

9. If $\alpha x + \beta y = 109$ is the equation of the chord of the ellipse $\frac{x^2}{9} + \frac{y^2}{4} = 1$, whose mid point is $\left(\frac{5}{2}, \frac{1}{2}\right)$,

then $\alpha + \beta$ is equal to

- (1) 37 (2) 46
(3) 58 (4) 72

Ans. (3)

10. If all the words with or without meaning made using all the letters of the word "KANPUR" are arranged as in a dictionary, then the word at 440th position in this arrangement, is :

- (1) PRNAKU (2) PRKANU
(3) PRKAUN (4) PRNAUK

Ans. (3)

11. Let α, β ($\alpha \neq \beta$) be the values of m , for which the equations $x + y + z = 1$; $x + 2y + 4z = m$ and $x + 4y + 10z = m^2$ have infinitely many solutions.

Then the value of $\sum_{n=1}^{10} (n^\alpha + n^\beta)$ is equal to :

- (1) 440 (2) 3080
(3) 3410 (4) 560

Ans. (1)

12. Let $A = [a_{ij}]$ be a matrix of order 3×3 , with $a_{ij} = (\sqrt{2})^{i+j}$. If the sum of all the elements in the third row of A^2 is $\alpha + \beta\sqrt{2}$, $\alpha, \beta \in \mathbf{Z}$, then $\alpha + \beta$ is equal to

- (1) 280 (2) 168
(3) 210 (4) 224

Ans. (4)

13. Let P be the foot of the perpendicular from the point $(1, 2, 2)$ on the line $L : \frac{x-1}{1} = \frac{y+1}{-1} = \frac{z-2}{2}$.

Let the line $\vec{r} = (-\hat{i} + \hat{j} - 2\hat{k}) + \lambda(\hat{i} - \hat{j} + \hat{k})$, $\lambda \in \mathbf{R}$,

intersect the line L at Q. Then $2(PQ)^2$ is equal to:

- (1) 27 (2) 25
(3) 29 (4) 19

Ans. (1)

14. Let a circle C pass through the points $(4, 2)$ and $(0, 2)$, and its centre lie on $3x + 2y + 2 = 0$. Then the length of the chord, of the circle C, whose mid-point is $(1, 2)$, is:

- (1) $\sqrt{3}$ (2) $2\sqrt{3}$
(3) $4\sqrt{2}$ (4) $2\sqrt{2}$

Ans. (2)

15. Let $A = [a_{ij}]$ be a 2×2 matrix such that $a_{ij} \in \{0, 1\}$ for all i and j . Let the random variable X denote the possible values of the determinant of the matrix A. Then, the variance of X is:

- (1) $\frac{1}{4}$ (2) $\frac{3}{8}$
(3) $\frac{5}{8}$ (4) $\frac{3}{4}$

Ans. (2)

16. Bag 1 contains 4 white balls and 5 black balls, and Bag 2 contains n white balls and 3 black balls. One ball is drawn randomly from Bag 1 and transferred to Bag 2. A ball is then drawn randomly from Bag 2. If the probability, that the ball drawn is white, is $\frac{29}{45}$, then n is equal to:

- (1) 3 (2) 4
(3) 5 (4) 6

Ans. (4)

17. The remainder, when 7^{103} is divided by 23, is equal to:

- (1) 14 (2) 9
(3) 17 (4) 6

Ans. (1)

18. Let $f(x) = \int_0^x t(t^2 - 9t + 20) dt$, $1 \leq x \leq 5$. If the range of f is $[\alpha, \beta]$, then $4(\alpha + \beta)$ equals:

- (1) 157 (2) 253
(3) 125 (4) 154

Ans. (1)

19. Let \hat{a} be a unit vector perpendicular to the vectors $\vec{b} = \hat{i} - 2\hat{j} + 3\hat{k}$ and $\vec{c} = 2\hat{i} + 3\hat{j} - \hat{k}$, and makes an angle of $\cos^{-1}\left(-\frac{1}{3}\right)$ with the vector $\hat{i} + \hat{j} + \hat{k}$. If \hat{a} makes an angle of $\frac{\pi}{3}$ with the vector $\hat{i} + \alpha\hat{j} + \hat{k}$, then the value of α is :

- (1) $-\sqrt{3}$ (2) $\sqrt{6}$
 (3) $-\sqrt{6}$ (4) $\sqrt{3}$

Ans. (3)

20. If for the solution curve $y = f(x)$ of the differential equation $\frac{dy}{dx} + (\tan x)y = \frac{2 + \sec x}{(1 + 2\sec x)^2}$,

$x \in \left(\frac{-\pi}{2}, \frac{\pi}{2}\right)$, $f\left(\frac{\pi}{3}\right) = \frac{\sqrt{3}}{10}$, then $f\left(\frac{\pi}{4}\right)$ is equal to:

- (1) $\frac{9\sqrt{3} + 3}{10(4 + \sqrt{3})}$ (2) $\frac{\sqrt{3} + 1}{10(4 + \sqrt{3})}$
 (3) $\frac{5 - \sqrt{3}}{2\sqrt{2}}$ (4) $\frac{4 - \sqrt{2}}{14}$

Ans. (4)

SECTION-B

21. If $24 \int_0^{\frac{\pi}{4}} \left(\sin \left\lfloor 4x - \frac{\pi}{12} \right\rfloor + [2 \sin x] \right) dx = 2\pi + \alpha$, where

$\lfloor \cdot \rfloor$ denotes the greatest integer function, then α is equal to _____.

Ans. (12)

22. If $\lim_{t \rightarrow 0} \left(\int_0^1 (3x + 5)^t dx \right)^{\frac{1}{t}} = \frac{\alpha}{5e} \left(\frac{8}{5} \right)^{\frac{2}{3}}$, then α is equal

to _____.

Ans. (64)

23. Let $a_1, a_2, \dots, a_{2024}$ be an Arithmetic Progression such that $a_1 + (a_5 + a_{10} + a_{15} + \dots + a_{2020}) + a_{2024} = 2233$. Then $a_1 + a_2 + a_3 + \dots + a_{2024}$ is equal to _____.

Ans. (11132)

24. Let integers $a, b \in [-3, 3]$ be such that $a + b \neq 0$. Then the number of all possible ordered pairs

(a, b) , for which $\left| \frac{z-a}{z+b} \right| = 1$ and $\begin{vmatrix} z+1 & \omega & \omega^2 \\ \omega & z+\omega^2 & 1 \\ \omega^2 & 1 & z+\omega \end{vmatrix}$

$= 1$, $z \in \mathbb{C}$, where ω and ω^2 are the roots of $x^2 + x + 1 = 0$, is equal to _____.

Ans. (10)

25. Let $y^2 = 12x$ the parabola and S be its focus. Let PQ be a focal chord of the parabola such that $(SP)(SQ) = \frac{147}{4}$. Let C be the circle described taking PQ as a diameter. If the equation of a circle C is $64x^2 + 64y^2 - \alpha x - 64\sqrt{3}y = \beta$, then $\beta - \alpha$ is equal to _____.

Ans. (1328)

JEE–MAIN EXAMINATION – JANUARY 2025

(HELD ON WEDNESDAY 29th JANUARY 2025)

TIME : 3 : 00 PM TO 6 : 00 PM

PHYSICS

TEST PAPER WITH ANSWER

SECTION-A

26. The difference of temperature in a material can convert heat energy into electrical energy. To harvest the heat energy, the material should have
- (1) low thermal conductivity and low electrical conductivity
 - (2) high thermal conductivity and high electrical conductivity
 - (3) low thermal conductivity and high electrical conductivity
 - (4) high thermal conductivity and low electrical conductivity

Ans. (3)

27. Given below are two statements. One is labelled as **Assertion (A)** and the other is labelled as **Reason (R)**.

Assertion (A) : With the increase in the pressure of an ideal gas, the volume falls off more rapidly in an isothermal process in comparison to the adiabatic process.

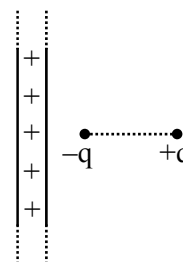
Reason (R) : In isothermal process, $PV = \text{constant}$, while in adiabatic process $PV^\gamma = \text{constant}$. Here γ is the ratio of specific heats, P is the pressure and V is the volume of the ideal gas.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both (A) and (R) are true but (R) is NOT the correct explanation of (A)
- (2) (A) is true but (R) is false
- (3) Both (A) and (R) are true and (R) is the correct explanation of (A).
- (4) (A) is false but (R) is true

Ans. (4)

28. An electric dipole is placed at a distance of 2 cm from an infinite plane sheet having positive charge density σ_0 . Choose the correct option from the following.



- (1) Torque on dipole is zero and net force is directed away from the sheet.
- (2) Torque on dipole is zero and net force acts towards the sheet.
- (3) Potential energy of dipole is minimum and torque is zero.
- (4) Potential energy and torque both are maximum

Ans. (3)

29. In an experiment with photoelectric effect, the stopping potential.

- (1) increases with increase in the wavelength of the incident light
- (2) increases with increase in the intensity of the incident light
- (3) is $\left(\frac{1}{e}\right)$ times the maximum kinetic energy of the emitted photoelectrons
- (4) decreases with increase in the intensity of the incident light

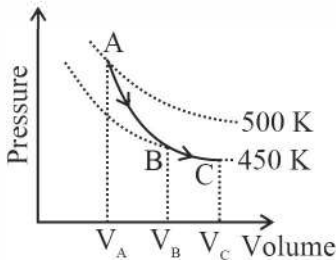
Ans. (3)

30. A point charge causes an electric flux of $-2 \times 10^4 \text{ Nm}^2\text{C}^{-1}$ to pass through a spherical Gaussian surface of 8.0 cm radius, centred on the charge. The value of the point charge is :

(Given $\epsilon_0 = 8.85 \times 10^{-12} \text{ C}^2\text{N}^{-1}\text{m}^{-2}$)

- (1) $-17.7 \times 10^{-8} \text{ C}$ (2) $-15.7 \times 10^{-8} \text{ C}$
 (3) $17.7 \times 10^{-8} \text{ C}$ (4) $15.7 \times 10^{-8} \text{ C}$

Ans. (1)



31.

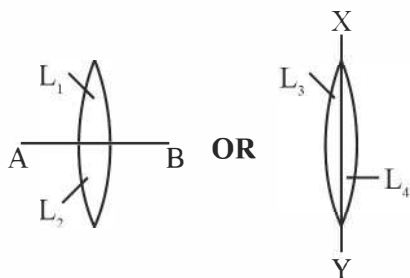
A poly-atomic molecule ($C_v = 3R$, $C_p = 4R$, where R is gas constant) goes from phase space point A ($P_A = 10^5 \text{ Pa}$, $V_A = 4 \times 10^{-6} \text{ m}^3$) to point B ($P_B = 5 \times 10^4 \text{ Pa}$, $V_B = 6 \times 10^{-6} \text{ m}^3$) to point C ($P_C = 10^4 \text{ Pa}$, $V_C = 8 \times 10^{-6} \text{ m}^3$). A to B is an adiabatic path and B to C is an isothermal path.

The net heat absorbed per unit mole by the system is :

- (1) $500R(\ln 3 + \ln 4)$ (2) $450R(\ln 4 - \ln 3)$
 (3) $500R \ln 2$ (4) $400R \ln 4$

Ans. (2)

32. Two identical symmetric double convex lenses of focal length f are cut into two equal parts L_1, L_2 by AB plane and L_3, L_4 by XY plane as shown in figure respectively. The ratio of focal lengths of lenses L_1 and L_3 is



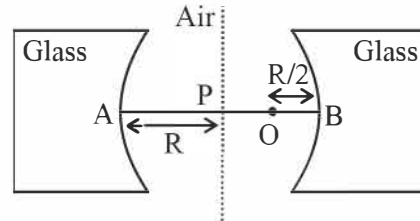
- (1) 1 : 4 (2) 1 : 1
 (3) 2 : 1 (4) 1 : 2

Ans. (4)

33. A plane electromagnetic wave propagates along the + x direction in free space. The components of the electric field, \vec{E} and magnetic field, \vec{B} vectors associated with the wave in Cartesian frame are :

- (1) E_y, B_x (2) E_y, B_z
 (3) E_x, B_y (4) E_z, B_y

Ans. (2)



34.

Two concave refracting surfaces of equal radii of curvature and refractive index 1.5 face each other in air as shown in figure. A point object O is placed midway, between P and B. The separation between the images of O, formed by each refracting surface is :

- (1) $0.214R$
 (2) $0.114R$
 (3) $0.411R$
 (4) $0.124R$

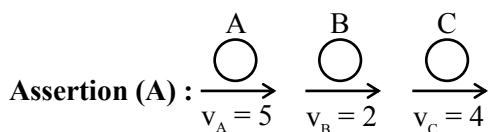
Ans. (2)

35. Two bodies A and B of equal mass are suspended from two massless springs of spring constant k_1 and k_2 , respectively. If the bodies oscillate vertically such that their amplitudes are equal, the ratio of the maximum velocity of A to the maximum velocity of B is

- (1) $\sqrt{\frac{k_1}{k_2}}$
 (2) $\frac{k_1}{k_2}$
 (3) $\frac{k_2}{k_1}$
 (4) $\sqrt{\frac{k_2}{k_1}}$

Ans. (1)

36. Given below are two statements. One is labelled as **Assertion (A)** and the other is labelled as **Reason (R)**.



Three identical spheres of same mass undergo one dimensional motion as shown in figure with initial velocities $v_A = 5$ m/s, $v_B = 2$ m/s, $v_C = 4$ m/s. If we wait sufficiently long for elastic collision to happen, then $v_A = 4$ m/s, $v_B = 2$ m/s, $v_C = 5$ m/s will be the final velocities.

Reason (R) : In an elastic collision between identical masses, two objects exchange their velocities.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both **(A)** and **(R)** are true but **(R)** is **NOT** the correct explanation of **(A)**
- (2) **(A)** is true but **(R)** is false
- (3) Both **(A)** and **(R)** are true and **(R)** is the correct explanation of **(A)**.
- (4) **(A)** is false but **(R)** is true

Ans. (4)

37. A sand dropper drops sand of mass $m(t)$ on a conveyer belt at a rate proportional to the square root of speed (v) of the belt, i.e. $\frac{dm}{dt} \propto \sqrt{v}$. If P is the power delivered to run the belt at constant speed then which of the following relationship is true ?

- (1) $P^2 \propto v^3$
- (2) $P \propto \sqrt{v}$
- (3) $P \propto v$
- (4) $P^2 \propto v^5$

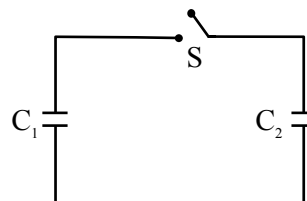
Ans. (4)

38. A convex lens made of glass (refractive index = 1.5) has focal length 24 cm in air. When it is totally immersed in water (refractive index = 1.33), its focal length changes to

- (1) 72 cm
- (2) 96 cm
- (3) 24 cm
- (4) 48 cm

Ans. (2)

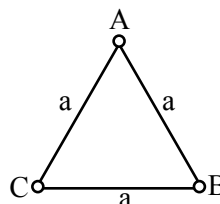
39. A capacitor, $C_1 = 6\mu\text{F}$ is charged to a potential difference of $V_0 = 5\text{V}$ using a 5V battery. The battery is removed and another capacitor, $C_2 = 12\mu\text{F}$ is inserted in place of the battery. When the switch 'S' is closed, the charge flows between the capacitors for some time until equilibrium condition is reached. What are the charges (q_1 and q_2) on the capacitors C_1 and C_2 when equilibrium condition is reached.



- (1) $q_1 = 15\mu\text{C}$, $q_2 = 30\mu\text{C}$
- (2) $q_1 = 30\mu\text{C}$, $q_2 = 15\mu\text{C}$
- (3) $q_1 = 10\mu\text{C}$, $q_2 = 20\mu\text{C}$
- (4) $q_1 = 20\mu\text{C}$, $q_2 = 10\mu\text{C}$

Ans. (3)

40.



Three equal masses m are kept at vertices (A, B, C) of an equilateral triangle of side a in free space. At $t = 0$, they are given an initial velocity $\vec{V}_A = V_0 \vec{AC}$, $\vec{V}_B = V_0 \vec{BA}$ and $\vec{V}_C = V_0 \vec{CB}$. Here, \vec{AC} , \vec{CB} and \vec{BA} are unit vectors along the edges of the triangle. If the three masses interact gravitationally, then the magnitude of the net angular momentum of the system at the point of collision is :

- (1) $\frac{1}{2} a m V_0$
- (2) $3 a m V_0$
- (3) $\frac{\sqrt{3}}{2} a m V_0$
- (4) $\frac{3}{2} a m V_0$

Ans. (3)

SECTION-B

46. The magnetic field inside a 200 turns solenoid of radius 10 cm is 2.9×10^{-4} Tesla. If the solenoid carries a current of 0.29 A, then the length of the solenoid is _____ π cm.

Ans. (8)

47. A parallel plate capacitor consisting of two circular plates of radius 10 cm is being charged by a constant current of 0.15 A. If the rate of change of potential difference between the plates is 7×10^8 V/s then the integer value of the distance between the parallel plates is –

$$\left(\text{Take, } \epsilon_0 = 9 \times 10^{-12} \frac{\text{F}}{\text{m}}, \pi = \frac{22}{7} \right) \text{ _____ } \mu\text{m.}$$

Ans. (1320)

48. A physical quantity Q is related to four observables

$$a, b, c, d \text{ as follows : } Q = \frac{ab^4}{cd}$$

where, $a = (60 \pm 3)\text{Pa}$; $b = (20 \pm 0.1)\text{m}$;

$c = (40 \pm 0.2)\text{Nsm}^{-2}$ and $d = (50 \pm 0.1)\text{m}$,

then the percentage error in Q is $\frac{x}{1000}$,

where $x = \text{_____}$.

NTA Ans. (77)

Allen Ans.(7700)

49. Two planets, A and B are orbiting a common star in circular orbits of radii R_A and R_B , respectively, with $R_B = 2R_A$. The planet B is $4\sqrt{2}$ times more massive than planet A. The ratio $\left(\frac{L_B}{L_A} \right)$ of angular momentum (L_B) of planet B to that of planet A (L_A) is closest to integer _____.

Ans. (8)

50. Two cars P and Q are moving on a road in the same direction. Acceleration of car P increases linearly with time whereas car Q moves with a constant acceleration. Both cars cross each other at time $t = 0$, for the first time. The maximum possible number of crossing(s) (including the crossing at $t = 0$) is _____.

Ans. (3)

JEE–MAIN EXAMINATION – JANUARY 2025

(HELD ON WEDNESDAY 29th JANUARY 2025)

TIME : 3 : 00 PM TO 6 : 00 PM

CHEMISTRY

TEST PAPER WITH ANSWER

SECTION-A

51. The calculated spin-only magnetic moments of $K_3[Fe(OH)_6]$ and $K_4[Fe(OH)_6]$ respectively are :

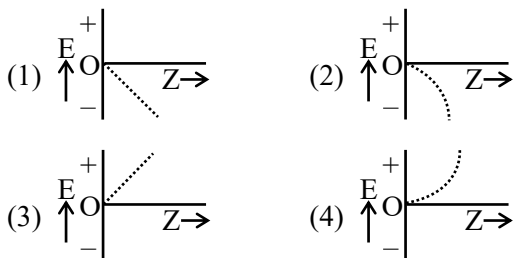
- (1) 4.90 and 4.90 B.M.
- (2) 5.92 and 4.90 B.M.
- (3) 3.87 and 4.90 B.M.
- (4) 4.90 and 5.92 B.M.

Ans. (2)

52. For hydrogen like species, which of the following graphs provides the most appropriate representation of E vs Z plot for a constant n ?

[E : Energy of the stationary state,

Z : atomic number, n = principal quantum number]



Ans. (2)

53. Given below are two statements :

Statement (I) : In partition chromatography, stationary phase is thin film of liquid present in the inert support.

Statement (II) : In paper chromatography, the material of paper acts as a stationary phase.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both **Statement I** and **Statement II** are false
- (2) **Statement I** is true but **Statement II** is false
- (3) Both **Statement I** and **Statement II** are true
- (4) **Statement I** is false but **Statement II** is true

Ans. (2)

54. Identify the essential amino acids from below :

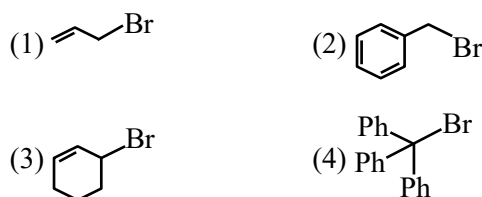
- | | |
|--------------|---------------|
| (A) Valine | (B) Proline |
| (C) Lysine | (D) Threonine |
| (E) Tyrosine | |

Choose the **correct** answer from the options given below :

- (1) (A),(C) and (D) only
- (2) (A),(C) and (E) only
- (3) (B),(C) and (E) only
- (4) (C),(D) and (E) only

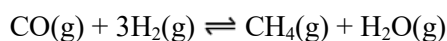
Ans. (1)

55. Which among the following halides will generate the most stable carbocation in Nucleophilic substitution reaction ?



Ans. (4)

56. Consider the equilibrium



If the pressure applied over the system increases by two fold at constant temperature then

- (A) Concentration of reactants and products increases.
- (B) Equilibrium will shift in forward direction.
- (C) Equilibrium constant increases since concentration of products increases.
- (D) Equilibrium constant remains unchanged as concentration of reactants and products remain same.

Choose the **correct** answer from the options given below :

- (1) (A) and (B) only
- (2) (A), (B) and (D) only
- (3) (B) and (C) only
- (4) (A), (B) and (C) only

Ans. (1)

57. Given below are two statements :

Statement (I) : NaCl is added to the ice at 0°C, present in the ice cream box to prevent the melting of ice cream.

Statement (II) : On addition of NaCl to ice at 0°C, there is a depression in freezing point.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) **Statement I** is false but **Statement II** is true
- (2) Both **Statement I** and **Statement II** are true
- (3) Both **Statement I** and **Statement II** are false
- (4) **Statement I** is true but **Statement II** is false

Ans. (2)

58. Given below are two statements :

Statement (I) : On nitration of m-xylene with HNO₃, H₂SO₄ followed by oxidation, 4-nitrobenzene-1, 3-dicarboxylic acid is obtained as the major product.

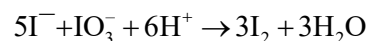
Statement (II) : CH₃ group is o/p-directing while-NO₂ group is m-directing group.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both **Statement I** and **Statement II** are false
- (2) **Statement I** is false but **Statement II** is true
- (3) Both **Statement I** and **Statement II** are true
- (4) **Statement I** is true but **Statement II** is false

Ans. (3)

59. 0.1 M solution of KI reacts with excess of H₂SO₄ and KIO₃ solution. According to equation



Identify the **correct** statements :

- (A) 200 mL of KI solution reacts with 0.004 mol of KIO₃
- (B) 200 mL of KI solution reacts with 0.006 mol of H₂SO₄
- (3) 0.5 L of KI solution produced 0.005 mol of I₂
- (4) Equivalent weight of KIO₃ is equal to

$$\left(\frac{\text{Molecular weight}}{5} \right)$$

Choose the **correct** answer from the options given below :

- (1) (A) and (D) only (2) (B) and (C) only
- (3) (A) and (B) only (4) (C) and (D) only

Ans. (1)

60. Match List-I with List-II :

List-I Applications		List-II Batteries/Cell	
(A)	Transistors	(I)	Anode - Zn/Hg ; Cathode - HgO + C
(B)	Hearing aids	(II)	Hydrogen fuel cell
(C)	Invertors	(III)	Anode - Zn; Cathode - Carbon
(D)	Apollo space ship	(IV)	Anode - Pb ; Cathode - Pb PbO ₂

Choose the **correct** answer from the options given below :

- (1) (A)-(III), (B)-(I), (C)-(IV), (D)-(II)
- (2) (A)-(III), (B)-(II), (C)-(IV), (D)-(I)
- (3) (A)-(IV), (B)-(III), (C)-(II), (D)-(I)
- (4) (A)-(II), (B)-(III), (C)-(IV), (D)-(I)

Ans. (1)

61. O₂ gas will be evolved as a product of electrolysis of :

- (A) an aqueous solution of AgNO₃ using silver electrodes.
- (B) an aqueous solution of AgNO₃ using platinum electrodes.
- (C) a dilute solution of H₂SO₄ using platinum electrodes.
- (D) a high concentration solution of H₂SO₄ using platinum electrodes.

Choose the **correct** answer from the options given below :

- (1) (B) and (C) only (2) (A) and (D) only
- (3) (B) and (D) only (4) (A) and (C) only

Ans. (1)

62. Identify the homoleptic complexes with odd number of d electrons in the central metal.

- (A) [FeO₄]²⁻ (B) [Fe(CN)₆]³⁻
- (C) [Fe(CN)₅NO]²⁻ (D) [CoCl₄]²⁻
- (E) [Co(H₂O)₃F₃]

Choose the **correct** answer from the options given below :

- (1) (B) and (D) only
- (2) (C) and (E) only
- (3) (A), (B) and (D) only
- (4) (A), (C) and (E) only

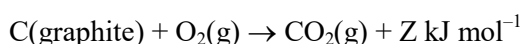
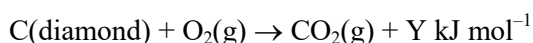
Ans. (1)

63. Total number of sigma (σ) _____ and pi (π) _____ bonds respectively present in hex-1-en-4-yne are :

- (1) 13 and 3
- (2) 11 and 3
- (3) 3 and 13
- (4) 14 and 3

Ans. (1)

64. If $C(\text{diamond}) \rightarrow C(\text{graphite}) + X \text{ kJ mol}^{-1}$



At constant temperature. Then

- (1) $X = Y + Z$
- (2) $-X = Y + Z$
- (3) $X = -Y + Z$
- (4) $X = Y - Z$

Ans. (4)

65. Given below are two statements :

Statement (I): It is impossible to specify simultaneously with arbitrary precision, both the linear momentum and the position of a particle.

Statement (II) : If the uncertainty in the measurement of position and uncertainty in measurement of momentum are equal for an electron, then the uncertainty in the measurement

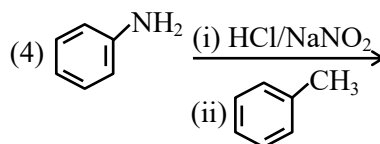
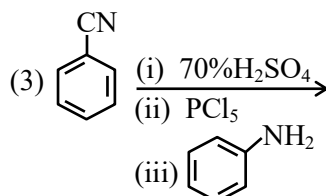
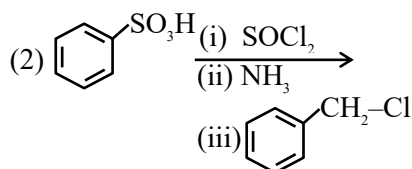
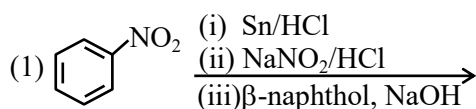
$$\text{of velocity is } \geq \sqrt{\frac{h}{\pi}} \times \frac{1}{2m}.$$

In the light of the above statements, choose the correct answer from the options given below :

- (1) **Statement I** is true but **Statement II** is false.
- (2) Both **Statement I** and **Statement II** are true.
- (3) **Statement I** is false but **Statement II** is true.
- (4) Both **Statement I** and **Statement II** are false.

Ans. (2)

66. Which one of the following reaction sequences will give an azo dye ?



Ans. (1)

67. Drug X becomes ineffective after 50% decomposition. The original concentration of drug in a bottle was 16 mg/mL which becomes 4 mg/mL in 12 months. The expiry time of the drug in months is _____.

Assume that the decomposition of the drug follows first order kinetics.

- (1) 12
- (2) 2
- (3) 3
- (4) 6

Ans. (4)

68. The type of oxide formed by the element among Li, Na, Be, Mg, B and Al that has the least atomic radius is :

- (1) A_2O_3
- (2) AO_2
- (3) AO
- (4) A_2O

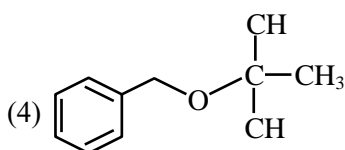
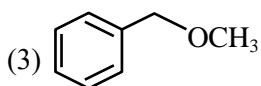
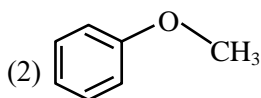
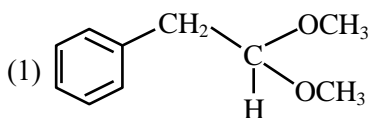
Ans. (1)

69. First ionisation enthalpy values of first four group 15 elements are given below. Choose the correct value for the element that is a main component of apatite family :

- (1) 1012 kJ mol^{-1}
- (2) 1402 kJ mol^{-1}
- (3) 834 kJ mol^{-1}
- (4) 947 kJ mol^{-1}

Ans. (1)

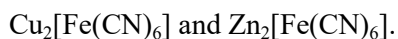
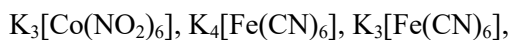
70. Which one of the following, with HBr will give a phenol ?



Ans. (2)

SECTION-B

71. Consider the following low-spin complexes



The sum of the spin-only magnetic moment values of complexes having yellow colour is _____

B.M. (answer is nearest integer)

Ans. (0)

72. Isomeric hydrocarbons \rightarrow negative Baeyer's test

(Molecular formula C_9H_{12})

The total number of isomers from above with four different non-aliphatic substitution sites is -

Ans. (2)

73. In the Claisen-Schmidt reaction to prepare, dibenzalacetone from 5.3 g benzaldehyde, a total of 3.51 g of product was obtained. The percentage yield in this reaction was _____ %.

Ans. (60)

74. In the sulphur estimation, 0.20 g of a pure organic compound gave 0.40 g of barium sulphate. The percentage of sulphur in the compound is _____ $\times 10^{-1}$ %.

(Molar mass : O = 16, S = 32, Ba = 137 in $g\ mol^{-1}$)

Ans. (275)

75. Total number of non bonded electrons present in NO_2^- ion based on Lewis theory is _____ .

Ans. (12)