

8. Let the equation of the circle, which touches x-axis at the point (a, 0), $a > 0$ and cuts off an intercept of length b on y-axis be $x^2 + y^2 - \alpha x + \beta y + \gamma = 0$. If the circle lies below x-axis, then the ordered pair $(2a, b^2)$ is equal to :

- (1) $(\alpha, \beta^2 + 4\gamma)$ (2) $(\gamma, \beta^2 - 4\alpha)$
 (3) $(\gamma, \beta^2 + 4\alpha)$ (4) $(\alpha, \beta^2 - 4\gamma)$

Ans. (4)

9. Let $\langle a_n \rangle$ be a sequence such that $a_0 = 0$, $a_1 = \frac{1}{2}$ and $2a_{n+2} = 5a_{n+1} - 3a_n$, $n = 0, 1, 2, 3, \dots$. Then

$\sum_{k=1}^{100} a_k$ is equal to :

- (1) $3a_{99} - 100$ (2) $3a_{100} - 100$
 (3) $3a_{100} + 100$ (4) $3a_{99} + 100$

Ans. (2)

10. $\cos\left(\sin^{-1}\frac{3}{5} + \sin^{-1}\frac{5}{13} + \sin^{-1}\frac{33}{65}\right)$ is equal to :

- (1) 1 (2) 0
 (3) $\frac{33}{65}$ (4) $\frac{32}{65}$

Ans. (2)

11. Let T_r be the r^{th} term of an A.P. If for some m,

$$T_m = \frac{1}{25}, T_{25} = \frac{1}{20} \text{ and } 20 \sum_{r=1}^{25} T_r = 13, \text{ then}$$

$5m \sum_{r=m}^{2m} T_r$ is equal to :

- (1) 112 (2) 126
 (3) 98 (4) 142

Ans. (2)

12. If the image of the point (4, 4, 3) in the line

$$\frac{x-1}{2} = \frac{y-2}{1} = \frac{z-1}{3} \text{ is } (\alpha, \beta, \gamma), \text{ then } \alpha + \beta + \gamma \text{ is}$$

equal to

- (1) 9 (2) 12
 (3) 8 (4) 7

Ans. (1)

13. If $\int_{-\frac{\pi}{2}}^{\frac{\pi}{2}} \frac{96x^2 \cos^2 x}{(1+e^x)} dx = \pi(\alpha\pi^2 + \beta)$, $\alpha, \beta \in Z$, then

$(\alpha + \beta)^2$ equals :

- (1) 144 (2) 196
 (3) 100 (4) 64

Ans. (3)

14. The sum of all local minimum values of the

$$\text{function } f(x) = \begin{cases} 1-2x, & x < -1 \\ \frac{1}{3}(7+2|x|), & -1 \leq x \leq 2 \\ \frac{11}{18}(x-4)(x-5), & x > 2 \end{cases}$$

- (1) $\frac{171}{72}$ (2) $\frac{131}{72}$
 (3) $\frac{157}{72}$ (4) $\frac{167}{72}$

Ans. (3)

15. The sum, of the squares of all the roots of the equation $x^2 + |2x - 3| - 4 = 0$, is :

- (1) $3(3 - \sqrt{2})$ (2) $6(3 - \sqrt{2})$
 (3) $6(2 - \sqrt{2})$ (4) $3(2 - \sqrt{2})$

Ans. (3)

16. Let for some function $y = f(x)$, $\int_0^x t f(t) dt = x^2 f(x)$,

$x > 0$ and $f(2) = 3$. Then $f(6)$ is equal to :

- (1) 1 (2) 2
 (3) 6 (4) 3

Ans. (1)

17. Let ${}^n C_{r-1} = 28$, ${}^n C_r = 56$ and ${}^n C_{r+1} = 70$. Let $A(4\cos t, 4\sin t)$, $B(2\sin t, -2\cos t)$ and $C(3r - n, r^2 - n - 1)$ be the vertices of a triangle ABC, where t is a parameter. If $(3x - 1)^2 + (3y)^2 = \alpha$, is the locus of the centroid of triangle ABC, then α equals :

- (1) 20 (2) 8
 (3) 6 (4) 18

Ans. (1)

JEE–MAIN EXAMINATION – JANUARY 2025

(HELD ON TUESDAY 28th JANUARY 2025)

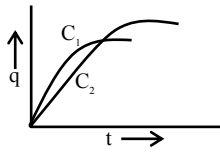
TIME : 9 : 00 AM TO 12 : 00 NOON

PHYSICS

TEST PAPER WITH ANSWER

SECTION-A

26. Two capacitors C_1 and C_2 are connected in parallel to a battery. Charge-time graph is shown below for the two capacitors. The energy stored with them are U_1 and U_2 , respectively. Which of the given statements is true ?



- (1) $C_1 > C_2, U_1 > U_2$ (2) $C_2 > C_1, U_2 < U_1$
 (3) $C_1 > C_2, U_1 < U_2$ (4) $C_2 > C_1, U_2 > U_1$

Ans. (4)

27. In the experiment for measurement of viscosity ' η ' of given liquid with a ball having radius R , consider following statements.
- Graph between terminal velocity V and R will be a parabola
 - The terminal velocities of different diameter balls are constant for a given liquid.
 - Measurement of terminal velocity is dependent on the temperature.
 - This experiment can be utilized to assess the density of a given liquid.
 - If balls are dropped with some initial speed, the value of η will change.

Choose the correct answer from the options given below:

- (1) B, D and E only
 (2) A, C and D only
 (3) C, D and E only
 (4) A, B and E only

Ans. (2)

28. Consider following statements:
- Surface tension arises due to extra energy of the molecules at the interior as compared to the molecules at the surface, of a liquid.
 - As the temperature of liquid rises, the coefficient of viscosity increases.
 - As the temperature of gas increases, the coefficient of viscosity increases.
 - The onset of turbulence is determined by Reynold's number.
 - In a steady flow two stream lines never intersect.

Choose the correct answer from the options given below :

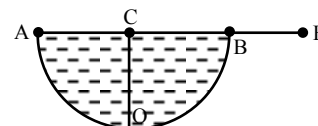
- (1) A, D, E only
 (2) C, D, E only
 (3) B, C, D only
 (4) A, B, C only

Ans. (2)

29. Three infinitely long wires with linear charge density λ are placed along the x -axis, y -axis and z -axis respectively. Which of the following denotes an equipotential surface ?
- $xy + yz + zx = \text{constant}$
 - $(x + y)(y + z)(z + x) = \text{constant}$
 - $(x^2 + y^2)(y^2 + z^2)(z^2 + x^2) = \text{constant}$
 - $xyz = \text{constant}$

Ans. (3)

30. A hemispherical vessel is completely filled with a liquid of refractive index μ . A small coin is kept at the lowest point (O) of the vessel as shown in figure. The minimum value of the refractive index of the liquid so that a person can see the coin from point E (at the level of the vessel) is _____.



- (1) $\sqrt{3}$ (2) $\frac{3}{2}$
 (3) $\sqrt{2}$ (4) $\frac{\sqrt{3}}{2}$

Ans. (3)

31. Consider a long thin conducting wire carrying a uniform current I . A particle having mass " M " and charge " q " is released at a distance " a " from the wire with a speed v_0 along the direction of current in the wire. The particle gets attracted to the wire due to magnetic force. The particle turns round when it is at distance x from the wire. The value of x is [μ_0 is vacuum permeability]

(1) $a \left[1 - \frac{mv_0}{2q\mu_0 I} \right]$ (2) $\frac{a}{2}$

(3) $a \left[1 - \frac{mv_0}{q\mu_0 I} \right]$ (4) $ae^{-\frac{4\pi mv_0}{q\mu_0 I}}$

Ans. (4)

32. A Carnot engine (E) is working between two temperatures 473K and 273K. In a new system two engines – engine E_1 works between 473K to 373K and engine E_2 works between 373K to 273K. If η_{12} , η_1 and η_2 are the efficiencies of the engines E, E_1 and E_2 , respectively, then

(1) $\eta_{12} < \eta_1 + \eta_2$ (2) $\eta_{12} = \eta_1 \eta_2$
 (3) $\eta_{12} = \eta_1 + \eta_2$ (4) $\eta_{12} \geq \eta_1 + \eta_2$

Ans. (1)

33. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R**
Assertion A: A sound wave has higher speed in solids than gases.

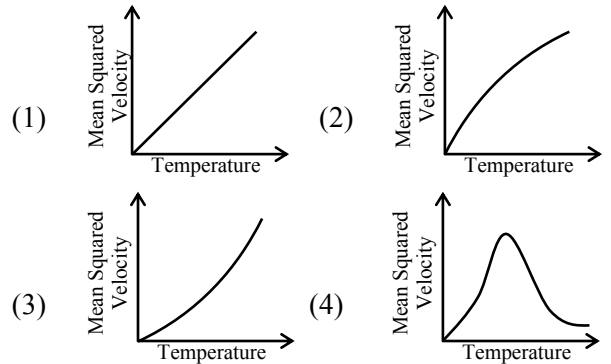
Reason R: Gases have higher value of Bulk modulus than solids.

In the light of the above statements, choose the **correct** answer from the options given below

- (1) Both **A** and **R** are true and **R** is the correct explanation of **A**
- (2) **A** is false but **R** is true
- (3) Both **A** and **R** are true but **R** is **NOT** the correct explanation of **A**
- (4) **A** is true but **R** is false.

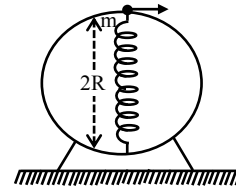
Ans. (4)

34. For a particular ideal gas which of the following graphs represents the variation of mean squared velocity of the gas molecules with temperature ?



Ans. (1)

35. A bead of mass ' m ' slides without friction on the wall of a vertical circular hoop of radius ' R ' as shown in figure. The bead moves under the combined action of gravity and a massless spring (k) attached to the bottom of the hoop. The equilibrium length of the spring is ' R '. If the bead is released from top of the hoop with (negligible) zero initial speed, velocity of bead, when the length of spring becomes ' R ', would be (spring constant is ' k ', g is acceleration due to gravity)



(1) $2\sqrt{gR + \frac{kR^2}{m}}$ (2) $\sqrt{2Rg + \frac{4kR^2}{m}}$
 (3) $\sqrt{2Rg + \frac{kR^2}{m}}$ (4) $\sqrt{3Rg + \frac{kR^2}{m}}$

Ans. (4)

36. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R**
Assertion A: In a central force field, the work done is independent of the path chosen

Reason R: Every force encountered in mechanics does not have an associated potential energy.

In the light of the above statements, choose the **most appropriate** answer from the options given below

- (1) **A** is true but **R** is false
- (2) Both **A** and **R** are true but **R** is **NOT** the correct explanation of **A**
- (3) Both **A** and **R** are true and **R** is the correct explanation of **A**
- (4) **A** is false but **R** is true

Ans. (2)

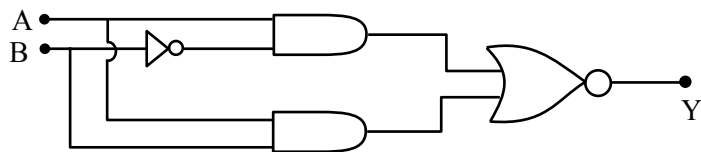
37. Choose the correct nuclear process from the below options

[p: proton, n: neutron, e^- : electron, e^+ : positron, ν : neutrino, $\bar{\nu}$: antineutrino]

- (1) $n \rightarrow p + e^- + \bar{\nu}$ (2) $n \rightarrow p + e^- + \nu$
 (3) $n \rightarrow p + e^+ + \bar{\nu}$ (4) $n \rightarrow p + e^+ + \nu$

Ans. (1)

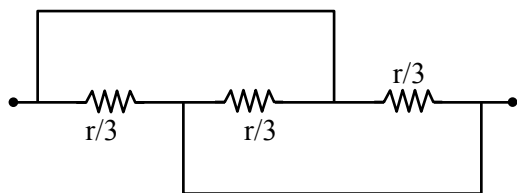
38. Which of the following circuits has the same output as that of the given circuit?



- (1) (2)
 (3) (4)

Ans. (1)

39. Find the equivalent resistance between two ends of the following circuit.



- (1) r (2) $\frac{r}{6}$
 (3) $\frac{r}{9}$ (4) $\frac{r}{3}$

Ans. (3)

40. A wire of resistance R is bent into an equilateral triangle and an identical wire is bent into a square. The ratio of resistance between the two end points of an edge of the triangle to that of the square is

- (1) $9/8$ (2) $8/9$
 (3) $27/32$ (4) $32/27$

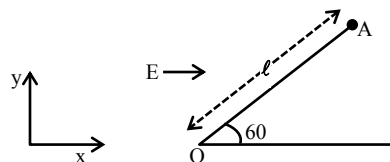
Ans. (4)

41. Due to presence of an em-wave whose electric component is given by $E = 100 \sin(\omega t - kx) \text{ NC}^{-1}$, a cylinder of length 200 cm holds certain amount of em-energy inside it. If another cylinder of same length but half diameter than previous one holds same amount of em-energy, the magnitude of the electric field of the corresponding em-wave should be modified as

- (1) $25 \sin(\omega t - kx) \text{ NC}^{-1}$
 (2) $200 \sin(\omega t - kx) \text{ NC}^{-1}$
 (3) $400 \sin(\omega t - kx) \text{ NC}^{-1}$
 (4) $50 \sin(\omega t - kx) \text{ NC}^{-1}$

Ans. (2)

42. A particle of mass ' m ' and charge ' q ' is fastened to one end 'A' of a massless string having equilibrium length ℓ , whose other end is fixed at point 'O'. The whole system is placed on a frictionless horizontal plane and is initially at rest. If uniform electric field is switched on along the direction as shown in figure, then the speed of the particle when it crosses the x-axis is



- (1) $\sqrt{\frac{2qE\ell}{m}}$ (2) $\sqrt{\frac{qE\ell}{4m}}$
 (3) $\sqrt{\frac{qE\ell}{m}}$ (4) $\sqrt{\frac{qE\ell}{2m}}$

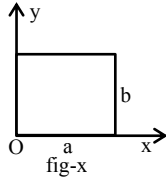
Ans. (3)

43. A proton of mass ' m_p ' has same energy as that of a photon of wavelength ' λ '. If the proton is moving at non-relativistic speed, then ratio of its de Broglie wavelength to the wavelength of photon is.

- (1) $\frac{1}{c} \sqrt{\frac{2E}{m_p}}$ (2) $\frac{1}{c} \sqrt{\frac{E}{m_p}}$
 (3) $\frac{1}{c} \sqrt{\frac{E}{2m_p}}$ (4) $\frac{1}{2c} \sqrt{\frac{E}{m_p}}$

Ans. (3)

44. The centre of mass of a thin rectangular plate (fig - x) with sides of length a and b, whose mass per unit area (σ) varies as $\sigma = \frac{\sigma_0 x}{ab}$ (where σ_0 is a constant), would be _____



- (1) $\left(\frac{2}{3}a, \frac{b}{2}\right)$ (2) $\left(\frac{2}{3}a, \frac{2}{3}b\right)$
 (3) $\left(\frac{a}{2}, \frac{b}{2}\right)$ (4) $\left(\frac{1}{3}a, \frac{b}{2}\right)$

Ans. (1)

45. A thin prism P_1 with angle 4° made of glass having refractive index 1.54, is combined with another thin prism P_2 made of glass having refractive index 1.72 to get dispersion without deviation. The angle of the prism P_2 in degrees is

- (1) 4 (2) 3
 (3) 16/3 (4) 1.5

Ans. (2)

46. A tiny metallic rectangular sheet has length and breadth of 5 mm and 2.5mm, respectively. Using a specially designed screw gauge which has pitch of 0.75 mm and 15 divisions in the circular scale, you are asked to find the area of the sheet. In this measurement, the maximum fractional error will

be $\frac{x}{100}$ where x is _____

Ans. (3)

47. The moment of inertia of a solid disc rotating along its diameter is 2.5 times higher than the moment of inertia of a ring rotating in similar way. The moment of inertia of a solid sphere which has same radius as the disc and rotating in similar way, is n times higher than the moment of inertia of the given ring. Here, n = _____.

Consider all the bodies have equal masses.

Ans. (4)

48. In a measurement, it is asked to find modulus of elasticity per unit torque applied on the system. The measured quantity has dimension of $[M^a L^b T^c]$. If $b = 3$, the value of c is _____

NTA Ans. (4)

ALLEN Ans. (0)

49. Two iron solid discs of negligible thickness have radii R_1 and R_2 and moment of inertia I_1 and I_2 , respectively. For $R_2 = 2R_1$, the ratio of I_1 and I_2 would be $1/x$, where $x =$ _____

Ans. (16)

50. A double slit interference experiment performed with a light of wavelength 600 nm forms an interference fringe pattern on a screen with 10^{th} bright fringe having its centre at a distance of 10 mm from the central maximum. Distance of the centre of the same 10^{th} bright fringe from the central maximum when the source of light is replaced by another source of wavelength 660 nm would be _____ mm.

Ans. (11)

JEE–MAIN EXAMINATION – JANUARY 2025

 (HELD ON TUESDAY 28th JANUARY 2025)

TIME : 9 : 00 AM TO 12 : 00 NOON

CHEMISTRY
SECTION-A
51. The incorrect decreasing order of atomic radii is :

- (1) $Mg > Al > C > O$
- (2) $Al > B > N > F$
- (3) $Be > Mg > Al > Si$
- (4) $Si > P > Cl > F$

Ans. (3)
52. Given below are two statements :

Statement I : In the oxalic acid vs $KMnO_4$ (in the presence of dil H_2SO_4) titration the solution needs to be heated initially to $60^\circ C$, but no heating is required in Ferrous ammonium sulphate (FAS) vs $KMnO_4$ titration (in the presence of dil H_2SO_4)

Statement II : In oxalic acid vs $KMnO_4$ titration, the initial formation of $MnSO_4$ takes place at high temperature, which then acts as catalyst for further reaction. In the case of FAS vs $KMnO_4$, heating oxidizes Fe^{2+} into Fe^{3+} by oxygen of air and error may be introduced in the experiment.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Statement I is false but Statement II is true
- (2) Both Statement I and Statement II are true
- (3) Statement I is true but Statement II is false
- (4) Both Statement I and Statement II are false

Ans. (2)
53. Match the **List-I** with **List-II**

| List-I (Redox Reaction) | | List-II (Type of Redox Reaction) | |
|----------------------------|---|-------------------------------------|-----------------------------|
| A | $CH_4(g) + 2O_2(g) \xrightarrow{\Delta} CO_2(g) + 2H_2O(l)$ | (I) | Disproportionation reaction |

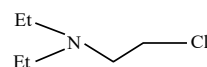
TEST PAPER WITH ANSWER

| | | | |
|---|---|-------|------------------------|
| B | $2NaH(s) \xrightarrow{\Delta} 2Na(s) + H_2(g)$ | (II) | Combination reaction |
| C | $V_2O_5(s) + 5Ca(s) \xrightarrow{\Delta} 2V(s) + 5CaO(s)$ | (III) | Decomposition reaction |
| D | $2H_2O_2(aq) \xrightarrow{\Delta} 2H_2O(l) + O_2(g)$ | (IV) | Displacement reaction |

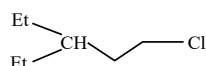
Choose the **correct** answer from the options given below :

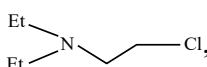
- (1) A-II, B-III, C-IV, D-I
- (2) A-II, B-III, C-I, D-IV
- (3) A-III, B-IV, C-I, D-II
- (4) A-IV, B-I, C-II, D-III

Ans. (1)
54. Given below are two statements :

Statement I :  will undergo

alkaline hydrolysis at a faster rate than



Statement II : , intramolecular

substitution takes place first by involving lone pair of electrons on nitrogen.

In the light of the above statements, choose the **most appropriate** answer from the options given below :

- (1) Both Statement I and Statement II are incorrect
- (2) Statement I is incorrect but statement II is correct
- (3) Both Statement I and Statement II are correct
- (4) Statement I is correct but Statement II is incorrect

Ans. (3)

55. A weak acid HA has degree of dissociation x . Which option gives the correct expression of $\text{pH} = \text{pK}_a$?

- (1) $\log(1 + 2x)$ (2) $\log\left(\frac{1-x}{x}\right)$
 (3) 0 (4) $\log\left(\frac{x}{1-x}\right)$

Ans. (4)

56. Consider 'n' is the number of lone pair of electrons present in the equatorial position of the most stable structure of ClF_3 . The ions from the following with 'n' number of unpaired electrons are :

- A. V^{3+}
 B. Ti^{3+}
 C. Cu^{2+}
 D. Ni^{2+}
 E. Ti^{2+}

Choose the **correct** answer from the options given below :

- (1) A and C only (2) A, D and E only
 (3) B and C only (4) B and D only

Ans. (2)

| $[\text{A}]_0 / \text{molL}^{-1}$ | $t_{1/2} / \text{min}$ |
|-----------------------------------|------------------------|
| 0.100 | 200 |
| 0.025 | 100 |

57.

For a given reaction $\text{R} \rightarrow \text{P}$, $t_{1/2}$ is related to $[\text{A}]_0$ as given in table :

Given : $\log 2 = 0.30$

Which of the following is *true* ?

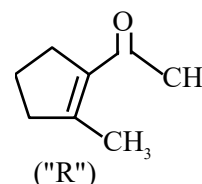
- A. The order of the reaction is $\frac{1}{2}$.
 B. If $[\text{A}]_0$ is 1M, then $t_{1/2}$ is $200\sqrt{10}$ min
 C. The order of the reaction changes to 1 if the concentration of reactant changes from 0.100 M to 0.500 M.
 D. $t_{1/2}$ is 800 min for $[\text{A}]_0 = 1.6$ M

Choose the **correct** answer from the options given below :

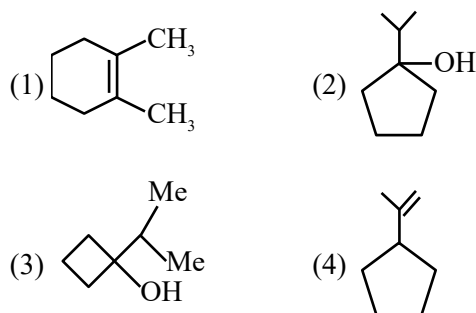
- (1) A and C only (2) A and B only
 (3) A, B and D only (4) C and D only

Ans. (3)

58. A molecule ("P") on treatment with acid undergoes rearrangement and gives ("Q") ("Q") on ozonolysis followed by reflux under alkaline condition gives ("R"). The structure of ("R") is given below :



The structure of ("P") is



Allen Ans. (2 or 4)

NTA Ans. (2)

59. Ice and water are placed in a closed container at a pressure of 1 atm and temperature 273.15 K. If pressure of the system is increased 2 times, keeping temperature constant, then identify correct observation from following :

- (1) Volume of system increases.
 (2) Liquid phase disappears completely.
 (3) The amount of ice decreases.
 (4) The solid phase (ice) disappears completely.

Ans. (4)

60. The molecules having square pyramidal geometry are

- (1) BrF_5 & XeOF_4 (2) SbF_5 & XeOF_4
 (3) SbF_5 & PCl_5 (4) BrF_5 & PCl_5

Ans. (1)

61. The metal ion whose electronic configuration is not affected by the nature of the ligand and which gives a violet colour in non-luminous flame under hot condition in borax bead test is

- (1) Ti^{3+} (2) Ni^{2+}
 (3) Mn^{2+} (4) Cr^{3+}

Ans. (2)

62. Both acetaldehyde and acetone (individually) undergo which of the following reactions?

- A. Iodoform Reaction
 B. Cannizaro Reaction
 C. Aldol condensation
 D. Tollen's Test
 E. Clemmensen Reduction

Choose the **correct** answer from the options given below :

- (1) A, B and D only (2) A, C and E only
 (3) C and E only (4) B, C and D only

Ans. (2)

63. In a multielectron atom, which of the following orbitals described by three quantum numbers with have same energy in absence of electric and magnetic fields?

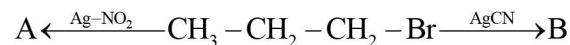
- A. $n = 1, l = 0, m_l = 0$
 B. $n = 2, l = 0, m_l = 0$
 C. $n = 2, l = 1, m_l = 1$
 D. $n = 3, l = 2, m_l = 1$
 E. $n = 3, l = 2, m_l = 0$

Choose the **correct** answer from the options given below :

- (1) A and B only (2) B and C only
 (3) C and D only (4) D and E only

Ans. (4)

64. The products A and B in the following reactions, respectively are



- (1) $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{ONO}$, $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{NC}$
 (2) $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{ONO}$, $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CN}$
 (3) $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{NO}_2$, $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CN}$
 (4) $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{NO}_2$, $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{NC}$

Ans. (4)

65. What is the freezing point depression constant of a solvent, 50 g of which contain 1 g non volatile solute (molar mass 256 g mol^{-1}) and the decrease in freezing point is 0.40 K?

- (1) $5.12 \text{ K kg mol}^{-1}$ (2) $4.43 \text{ K kg mol}^{-1}$
 (3) $1.86 \text{ K kg mol}^{-1}$ (4) $3.72 \text{ K kg mol}^{-1}$

Ans. (1)

66. Consider the following elements In, Tl, Al, Pb, Sn and Ge.

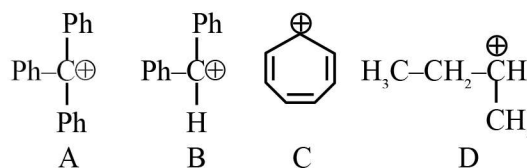
The most stable oxidation states of elements with highest and lowest first ionisation enthalpies, respectively, are

- (1) +2 and +3 (2) +4 and +3
 (3) +4 and +1 (4) +1 and +4

Allen Ans. (2)

NTA Ans. (3)

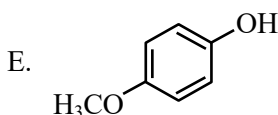
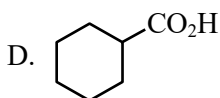
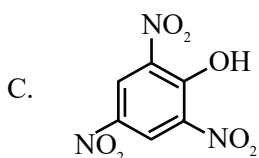
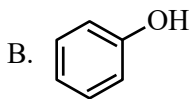
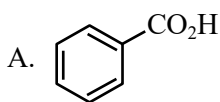
67. The correct order of stability of following carbocations is :



- (1) $A > B > C > D$ (2) $B > C > A > D$
 (3) $C > B > A > D$ (4) $C > A > B > D$

Ans. (4)

68. The compounds that produce CO_2 with aqueous NaHCO_3 solution are :



Choose the **correct** answer from the options given below :

- (1) A and C only (2) A, B and E only
 (3) A, C and D only (4) A and B only

Ans. (3)

69. Which of the following oxidation reactions are carried out by both $\text{K}_2\text{Cr}_2\text{O}_7$ and KMnO_4 in acidic medium ?

- A. $\text{I}^- \rightarrow \text{I}_2$
 B. $\text{S}^{2-} \rightarrow \text{S}$
 C. $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+}$
 D. $\text{I}^- \rightarrow \text{IO}_3^-$
 E. $\text{S}_2\text{O}_3^{2-} \rightarrow \text{SO}_4^{2-}$

Choose the **correct** answer from the options given below :

- (1) B, C and D only (2) A, D and E only
 (3) A, B and C only (4) C, D and E only

Ans. (3)

70. Given below are two statements :

Statement I : D-glucose pentaacetate reacts with 2, 4-dinitrophenylhydrazine.

Statement II : Starch, on heating with concentrated sulfuric acid at 100°C and 2-3 atmosphere pressure produces glucose.

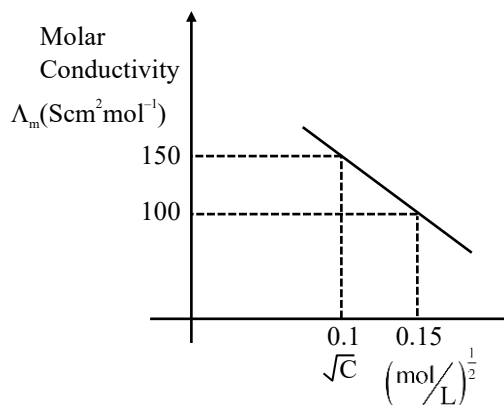
In the light of the above statements, choose the **correct** answer from the options given below

- (1) Both Statement I and Statement II are false
 (2) Statement I is false but Statement II is true
 (3) Statement I is true but Statement II is false
 (4) Both Statement I and Statement II are true

Ans. (2)

SECTION-B

71. Given below is the plot of the molar conductivity vs $\sqrt{\text{concentration}}$ for KCl in aqueous solution.



If, for the higher concentration of KCl solution, the resistance of the conductivity cell is 100Ω , then the resistance of the same cell with the dilute solution is 'x'Ω.

The value of x is _____ (Nearest integer)

Ans. 150

72. Quantitative analysis of an organic compound (X) shows following % composition.

C : 14.5% Cl : 64.46%
 H : 1.8%

(Empirical formula mass of the compound (X) is _____ $\times 10^{-1}$)

(Given molar mass in g mol^{-1} of C : 12, H : 1, O : 16, Cl : 35.5)

Ans. 1655

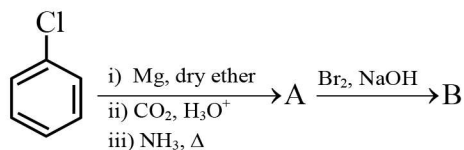
73. The molarity of a 70% (mass/mass) aqueous solution of a monobasic acid (X) is _____ M (Nearest integer)

[Given : Density of aqueous solution of (X) is 1.25 g mL^{-1}

Molar mass of the acid is 70 g mol^{-1}]

Ans. 125

74. Consider the following sequence of reactions :



Chlorobenzene

11.25 mg of chlorobenzene will produce _____ $\times 10^{-1}$ mg of product B.

(Consider the reactions result in complete conversion.)

[Given molar mass of C, H, O, N and Cl as 12, 1, 16, 14 and 35.5 g mol^{-1} respectively]

Ans. 93

75. The formation enthalpies, ΔH_f^\ominus for $\text{H}_{(\text{g})}$ and $\text{O}_{(\text{g})}$ are 220.0 and 250.0 kJ mol^{-1} , respectively, at 298.15 K, and ΔH_f^- for $\text{H}_2\text{O}_{(\text{g})}$ is $-242.0 \text{ kJ mol}^{-1}$ at the same temperature. The average bond enthalpy of the O–H bond in water at 298.15 K is _____ kJ mol^{-1} (nearest integer).

Ans. 466