molecule is 2-bromo-3-ethyl-1, 4-pentadiene.

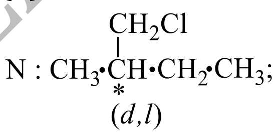
355 **(c)**

N is pentavalent which is not possible.

356 **(a)**

Follow inductive effect.

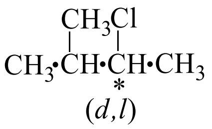
357 **(b)**



N : $\mathrm{CH_3{\cdot}\underset{*}{\overset{CH_2Cl}{C}}H{\cdot}CH_2{\cdot}CH_3}$;

(*d,l*)

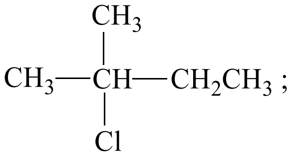
(I)



$\mathrm{CH_3{\cdot}\overset{CH_3}{C}H{\cdot}\underset{*}{\overset{Cl}{C}}H{\cdot}CH_3}$

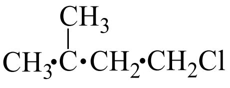
(*d,l*)

(II)



$\mathrm{CH_3{-}\underset{Cl}{\overset{CH_3}{C}}H{-}CH_2CH_3}$;

(III)



$\mathrm{CH_3{\cdot}\overset{CH_3}{C}{\cdot}CH_2{\cdot}CH_2Cl}$

(IV)

M : (I), (II), (III), ; (I) and (II)

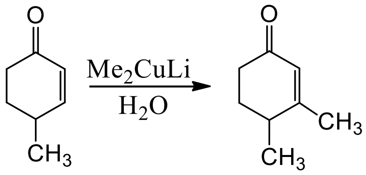
Cannot be separated by fractional distillation.

358 **(c)**

It is a fact.

359 **(a)**

Replacement of an atom or group by other atom or group is known as substitution reaction

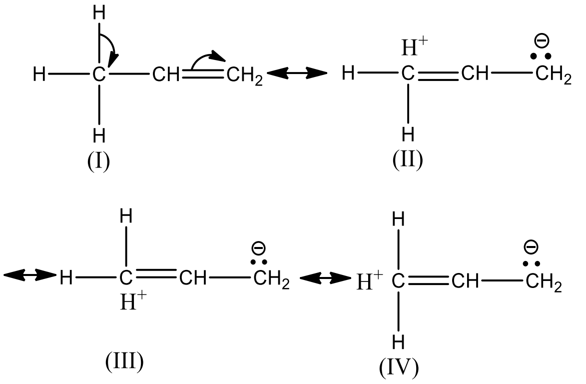


360 **(b)**

It is a fact.

361 **(a)**

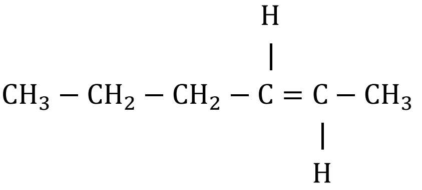
Orbital interaction between the σ- bonds of a substituent group and a neighbouring π-orbital is known as hyperconjugation.



362 **(c)**
sp^3-hybridization with one position occupied by lone pair like NH_3.

364 **(a)**
2-hexyne gives *trans*-2-hexene on treatment Li/NH_3

$$CH_3-CH_2-CH_2-C\equiv C-CH_3+H_2\xrightarrow{Li/NH_3}$$

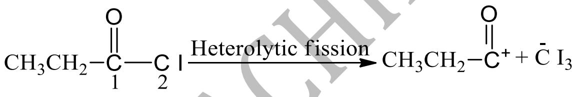


$$CH_3-CH_2-CH_2-\overset{\overset{\displaystyle H}{|}}{C}=\underset{\underset{\displaystyle H}{|}}{C}-CH_3$$

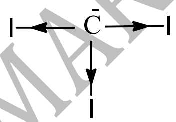
trans-2-hexene

365 **(b)**
Lassaigne's test is given by those nitrogenous compounds in which carbon is also present along with nitrogen.
In $NH_2.NH_2.HCl$, carbon is absent, so it does not given Lassaigne's test.

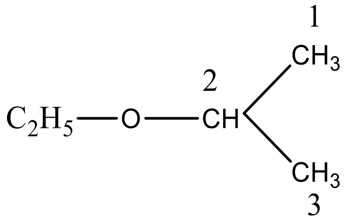
366 **(b)**



$\bar{C}I_3$ is most stable carbanion because of $-I$ effect of I which disperses negative charge on $\bar{C}$. Center



367 **(d)**



2-ethoxy propane

The above compound is an ether and its name is written as alkoxy alkane. Oxy is attached with the lower group. Hence, the IUPAC name of above compound is 2-ethoxy propane.

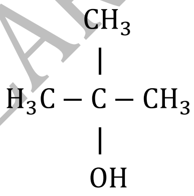
368 **(b)**
In thin layer chromatography the relative adsorption of each component of the mixture is expressed in terms of retention factor (R_f)

$$R_f=\frac{\text{distance moved by the spot centre from base line}}{\text{distance moved by the solvent from the base line}}$$

369 **(c)**
$(CH_3)_3CBr+H_2O\rightarrow(CH_3)_3C-OH+HBr$
Br is subsituted by $-OH^-$(nucleophile)
S_N1(unimolecular nuclerophilic substitution reaction)

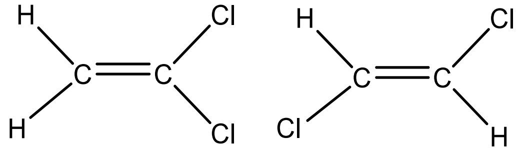
370 **(c)**
$C_6H_5\overset{+}{C}HC_6H_5$ is the most stable since the +ve charge can be delocalized on both phenyl rings

371 **(a)**
2-methyl propane-2-ol is *tert*-butyl alcohol.



$$H_3C-\underset{\underset{\displaystyle OH}{|}}{\overset{\overset{\displaystyle CH_3}{|}}{C}}-CH_3$$

372 **(c)**



Since, in the above structures, position of Cl is different, these are position isomers, which is a type of structural isomerism.

374 **(b)**
More directionally concentrated orbitals show more overlapping. Also more closer are shells to the nucleus more is overlapping.

375 **(b)**
Removal of H from ketone gives resonance stabilized carbanion.

376 **(d)**
$CH_3CH_2X\longrightarrow CH_2=CH_2(sp^3\text{ to }sp^2)$;
$CH_2=CHX\longrightarrow CH\equiv CH\ (sp^2\text{ to }sp)$;
$CH_2XCH_2CH_2X\longrightarrow\Delta$ (No change).

377 **(b)**
Follow IUPAC rules.

378 **(c)**
It is a fact.

379 **(d)**

S_N1(Unimolecular nucleophilic substitution reactions)

Rate∝ (substrate)

Rate determining step in the formation of carbocation depends on the stability of carbocation formed. The stability of carbocations follow the order

$(C_6H_5)_2\,C^+(CH_3) > (C_6H_5)_2\overset{+}{C}H > (CH_3)_2\overset{+}{C}H > CH_3CH_2^+$

∵ Order of S_N1 reacticity is

$(C_6H_5)_2C(CH_3)Br > (C_6H_5)_2CHBr > (CH_3)_3CBr$
$> (CH_3)_2CHBr > C_2H_5Br$

i.e., iii>ii>i>iv>v

380 **(c)**

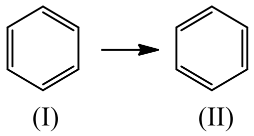
$CH_3CH_2CH_3 \xrightarrow{\text{Hetrolysis}} \bar{C}H_3 + CH_3\,\overset{+}{C}H_2$, $CH_3\,\overset{+}{C}H_2$ is more stable than $CH_3—\bar{C}H_2$

due to dispersal of +ve charge on ethylium ion on account of +ve inductive effect. Thus,

propane will not give $\overset{+}{C}H_3$ and $CH_3\,\bar{C}H_2$.

381 **(d)**

Delocalised electrons are present in benzene, 1, 3-butadiene and 1,3,5-hexatriene



382 **(a)**

A characteristic of *dextrorotatory*.

383 **(c)**

It is a fact.

384 **(c)**

If a liquid decomposes at or below its boiling point, it is purified by vacuum distillation, impure glycerine is purified by this method

386 **(a)**

Distillation process is not used for purification of solid impurities. It is used for the purification of liquids which boils without decomposition and contains non-volatile impurities.

387 **(b)**

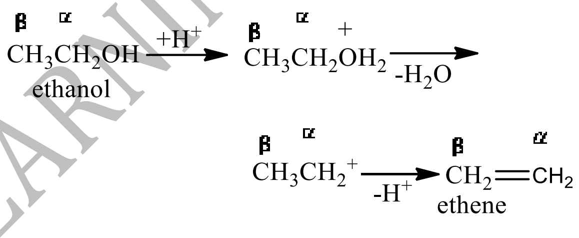
sp^3, sp^2 and sp
– orbitals are at 109°28′, 120° and 180°.

388 **(b)**

2°H is more reactive than 1°.

389 **(b)**

Dehydration of alcohol involves the loss of two atoms or groups from the adjacent carbon atoms, hence it is an example of β-elimination reaction.



$\overset{\beta\ \ \alpha}{CH_3CH_2OH} \xrightarrow{+H^+} \overset{\beta\ \ \alpha}{CH_3CH_2\overset{+}{O}H_2} \xrightarrow[-H_2O]{}$
ethanol

$\overset{\beta\ \ \alpha}{CH_3CH_2^+} \xrightarrow[-H^+]{} \overset{\beta\ \ \alpha}{CH_2{=}CH_2}$
ethene

391 **(d)**

(i) Eantiomers are pair of optical isomers which are related as non-superimposable mirror images of each other.

(ii) Diasteriomers are pair of optical isomers which cannot be related as non-superimposable mirror images of each other.

∴ The only correct statement about given structures is that (*A*) and (*B*) are enantiomers.

392 **(a)**

Two positive charges present at the adjacent place, elevates the energy, thus lowers the stability most.

393 **(d)**

Due to H-bonding.

394 **(a)**

Follow inductive effect.

395 **(b)**

Naphthalene and benzoic acid cannot be separated by the sublimation method because the naphthalene and benzoic acid both are sublimes on heating. They are separated by hot water in which benzoic acid dissolves but naphthalene does not

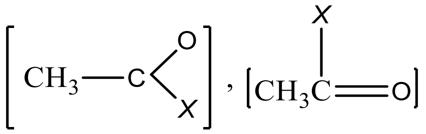
396 **(d)**

All are used as dehydrohalogenating agent.

$$C_2H_5X \xrightarrow[\text{or (ii) } C_2H_5ONa \text{ or (iii) } NaNH_2]{\text{(i) KOH alc.}} C_2H_4 + HX.$$

397 **(c)**

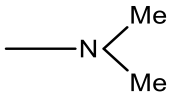
In the given electrophile



Group is the same. So, only X affects their activity, *i.e.*, we have to discuss activity due to

(a) $-OCH_3$ (b) $-Cl$

(c)



(d) $-S-CH_3$

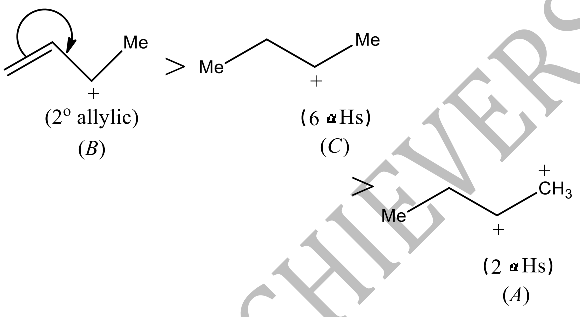
Since, amines are less actives, therefore, electrophile (c) will be least active.

398 **(c)**

The positions of Cl are different.

399 **(a)**

Higher the stability of carbocation, faster is the reaction because S_N1 reactions involve the formation of carbocation intermediate.



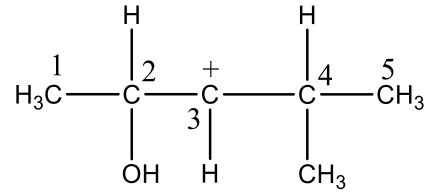
400 **(b)**

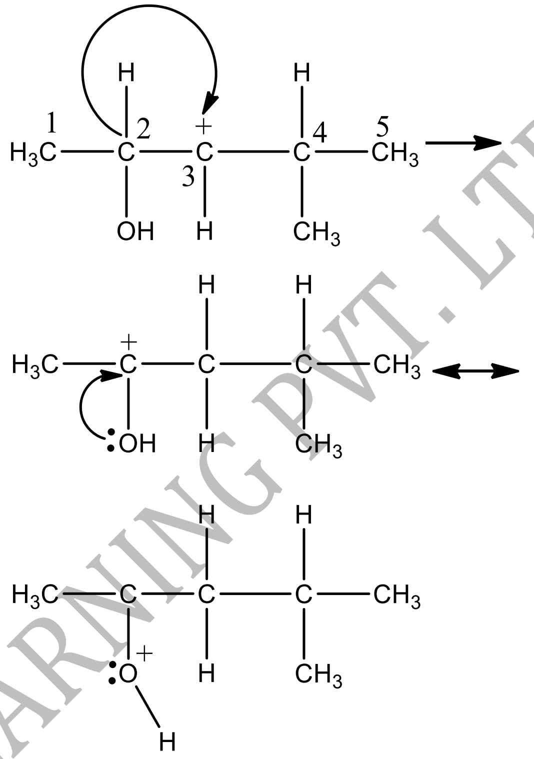
Both Wolff-Kishner and Clemmensen reduction are used to convert $>$CO to $>CH_2$.

The later is not suitable as it will also attack –OH group of ring.

401 **(d)**

In the following carbocation; H/CH_3 that is most likely to migrate to the positively charged carbon is

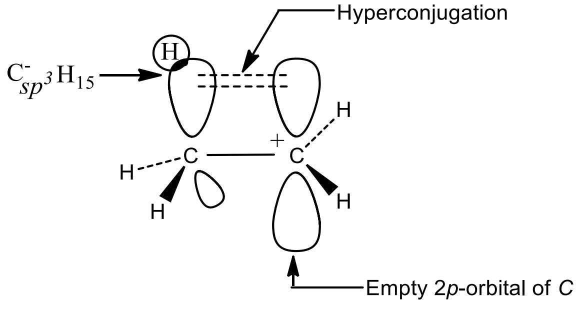




(A resonance stabilised carbocation)

402 **(b)**

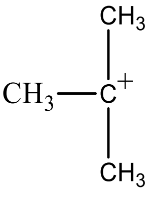
Hyperconjugation arises due to the partial overlap of a $sp^3 - s$ (a C-H bond) with the empty p –orbital of an adjacent positively charged carbon atom.



Hyperconjugation in ethyl cation

403 **(c)**

In the triphenyl methyl carbonium ion the π-electrons of all the three benzene rings are delocalised with the vacant p-orbital of central carbon atom. So, it is resonance stabilised. It is the most stable of all the carbonium ions given

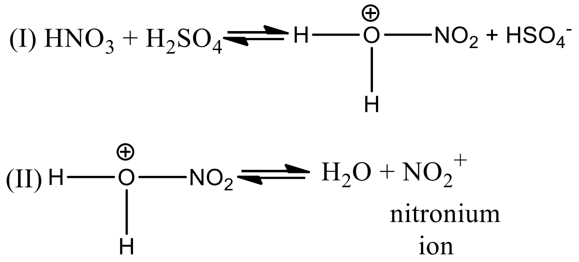


is stabilised by hyperconjugattion, a

second order resonance.

404 **(c)**

During nitration of benzene the attacking electrophile is NO_2^+. It is formed as follows by reaction between HNO_3 and H_2SO_4.

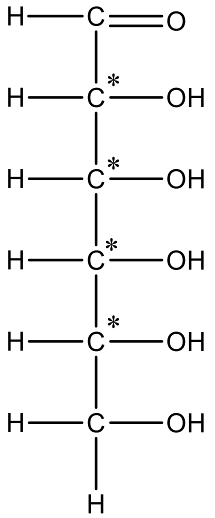


405 **(b)**

The R and S enantiomers of an optically active compound differ in their optical rotation of plane polarised light.

406 **(a)**

A carbon atom which is attached by four different group is called chiral centre of asymmetric carbon atom. (+)-glucose has four chiral centres.



408 **(b)**

C_2H_2 has two π-bonds.

409 **(b)**

The various structure derived for a molecule but none of them truly represents all the properties of that molecule are said to be canonical forms and the molecule is said to show resonance.

410 **(d)**

The possible primary amine with the formula $C_4H_{11}N$ are

(i) $CH_3CH_2CH_2CH_2NH_2$

(ii) $CH_3 - \overset{\overset{CH_3}{|}}{CH} - CH_2 - NH_2$

(iii) $CH_3 - \overset{\overset{CH_3}{|}}{\underset{\underset{NH_2}{|}}{C}} - CH_3$

(iv) $CH_3 - CH_2 - \underset{\underset{CH_3}{|}}{CH} - NH_2$

411 **(c)**

Stereoisomerism is of two types-optical and geometrical.

412 **(a)**

Catenation is the tendency to unite atoms of an element to form a long carbon chain or ring.

414 **(b)**

Electron withdrawing group has $-I$ effect while electron donating group has $+I$ effect.

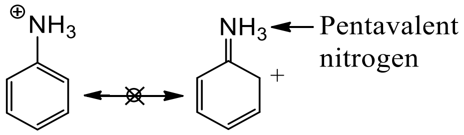
In CH_3COOH, the alkyl group $(-CH_3)$ due to its greater $+I$ effect increases the electron density on oxygen atom of the O-H bond. Due to this the release of H^+ ion in acetic acid will be more difficult as compared to formic acid.

415 **(b)**

It is a fact.

416 **(b)**

If positive charge is present on nitrogen then positive charge will not be in conjugation to the ring because in this case nitrogen will become pentavalent



418 **(c)**

HNO_3 is added to decompose Na_2S and NaCN if present

$Na_2S + 2HNO_3 \rightarrow 2NaNO_3 + H_2S \uparrow$

$NaCN + HNO_3 \rightarrow NaNO_3 + HCN \uparrow$

419 **(c)**

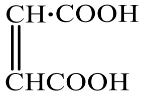
It is a fact.

420 **(c)**

The definition of electromeric effect.

421 **(c)**

Maleic acid is



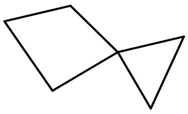
$\begin{array}{c} CH \cdot COOH \\ \| \\ CHCOOH \end{array}$

It has no asymmetric carbon, however it shows geometrical isomerism.

422 **(a)**

Sbstituion of $-OH$ group from a substrate can be easily made by PCl_5.

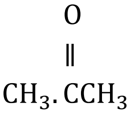
423 **(c)**



Chair form is unsymmetrical due to absence of any element of symmetry.

424 **(a)**

When two compounds have similar molecular formula but differ in the functional group then the isomerism is called functional group isomerism *i.e.*,



CH_3CH_2CHO and $CH_3\overset{\overset{O}{\|}}{C}CH_3$

426 **(d)**

Count σ-and π-bonds on each carbon and report hybridization.

427 **(d)**

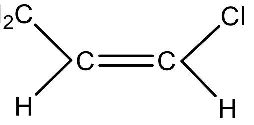
Carbonyl compounds show nucleophilic addition.

428 **(a)**

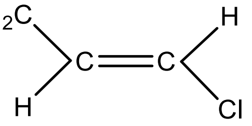
Twelve in all

(a) Six geometrical isomers

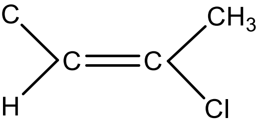
(i)



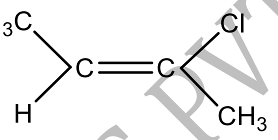
(ii)



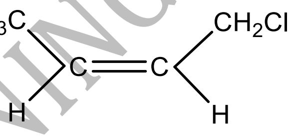
(iii)



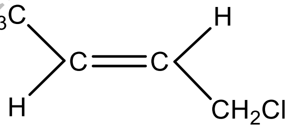
(iv)



(v)

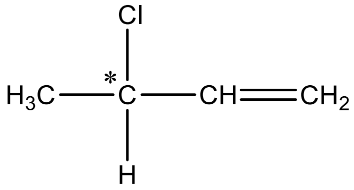


(vi)

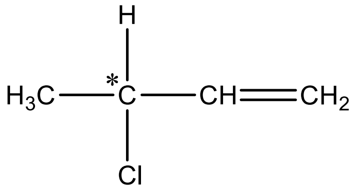


(b) Two optical isomers

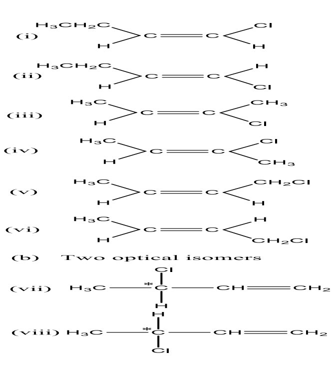
(vii)



(viii)



(c)



429 **(c)**

Ortho and *para*- nitrohenol is separated by distillation because p −nitrophenol has higher boiling point than o-nitrophenol due to H-bonding.

430 **(b)**

Electromeric effect implies complete transfer of π electrons in presence of a reagent. Since, simple ethers do not contain a multiple bond, therefore, they do not show electromeric effect

431 **(b)**

It is definition of electromeric effect.

432 **(a)**

Mass of silver salt=0.4 g

Mass of silver =0.26 g

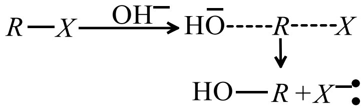
Eq. mass of silver salt/Eq. mass of Ag= $\frac{\text{wt.of silver salt}}{\text{wt.of silver}}$

Eq. mass of silver salt $= \frac{108\times0.4}{0.26} = 166$

Eq. mass of acid $= 166 - 108 = 58$

433 **(a)**

The S_N2 mechanism always involves 100% inversion since nucleophile attacks from back side of leaving group.

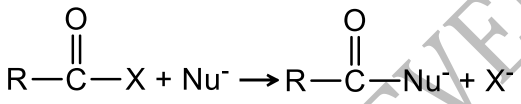


$$R—X \xrightarrow{OH^-} H\bar{O}\text{-----}R\text{-----}X \downarrow HO—R + X^-$$

434 **(a)**

Follow IUPAC rules.

435 **(d)**



$$R—\overset{O}{\overset{||}{C}}—X + Nu^- \longrightarrow R—\overset{O}{\overset{||}{C}}—Nu^- + X^-$$

Best leaving group (poorest nucleophile) is $Cl^{\oplus}$, thus fastest reaction is with Cl.

436 **(c)**

Note that propyl (propan-l-ol) and isopropyl alcohol (propan-2-ol) are position isomers.

437 **(d)**

Empirical formula of acid $= CH_2O_2$

We know that molecular formula $= n$ (empirical formula)

If $n = 1$ molecular formula $= (CH_2O_2)_1 = CH_2O_2$

If $n = 2$ molecular formula $= (CH_2O_2)_2 = C_2H_4O_4$

If $n = 3$ molecular formula $= (CH_2O_2)_3 = C_3H_6O_6$

Thus, the probable molecular formula $= CH_2O_2$

438 **(c)**

Excited state of carbon is $2s^12p^3$.

439 **(a)**

$$C = 10.5\text{ g} = \frac{10.5}{12}\text{mol} = 0.87\text{ mol}$$

$$H = 1\text{ g} = \frac{1}{1}\text{mol} = 1\text{ mol}$$

$$\therefore (C_{0.87}H_1)_7 = C_{6.09}H_7 \approx C_6H_7$$

$$pV = nRT = \frac{w}{m}RT$$

$$1 \times 1 = \frac{2.4}{m} \times 0.082 \times 400$$

$$m = 79$$

Hence, the hydrocarbon is C_6H_7

440 **(c)**

There are four stereoisomers

$cis - R \quad cis - S$

$trans - R \quad trans - S$

441 **(c)**

IUPAC name is N-methyl methanamine.

442 **(d)**

Oil and water are immiscible liquids thus are separated by using separating funnel

443 **(a)**

Can be solved on the basis of hyperconjugative structures

444 **(d)**

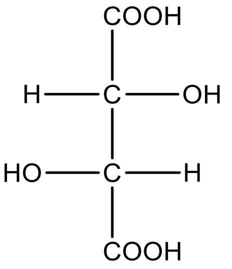
Total number of optical isomers=$(2)^n$

(where n=number of asymmetric carbon atom).

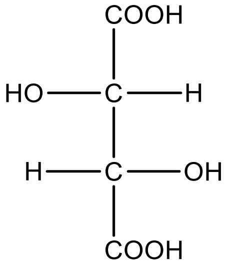
$= (2)^2 = 4$

Out of these four optical isomers two are *meso* structures which are optically in active.

∴ Only two structures d and l- tartaric acid are optically active.



d-tartaric acid



l-tartaric acid

446 **(d)**

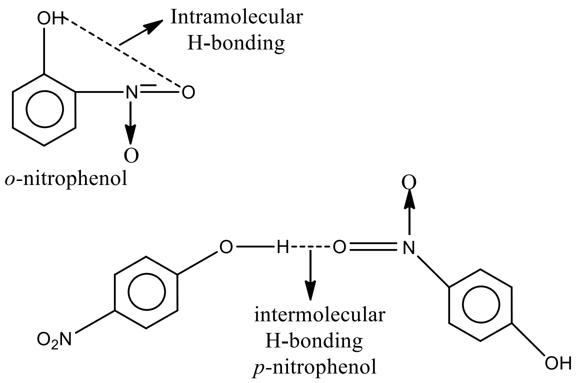
Three coordinate bond on O atom.

447 **(c)**

Furan is heterocyclic and aromatic due to Huckel's rule of aromaticity, $i.e.$, 6π-electrons.

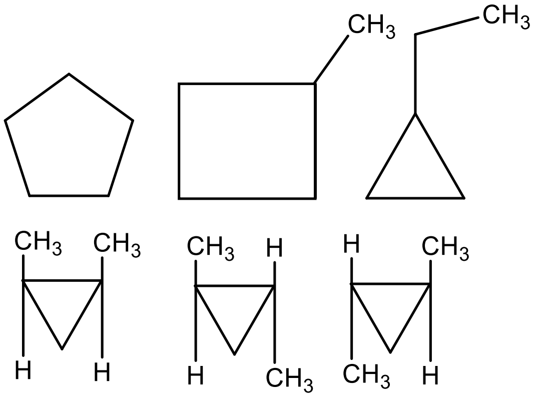
448 **(d)**

In mixture of o-nitrophenol, p-nitrophenol, o-nitrophenol is steam volatile due to intra molecular hydrogen bonding whereas p-nitrophenol is less volatile due to inter molecular hydrogen bonding.



449 **(c)**

The total number of cyclic isomers are six as shown below.



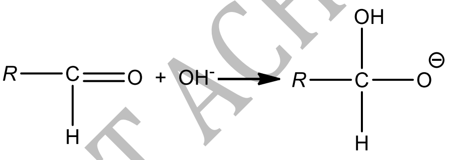
451 **(d)**

Electromeric effect involves complete transfer of π-electron pair to more electronegative atom on the need of attacking reagent.

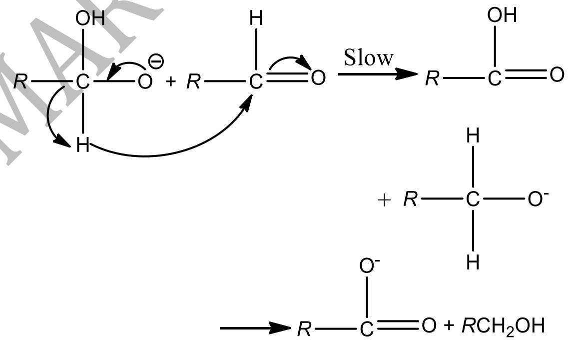
452 **(c)**

Cannizzaro reaction involves oxidation as well as reduction of aldehydes having lack of $\alpha - H$ atom. The mechanism of this reaction is as

(I) Attack of OH^- on carbonyl carbon



(ii) Transfer of hydride ion



(II) Transfer of hydride ion

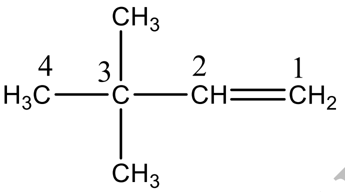
453 **(d)**

$$C:H:Br = \frac{2.27}{12} : \frac{5.69}{1} : \frac{65.04}{80}$$
$= 2.43 : 5.69 : 0.813$
$= 3 : 7 : 1$
or empirical formula $= C_3H_7Br$

454 **(b)**

The terminal $'e'$ of ene is retained if suffix name starts with consonant.

455 **(d)**



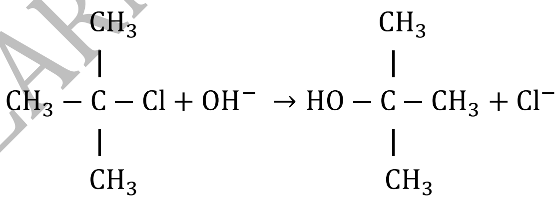
IUPAC name=3, 3-dimethyl-1-butene.

456 **(b)**

CaO is added to NaOH to retard activity of NaOH, otherwise decarboxylation of acids will occur more violently.

$$RCOONa \xrightarrow{\text{NaOH + CaO}} R - H + Na_2CO_3$$

457 **(b)**



$$CH_3 - \underset{CH_3}{\overset{CH_3}{C}} - Cl + OH^- \rightarrow HO - \underset{CH_3}{\overset{CH_3}{C}} - CH_3 + Cl^-$$

Rate $\propto$ [t-butyl chloride]

Tertiary butyl carbocation is first formed which is more stable

458 **(c)**

In heterolysis, the covalent bond is broken in such a way that one species (less electronegative) is deprived if its own electron, while the other species gain both the electrons

$$CH_3CH_2CH_3 \longrightarrow \overset{\ominus}{C}H_3 + \overset{\oplus}{C}_2H_5$$

459 **(c)**

Both have different functional groups, $i.e.$, $-CN$ and $-NC$.

460 **(b)**

A mixture of 50-50% of d and its l form is called racemic mixture.

461 **(d)**

Sublimation is the process employed for those solids which convert directly into vapours on heating without converting into liquid phase

462 **(c)**

It is the definition of optical activity.

463 **(a)**

Stability of alkyl carbanion $\propto$

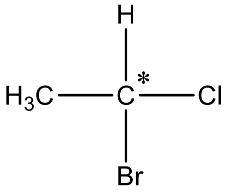
$\xrightarrow[\text{magnitude of negative charge}]{1}$ and magnitude of negative charge $\propto +I$ power of the group. Hence, acetylenic carbanion is more stable than vinylic carbanion which is more stable than alkyl carbanion

465 **(b)**

Follow IUPAC rules.

466 **(b)**

A carbon atom which is attached by four different groups is called an asymmetric carbon atom or chiral centre.



468 **(a)**

A carboxylic acid is stronger acid than phenol, hence both III and IV are stronger acids than both I and II. Also IV has a methyl group that gives electrons donating inductive effect and decreases the acid strength. Therefore, III is stronger acid than IV. Between I and II, the dominate electron withdrawing inductive effect of chlorine increases acid strength of phenol slightly, hence II is stronger of phenol slightly, hence, II is stronger acid than I.

Thus, the overall order is: (a) III>IV>II>I.

469 **(a)**

All neutral covalent compound in which central atom has incomplete octet are electrophile. For example $BeCl_2, BH_3, ZnCl_2, AlCl_3$

470 **(d)**

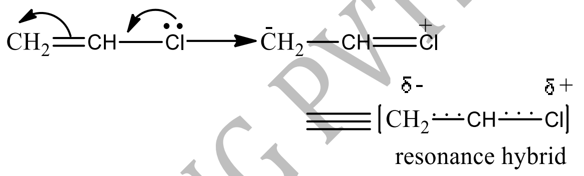
Planar hexagon conformer has considerable angle strain due to the fact that its bonds are not 109.5°. It also has torsional strain. Due to presence of these strains planar hexagon conformer of cyclohexane is least stable.

471 **(a)**

Both have same molecular formula.

472 **(a)**

Due to resonance partial double bond character is created on vinyl chloride. So, chlorine atom is not replaced easily

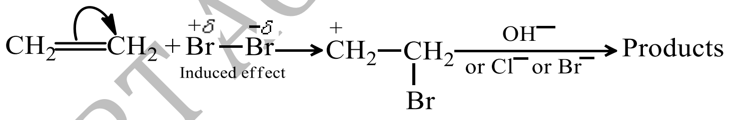


473 **(a)**

Glycerol is a trihydric alcohol. It is $CH_2OH - CHOH - CH_2OH$. It contains three hydroxyl group. It is present in nature in oils and fats as its carboxylic esters.

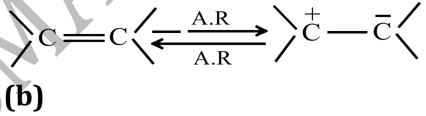
474 **(d)**

Once the carbocation is formed as an intermediate, the nucleophile Cl^- and OH^- present in solution also attach it in addition of Br^-.



475 **(a)**

It is arised on the need of attacking reagent, *e*. g.,



478 **(b)**

An organic reaction intermediate, neutral species having divalent carbon atom with six valence electron out of which two are present in same orbital with opposite spin is called singlet carbene.

479 **(a)**

Nitration or aromatic compounds takes place by an electrophile. The electrophile will be more attracted towards electron rich position in benzene ring. Hence, electron donating groups will be easily nitrated.

Toluene will be most easily nitrated among these compounds due to presence of electron donating group $(i.e., CH_3)$.

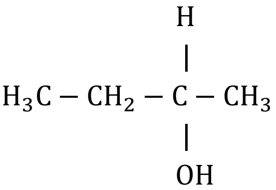
Nitrobenzene will be most slowly nitrated due to the presence of electron withdrawing group $(i.e., NO_2)$.

CH_3NO_2 will be formed by free radical substitution of CH_4

480 **(b)**

Optical isomerism is shown by an asymmetric carbon atom which has a carbon atom attached to four different atoms or groups.

So, butanol-2 is optically active.



$$H_3C-CH_2-\underset{\underset{OH}{|}}{\overset{\overset{H}{|}}{C}}-CH_3$$

butanol-2

481 **(c)**

When sodium or potassium alkoxide is heated with an alkyl halide to give ether, this reaction is known as Williamson's synthesis.

$R\text{ONa}+\text{R}'\text{X}\rightarrow \text{R}-\text{O}-\text{R}'+\text{NaX}$

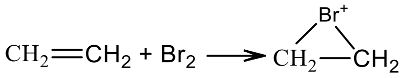
This is an example of nucleophilic substitution and follow S_N2 mechanism.

482 **(d)**

These are characteristics of free radicals.

483 **(a)**

Addition of Br_2 on ethane follow electrophilic addition



Intermediate is cyclic bromonium ion

484 **(d)**

Metamers of ethyl propionate are as

$CH_3COOC_3H_7, C_3H_7COOCH_3$

485 **(c)**

Each carbon has two σ-and two π-bonds.

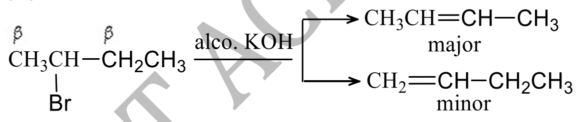
486 **(b)**

It should be 4-ethyl-3-methyl heptane.

487 **(c)**

Due to asymmetric carbon atom in it.

488 **(b)**



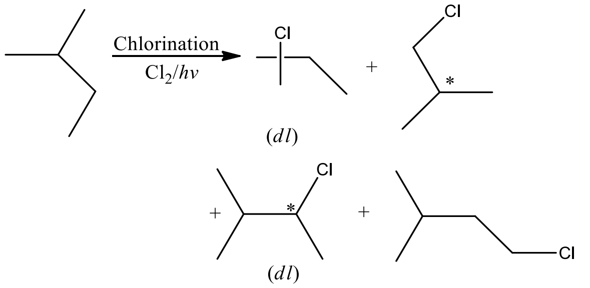
This reaction is governed by Saytzeff's rule. According to this rule the elimination of β-hydrogen atom take place from the carbon having the lesser number of H-atoms or in other words a stable alkene is formed. (More substituted alkene is more stable)

489 **(a)**

Due to same functional group.

490 **(a)**

On chlorination of 2-methyl butane



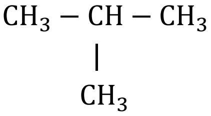
2-chiral compound are formed.

491 **(d)**

Due to the presence of asymmetric carbon atom.

492 **(c)**

Two isomers



$$CH_3-\underset{\underset{CH_3}{|}}{CH}-CH_3$$

Iso-butane

$CH_3-CH_2-CH_2-CH_3$

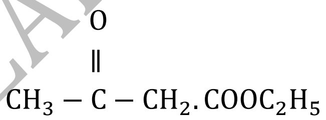
n-butane

493 **(c)**

−do −

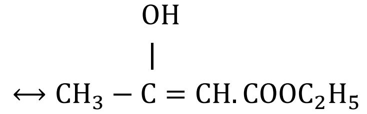
494 **(d)**

Ethyl acetoacetate shows tautomerism.



$$CH_3-\overset{\overset{O}{\|}}{C}-CH_2.COOC_2H_5$$

Keto form

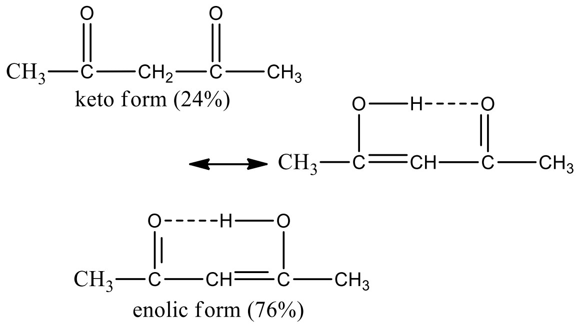


$$\leftrightarrow CH_3-\overset{\overset{OH}{|}}{C}=CH.COOC_2H_5$$

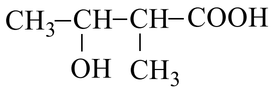
enol form

495 **(d)**

The amount of enolic form is highest in acetyl due to the stabilisation of enolic form by hydrogen bonding



496 **(d)**

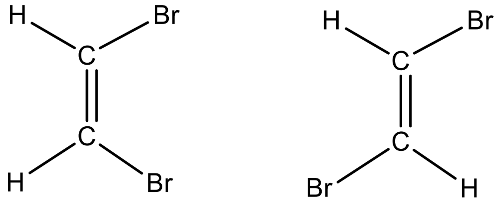


$$CH_3-\underset{\underset{OH}{|}}{CH}-\underset{\underset{CH_3}{|}}{CH}-COOH$$ has two asymmetric carbon atoms and

molecule has no symmetry. Thus, number of optical isomers = $2^n = 2^2 = 4$.

497 **(c)**

Due to restricted rotation about double bond, the alkene shows geometrical isomerism because the relative position of atoms or groups attached to the carbon atoms of the double bond get fixed. If same groups or atoms attached with double bond bearing carbon, then alkene doesn't show geometrical isomerism.



cis form *trans* form

498 **(a)**

Due to $+I$ effect of CH_3 in toulene, it is more reactive than bezene. Due to electron withdrawing nature of – COOH group in benzoic acid and – NO_2 group in nitrobenzene, both benzoic acid and nitrobenzene are less reactive than benzene.

499 **(b)**

These are isopentane, neopentane and n-pentane.

500 **(d)**

(a) Ethylene dichloride and ethylidine chloride both react with alc. KOH to produce ethyne.

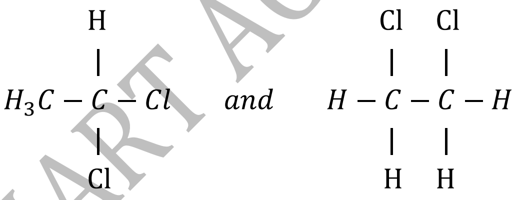
$H_3C - CHCl_2 + 2KOH(alc.) \rightarrow$

$HC \equiv CH + 2KCl + 2H_2O$

$ClCH_2 - CH_2Cl + 2KOH$ alc. $\rightarrow$

$HC \equiv CH + 2KCl + 2H_2O$

(b) They are position isomers because they have same molecular formula but different position of chlorine atom.



$H_3C - \underset{\substack{| \\ Cl}}{\overset{\substack{H \\ |}}{C}} - Cl$ and $H - \underset{\substack{| \\ H}}{\overset{\substack{Cl \\ |}}{C}} - \underset{\substack{| \\ H}}{\overset{\substack{Cl \\ |}}{C}} - H$

ethylidene dichloride ethylene dichloride

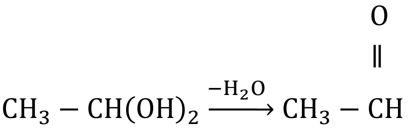
(c) ∵Their molecular formula is same

∴ They have same percentage of chlorine.

(d) They give different product on hydrolysis

$CH_3CHCl_2 + 2KOH\ (aq) \longrightarrow$

ethyledene chloride



$CH_3 - CH(OH)_2 \xrightarrow{-H_2O} CH_3 - \overset{\substack{O \\ \|}}{C}H$

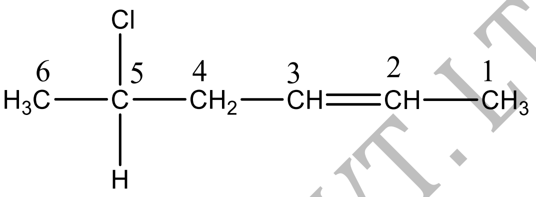
Unstable acetaldehyde

$CH_2Cl - CH_2Cl + 2KOH\ (aq) \longrightarrow \underset{\substack{| \\ OH}}{CH_2} - \underset{\substack{| \\ OH}}{CH_2}$

ethylene dichloride ethylene glycol

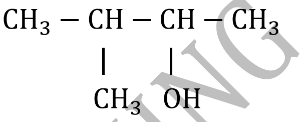
501 **(a)**

First the longest continuous chain of carbon atoms is selected. Now numbered the chain from the side containing senior functional group ($i.e.$, the group placed above in the seniority table).



5-chloro-hex-2-ene

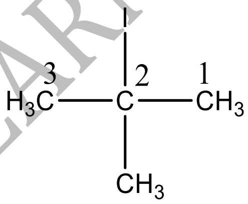
502 **(d)**



$CH_3 - \underset{\substack{| \\ CH_3}}{CH} - \underset{\substack{| \\ OH}}{CH} - CH_3$

2-methyl butan-2-ol

503 **(d)**



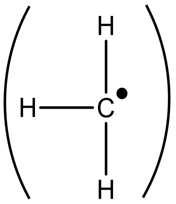
t- butyl iodide

Its IUPAC name is 2-iodo-2-methyl propane.

504 **(b)**

Number of hybrid orbitals =number of σ-bonds+number of *lps*

In $\dot{C}H_3$



; σ-bonds = 3

lps = 3

∴Number of hybrid orbitals =3+0=3

Hence, hybridisation is sp^2 and geometry is planar.

506 **(b)**

Follow the concept of hyperconjugation.

507 **(c)**

On the basis of stability of carbocation formed.

508 **(a)**

ϕ is phenyl group, the question can be solved on the basis of number of conjugative structures

509 **(a)**

S_N2 order : methyl > 1° > 2° > 3°.

510 **(c)**

Activating groups like – OCH_3, $-OH$ etc activates the benzene ring towards electrophilic substitution while deactivating groups like$-NO_2, -COOH$ etc. deactivates the benzene ring towards electrophilic substitution. Thus, order of reaction towards electrophile (of the given compounds) is as

I>II>III.

511 **(a)**

Electron deficient species or electron acceptor is electrophile. For example

$\dot{C}H_3, \ddot{C}H_2, \ddot{C}X_2$

513 **(b)**

As the min mol wt. must have at least one S-atom so

$$S\ \% = \frac{\text{wt. of one S} - \text{atom}}{\text{min. mol. wt}} \times 100$$

$$4 = \frac{32}{\text{min mol. wt}} \times 100$$

Min mol wt. $\frac{32 \times 100}{4} = 800$

514 **(b)**

A carbonium ion is sp^2-hybridized.

515 **(a)**

It is a fact.

517 **(c)**

Ortho and *para* directing groups facilitate the ring for electrophilic substitution reaction. – NH_2 group increase electron density in ring, hece activite it is to electrophilic substitutipon.

518 **(d)**

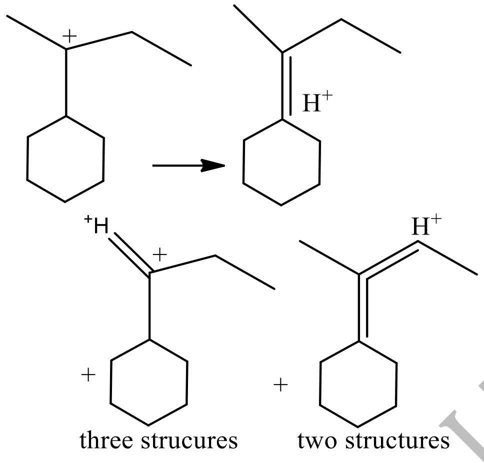
n-pentanol, 2-pentanol, 3-pentanol, 2-methylbutanol, 2-methylbutan-2-ol, 3-methylbutanol, 2, 2-dimethypropanol, and 3-methylbutan-2-ol (8 isomers)

519 **(a)**

Substances which sublimes on heating is usually purified by sublimation. Hence, naphthalene is purified by sublimation.

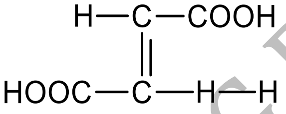
520 **(d)**

There are total $6\alpha - H$ to sp^2 carbon and they all can participate in hyperconjugation.



521 **(d)**

The structural formula of fumaric acid is

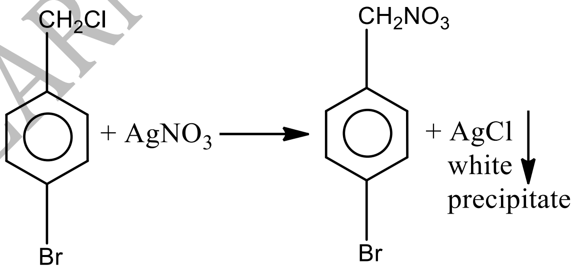


$$\begin{array}{l} \mathrm{H-C-COOH} \\ \quad \ \ \| \\ \mathrm{HOOC-C-H-H} \end{array}$$

(*trans* butane1,4-dioic acid)

523 **(b)**

A chloride linked with alkyl group is replaced with $AgNO_3$ and give white precipitate of AgCl.



524 **(c)**

Br is replaced by – OH.

525 **(c)**

Free radicals are represented by putting dot on entity.

527 **(d)**

Kjeldahl and Duma's methods are used for the quantitative estimation of nitrogen in an organic compound. In the Kjeldahl method, the nitrogen element of organic compound is changed to the ammonia.

528 **(d)**

Lesser is bond energy of $2p - 2p$ overlapping in C—C, more is its reactivity than C—H bond showing $2p$-$1s$ overlapping.

529 **(a)**

Carbon atom in singlet carbene is sp^2-hybridized.

531 **(b)**

$C_6H_5CH_2^+$ is stabilized by conjugation while intermediates of rest of the compounds given are stabilized by hyperconjugation

532 **(a)**

$CH_3CH_2CH_2CH_2CH_3$ is n-pentane.

533 **(b)**

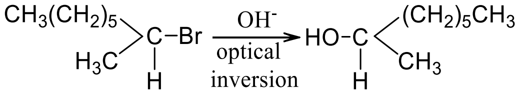
Resonance in benzene gives rise to identical C— C bond lengths.

534 **(a)**

During nitration benzene ring is attacked by NO_2^+ and hydrogen of benzene ring is replaced by NO_2 group.

∴ Nitration of benzene is electrophilic substitution because NO_2^+ is an electrophile.

535 **(b)**



In this reaction inversion takes place. Hence, it is an example of S_N2 reaction. In this mechanism the attack of OH^- ions take place from the back side while the Br^- ion leaves from the front side

536 **(b)**

Organic compound containing nitrogen is fused with a small piece of sodium metal to form NaCN.

$$Na + \underbrace{C + N}_{\text{from organic compound}} \rightarrow NaCN$$

537 **(d)**

The order of nucleophilicity depends upon the nature of alkyl group 'R' on which a nucleophile to attack as well as on nature of solvent. However, if these are same, then weaker is acid, stronger is base, *i.e.*, stronger is nucleophilicity. This acidic character is.

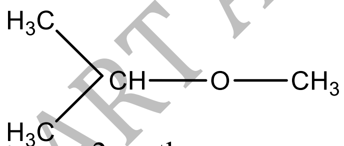
$HI > HBr > HCl > HCN > H_2O > EtOH$

540 **(d)**

Only two unsymmetrical ether is possible from the formula $C_4H_{10}O$ *viz*

$CH_3CH_2CH_2 - O - CH_3$

methoxy propane



2-methoxy propane

541 **(c)**

S_N1 mechanism gives rise to 50% inversion as it involves front seat as well as back seat substitution. This leads to racemic products.

542 **(c)**

Canonical structures proposed in resonance are not the real structure of compound. The compound showing resonance has a definite structure which can however be not drawn on paper.

543 **(d)**

All the names are correct. Options (d) is IUPAC name.

544 **(b)**

On adding SCN^- to an aqueous solution of $Fe(NO_3)_3$, a blood red colour, due to formation of $[Fe(H_2O)_5(SCN]^{2+}$complex is obtained. This test is used for the detection of Fe^{3+}ion.

$$SCN^- + Fe(NO_3)_3 + 5H_2O \rightarrow$$
$$[Fe(OH_2)_5(SCN)]^{2+} + 3NO_3^-$$

Blood red colour

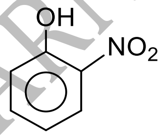
545 **(d)**

In case of alkyl carbocations as the number of R group decreases stability decreases. Thus, the correct order of stability of carbocation is

$$R_3C^+ > R_2\overset{+}{C}H > R\overset{+}{C}H_2 > \overset{+}{C}H_3$$

546 **(b)**

Ortho nitrophenol is the most acidic because electron withdraeing group increases acidic character due to $-I$ effect of NO_2



o - nitrophenol

548 **(a)**

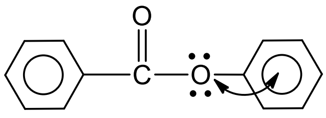
$CCl_2 = CCl_2$ has ethene like structure (*i.e.*, sp^2-hybridization); CCl_4 has CH_4 like structure, *i.e.*, sp^3-hybridization.

549 **(b)**

$C_6H_5C \equiv N$ and $C_6H_5N \equiv C$ are functional isomers.

551 **(c)**

Second ring is in conjugation with lone pair of oxygen



conjugation

552 **(d)**

$$\%\text{ of N} \frac{28}{22400} \times \frac{\text{volume of } N_2 \text{at NTP}}{\text{wt. of compound}} \times 100$$

$$= \frac{28}{22400} \times \frac{224}{1.18} \times 100$$

$$= \frac{28}{1.18} = 23.728$$

553 **(d)**

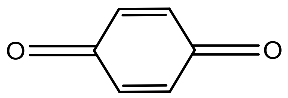
In (II) and (IV) lone pair is involved in resonance.

554 **(d)**

A primary carbon is one which is joined to 1 carbon atom. A secondary carbon atom is joined to two carbon atoms and the tertiary carbon is attached to three carbon atoms. A quaternary carbon has all its four valencies attached to carbon atoms.

555 **(a)**

The compounds which contain active methylene group at the adjacent position of carboxyl group show tautormerism.



This compound does not contain active methylene group, hence does not exhibit tautomerism. Moreover, this compound is highly stable due to extensive cross-conjugation.

557 **(b)**

CH_3CH_2OH and CH_3OCH_3 have different functional groups.

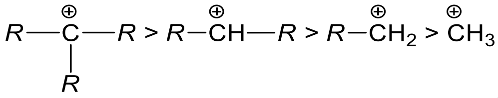
(*ie*., -OH in a alcohol and –O-in ether),hence they are the example of functional isomerism.

558 **(a)**

Two pairs of *cis* and *trans* forms.

559 **(b)**

Stability of alkyl carbocations can be explained by inductive effect and hyperconjugation. According to these two effect the stability order is



$$R-\overset{\oplus}{\underset{R}{C}}-R > R-\overset{\oplus}{C}H-R > R-\overset{\oplus}{C}H_2 > \overset{\oplus}{C}H_3$$

560 **(d)**

In the given compound four π- electrons of double bond and 1 lone pair on N atom leads to delocalization of six electrons

562 **(a)**

Due to resonance, extra stability in 3° carbocation.

564 **(b)**

Follow IUPAC rules.

565 **(b)**

Na reacts with water less violently than K and Rb.

566 **(a)**

$Na_2S + Na_2[Fe(CN)_5NO] \rightarrow Na_4[Fe(CN)_5NOS]$

Sod. extract pink/violet colour

567 **(b)**

Heterolytic bond fission produces +ve and –ve ions.

568 **(b)**

Racemic mixture is formed by mixing two chiral compounds.

569 **(c)**

Distillation is applied if organic liquid is stable at its boiling point and contains a non-volatile impurity

570 **(b)**

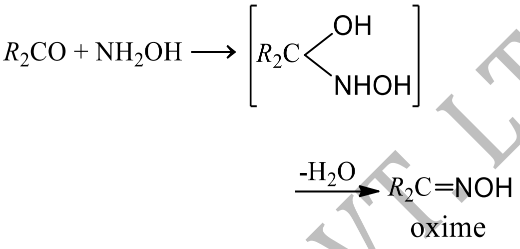
Two double bonds are treated as different functional group with a triple bond.

571 **(b)**

Follow priority rule.

572 **(c)**

Aldehydes and ketones combine with a variety of compounds of the $Z - NH_2$ to form oxime



573 **(b)**

Addition of HCl is not peroxide effect and it occurs via electrophilic addition.

574 **(b)**

$(CH_3)_3C - CH_2OH$ is *neo* -pentyl alcohol.

575 **(d)**

−do −

576 **(d)**

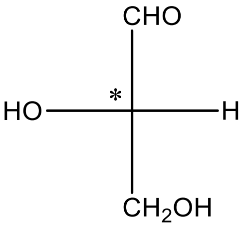
A lot of plant kingdom is made up of cellulose.

579 **(b)**

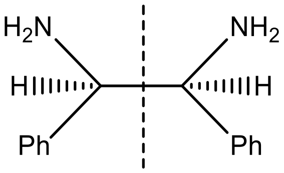
The different arrangement of atoms in space that results from the carbon-carbon single bond free rotation by 360° are called conformations or conformational isomers and this phenomenon is called conformational isomerism.

580 **(a)**

The molecule, which is optically active, has chiral centre, is expected to rotate the plane of polarised light.



One chiral center⟹optically active



Two chiral centres, but plane of symmetry within molecule⟹ optically inactive

583 **(b)**

Note the fact.

585 **(d)**

Alkanes do not show resonance.

586 **(d)**

Both $CHCl_3$ and CCl_4 burn with smoky flame although both are aliphatic. $C_6H_5CH_2OH$, aromatic but burns with non smoky flame. These are exceptions.

587 **(b)**

BF_3 is an electron deficient compound.

590 **(a)**

Isomers of propionic acid are as

$$\overset{O}{\overset{\|}{CH_3-C-OCH_3}} \quad \overset{O}{\overset{\|}{H-C-OC_2H_5}}$$

methyl ethanoate ethyl formate

591 **(c)**

Heterolytic bond fission give rise to formation of ions.

592 **(a)**

This can be judged by comparing the stabilities of carboxylate ions formed. The most stable carboxylate ion is formed by strongest acid

593 **(c)**

CH_3NH_2 and CH_3OH are nucleophiles, CH_3-Cl is an electrophile. But $CH_3-\overset{\delta+}{C}\equiv\overset{\delta-}{N}\colon$ is a nucleophile due to the presence of a lone pair of electrons on N and is an electrophile due to the presence of a partial positive charge on C

594 **(a)**

No bond around chiral carbon is broken and so configuration will be retained.

595 **(b)**

Due to resonance in benzene.

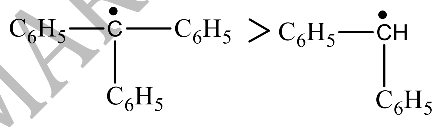
596 **(c)**

Kjeldahl's method is used for estimation of nitrogen

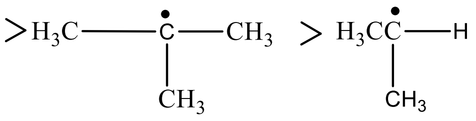
597 **(a)**

Free radicals stability

Free radicals stability



Highly stable by delocalisation



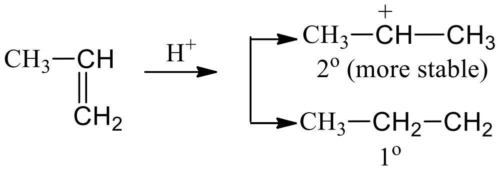
9-hyperconjugative
hydrogens and +I effect

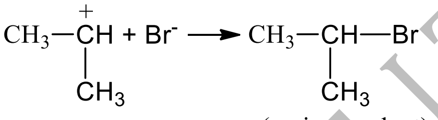
598 **(b)**

HBr being better source of proton. It gives a H^+ and a Br^- ion

$HBr \longrightarrow H^+ + Br^-$

Thus, H^+attack the π bond of propene to form carbonium ion as





(major product)

599 **(c)**

Follow IUPAC rules.

600 **(d)**

Report prefixes in alphabetic order.

601 **(c)**

Sulphur is present in the sodium extract in the form of sodium sulphide (Na_2S)

$FeCl_3$ gives blood red colour with sodium extract contain N and S

602 **(a)**

The name of the compound

(2Z,4Z)-2, 4-hexadiene.

603 **(c)**

It is a fact.

604 **(a)**

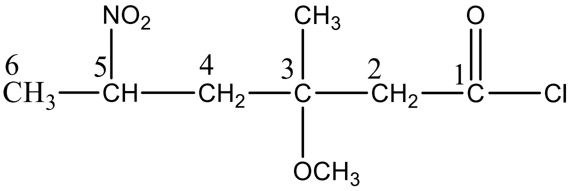
The number of stereoisomers=2’

(Here, n =chiral carbon atom)

Thus, number of stereoisomers=2^3=8

605 **(a)**

The structural formula of the compound 5-nitro-3-methoxy-3-methyl hexanoyl chloride is as



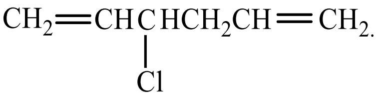
606 **(d)**

Presence of methyl group on NH_3 molecule increases the tendency of N atom to lose electron pair. However, tertiary, amines are less basic due to steric hindrance.

607 **(a)**

$CHCl = CHCH_2CH_2CH = CH_2$;

$CH_2 = CClCH_2CH_2CH = CH_2$;



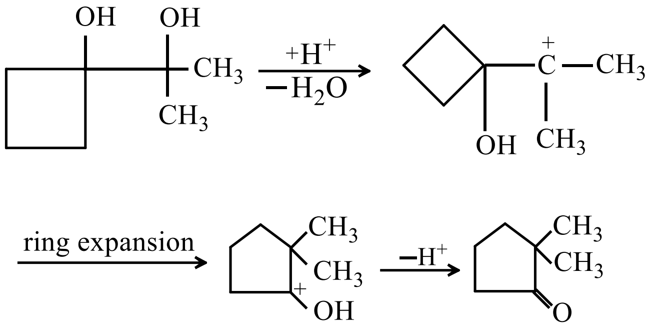
$CH_2{=}CH\underset{\displaystyle Cl}{\underset{|}{C}}HCH_2CH{=}CH_2.$

608 **(c)**

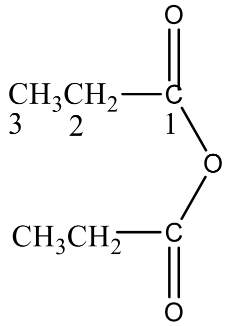
2° carbocation is more stable.

609 **(b)**

In such cases where migrating group is cycloalkyl group, ring expansion may occur.



610 **(d)**



propanoic anhydride

In IUPAC system, anhydrides are named as alkanoic anhydride.

611 **(c)**

Stronger is an acid, weaker is its conjugate base or weaker is nucleophile. The acidic character order is :

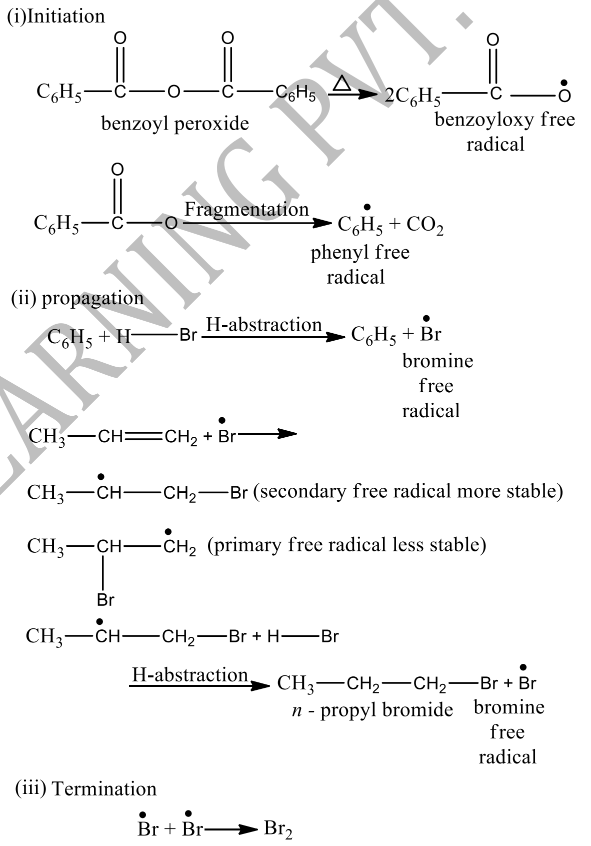
$HF > H_2O > NH_3 > CH_4$.

613 **(c)**

$$CH_3-CH=CH_2+HBr\xrightarrow[\text{peroxide}]{\text{Organic}}CH_3-CH_2-CH_2-Br$$

requires radical intermediate.

Mechanism



614 **(a)**

Chiral molecules should not contain any kind of symmetry.

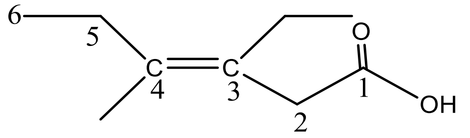
615 **(d)**

It is honour to Lavoisier.

616 **(d)**

Stability of carbanions increase with increasing in s-character of hybrid orbitals of carbon bearing charge therefore, the order is $sp^3 < sp^2 < sp$

617 **(c)**



E-3-ethyl-4-methyl hex-3-en-1-oic acid.

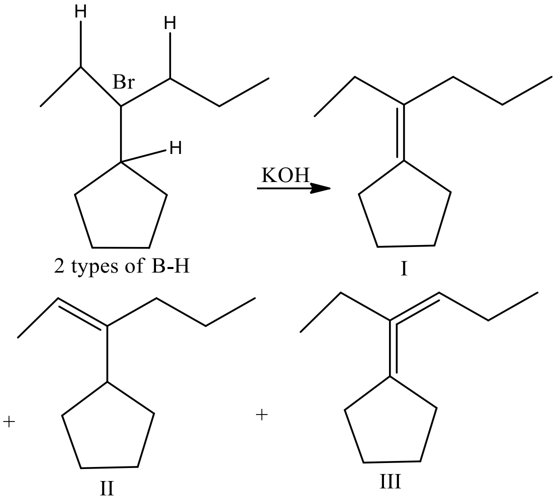
[The configuration of this compound is E because bulkier groups are present at opposite of the double bond.]

618 **(a)**

−do −

619 **(c)**

The substrate has three different types of B − H, therefore, first, three structural isomers of alkenes are expected as



The last two alkenes II and III are also a capable of showing geometrical isomerism hence two geometrical isomers for each of them will be counted giving a total of five isomers.

620 **(c)**

CH_2Cl_2 has sp^3 -hybridization and tetrahedral nature.

621 **(a)**

Chain initiation step involves formation of free radicals only.

622 **(d)**

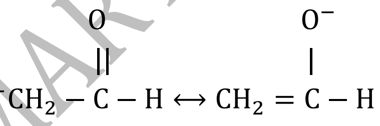
Follow characteristics of S_N2 mechanism.

623 **(a)**

Follow elimination of HBr from two ends.

624 **(d)**

$^{\ominus}CH_2CHO$ is the most stable carbanion since it is stabilize by resonance



$$^{-}CH_2 - \overset{\overset{\displaystyle O}{\|}}{C} - H \leftrightarrow CH_2 = \overset{\overset{\displaystyle O^-}{|}}{C} - H$$

625 **(b)**

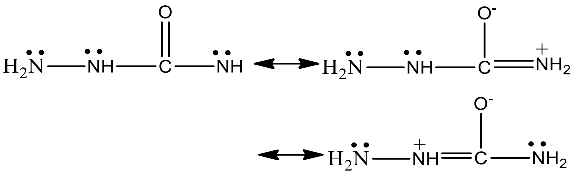
X-crown-Y, 18-crown-6

First number X is the total number of ‘C’ and ‘O’ atoms in the ring and second number Y is the number of oxygen atom in ring.

626 **(b)**

When the nucleophilic site is the same, nucleophilicity parallels basicity. It means more basic the nucleophile, stronger is the nucleophile.

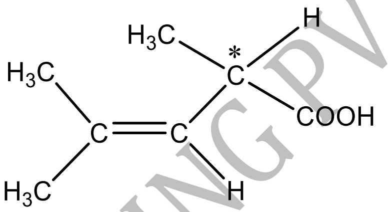
$H_2\ddot{N}$ (I) is the most nucleophilic



Furthermore the NH_2 group away from the $-C-$group is not involved in resonance. Hence, its lone pair is reading available.

627 **(c)**

A compound could be optically active only when it contains, at least one asymmetric carbon atom or a chiral centre.



629 **(d)**

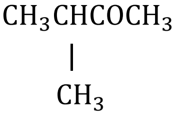
$$C:H:Cl:O = \frac{18.5}{12} : \frac{1.55}{1} : \frac{55.04}{35.5} : \frac{24.81}{16} = 1:1:1:1$$

630 **(b)**

2-pentanone and 3-methyl-2-butanone are chain isomers because they differ in carbon skeleton.

$CH_3CH_2CH_2COCH_3$

2-pentanone



$$CH_3\underset{\underset{\displaystyle CH_3}{|}}{C}HCOCH_3$$

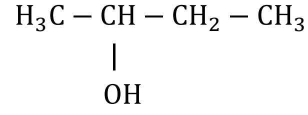
3-methyl-2-butanone

631 **(c)**

Isomers of $C_4H_{10}O$ are as follows

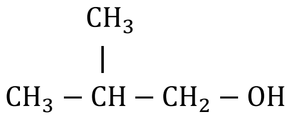
(i) $H_3C - CH_2 - CH_2 - CH_2 - OH$

butanol-1



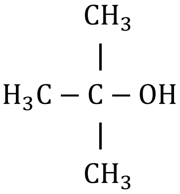
(ii) $$H_3C - \underset{\underset{\displaystyle OH}{|}}{CH} - CH_2 - CH_3$$

butanol-2



(iii) $$CH_3 - \overset{\overset{\displaystyle CH_3}{|}}{CH} - CH_2 - OH$$

2-methyl propanol -1

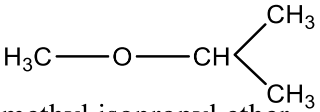


(iv) $$H_3C - \overset{\overset{\displaystyle CH_3}{|}}{\underset{\underset{\displaystyle CH_3}{|}}{C}} - OH$$

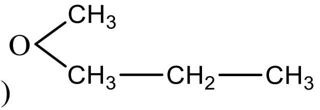
2-methyl propan -2-ol

(v) $H_3C - CH_2 - O - CH_2 - CH_3$

Diethyl ether

(vi) 

methyl isopropyl ether

(vii) 

methyl propyl ether

Hence, three isomeric ethers are possible.

632 **(c)**

Liebig's method is used to estimate carbon and hydrogen.

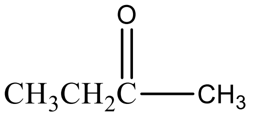
C and H $\xrightarrow{[O]} CO_2 + H_2O$

$$\%C = \frac{12}{44} \times \frac{\text{weight of } CO_2}{\text{weight of compound}} \times 100$$

$$\%H = \frac{2}{18} \times \frac{\text{weight of } H_2O}{\text{weight of compound}} \times 100$$

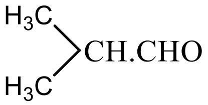
633 **(d)**

There are six isomers possible for the compounds having molecular formula C_4H_8O, which are as follows

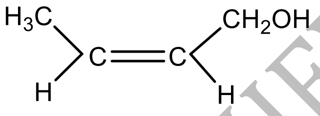


$CH_3CH_2\overset{O}{\overset{\|}{C}}-CH_3$, $CH_3CH_2CH_2CHO$,

(i) (ii)

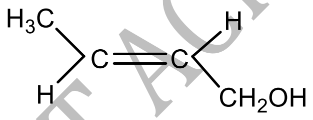


(iii)



(*cis*)

(iv)



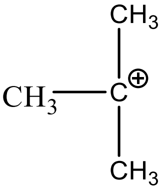
(*trans*)

(v)

and $CH_2=CH-CH_2CH_2OH$

(vi)

634 **(c)**

C_2H_5 and  are stabilised hyperconjugation. In $C_6H_5\overset{\oplus}{C}H_2$ benzyl group is resonance stabilised. In the triphenyl methyl carbonium ion, the π-electrons of all the three benzene rings are delocalised with the vacant p-orbital of central carbon atom. So, it is resonance stabilised. Therefore, it is most stable of the given carbonium ions.

More the number of resonatic structures more will be the stability.

635 **(d)**

Chlorine atoms are strongly electro negative (show negative inductive effect i.e., $-I$ effect).

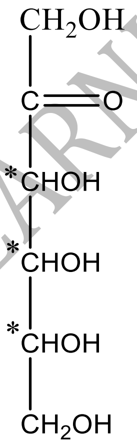
They deactivate the ring towards electrophilic reaction.

The increasing order of substituent-E towards electrophilic substitution is

$-CCl_3 < -CHCl_2 < -CH_2Cl < -CH_3$

636 **(b)**

Fructose has three chiral carbon atoms, hence the number of optical isomerism $= 2^3 = 8$



637 **(d)**

O has two unpaired electrons in ground state.

638 **(d)**

An organic ion with a pair of available electrons and a negative charge on the central carbon atom is called a carbanion.

Electron attracting group $\left(-CN, >C=O\right)$ increases stability and electron releasing group $(-CH_3 etc)$ decreases the stability of carbanion. In $(CH_3)_3C^-$, three – CH_3 groups (electron releasing group) are present, so it least stable.

639 **(d)**

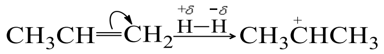
The stability order of conformation of cyclohexane is chair>twist boat>boat>half chair. Hence, half chair is less stable due to torsional and angle strain.

640 **(a)**

It is a fact.

642 **(a)**

Formation of 2° carbocation, $i.e.$,



$CH_3CH=CH_2 \xrightarrow{\overset{+\delta}{H}-\overset{-\delta}{H}} CH_3\overset{+}{C}HCH_3$

643 **(d)**

In CH_3CH_2OH, there is intermolecular H-bonding, while it is absent in isomeric ether CH_3OCH_3

1. Larger heat is required to vaporise CH_3CH_2OH as compared to CH_3, OCH_3, thus (a) is incorrect.

2. CH_3CH_2OH is less volatile than CH_3OCH_3, thus vapour pressures are different, thus (b) is incorrect.

3. Boiling point of $CH_3CH_2OH > CH_3OCH_3$, thus (c) is incorrect.

Density $= \frac{\text{mass}}{\text{volume}}$, due to ideal behaviour at a given temperature and pressure volume and molar mass are same.

Hence, they have same vapour density.

644 **(a)**

Follow IUPAC rules.

646 **(b)**

% of chlorine $= \frac{35.5}{143.5} \times \frac{\text{mass of AgCl}}{\text{mass of the compound}} \times 100$

$= \frac{35.5}{143.5} \times \frac{0.287}{0.099} \times 100$

$= 71.71\ \%$

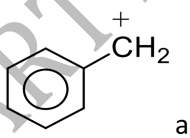
647 **(c)**

Due to presence of delocalised π-electrons in the aromatic compounds, the electron density is maximum inside the ring. Therefore, aromatic compounds undergo electrophilic substitution reaction and resistance to addition reactions.

648 **(a)**

The acidic nature is $H_2O > C_2H_2 > C_2H_4 > C_2H_6$.

649 **(b)**

Benzyl  allyl $\overset{+}{CH_2}=CH-\overset{+}{CH_2}$

methyl CH_3,vinyl $CH_2=\overset{+}{CH}$ given options can be solved on the basis of conjugative and hyperconjugative structures

650 **(a)**

Follow IUPAC rules.

652 **(a)**

$(CH_3)_3C-C(CH_3)_3$ has maximum number of alkyl groups in it.

653 **(d)**

Halogenation of methane is chain reaction, propagate through free radical

655 **(a)**

Follow IUPAC nomenclature.

658 **(c)**

∵ C – H bond in toluene has partial double bond character due to resonance.

∴ C – H bond in toluene has less energy as compared to others.

659 **(b)**

The reaction $(CH_3)_3C-Br \xrightarrow{H_2O} (CH_3)_3COH$ is an example of substitution reaction

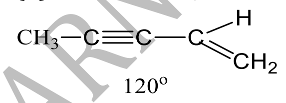
660 **(b)**

Draw the isomers.

661 **(b)**

Distillation is used to purify liquids and based on difference in their boiling points. When the boiling points of liquids are very closed to each other, then fractional distillation is used.

662 **(c)**



The C-atoms attached to the triple bond lie in a straight line while the carbon of the CH_2 group is inclined at an angle of 120°. Therefore, only 2 carbon atoms are linearly arranged

663 **(c)**

The conversion of α-glucose to β-glucose is called mutarotation.

664 **(c)**

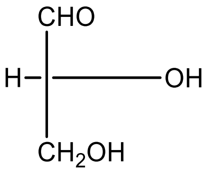
Positively charged species in which central atom has incomplete octet is called electrophile, H^+, X^+, R^+ are electrophile

665 **(b)**

An organic reaction intermediate, neutral species having divalent carbon atom with six valence electron out of which two are present in same orbital with opposite spin is called singlet carbene

666 **(a)**

The structure of D-glyceraldehyde is as

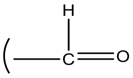


The priority of groups is decided by the following rules

(i) Atom having higher atomic number gets higher priority

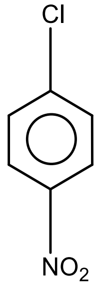
(ii) If the priority cannot be decided by rule 1 then the next atoms are considered for priority assignment.

(ii) Where there is a = bond or=bond both atoms are considered to be duplicated or triplicated

( has higher priority than $-CH_2OH$).
Hence, the correct order of priority of groups in D-glyceraldehyde is as:

$\mathrm{OH}(1), \mathrm{CHO}(2), \mathrm{CH_2OH}\ (3)$and $\mathrm{H}(4)$

668 **(c)**
Nucleophile always attacks on electron deficient site. Presence of electron withdrawing groups such as NO_2, CHO etc decreases the electron density on benzene nucleus, hence such groups activate the ring towards nucleophilic attack.
While presence of electron releasing groups such as R or OR increases the electron density, thus deactivates the nucleus towards nucleophilic attack.
NO_2 group activates the ring more than Cl towards nucleophilic attack,
Hence reacts readily with nucleophile.



669 **(d)**
The increasing order is :
$-CH_3 < CH_3-CH_2- < (CH_3)_2CH- < (CH_3)_3C-$

670 **(a)**
The treatment with $FeCl_3$ yield ferric ferrocyanide which has bulish green or prussian blue colouration
$3Na_4[Fe(CN)_6] + 4FeCl_3 \rightarrow Fe_4[Fe(CN)_6]_3$

671 **(c)**
Each π-bond contributes two πelectrons and the two electrons of the lone pair are present in a p-orbital. Therefore, total number of π electrons is six

672 **(a)**
Allyl carbocations are more stable than the alkyl carbocations due to the resonance stabilization

673 **(b)**
The middle carbon has 2σ-and 2π-bonds.

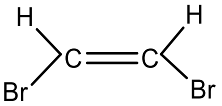
675 **(c)**
Di-chloro acetic acid due to presence of two electron with drawing chloro groups
($-I$ showing group)is more acidic than acetic acid($+I$ showing $-CH_3$group).

677 **(b)**
$Na_3PO_4 + 3HNO_3 \rightarrow H_3PO_4 + 3NaNO_3$
sod. extract
$H_3PO_4 + 12(NH_4)_2MoO_4 + 21HNO_3$
$\rightarrow (NH_4)_3PO_4.12MoO_3 + 21NH_4NO_3 + 12H_2O$
yellow ppt.

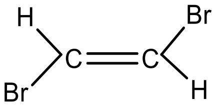
678 **(b)**
1,2,3; 1,2,4 and 1,3,5-trimethyl benzene.

679 **(b)**
Hydride ion is formed when hydrogen accept a proton, so it has a tendency to donate electron. Since, hydride ion (H^-) has a tendency to donate electron, it functions as nucleophile.

680 **(b)**
The structure of isomers from $C_2H_2Br_2$ are
$CH_2 = CBr_2$;
$CH_2{=}CBr$;
I



cis-isomer II

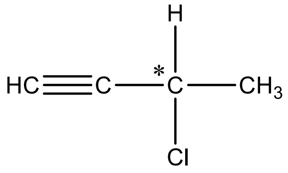


trans-isomer III

681 **(a)**
$CH_3MgX + CH_3C \equiv \underset{\text{Acidic}}{C-H} \longrightarrow CH_4 + CH_3C \equiv C.MgX.$

682 **(a)**
The correct decreasing order of priority for the functional groups of organic compounds in the IUPAC system is
$-COOH > SO_3H > -COOR > COCl > -CONH_2 > -CN > -CH=O$

683 **(b)**
Optical isomerism is shown by compounds which have one or more chiral carbon atoms.



(b)

∵ It has asymmetric or chiral carbon atom.

∴ It shows optical isomerism.

684 **(d)**

S_N2 reaction does not involve ion formation, these infact involve formation of transition state

685 **(a)**

pK_a value of carboxylic group is less than pK_a of $\overset{+}{NH_3}$ in amino acid and $—\overset{+}{NH_3}$ (Z) will have comparatively less pK_a than $—\overset{+}{NH_3}$ (Y) due to –I effect of carboxylic group. We know that acidic strength in inversely proportional to pK_a. Hence, correct order of acidic strength is

$$—CO_2H(X) > —\overset{+}{NH_3}(Z) > —\overset{+}{NH_3}(Y)$$

686 **(a)**

Chromatography is a modern technique used for the separation of mixtures into its components, purification of compounds and also to test the purity of compounds.

688 **(d)**

Kharasch effect involves addition of HBr.

689 **(d)**

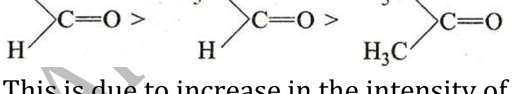
These are characteristics of carbocations.

691 **(b)**

Symmetrical *trans*-form has non-polar nature.

694 **(c)**

The nucleophilic addition reaction is the characteristic addition of carbonyl compounds. Reactivity order of carbonyl compounds is in the order.



This is due to increase in the intensity of charge on carbon of carbonyl group due to $+I$ effect of alkyl groups.

695 **(c)**

Chlorobenzene has only one deactivating group, .*e*., −Cl. In 2, 4-dinitrochlorobenzene three deactivating group, *i.e.*, two – NO_2 and one – Cl are present and p −nitrochlorobenzene two deactivation groups, *i.e.*, one $NO_2$and one Cl is present. So, the order of reactivity is *A>C>B.*

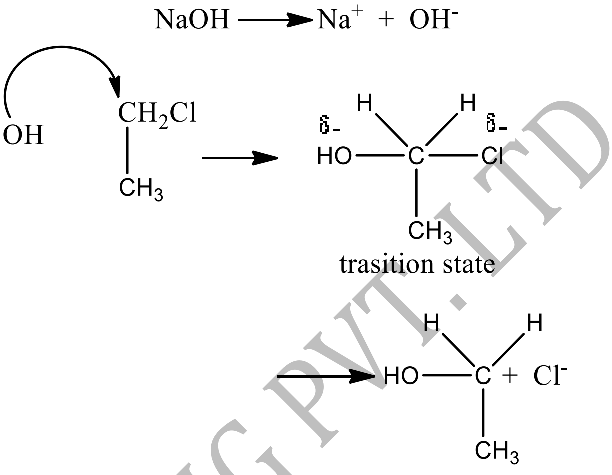
696 **(a)**

Acid has —COOH group whereas, ester has —COO*R* group.

697 **(d)**

The given reaction can be represented as

$NaOH \rightarrow Na^+ + OH^-$



Since in this reaction, a nucleophile replaces the other group, it is a example of nucleophilic substitution reaction.

The mechanism shows that the rate depends on the concentration of both alkyl halide and nucleophile. So, it is an example of S_N2(nucleophilic substitution of II order) reaction.

698 **(a)**

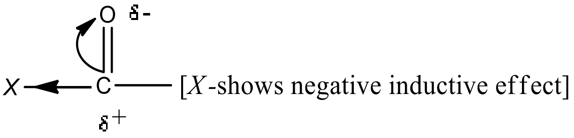
Follow IUPAC rules.

699 **(d)**

Follow IUPAC rules.

702 **(a)**

Carbonyl compounds undergoes nucleophilic addition reaction.



If group or atom attached with carbonyl carbon shows negative inductive effect, then it decreases electron density on carbonyl carbon and facilitate the attack of nucleophile, hence reactivity of carbonyl compound increases. The aromatic aldehydes and ketones are less reactive than their aliphatic analogues due to +R effect of benzene ring. The increasing order of the nucleophilic addition reaction in the following compounds will be.

$CH_3CHO > CH_3COCH_3 > PhCOCH_3 > PhCOPh$

703 **(b)**

Chiral compounds which have one chiral centre. All four atoms or groups attached to carbon are different.

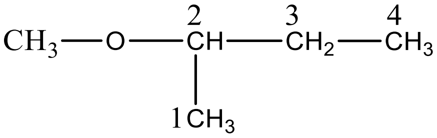
704 **(b)**

Due to the presence of methyl group positive inductive effect increases and the stability of carbocation also increases. The stability order of carbocation is

Tertiary > Secondary > Primary

705 **(b)**

According to IUPAC system ethers are named as alkoxy alkanes. The larger alkyl group forms the parent chain while lower alkyl group is taken with the ethereal oxygen and forms a part of alkoxy group.



$CH_3-O-\underset{\underset{CH_3}{|}}{CH}-CH_2-CH_3$

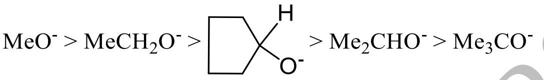
2-methoxybutane

706 **(c)**

H_3^+O cannot accept electron pair.

707 **(a)**

Epoxide is ambident substrate for nucleophilic substitution reactions. In protonated epoxide carbon-2 and carbon-3 both aquire some positive charge due to the highly electronegative atom



708 **(d)**

Beilstein test is used to detect halogens in organic compounds.

709 **(b)**

Free radical chain reaction is initiated by UV light. It proceeds in three main steps like initiation, propagation and termination. It gives major products derived from most stable free radical

711 **(b)**

Follow IUPAC rules.

713 **(d)**

Elimination reactions involves removal of a molecule (HBr here) from a substrate.

714 **(c)**

O OH

‖ |

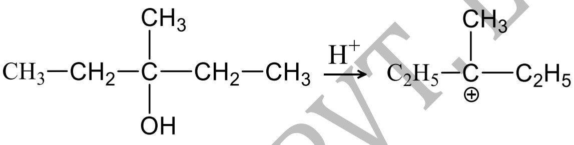
$CH_2CH_2-CH_2-C-CH_3 \rightleftharpoons CH_3CH_2CH_2-C$
$= CH_2$

ketoform enol form

715 **(b)**

The species which are electron deficient and accept a pair of electron are called electrophile. Hence, SO_3 is a electrophile as it contains an electron deficient centre. While H_2O, NH_3 *and* $R-O-R$ are nucleophiles.

718 **(a)**



In the above reaction more stable carbocation is generated hence, the compound dehydrated very easily

719 **(a)**

In case of kjeldahl's method the percentage of N_2 is then calculated from the amount of NH_3

721 **(b)**

Such dehydrohalogenation follow E_2 mechanism. The driving force of such reaction is the stability of alkene produced. Since, tertiary alkyl halide can give more substituted alkene, it reacts fastest followed by secondary and primary *i.e.*, 3° > 2° > 1°.

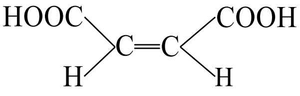
722 **(c)**

Central carbon atom is chiral carbon.

723 **(d)**

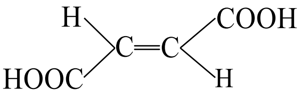
Those organic compounds, which are volatile in steam are purified by steam distillation. Since, aniline is a steam volatile compound, hence it is purified by steam distillation.

724 **(b)**



Maleic acid

and



Fumaric acid

are geometrical isomers.

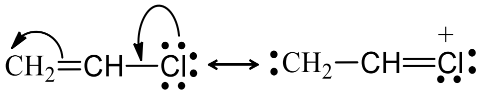
725 **(b)**

$C_6H_5O^-$ possess less nucleophilicity due to stabilized nature of phenoxide ion. CH_3OH is weaker acid than CH_3COOH and thus CH_3O^- is stronger base.

Acidic order : $CH_3COOH > H_2O > CH_3OH$

726 **(c)**

Vinyl chloride is least reactive for S_N reaction due to resonance



729 **(b)**

The chain propagation step involves the use of free radical and regeneration of another free radical.

730 **(b)**

CH_3NC is methaneisonitrile.

732 **(a)**

Benzyl carbonium is more stable due to resonance and thus, benzyl chloride is more reactive.

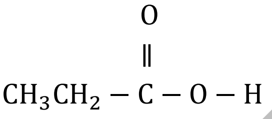
733 **(a)**

It is a fact.

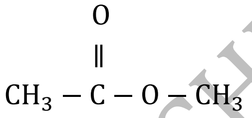
734 **(b)**

$C_nH_{2n}O_2$ is general formula for open chain acid and ester.

$n = 3 \quad C_3H_6O_2$

Acid 

$CH_3CH_2-\overset{\overset{O}{\|}}{C}-O-H$

Easter 

$CH_3-\overset{\overset{O}{\|}}{C}-O-CH_3$

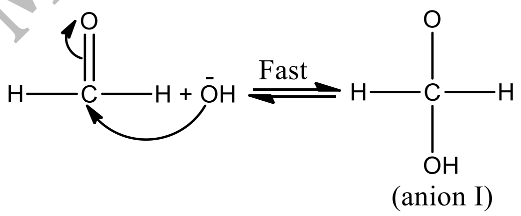
735 **(a)**

The Cannizzaro reaction is as

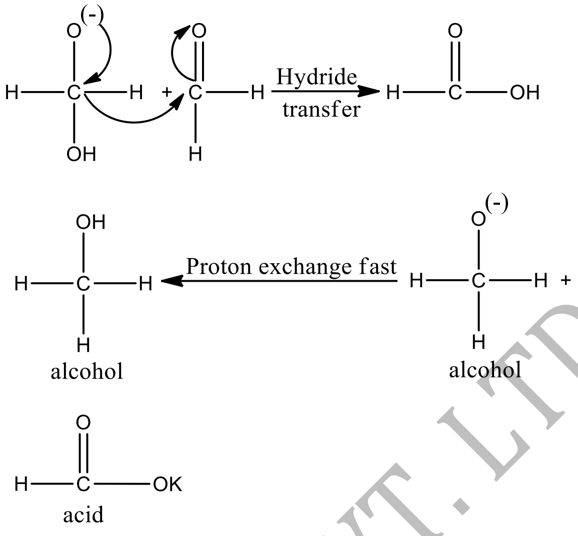
$$HCHO + HCHO \xrightarrow{KOH(conc.)} \underset{\text{methyl alcohol}}{CH_3OH} + \underset{\text{acetic acid}}{HCOOK^+}$$

The mechanism of Cannizzaro reaction is as

Step I Attack of nucleophile OH^- to the carbonyl carbon

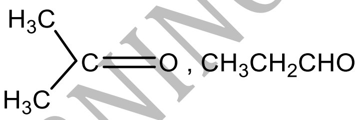


Step II The transfer of hydride ion from anion (I) to second molecule of aldehyde and finally rapid transfer of proton takes place.



737 **(a)**

Propanal and propanone are functional isomers



738 **(a)**

It is a fact.

739 **(a)**

Angle strain, $\alpha = \frac{1}{2}[109°28' - \theta]$

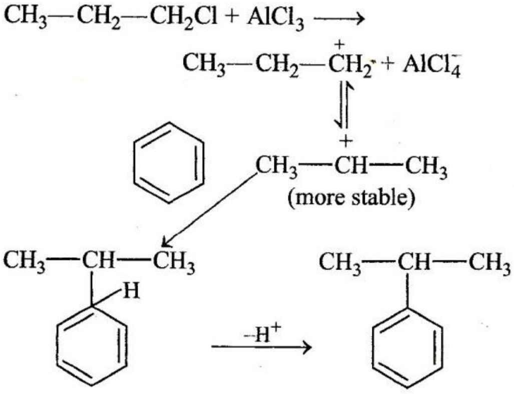
In case of cyclopropane,

$\theta = 60°$

$\therefore \alpha = \dfrac{1}{2}(109°28' - 60°) = 24°44'$

740 **(d)**

The function of $AlCl_3$, in Friedel-Craft reaction, is to produce electrophile, which later add to benzene nucleus



741 **(c)**

Kejldahl's method is used for the estimation of nitrogen. The organic compound is heated with conc. H_2SO_4 in presence of K_2SO_4 (used to elevate boiling point of H_2SO_4) and $CuSO_4$ (used as catalyst) to convert all the nitrogen into $(NH_4)_2SO_4$.

742 **(c)**

Acetone and methanol have nearly equal boiling point. thus, they are separated by fractional distillation

743 **(b)**

Follow IUPAC rules.

744 **(c)**

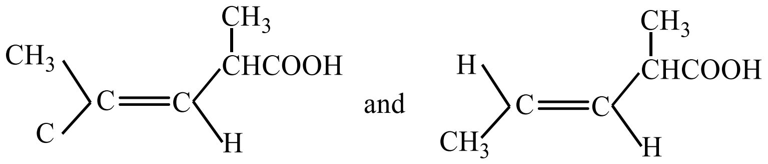
The oxygen atom in phenol has more dominating resonance effect than inductive effect. Increase in charge separation decreases the stability of a resonating structure

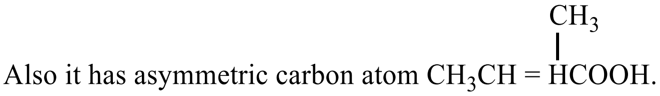
Stability of resonating structure in decreasing order will be

I > II ≡ IV > III

745 **(c)**

The acid exist in *cis* and *trans* forms :

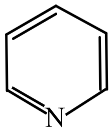



Also it has asymmetric carbon atom $CH_3CH = \overset{CH_3}{\overset{|}{H}}COOH$.

746 **(a)**

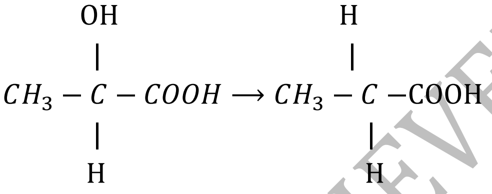
Follow the mechanism of esterification.

747 **(b)**

Pyridine  is a heterocyclic compound having six membered ring formed with C and N-atoms.

748 **(a)**

When – OH group of lactic acid is replaced by H, then chiral carbon is lost.


$$\underset{\text{lactic acid}}{CH_3 - \overset{OH}{\overset{|}{\underset{H}{\underset{|}{C}}}} - COOH} \longrightarrow CH_3 - \overset{H}{\overset{|}{\underset{H}{\underset{|}{C}}}} - COOH$$

lactic acid

∴Its optical activity is lost.

750 **(a)**

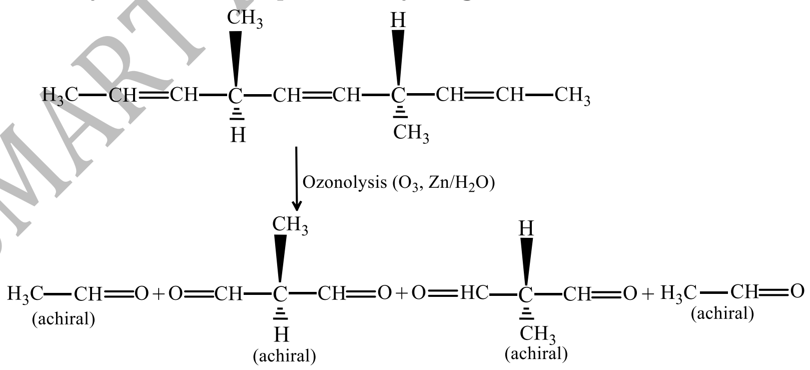
2-butene exhibit rotamers. Rotamers are the isomers formed by restricted rotation.

751 **(d)**

It contains lone pair electron on N atom.

752 **(a)**

Ozonolysis of the compound may be given as :



753 **(b)**

2-aminopentane and 3-aminopentane; Position is different.

755 **(c)**

$$\% \text{ of H} = \frac{2}{18} \times \frac{\text{weight of } H_2O}{\text{weight of organic compound}} \times 100$$

$$= \frac{2}{18} \times \frac{0.9}{0.5} \times 100 = 20\%$$

$\therefore$ The percentage of carbon $= 100 - 20 = 80\ \%$

756 **(b)**

$o-$ and p-directing groups facilitate S_E reactions whereas m-directing groups deactivate benzene ring for S_E reactions.

757 **(a)**

(+) and (-) tartaric acid does not possess any element of symmetry.

758 **(b)**

A molecule having doubly bonded carbon atoms shows geometrical isomerism only if both the doubly bonded carbon have altogether different group, *i.e.*, $_{ba}C \equiv C_{ab}$ or $_{ab}C = C_{ac}$ or $_{dc}C = C_{ab}$.

759 **(c)**

The chemical formula of thiourea is NH_2CSNH_2 so here Na_2S, NaCN and NaCNS will be formed but not Na_2SO_4

760 **(a)**

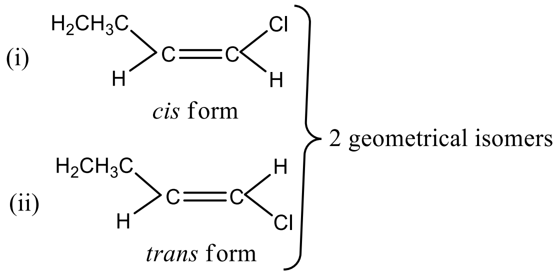
A white precipitate with am. $AgNO_3$ confirms the presence of terminal alkyne.

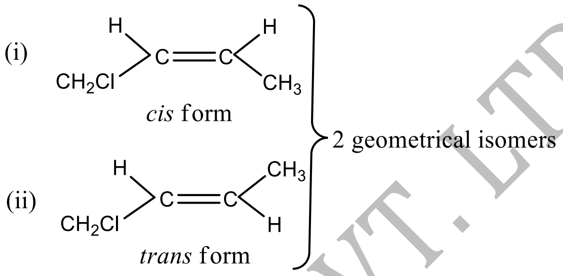
761 **(a)**

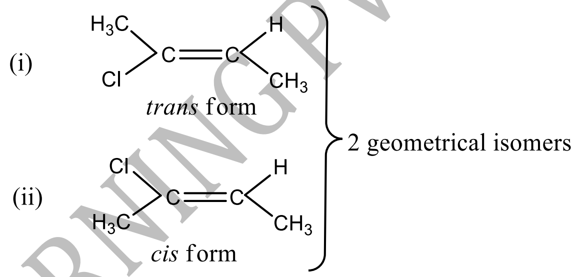
Racemisation involves change in entropy, *i.e.*, change in arrangement of groups position leading to a change in entropy of disorderness.

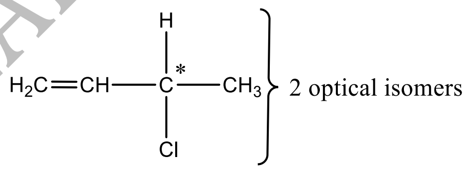
762 **(a)**

The acyclic stereoisomers of C_4H_7Cl are







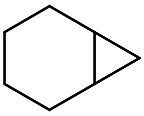


Number of optical isomers $=2^n = 2^1 = 2$

Hence, total number of geometrical isomers=6

Total number of optical isomers =2.

763 **(c)**



bicyclo (4, 1, 0) heptane

This compound contains 7 carbon atoms, so the corresponding alkane is heptane. Two bridges contain 4 and 1 carbon atom respectively and one bridge does not contain any carbon atom. So the name of the compound is bicyclo (4,1,0) heptane.

764 **(d)**

Stability of alkyl free radicals can be explained by hyperconjugation and number of resonating structure due to the hyperconjugation. The decreasing order of stability of alkyl free radical is as follows

3° free radical $>$ 2° free radical $>$ 1° free radical $>$ $\dot{C}H_3$

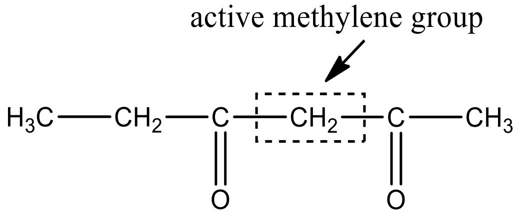
766 **(b)**

Inductive effect involves only displacement (and not delocalisation) of σ −electrons.

767 **(c)**
Meso forms are optically inactive as they are superimposable to their mirror images.

768 **(b)**
CH_4 has highest ratio of H to C

769 **(b)**



$H_3C—CH_2—\overset{\overset{O}{\|}}{C}—CH_2—\overset{\overset{O}{\|}}{C}—CH_3$

2, 4-hexanedione

When methylene group $(-CH_2)$ is attached with two electron withdrawing groups (like, $-CHO, > C = O, -COOH, -CN, -X, etc$), its acidity will increase due to $-I$ effect of the electron withdrawing groups.

770 **(a)**
Follow IUPAC rules.

771 **(c)**
The reactivity order for H atom is $3° > 2° > 1°$; Neocarbon does not have H atom.

773 **(b)**
— do—

774 **(a)**
Organic compound which are volatile in steam can be purified by steam distillation. It is based on the fact that vaporisation of organic liquid takes place at lower temperature than its boiling point

775 **(b)**
Follow IUPAC rules.

777 **(d)**
$CH_3\bar{O}$ is nucleophile;

$$CH_3OH + Na \longrightarrow CH_3\overset{-}{O}\overset{+}{Na} + (1/2)H_2.$$

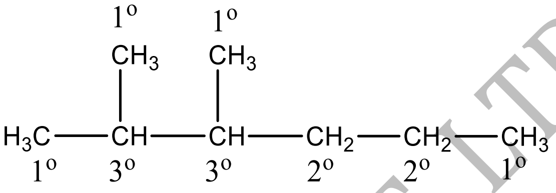
778 **(a)**
Inductive effect is the permanent effect on σ −electrons. It involve the electron displacement along the chain of saturated carbon atoms due to the presence of a polar covalent bond at one end of the chain.

779 **(c)**
Homologous differ by a group $—CH_2$and cannot be isomer.

780 **(c)**
The reagent selected should be such that only one of components to be separated, reacts with it.
Aniline+aq. HCl→salt, which is water soluble
Nitrobenzene +aq. HCl→no reaction
∴ aq. HCl is used to separate aniline and nitrobenzene.

781 **(b)**
Formic acid was obtained from ant (fromica in greek). This is trivial name for HCOOH.

782 **(a)**
The structure of 2, 3-dimethyl hexane is



$\underset{1°}{H_3C}—\underset{3°}{\overset{\overset{1°}{CH_3}}{\overset{|}{CH}}}—\underset{3°}{\overset{\overset{1°}{CH_3}}{\overset{|}{CH}}}—\underset{2°}{CH_2}—\underset{2°}{CH_2}—\underset{1°}{CH_3}$

So, the number of tertiary carbon atoms=2
The number of secondary carbon atoms=2
The number of primary carbon atoms=4

783 **(a)**
Follow IUPAC rules.

786 **(c)**
CH_3^+ has planar structure.

787 **(d)**
These are characteristics of carbanion.

788 **(a)**
Follow Saytzeff rule for elimination. 3-halopentane will give only pentene-2.

789 **(b)**

Atom	**Atomic Mass (a)**	**Percentage (b)**	$\frac{b}{a} = x$	**Ratio**
C	12	10.06	$\frac{10.06}{12}$	1
H	1	0.84	$\frac{0.84}{1}$	1
Cl	35.5	89.10	$\frac{89.10}{35.5}$	3

Empirical formula $= CHCl_3$
Empirical formula mass $= 12 + 1 + 106.5 = 119.5 \approx 120$
Molecular mass $= 2 \times V.D = 2 \times 60 = 120$

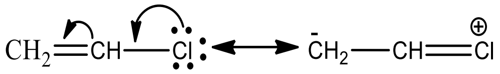
$$n = \frac{\text{molar mass}}{\text{empirical formula mass}}$$

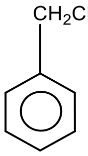
$$= \frac{120}{120} = 1$$

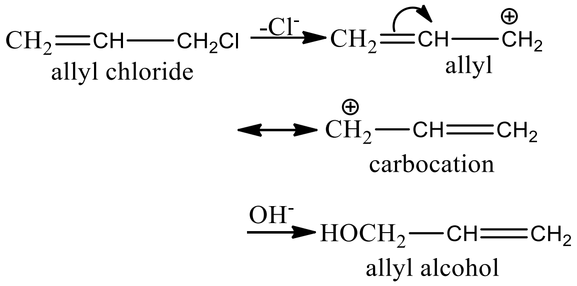
Molecular formula $= (CHCl_3)_1 = CHCl_3$

790 **(d)**
During nucleophilic substitution weaker nucleophile is replaced by stronger nucleophile. The compound having C-Cl bond which can be most easily broken will be most reactive towards nuclophilic substitution reaction.
In vinyl chloride $CH_2 = CH - Cl$ and chlorobenzene C_6H_5Cl the $C - Cl$ bond has partial double bond character due to resonance.

∴ They do not give nucleophilic substitution reaction easily

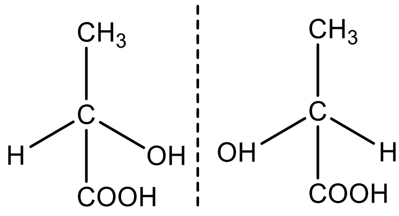


Benzyl chloride,  give nucleophilic substitution easily because they carbocation formed is stabilised due to resonance.



791 **(a)**

Enantiomers are non-superimposable mirror images, *e.g,* lactic acid



Diastereomers are non-superimposable and are not the mirror images of each other. Moreover, *meso* form has plane of symmetry.

792 **(b)**

Nucleophilic strength increases down a column of the Periodic Table (in solvents that can have hydrogen bonds, such as water, alcohols, thio alcohols).

Nucleophilic strength $RO^- < RS^-$

Base strength $RO^\ominus > RS^-$

Thus, $RO^\ominus$ is more nucleophilic but less basic than RO^-

795 **(a)**

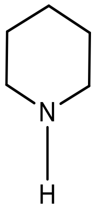
We know that there are seven isomers in $C_4H_{10}O$. Out of these seven isomers, four are of alcohol and three are of ether.

796 **(a)**

Tertiary halide always favours S_N1 mechanism (as they give comparatively stabler carbocation) white primary halide favours S_N2 mechanism.

797 **(d)**

Electron donors are bases. Since, electron density is highest at



(Piperidine), hence, it is most basic.

798 **(d)**

Follow IUPAC rules.

800 **(c)**

To be optically active, compound or structure should posses a chiral or asymmetric carbon atom. 1-chloropentane is not chiral.

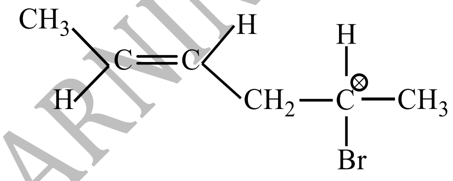
801 **(c)**

Stearic hinderance in tertiary halides give rise to less reactivity for S_N2.

802 **(b)**

Addition of Br_2 gives altogether different products units *cis* and *trans* butene-2.

803 **(d)**



It has one chiral centre (two enantiomer) and two geometrical isomers

cis–d, *trans–d*, *cis– and trans–l*.

804 **(d)**

Glucose contains four chiral carbon atoms hence number of possible optical isomers are $2^4 = 16$.

805 **(c)**

Markownikoff's rule is for addition of unsymmetrical additive on unsymmetrical alkene.

806 **(d)**

Presence of halogen in organic compound can be detected by Beilstein's test.

807 **(c)**

The bond energy of catenation order is C > Si > S > P.

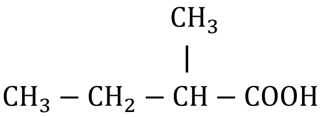
808 **(d)**

$R - X \xrightarrow{\text{NaOH}} R - \text{OH} + \text{Na}X$

$R - X \xrightarrow{\text{OH}^-} R - \text{OH} + X^-$. This is nucleophilic substitution.

809 **(c)**

2-methyl butanoic acid exhibits stereo isomerism.



$$CH_3 - CH_2 - \overset{\overset{\displaystyle CH_3}{|}}{CH} - COOH$$

It shows optical isomerism because it contains asymmetric carbon atom.

810 **(a)**

— CH_3 is electron repelling group.

812 **(c)**

Metamerism is found in molecules having polyvalent functional group.

813 **(d)**

There are IUPAC rules.

814 **(c)**

$a = 2^n$; where n is no. of dissimilar asymmetric carbon atoms and a is no. of optically active isomers.

816 **(a)**

Follow IUPAC rules.

817 **(d)**

Nucleophile $(-NH_3)$ replaces other nucleophile $(-Br)$ in the reaction.

818 **(a)**

Meso form is optically inactive.

819 **(d)**

Formation of ethylene from acetylene is an example of addition reaction

$CH \qquad CH_2$

$||| \ + H_2 \xrightarrow{Ni} \ ||$

$CH \qquad CH_2$

Ethyne ethene

820 **(b)**

Sodium hydrogen sulphite adds to aldehydes and ketones to form crystalline bisulphite addition products. The product is water soluble and can be converted back to the original carbonyl compound by treating it with dilute mineral acid or alkali. Therefore, these are useful for separation and purification of aldehydes like acetaldehydes.

821 **(b)**

Zn dust is used for dehalogenation,

$$CH_2X.CH_2X \xrightarrow[\text{alc.}]{\text{Zn dust}} CH_2{=\!=}CH_2.$$

823 **(d)**

Resonance in a molecule is arised due to delocalisation of π-electrons.

824 **(d)**

$$\underset{sp^3}{CH_3}-\underset{sp^3}{CH_2}-\underset{sp^2}{CH}{=}\underset{sp}{\overset{\oplus}{CH}}$$

Electronegativity of different hybrid and unhybrid orbitals in decreasing order is as follows

$s > sp > sp^2 > sp^3 > p$

825 **(b)**

$CH_2Br \qquad\qquad CH$

$| \ + \ 2KOH \xrightarrow{\Delta} \ ||| + 2KBr + 2H_2O$

$CH_2Br \qquad\qquad CH$

ethylene dibromide acetylene

This is a dehydrohalogenation reaction.

826 **(d)**

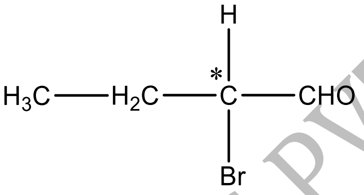
Stereoisomerism is of two types, geometrical and optical.

827 **(c)**

Follow IUPAC rules.

828 **(c)**

Compounds having asymmetric C-atom is optically active, *e.g.*,

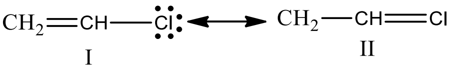


$$H_3C-H_2C-\overset{H}{\underset{Br}{\overset{*}{C}}}-CHO$$

The C-atom whose four valencies are satisfied by four different groups is asymmetric C-atom.

829 **(b)**

Chlorine of vinyl chloride $(CH_2 = CHCl)$ is non-reactive (less reactive) towards nucleophile in nucleophilic substitution reaction because it shows the following resonating structure due to $+M$ effect of – Cl atom.



In structure II, Cl-atom have positive charge and partial double bond character with C of vinyl group, so it is more tightly attracted towards the nucleus and it does not get replaced by nucleophile in S_N– reaction.

830 **(d)**

Follow mechanism of debromination.

831 **(c)**

Atom	**Atomic Mass (a)**	**Percentage (b)**	$\frac{b}{a} = x$	**Ratio**
C	12	40	$\frac{40}{12} = 3.33$	1
H	1	6.66	$\frac{6.66}{1} = 6.66$	2
O	16	53.34	$\frac{53.34}{16} = 3.33$	1

Hence, empirical formula $= CH_2O$

832 **(d)**

Grignard reagents can act as electrophile and nucleophile.

833 **(b)**

Both these carbon atoms have 3σ-and 1π-bond. Recall hybridized orbitals never from π-bonds.

834 **(c)**

S_N1 mechanism involves the formation of carbocation intermediate. Hence, the species which gives the most stable carbocation readily undergoes S_N1 mechanism. t-butyl bromide gives the most stable carbocation, $i.e.$, 3° carbocation, so it readily undergoes S_N1 reaction.

835 **(b)**

Follow IUPAC rules.

836 **(b)**

In the Lassaigne's test, a blue colour is obtained if the organic compound contains nitrogen. The blue colour is due to ferri-ferro cyanide $i.e.$, $Fe_4[Fe(CN)_6]_3$.

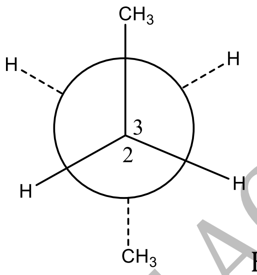
837 **(d)**

According to Cahn-Ingold-Prelog sequence rules, the priority of groups is decided by the atomic number of their atoms. When the atom (which is directly attached to the asymmetric carbon atom) of a group has higher atomic number, then the group gets higher priority. Groups which atoms of comparable atomic number having double or triple bond, have high priority than those have single bond.

Hence, the order of priority of group is

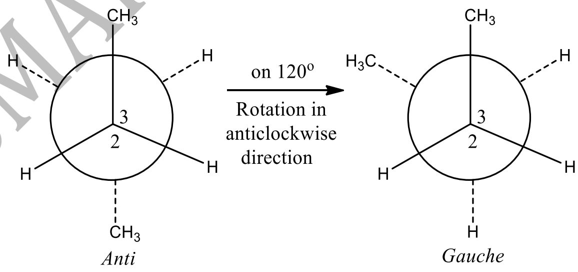
$-OH > -COOH > -CHO > -CH_2OH$

838 **(c)**

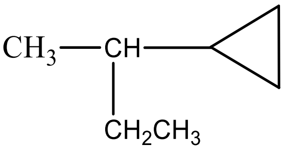


Here, when C_2 is rotated anticlockwise 120° about $C_2 - C_3$ bond the resulting conformer is *Gauche* conformer.

Hence,

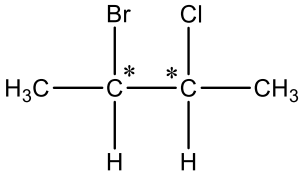


839 **(c)**



contains asymmetric carbon, thus optically active.

840 **(c)**



2-bromo 3-chloro butane

∴ Number of asymmetric carbon atoms=2

∴ Number of chiral isomers $= 2^n = 2^2 = 4$

841 **(c)**

Glycerol can be separated from spent lye in soap industry by the distillation under reduced pressure because it decomposes near its boiling point

843 **(b)**

In gas phase tertiary amines are more basic than secondary amines which are more basic than ammonia

$-I$ group present on central atom decreases electron density, hence decreases basicity

$CH_3NH_2 > NH_3 > NF_3$

844 **(a)**

Atom	**At mass ($\boldsymbol{a}$)**	**% ($\boldsymbol{b}$)**	$\boldsymbol{\frac{b}{a}}$	**Ratio**
C	12	49.3	$\frac{49.3}{12} = 4.10$	2
H	1	6.84	$\frac{6.84}{1} = 6.84$	3
O	16	43.86	$\frac{43.86}{16} = 2.74$	1

Hence, empirical formula $= (C_2H_3O)$

Molecular mass $= 2 \times VD = 2 \times 73$

$= 146$

$$n = \frac{\text{molar mass}}{\text{empirical formula mass}} = \frac{146}{43} \approx 3$$

So, formula$= (C_2H_3O)_3 \approx C_6H_9O_3$

845 **(c)**

Wöhler prepared urea from inorganic compounds and rejected the vital force theory that organic compounds can only be synthesised from living organisms.

846 **(c)**

Follow mechanism of addition of HCl and HI in presence of peroxide. One of the chain propagation step is endothermic in both cases.

847 **(c)**

All aromatic compounds are resonance hybrid.

848 **(a)**

It is the stability order for various conformers.

849 **(c)**

Glucose has aldehyde group and fructose keto group. The general formula for both is $C_6H_{12}O_6$.

851 **(b)**

Follow conformation.

852 **(b)**

In $o-, m-, p-$ derivatives vectors are at 60°, 120° and 180°. Thus, *para* has zero dipole moment. Also *ortho* form has more dipole moment than *meta* form.

853 **(c)**

The staggered form has lower energy than eclipsed form because of repulsive interaction between the H-atoms attached to two carbon atoms are minimum due to maximum distance between them.

854 **(c)**

Victor Mayer's method is applicable only for the determination of molecular mass of volatile substance

856 **(d)**

Hexane is non-polar molecule.

857 **(c)**

Nucleophilies may be neutral or negatively charged, whereas substrate undergoing nucleophilic substitution may be neutral or positively charged

$C_2H_5-I+OH^- \longrightarrow C_2H_5OH+I^-$

858 **(a)**

Nucleophilicity increases on going down in the group of the Periodic Table

$I^{\ominus} > Br^{\ominus} > Cl^{\ominus} > F^{\ominus}$

859 **(d)**

Free radicals have unpaired electrons, but are neutrals and are reactive.

$\dot{C}H_3 + \dot{C}H_3 \longrightarrow CH_3-CH_3$

861 **(b)**

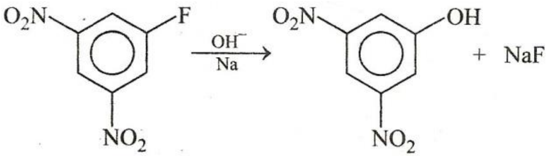
Follow mechanism of Kharasch effect.

862 **(d)**

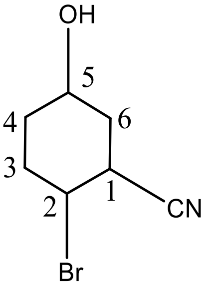
—do—

864 **(c)**

Reaction of NaOH with dinitrofluorobenzene represents nucleophilic aromatic substitution reaction because $-NO_2$ group is deactivating group. They make benzene nucleus electron deficient and facilitate the nucleophile to attack the ring.



865 **(b)**



Cyano group has the highest priority therefore, parent name must be benzonitrile. Br occurs at 2-position, and hydroxyl at 3-position, hence the IUPAC name is 2-bromo-5-hydroxy benzonitrile.

866 **(d)**

Ethers show metamerism.

867 **(b)**

Due to resonance; the carbonyl group of benzoic acid is coplanar with the ring. If the electron withdrawing substituent (i.e., $-I$ showing) is present at *ortho* position, it prevents the coplanarity and thus, the resonance. Hence, makes the acid more stronger.

Thus, among the given acids, *ortho* hydroxy benzene acid is the most acidic.

868 **(a)**

Diamond (sp^3), Graphite (sp^2), Acetylene (sp).

869 **(d)**

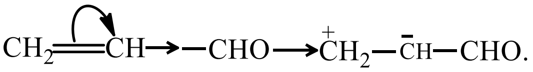
$CH_3CHClCOOH$ contains asymmetric carbon atom.

870 **(d)**

Statement (c) is wrong.

871 **(b)**

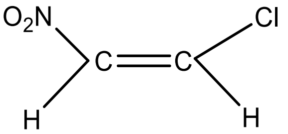
The –ve inductive effect of –CHO group play role to give anti Markownikoff's addition.



$CH_2{=}CH \rightarrow -CHO \longrightarrow \overset{+}{C}H_2-\overset{-}{C}H-CHO.$

872 **(c)**

The structure of 1-chloro-2-nitroethene is as



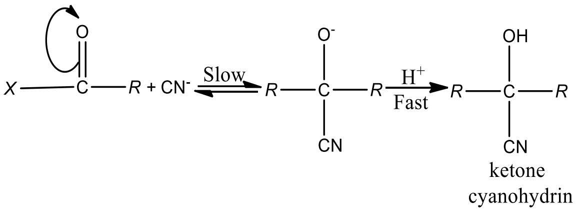
In this compound E-Z isomerism is possible because it is highly substituted alkene. The E-Z system of nomenclature is developed by Cahn, Inglod and Prelog.

873 **(a)**

$CH_3C=N$ is known as acetonitrile or methyl cyanide.

874 **(b)**

Ketone undergoes nucleophilic addition reaction because nucleophilic end of reagent attack first followed by electrophilic end of reagent.



875 **(d)**

Halogen containing compounds (C_6H_5Cl)

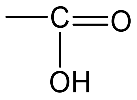
When placed in a flame, the presence of halogen is revealed by a green to blue flame.

876 **(a)**

Two similar asymmetric carbon atoms; $\therefore a = 2^{n-1}$. Also meso form $= 2^{\frac{n}{2}-1}$.

Total $= a + m$.

877 **(d)**

In C_6H_5 ring there are three π-bonds and one π-bond is present in  $-\underset{\large OH}{\underset{|}{C}}=O$ group.

Therefore, in all there are four π- bonds in C_6H_5COOH. In $CH_3CH_2COCH_3$ there is only one π-bond in $C=O$ group, in $CH_2=CH-CH=CH_2$ there are two π-bonds while in $HC\equiv C-CH=CH_2$ there are three π- bonds

878 **(c)**

Order of bond length

σ bond $(sp^3) > \sigma$ bond $(sp^2) > \sigma$ bond (sp)

879 **(d)**

It is a reason for the given fact.

880 **(c)**

The octet of all atoms are complete in structures a and b. The molecule in which all the atoms have completed octet is more stable than atom which have incomplete octet. Larger the number of resonating structures, larger will be the stability, thus structures a and b are stable.

In structure (d), the electron deficient of positive charged carbon is duly compensated by one pair electrons of adjacent oxygen atoms while such neighbour group support is not available in structure (c). Hence, structure (c) is least stable in comparison to structure (d).

883 **(a)**

CH_3CH_2Cl; CH_3CHCl_2; CH_2ClCH_2Cl; CH_3CCl_3; $CH_2CHCl_2CCl_3$; CCl_3CCl_3

884 **(d)**

(1) and (3) are enantiomeric forms to each other .

885 **(d)**

Methoxy group, due to $+I$ effect, increase electron density on OH- group, thus making it less acidic. Thus, o-methoxy phenol and acetylene are less than phenol.

p-nitrophenol is more acidic than phenol.

886 **(c)**

When organic compound is fused with sodium metal, nitrogen of the compound is converted into sodium cyanide as

$$Na + C + N \rightarrow NaCN$$

887 **(c)**

It is structure of furan, a heterocyclic compound.

888 **(a)**

Diazonium salts are highly reactive. In Sandmeyer reaction diazo group is replaced by chlorine or bromine in presence of CuCl or CuBr.(Substitution reaction)

$$C_6H_5\overset{\oplus}{N_2}\overset{\ominus}{Cl} \xrightarrow{CuCl} C_6H_5Cl + N_2$$

889 **(b)**

X^- is replaced by OH^-.

890 **(b)**

There are four structural isomers are possible for C_4H_9Cl

(a)$CH_3CH_2CH_2CH_2Cl$

(b)$CH_3-CH_2-\underset{Cl}{\underset{|}{CH}}-CH_3$

(c)$CH_3-\underset{CH_3}{\underset{|}{CH}}-CH_2Cl$

(d)$CH_3-\overset{Cl}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}-CH_3$

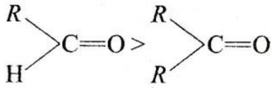
891 **(b)**

A carbanion or carboanion has –ve charge on it.

894 **(c)**

The case with which a nucleophile attacks the carbonyl groups depends upon the electron-deficiency, $i.e$, magnitude of the positive charge on the carbonyl carbon. Since, an alkyl groups has electron-donating inductive effect.

($+I$ effect), therefore, greater the number of alkyl groups attached to the carbonyl groups greater is the electron-density on the carbonyl carbon and hence, lower is its reactivity towards nucleophilic addition reactions.



896 **(c)**

n-pentane and isopentane or 2-methylbutane are chain isomers since both have different hydrocarbon chain.

897 **(a)**

$CH_2 = CH - CHO$

Prop -2-en-1-al

898 **(d)**

Free radicals have unpaired electrons but are neutrals and are reactive.

$\dot{C}H_3 + \dot{C}H_3 \longrightarrow CH_3 - CH_3$

900 **(a)**

The second carbon is asymmetric.

901 **(c)**

Both have different mode of linkage, $i.e.$, chain and ring.

902 **(c)**

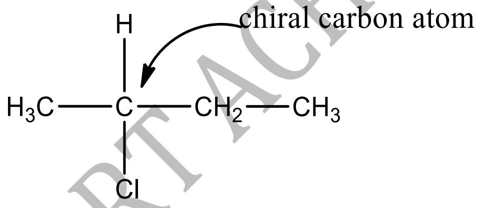
Carbanions contain even number of valence electrons and thus, show diamagnetic behaviour.

903 **(d)**

Molecules with two similar groups attached on either of the doubly bonded carbon do not show geometrical isomerism.

904 **(c)**

Compound $CH_3 - CHCl - CH_2 - CH_3$ shows optical isomerism due to the presence of chiral carbon atom.



905 **(c)**

The Kolbe's electrolysis proceeds via free radical mechanism. For example, when sodium propionate is electrolysed, n —butane, ethane, ethylene are obtained. The propionate ion discharge at the anode to form free radicals.

$C_2H_5COO^- \rightarrow C_2H_5COO^{\bullet} + e^-$

$C_2H_5COO^{\bullet} \rightarrow C_2H_5^{\bullet} + CO_2$

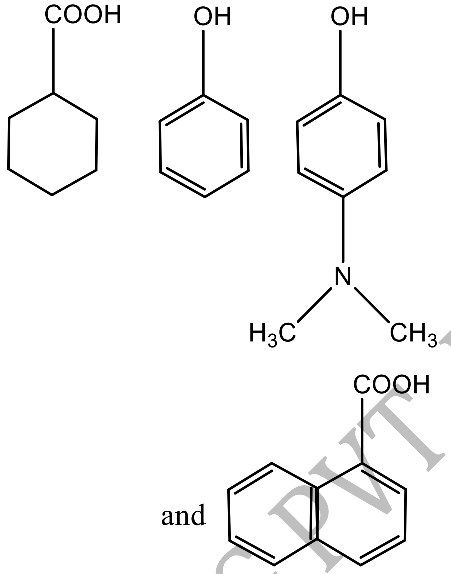
$2C_2H_5^{\bullet} \rightarrow C_4H_{10}$

$C_2H_5^{\bullet} + C_2H_5^{\bullet} \rightarrow C_2H_4 + C_2H_6$

907 **(c)**

In TLC, adsorbent is made of silica gel or alumina gel

908 **(d)**



are soluble in aq. NaOH. Benzylic alcohol is less acidic than water so not soluble in aq. NaOH,

909 **(c)**

Lead unreached 0.1 M (=0.2 N)$H_2SO_4 = V$ mL

20 Ml of 0.5 NaOH=V mL of 0.2 N H_2SO_4

$20 \times 0.5 = V \times 0.2$

$V = \frac{20 \times 0.5}{0.2} = 50$ mL

Used $H_2SO_4 = 100 - 50 = 50$ mL

$\%$ of N $= \frac{1.4\,NV}{w} = \frac{1.4 \times 0.2 \times 50}{0.30}$

$= 46.67\ \%$

% of nitrogen in

(a) $CH_3CONH_2 = \frac{14 \times 100}{59} = 23.73\ \%$

(b) $C_6H_5CONH_2 = \frac{14 \times 100}{121} = 11.57\ \%$

(c) $NH_2CONH_2 = \frac{28 \times 100}{60} = 46.67\ \%$

(d) $NH_2CSNH_2 = \frac{28 \times 100}{76} = 36.84\ \%$

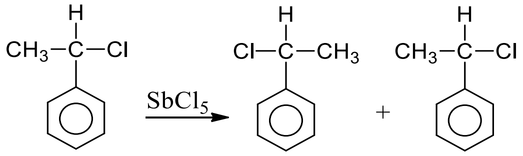
910 **(d)**

Removal of H from alkane, alkene and alkyne gives alkyl, alkenyl, alkynyl groups respectively.

911 **(d)**

There are two symmetrical hexenes as given in (a) and (b).

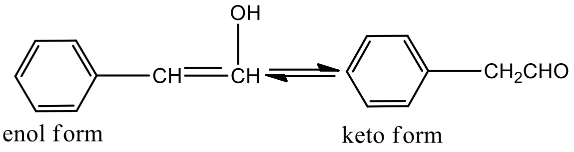
913 **(d)**



The solution of D(+)-2-chloro-2-phenyl ethane in toluene racemises slowly in the presence of $SbCl_5$ due to the formation in carbocation

914 **(a)**

Tautomerism it is functional isomerism in which the isomers are readily interchangeable and maintain a dynamic equilibrium with each other.



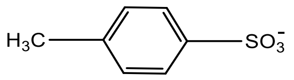
915 **(b)**

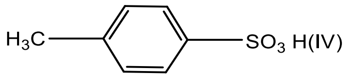
If acid is weak, its conjugate base (nucleophile) is strong and *vice versa*.

(A)$CH_3 - \overset{\overset{O}{\|}}{C} - O^-$ is conjugated base of $CH_3\overset{\overset{O}{\|}}{C}OH$ (I)

(B) CH_3O^- is a conjugate base of CH_3OH (II)

(C) CN^- is a conjugate base of HCN (III)

(D)H_3 
SO_3^- is a conjugate base

Of 

Acidic nature of IV>I>III>II

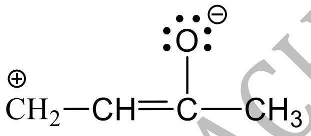
and nucleophilicity of B>C>A>D.

916 **(d)**

S_N2 reaction proceeds with inversion and a transition state is formed which does not carry any charge

918 **(b)**

The most stable one is that in which the positive and negative charges reside on the most electropositive and most electronegative atoms of the species respectively. Like



919 **(b)**

Organic compound $\xrightarrow{[O]} CO_2 + H_2O$

17.6 g 7.2g

$\%of\ C = \frac{12}{44} \times \frac{17.6}{5.6} \times 100 = 85.7\%$

$\%\ of\ H = \frac{2}{18} \times \frac{7.2}{5.6} \times 100 = 14.28\%$

Element	Percentage	Relative no. of atoms	Simplest ratio
C	85.7	85.7/12=7.14	7.14/7.14=1
H	14.28	14.28/1=14.28	14.28/7.14=2

Hence, empirical formula of compound of = CH_2

∴ Molecular formula of compound =C_4H_8

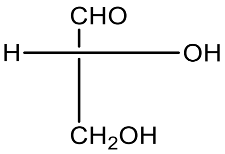
920 **(c)**

Stronger is acid, weaker is its conjugate base.

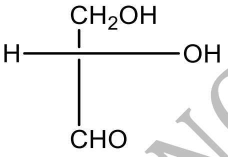
Acidic nature : $CH \equiv CH > CH_2 = CH_2 > CH_3 - CH_3$

Conjugate basic : $CH \equiv C^- < CH_2 = CH^- < CH_3 - CH_2^-$ nature.

921 **(b)**



D-glyceraldehyde has the above formula. So, the Fischer's projection formula which is identical to it, is



922 **(b)**

Percentage of N in an organic

$$\text{Compound} = \frac{1.4 \times N \times V}{w}$$

$$= \frac{1.4 \times 1 \times 30}{1.2} = 35$$

923 **(c)**

Dissociation of proton from ddd is very difficult due to $-I$ effect of $Cl^-$and N^+ while in C_6H_5OH due to the reasonance stabilization of phenoxide ion proton eliminates easily. Similarly due to H-bonding in $C_6H_5CH_2OH$ it can be eliminate easily and in $CH_3C \equiv CH$ the proton is acidic in nature hence, it can be dissociated

924 **(d)**

$$\% \text{ of S} = \frac{32}{233} \times \frac{\text{wt. of BaSO}_4}{\text{wt. of organic compound}} \times 100$$

$$= \frac{32}{233} \times \frac{1.158}{0.53} \times 100$$

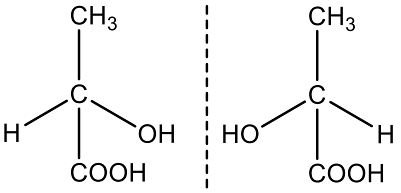
$$= 30\%$$

925 **(c)**

Tautomerism is a dynamic isomerism because two forms (keto and enol) of substance cannot be separated they are in dynamic equilibrium with each other.

926 **(a)**

Enantiomers are non-superimposable mirror images. *e.g.,* lactic acid



927 **(a)**

Vinyl carbocations are more stable than primary carbocation but less stable than secondary carbocation.

928 **(b)**

Duma's method involve the determination of nitrogen content in the organic compound in the form of N_2

$N_2O + Cu \rightarrow N_2 + Cu$

$$d\% \text{ of } N = \frac{28}{22400} \times \frac{\text{volume of } N_2 \text{at NTP}}{\text{weight of compound}} \times 100$$

930 **(a)**

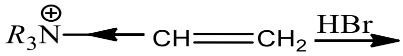
The correct order of reactivity is

$RI > RBr > RCl > RF$

It is due to fact, that weaker the base, better it will be the leaving group.

Hence, I^- is the best leaving group.

931 **(b)**



$R_3\overset{\oplus}{N} \longleftarrow CH{=}CH_2 \xrightarrow{HBr}$

Due to $R_3\overset{\oplus}{N}-$ (e^-withdrawing tendency) carbocation will appear farther to that (terminal). Hence, product is $R_3N - CH_2 - CH_2Br$.

932 **(d)**

n-pentane and 2-methyl butane are constitutional isomers or chain isomers or skeletal isomers.

934 **(a)**

Follow IUPAC rules.

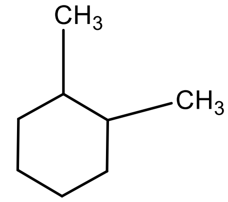
935 **(b)**

In toluene $(-CH_3)$ group is present which has $+I$ effect and increases electron density on *ortho* and *para* position. While in sulphonation – SO_3 acts as electrophile. Therefore, it $(-SO_3H)$ attacks on *ortho* and *para* position readily.

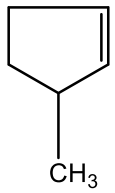
936 **(d)**

Follow elimination rules.

937 **(a)**

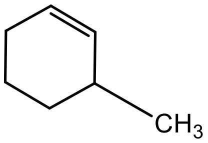


1, 2-dimethyl cyclohexane



3-methyl cyclo pentene

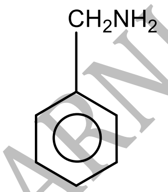
and



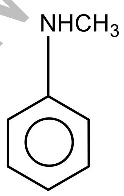
3-methyl cyclo hexene

contain two, one, one asymmetric carbon atoms respectively.

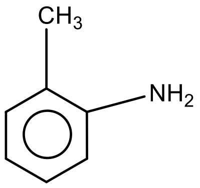
939 **(a)**



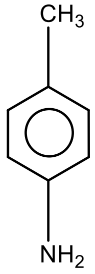
benzylamine



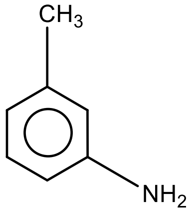
N-methyl aniline



o-toluidine



p-toluidine

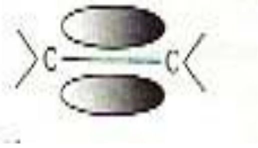


m-toluidine

∴ Total of five isomers are possible by formula C_7H_9N.

941 **(a)**

Such a condition is seen when π bond is formed between similar atoms *ie*,



943 **(b)**

From kjeldahl's method,

$$\text{Percent of nitrogen} = \frac{1.4 \times N \times V}{W}$$

$$= \frac{1.4 \times 0.5 \times 2 \times 10}{0.25} = 56\%$$

944 **(d)**

It is the definition of distereoisomers.

945 **(b)**

The central carbon is attached to four different substituents, hence it is chiral, therefore optically active.

946 **(b)**

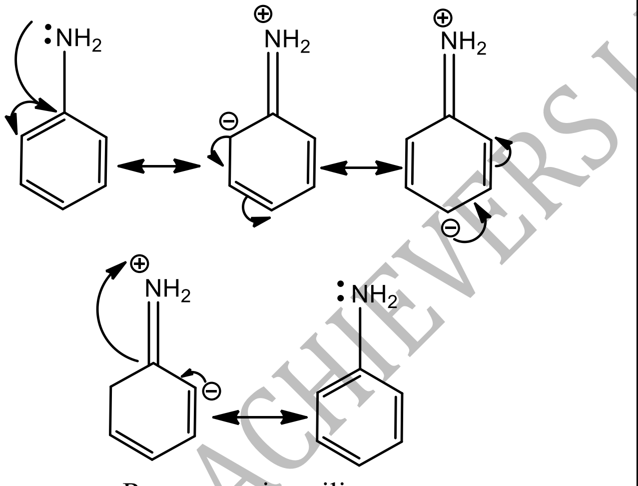
$HNO_2 + H_2SO_4 \longrightarrow \underset{\text{Nitrosonium ion}}{NO^+} + HSO_4^- + H_2O$

947 **(b)**

$-$Cl is an electron withdrawing (*i.e.,*$-I$ showing) group. It withdraws electrons when attached to the carboxylic acid and decreases the electron density on the oxygen atom. This will facilitate the release of H^+ by making O $-$ H bond more polar and thus – Cl increases the acidity of acetic acid when attached at, α position because of $-I$ effect.

948 **(a)**

$-NH_2$ has +R effect, it donates electrons to the benzene ring. As a result, the lone pair of electron on the N-atom gets delocalized over the benzene ring. As a result, the lone pair of electron on the N-atom gets delocalized over the benzene ring and thus it is less readily available for protonation. Hence, aniline is a weaker base than cyclohexylamine.



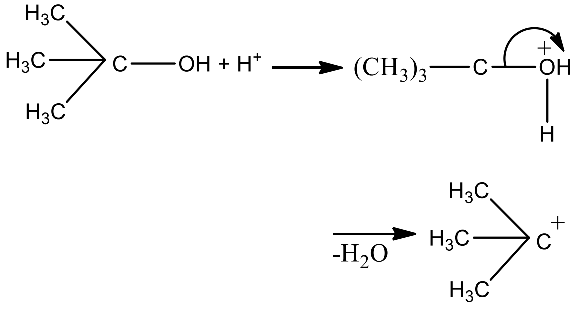
Resonance in aniline

949 **(c)**

This is annulene.

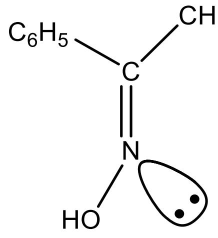
950 **(a)**

Alcohols undergo dehydration usually by E1 mechanism. This is because elimination is preferred in case of tertiary alcohols, *e. g.*,

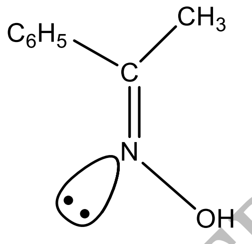


951 **(c)**

Acetophenone oxime can show geometrical isomerism.



syn-form



anti-form

952 **(d)**

Ethers show functional isomerism with molecular formula $C_nH_{2n+2}O$. For example, C_2H_6O

CH_3OCH_3, CH_3CH_2OH

dimethyl ether, ethyl alcohol

953 **(c)**

The alkyl halide which produce stable carbonium ion undergo S_N1 reaction. Tertiary alkyl halide is most stable. Thus, decreasing order of tendency of alkyl halides undergoes S_N1 mechanism.

t-alkyl halide>*sec*-alkyl halide> *primary*-alkyl halide

954 **(c)**

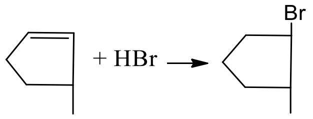
Vinyl chloride($CH_2 = CHCl$) undergoes addition and elimination reactions. Substitutions reaction is shown by compounds having single bonds only.

955 **(a)**

Fisher projections are for illustration of optical isomers.

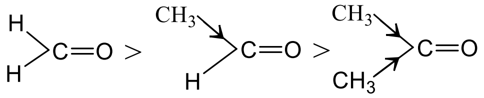
956 **(b)**

A hydrogen halide contain a highly polar H $-$ X bond can easily lose $\overset{\oplus}{\mathrm{H}}$ to the pi bond of an alkene. The result of the attack of $H^{\oplus}$ is an intermediate carbocation, which quickly undergoes reaction with the negative halide ion (X^-) to yield an alkyl halide



960 **(c)**

Aldehydes and ketones readily undergo nucleophilic addition reaction. The order of reactivity, is as the $+I$ effect of alkyl group increases



961 **(b)**

The structures of maleic and fumaric acids are given below

H − C − COOH　　　　　　H − C − COOH
∥　　　　　　　　　　　　∥
H − C − COOH　　　　　　HOOC − C − H
(maleic acid)　　　　　　(fumaric acid)
molecular formula ($C_4H_4O_4$)　molecular formula($C_4H_4O_4$)

The structures of fumaric and maleic acid suggest that they are geometrical isomers because they have same molecular formula but different spatial arrangement of atoms around a double bond.

962 **(b)**

Stability order of arenium ion

II>III>I

963 **(d)**

In condensed structure formula, paranthesis is used for identical group of atoms.

964 **(d)**

Compound which sublime on heating can be purified by sublimation method.

Benzoic acid, camphor and naphthalene sublime on heating hence, they are purified by sublimation method.

965 **(d)**

$CH_2 = CH.CH_2Cl$ compound undergoes nucleophilic substitution most readily.

966 **(a)**

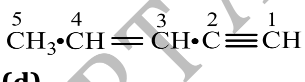
As the $-I$ group increases at the α-carbon, acidity increases

967 **(b)**

In S_N2 reactions, the nucleophile attaches itself from the direction opposite to that of the nucleophile already present in the second step, the previous nucleophile is removed and a single stereoisomer is obtained

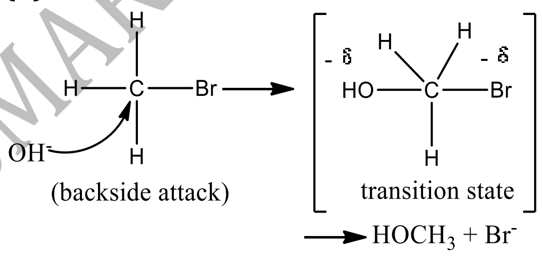
968 **(b)**

Follow IUPAC rules.



$\overset{5}{C}H_3{\bullet}\overset{4}{C}H=\overset{3}{C}H{\bullet}\overset{2}{C}\equiv\overset{1}{C}H$

969 **(d)**



Since, the reaction rate depends upon the concentration of both reactant and nucleophile, it is a S_N2 reaction. It involves inversion of configuration.

970 **(c)**

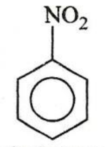
The given compound is nitrobenzene.

971 **(c)**

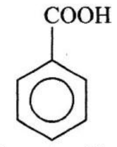
Both differs by a $-CH_2$ group.

972 **(d)**

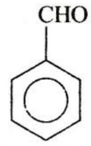
During electrophilic substitution electrophile attacks the double bond of benzene ring. The aromatic compounds having electron donating groups undergo electrophilic substitution more easily due to favouable effect of electron donating group.

(a)


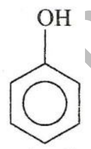
nitrobenzene

(b)


benzoic acid

(c)


benzaldehyde

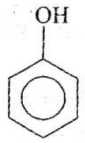
(d)


phenol

(i) NO_2, COOH and CHO groups are electron withdrawing groups so, they decrease the reactivity of organic compounds.

(ii) $-$ OHgroup is electron donating group, so it increases the electron density in benzene ring ad increases the rate of reaction

∴



undergoes substitution most easily.

973 **(b)**

Oxygen cannot be detected by direct test because oxygen is present in atmosphere and all tests are carried in atmosphere of oxygen

974 **(a)**

Follow IUPAC rules.

975 **(a)**

$o-, m-, p-$ isomers are position isomers.

976 **(b)**

Carbinol is trivial name for HCH_2OH. Thus, $C_6H_5CH_2OH$ is phenyl carbinol and chloral is CCl_3CHO.

977 **(a)**

A primary carbon is one which is attached with one carbon atom.

978 **(a)**

Follow IUPAC rules.

979 **(c)**

1,1-dibromoethane and 1,2-dibromoethane.

980 **(a)**

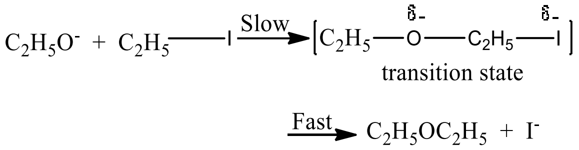
A solvent molecule lacking a polar X-H bond is called aprotic solvent. NH_3, SO_2 and CH_3CN are aprotic solvent while CH_3COOH is protic solvent.

981 **(b)**

When sodium ethoxide reacts with iodoethane, diethyl ether is obtained (**Williamson's synthesis**)

The mechanism of this reaction is as follows

$C_2H_5ONa \rightleftharpoons C_2H_5O^- + Na^+$



$$C_2H_5O^- + C_2H_5—I \xrightarrow{Slow} \left[\underset{\text{transition state}}{C_2H_5—\overset{\delta-}{O}—C_2H_5—\overset{\delta-}{I}}\right] \xrightarrow{Fast} C_2H_5OC_2H_5 + I^-$$

Since, the reaction involves substitution of a group by a nucleophile, it is an example of nucleophilic substitution reaction.

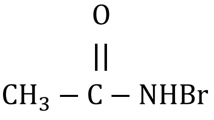
982 **(c)**

The organic compounds is fused with Na metal, Na_2S is formed which is tested as

It reacts with lead acetate and forms black ppt. of PbS

$Na_2S + (CH_3COO)_2Pb \rightarrow \underset{\text{black ppt}}{PbS} + 2CH_3COONa$

983 **(c)**



$$CH_3-\overset{\overset{\large O}{||}}{C}-NHBr$$

N-bromoethanamide

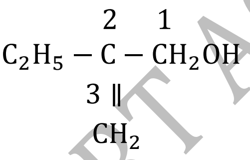
984 **(c)**

$$\frac{E}{108} = \frac{100}{60}$$

Eq. wt of the silver salt $E = 108 \times \frac{100}{60} = 180$

$\therefore$ Eq. wt. of the acid $= E - 108 + 1 = 73$

985 **(a)**

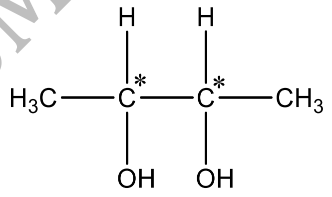


$$C_2H_5-\overset{2}{\underset{3\,\|\atop CH_2}{C}}-\overset{1}{CH_2}OH$$

2-ethylprop-2-en-1-ol

986 **(d)**

The structure of butane-2, 3-idol is as



$\because$ Optical isomers in compounds have similar asymmetric carbon atom, which are even in number $= 2^n - 1$

Here, $n = 2$

$\therefore$ Total number of optically active stereoisomers

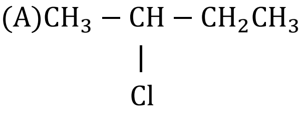
$= 2^2 - 1 = 3$

987 **(b)**

$CH_4 \longrightarrow CH_3^+ + H^-$;

CH_3^+ is methyl carbonium.

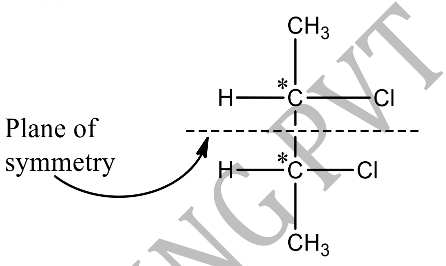
988 **(b)**



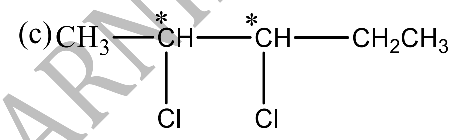
$$(A)CH_3-\underset{\underset{\large Cl}{|}}{CH}-CH_2CH_3$$

One asymmetric carbon atom, forms d, and l −optical and isomers.

(b) Two asymmetric carbon atoms, forms d, l and *meso* forms

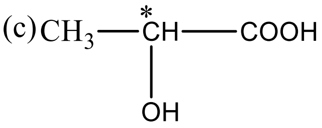


Meso due to internal compensation



$$(c)CH_3—\underset{\underset{\large Cl}{|}}{\overset{*}{C}H}—\underset{\underset{\large Cl}{|}}{\overset{*}{C}H}—CH_2CH_3$$

Two asymmetric carbon atoms but does not have symmetry. Hence, *meso* form is not formed.



$$(c)CH_3—\underset{\underset{\large OH}{|}}{\overset{*}{C}H}—COOH$$

One asymmetric carbon atom, *meso* form is not formed.

989 **(d)**

More the number of hyperconjugated structures, more will be electronegative chlorine atom.

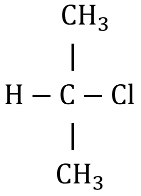
(a) CH_3-Cl

3-hyperconjugated structures

(b) CH_3-CH_2-Cl

2-hyperconjugated structures

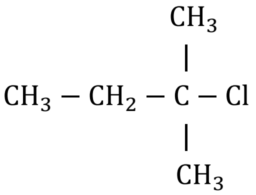
(c)



$$H-\underset{\underset{\large CH_3}{|}}{\overset{\overset{\large CH_3}{|}}{C}}-Cl$$

6- hyperconjugated structures

(d)



$$CH_3-CH_2-\underset{\underset{\large CH_3}{|}}{\overset{\overset{\large CH_3}{|}}{C}}-Cl$$

8-hyperconjugated structures

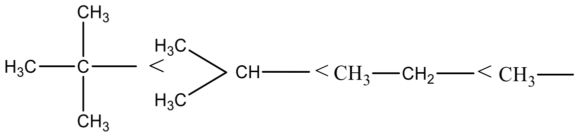
$\because$ 8-hyperconjugated structures are possible for

(d)
∴Chlorine in this is most electronegative.

990 **(a)**
Follow resonance characteristics.

991 **(b)**
Diastereoisomers are a pair of optical isomers which cannot be related as non-superimposable mirror images of each other.

994 **(a)**
The rate of reaction is influenced by the hyperconjugation effect of group R. it depends on the electron donating power of alkyl group (R). The electron releasing power of R group depends on the number of hydrogen present on α carbon. The increasing order of speed with R group in the reaction is



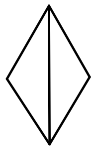
i.e., IV<III<II<I

995 **(b)**
The given compound form two geometrical isomers and two optical isomers.

996 **(b)**
Geometrical isomerism is shown by >C=C< only when identical groups are not present on the double bonded carbon atoms.

999 **(b)**
$RX + I^- \rightarrow R - I + X^-$
This reaction is an example of nucleophilic substitution.

1000 **(c)**



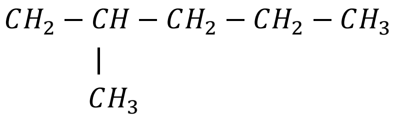
Its IUPAC name is bicyclo [1,1,0] butane.

1001 **(a)**
Weak base ⟶ strong conjugate acid
$CH \equiv C^- \rightarrow CH \equiv CH$
weakest base (strongest acid among the given)
$CH_3 - CH_2^- \rightarrow CH_3CH_3$
(strongest base) (weakest acid)

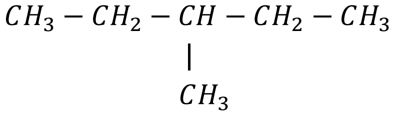
1002 **(d)**
With the increasing basicity of the added base, the rates of the elimination reactions have been found to increase. Thus, RO^- is most reactive

1003 **(b)**
Follow IUPAC rules.

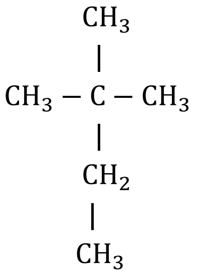
1004 **(a)**
Number of isomers in hexane are five as follows
$CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_3$



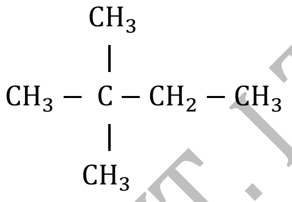
$CH_2 - \underset{\underset{\displaystyle CH_3}{|}}{CH} - CH_2 - CH_2 - CH_3$



$CH_3 - CH_2 - \underset{\underset{\displaystyle CH_3}{|}}{CH} - CH_2 - CH_3$

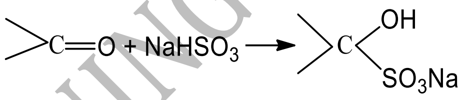


$CH_3 - \overset{\overset{\displaystyle CH_3}{|}}{\underset{\underset{\displaystyle CH_2 - CH_3}{|}}{C}} - CH_3$

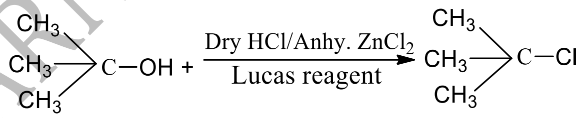


$CH_3 - \overset{\overset{\displaystyle CH_3}{|}}{\underset{\underset{\displaystyle CH_3}{|}}{C}} - CH_2 - CH_3$

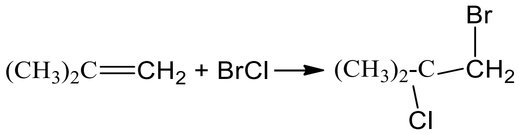
1005 **(c)**
(a) it is Diels Alder's reaction (cyclo addition)



(b)
It is nucleophilic addition reaction



(c)
It is nucleophilic substitution reaction



(d)
It is electrophilic addition reaction

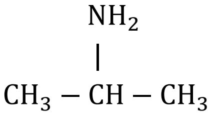
1006 **(d)**
(d) is with maximum conjugative structure among them

1007 **(a)**
Based on heat of hydrogenation.

1010 **(c)**
Butyne-2 is $CH_3 - C \equiv C - CH_3$; Two carbon of corner are sp^3-hybridized.

1011 **(c)**
Four π-electrons of double bond and 1 lone pair on N atom leads to delocalisation of six electrons.

1012 **(d)**



$CH_3 - \overset{\overset{\displaystyle NH_2}{|}}{CH} - CH_3$
2-propanamine

1013 **(b)**

Element	Percentage	$\frac{\text{Percentage}}{at.wt.}$	Simple ratio

C	20.0	$\frac{20.0}{12} = 1.66$	$\frac{1.66}{1.66} = 1$
H	6.67	$\frac{6.67}{1} = 6.67$	$\frac{6.67}{1.66} = 4$
N	46.67	$\frac{46.67}{14} = 3.33$	$\frac{3.33}{1.66} = 2$
O	26.66	$\frac{26.66}{16} = 1.66$	$\frac{1.66}{1.66} = 1$

Empirical formula= CH_4N_2O

Empirical formula weight

$$= 12 + (4 \times 1) + (2 \times 14) + 16 = 60$$

$$\therefore \ n = \frac{\text{mol. formula weight}}{\text{emp. formula weight}}$$

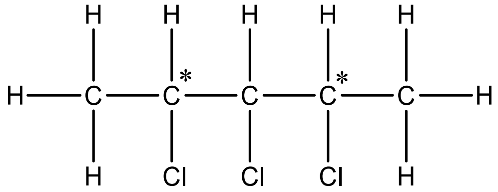
$$= \frac{60}{60} = 1$$

$\therefore$ Molecular formula=CH_4N_2O

Given compound gives biuret test.Thus, given compound is urea $(NH_2)_2CO$.

101 **(c)**

4 The structure of 2, 3, 4-trichloro pentane is



Hence, two chiral carbon atoms are present in 2, 3, 4-trichloropentane.

101 **(b)**

5 The electronegativity of F is maximum and thus, C— F bond is more polar.

101 **(a)**

6 1, 1-dichloro-1-pentene does not exhibit geometrical isomerism.

101 **(b)**

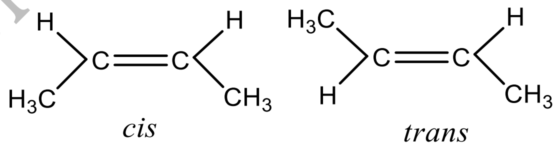
7 The hydrolysis of sugar solution (*dextrorotatory*) leads to formation of *laevorotatory* mixture due to formation of glucose (*dextrorotatory*) and fructose (*laevorotatory*).

101 **(c)**

8 First two and fourth will show stereoisomerism.

101 **(d)**

9 Geometrical isomers of $CH_3CH = CH - CH_3$ are



102 **(c)**

0 Ethyl alcohol shows functional isomerism with dimetyl ether.

C_2H_5OH $CH_3 - O - CH_3$

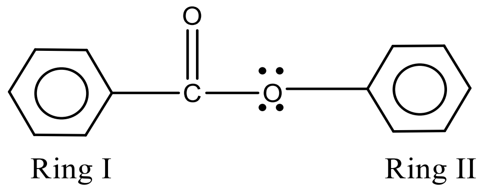
alcohol ether

102 **(a)**

1 Highest to lowest priority ($Br > Cl > CH_3$) is clockwise than *R*.

102 **(c)**

2

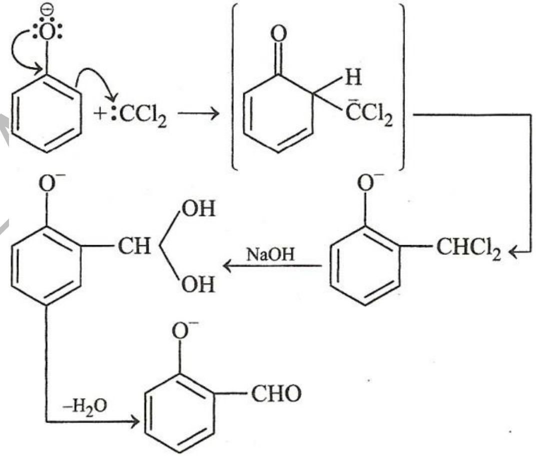


Electrophilic substitution reaction takes place in compounds in which π −electrons are highly delocalised. The electrophile attacks the region of high electron density, therefore, electrophilic substitution occurs at *ortho*/*para* position at ring II.

102 **(c)**

3 When phenol reacts with chlrtoform and aqueous NaOH solution, it give salicyladehyde.

$CHCl_3 + OH^- \leftrightharpoons H_2O + CCl_3^-$



102 **(c)**

4 It is definition of asymmetric synthesis.

102 **(a)**

5 Since, the compound on heating with CuO producedCO_2, it contains carbon. Again, it does not produce water, hence it does not contain hydrogen. So, the organic compound is carbon tetrachloride (CCl_4).

102 **(d)**

6 If organic compound gives blue or green colour at the tip of red hot copper wire, this indicates the presence of halogens in the compound. However, compounds like urea, thiourea etc also give blue or green colour in this test even in the absence of halogens, thus, it is not a sure test for halogens

102 **(a)**

8 Mesomeric effect involves complete transfer of π or lone pair of electrons to the adjacent atom or covalent bond. Hence, it involves delocalisation of pi (π) electrons.

1030 **(b)**

Species	**Valence electrons**	**Magnetic behavior**
Carbonium ion	6	Diamagnetic
Free radical	7	Paramagnetic
Carbene	6	Diamagnetic
Nitrane	6	Diamagnetic

1031 **(b)**

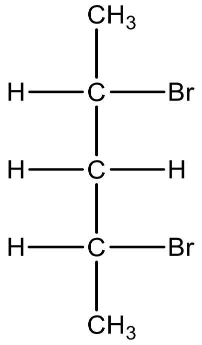
It should contain $(CH_3)_2$ CH − group to be named as iso.

1032 **(b)**

One propene and one cyclopropane.

1033 **(c)**

Those compounds which contain two or more asymmetric carbon atoms but are optically inactive due to presence of plane of symmetry, are called *meso* compounds. M*eso* compounds are optically inactive due to internal compensation. Out of the given compounds only 2, 4-dibromopentane have a plane of symmetry, so it is a *meso* compound.



2, 4-dibromopentane (*meso* compound)

1034 **(a)**

It is a fact.

1036 **(a)**

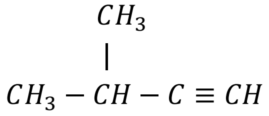
Electrophiles are the species having a tendency to accept a pair of electron, *e.g.*, NO_2^+, Br^+ etc. Nucleophiles are the species having a tendency to donate a pair of electron. *e.g.*, $CH_3OH. N_3^-$

1037 **(a)**

C_5H_8 has three possible alkynes. These are

$CH_3CH_2 - CH_2C \equiv CH, CH_3CH_2C \equiv C - CH_3$

pent-1-yne pent-2-yne



$$CH_3 - \underset{}{\overset{\overset{\displaystyle CH_3}{|}}{CH}} - C \equiv CH$$

3-methyl but-1-yne

1040 **(d)**

I can have maximum 3 hyperconjugative structures. II has maximum 5 hyperconjugative structure, III has 2 conjugative structure while IV has 1 conjugative structure

1041 **(a)**

$CH_3C \equiv CCH_2CH_3$; It is always pentyne-2 and never pentyne-3.

1043 **(d)**

The components of an azeotropic mixture are separated by special method,*ie*, fractional distillation. The simple fraction which distils at 337.8 K is a ternary azeotrope consisting of all water. Some alcohol and benzene

1044 **(a)**

Hence, an equimolar mixture of the enantiomers (*dextro or laevo* forms) is called racemic mixture. It is represented as *dl*-form or ± form and is optically inactive due to external compensation. Separation of racemic mixture into *d*- and *l*-forms is called resolution.

1045 **(a)**

Two successive homologous differ in their formula by CH_2 or have a difference of 14 units in their molecular weight.

1046 **(c)**

If a liquid decomposes at or below its boiling point, it is purified by vaccum distillation

1047 **(c)**

$CH_3CH_2NH_2$ and CH_3NHCH_3.

1048 **(a)**

IUPAC name of CH_3CH_2CHO is propan-1-al.

1049 **(b)**

Alkyl group (an electron releasing (+*I group*) increases electron density at N-atom, hence, basic nature is increased. In ammonia, no alkyl group is present, so it is least basic.

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

Assertion - Reasoning Type

This section contain(s) 0 questions numbered 1 to 0. Each question contains STATEMENT 1(Assertion) and STATEMENT 2(Reason). Each question has the 4 choices (a), (b), (c) and (d) out of which **ONLY ONE** is correct.

a) Statement 1 is True, Statement 2 is True; Statement 2 **is** correct explanation for Statement 1

b) Statement 1 is True, Statement 2 is True; Statement 2 **is not** correct explanation for Statement 1

c) Statement 1 is True, Statement 2 is False

d) Statement 1 is False, Statement 2 is True

1

Statement 1: Equivalent weight of ozone in the change $O_3 \rightarrow O_2$ is 8.

Statement 2: 1 mole O_3 of on decomposition gives $\frac{3}{2}$ moles of O_2.

2

Statement 1: Equivalent of $K_2 Cr_2O_7$ has 1 equivalent of K and Cr and O each.

Statement 2: A species contains same number of equivalents of its components.

3

Statement 1: The empirical formula of glucose is CH_2O which represents the relative number of atoms of each atom present in this molecule

Statement 2: In glucose, the elements C, H, O have combined in the ratio 1:2:1

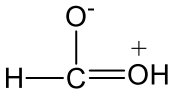
4

Statement 1: In Messenger's method, the colourless solution, the liquid is transferred to beaker and barium chloride added to estimate sulphuric acid as $BaSO_4$ in the usual way

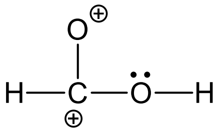
Statement 2: If an experiment 0.36 g of an organic compound gave 0.35 of $BaSO_4$. the percentage of sulphur in the compound is 13.35%

5

Statement 1:



is more stable than



Statement 2: Compound in which the positive and negative charges reside on the most electropositive and most electronegative atoms of the species respectively is more stable

6

Statement 1: Electrophile are electron rich in nature

Statement 2: H_3O^+, BF_3 and $AlCl_3$ are electrophile and can accept electron pair

7

Statement 1: Dehydration of alcohol is an example of elimination reaction

Statement 2: When H_2SO_4 or H_3PO_4(concentrated) are used as dehydrating agent, the mechanism is E_1

8

Statement 1: A solution which contains one gram equivalent of solute per litre of solutions is known as molar solution.

Statement 2: Normality=normality $\times \frac{\text{mol.wt.of solute}}{\text{eq.wt.of solute}}$

9

Statement 1: Lessaigne's test is the for nitrogen only

Statement 2: If halogens are present in organic compound, it react with sodium to form sodium halides which dissolves in water

10

Statement 1: The order of stability of carbocation are $R_3C^+ > R_2\overset{+}{C}H > R\overset{+}{C}H_2 > \overset{+}{C}H_3$.

Statement 2: The stability of carbocations is influenced by both resonance and inductive effects

11

Statement 1: Benzoic acid is purified by sublimation process

Statement 2: Sublimation process is very useful in separating a volatile solid from a non-volatile solid

12

Statement 1: The molality of the solution does not change with change in temperature.

Statement 2: The molality of the solution is expressed in units of moles per 1000 g of solvent.

13

Statement 1: 'A Victor Mayer tube' of hard glass, having a side-tube, leading to the arrangement for collection of displaced air over water.

Statement 2: An outer jacket of copper, containing a liquid boiling at nearly 30°C higher than the substance whose molecular mass is to be determined

14

Statement 1: Normally and molarity can be calculated from each other.

Statement 2: Normally is equal to the product of molarity and n.

15

Statement 1: Carbanion like ammonia have pyramidal shape

Statement 2: The carbon atom carrying negative charge has an octet of electrons

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

: ANSWER KEY :

1)	**b**	**2)**	**a**	**3)**	**a**	**4)**	**b**
5)	**a**	**6)**	**d**	**7)**	**b**	**8)**	**d**
9)	**d**	**10)**	**c**	**11)**	**a**	**12)**	**a**
13)	**b**	**14)**	**a**	**15)**	**b**		

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

: HINTS AND SOLUTIONS :

1 **(b)**

$$2O_3 \rightarrow 3O_2$$

$2 \text{ mol } O_3 \equiv 3 \text{ mol } O_2 = 3 \times 2 \text{ eq } O_2$

$$E_O = \frac{M}{6}$$

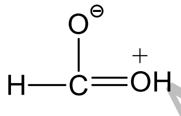
$$= \frac{48}{6} = 8$$

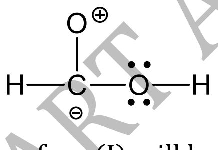
3 **(a)**
The empirical formula of compound in the simplest formula deduced from its percentage composition showing its composition by mass

4 **(b)**
Percentage of sulphur

$$= \frac{32}{233} \times \frac{0.35}{0.36} \times 100 = 13.35\ \%$$

5 **(a)**
Both the structures are resonating structure of formic acid

In  negative charge is on oxygen but

in  negative charge is on carbon therefore (I) will be more stable than the (II)

6 **(d)**
Electrophiles are electron deficient while nucleophiles are electron rich in nature, *ie*, electrophile can accept an electron pair while nucleophile donates an electron pair

7 **(b)**
Alcohols leading to conjugated alkenes are more easily dehydrated than the alcohols leading to non-conjugated alkenes

8 **(d)**
A solution; which contains one gram mole of solute per litre of solution is known as molar solution (M).

9 **(d)**
$\text{Na} + \underbrace{\text{C} + \text{N}} \rightarrow \text{NaCN}$

$2\text{Na} + \text{S} \rightarrow \text{Na}_2\text{S}$

$\text{Na} + \text{Cl} \rightarrow \text{NaCl}$

$\text{Na} + \text{Br} \rightarrow \text{NaBr}$

$\text{Na} + \text{I} \rightarrow \text{NaI}$

10 **(c)**
The stability of carbocation is explained on the basis of hyperconjugation and inductive effect hence the stability order of carbocation is $3° > 2° > 1° >$Methyl carbocation

11 **(a)**
In sublimation, certain substances when heated, first directly convert from the solid to the vapour state without melting. The vapour when cooled, give back the solid substance

12 **(a)**
Molality does not depend upon volume of the solution as molarity or normality. So, it does not depend upon temperature.

13 **(b)**
In Victor Mayer method, a known mass of the substance is converted into vapour by dropping in a hot tube. The vapour displaces its own volume of air which is collected over water and its volume measured at the observed temperature and pressure

14 **(a)**
Normality = molarity × n

(n =mol wt. of solute /eq. wt. of solute)

15 **(b)**

Carbon in carbanion is sp^3-hybridised with one orbital occupied by a lone pair

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

Matrix-Match Type

This section contain(s) 0 question(s). Each question contains Statements given in 2 columns which have to be matched. Statements (A, B, C, D) in **columns I** have to be matched with Statements (p, q, r, s) in **columns II**.

1. Compare List I and List II and choose the correct matching codes from the choices given.

	Column-I		Column- II
(A)	Glycerol	(1)	Sublimation
(B)	*o*-nitrophenol	(2)	Beilstein's test
(C)	Anthracene	(3)	Victor-Meyer's method
(D)	Halogens	(4)	Steam distillation
(E)	Molecular weight	(5)	Vacuum distillation
		(6)	Eudiometry

CODES :

	A	**B**	**C**	**D**	**E**
a)	5	4	1	2	3
b)	4	5	1	6	3
c)	6	4	1	3	3
d)	5	4	6	2	3
e)	4	6	2	3	3

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

: ANSWER KEY :

1)	**a**

ORGANIC CHEMISTRY - SOME BASIC PRINCIPLES AND TECHNIQUES

CHEMISTRY

: HINTS AND SOLUTIONS :

1 **(a)**

Sublimation conversion of solid directly into gaseous phase. The liquid state does not exist.
e. *g*., naphthalene, anthracene.
Beilstein test simple chemical test for halogens.
Victor-Meyer's method Standard laboratory method for determining the molecular weight of a volatile liquid.
Steam distillation vaporisation of the volatile constituents of a liquid. It is special type of distillation for a temperature sensitive materials like aromatic natural compounds. *e*. *g*., *ortho* nitrophenol, cinnamaldehyde.
Vacuum distillation at reduced pressure. It is generally used in case of high boiling liquids which decompose below their normal boiling points, cannot be purified by distillation at atmospheric pressure *e*. *g*., glycerol.
Eudiometry it is the process of determining the constituents of a gaseous mixture by eudiometer. It is used for ascertaining the purity of air or amount of O_2 in it.